

## 12.2 The Ideal Gas Law

Universal Law: Gas Behavior

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Where:

- $P$  = pressure (Pa)
- $V$  = volume ( $\text{m}^3$ )
- $N$  = number of particles
- $k$  =  $1.38 \times 10^{-23}$  J/K (Boltzmann constant)
- $T$  = absolute temperature (K)

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