Chemistry Question Paper

Section A — Short & Definition Questions

- 1. Write three differences between Analytical and Physical Chemistry.
- 2. Write three differences between Atom and Ion.
- 3. Write three differences between Compound and Mixture.
- 4. Define:
 - o Gram Formula Mass
 - Free Radical
 - Avogadro Number (with symbol)
 - o Empirical Formula
 - Molecular Formula
- 5. Define:
 - Ionization Energy
 - Electron Affinity
 - o Electronegativity
- 6. Define isotopes. Give any three applications of isotopes in daily life.
- 7. Give the statements of:
 - o Modern Periodic Law
 - o Mendeleev's Periodic Law
 - o Dobereiner's Law of Triads

Section B — Long / Descriptive Questions

- 8. Describe the discovery of cathode rays through the discharge tube experiment with a labeled diagram. Also write down the properties of cathode rays.
- 9. What is a period? How many periods are there in the modern periodic table? Describe them. Also determine the demarcation of the periodic table into s, p, d and f blocks.
- 10. Define chemistry. Write the names of three branches of chemistry and describe any five branches.

Section C — Chemical Equations, Formulas & Atomic Structure

Q.13) Balance the following chemical equations:

- a) $H_2 + O_2 \rightarrow H_2O$
- b) $N_2 + H_2 \rightarrow NH_3$
- c) $KClO_3 \rightarrow KCl + O_2$
- d) $NH_3 + O_2 \rightarrow NO + H_2O$
- e) NaHCO₃ \rightarrow Na₂CO₃ + H₂O + CO₂

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- (a) Write down the empirical and molecular formula for:
 - Glucose
 - Hydrogen peroxide
 - Benzene
- (b) Write the chemical formulae for:

- Sodium carbonate
- Sodium bicarbonate
- Caustic soda
- Sulfuric acid
- Sugar
- Glucose

Q.17) Find the number of protons, neutrons and electrons present in the following isotopes:

- a) Fe with Z = 26, A = 56
- b) O with Z = 8, A = 17
- c) U with Z = 92, A = 235
- d) C with Z = 6, A = 14

Instructions to Candidates

- Attempt all questions.
- Show all working where applicable.
- Diagrams should be neat and labeled.
- Write answers clearly.