Chapter 22: Atomic and Nuclear Physics

The Structure of Matter and Nuclear Processes

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NANMO Physics 12

Winter 2025

Outline

By the end of this section, you will be able to:

- Describe Rutherford's experiment and his model of the atom

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Learning Objectives: 22.1

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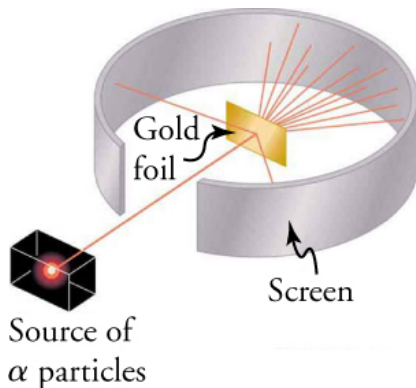- Describe the quantum model of the atom

22.1 Rutherford's Gold Foil Experiment
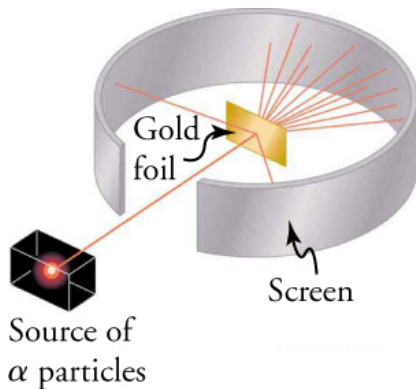
The Setup (1909):

- Radioactive source shoots alpha particles
- Thin gold foil target
- Phosphorescent screen detects scattered particles
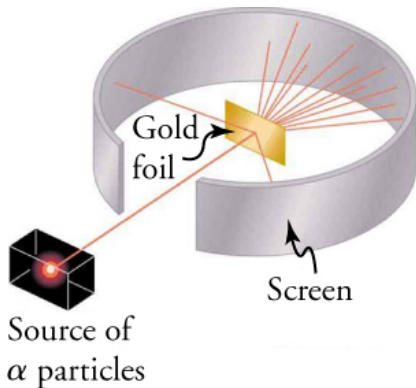
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Observed: Some particles bounced straight back!
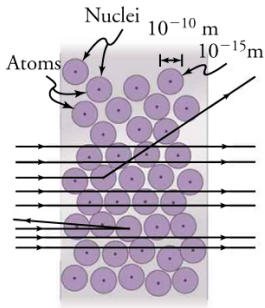
22.1 The Nuclear Atom

Rutherford's Revolutionary Conclusion

The atom has a tiny, dense nucleus
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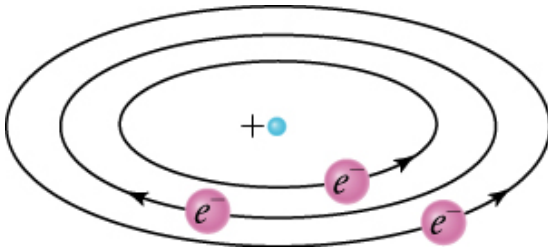
Rutherford's Revolutionary Conclusion

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Key Facts:

- Nucleus: $\sim 10^{-15}$ m
- Atom: $\sim 10^{-10}$ m
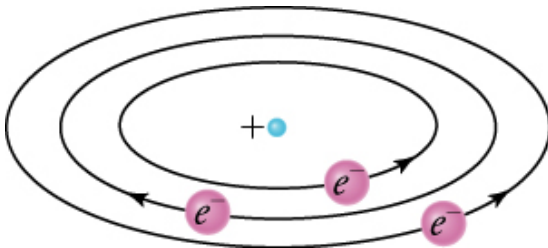- Nucleus is 100,000 times smaller!
- Density: 10^{15} g/cm^3

22.1 The Planetary Model

The Analogy

Low-mass electrons orbit massive nucleus
like planets orbit the Sun
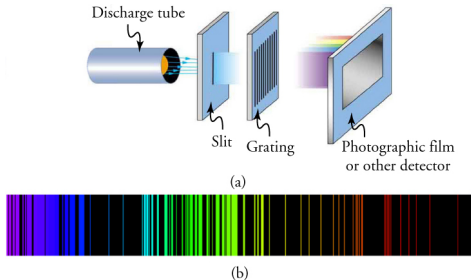
22.1 The Planetary Model

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Low-mass electrons orbit massive nucleus
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The Problem

Classical physics predicts orbiting electrons should radiate energy
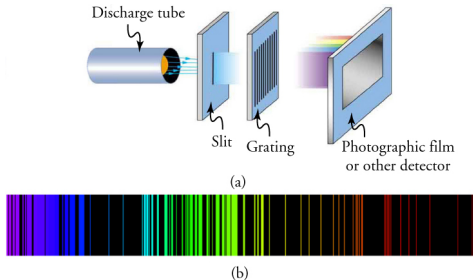and spiral into the nucleus in 10^{-10} seconds!

22.1 Emission and Absorption Spectra

Discharge tube

Slit Grating Photographic film or other detector

(a)

(b)

The Mystery:

- Heat a gas
- It emits discrete wavelengths
- Not continuous spectrum!

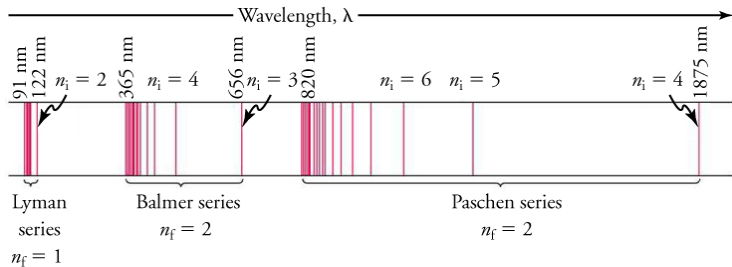# 22.1 Emission and Absorption Spectra

(a)

(b)

The Mystery:

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Each Element Has a Unique Spectral Fingerprint

Iron, hydrogen, helium - all produce different line patterns
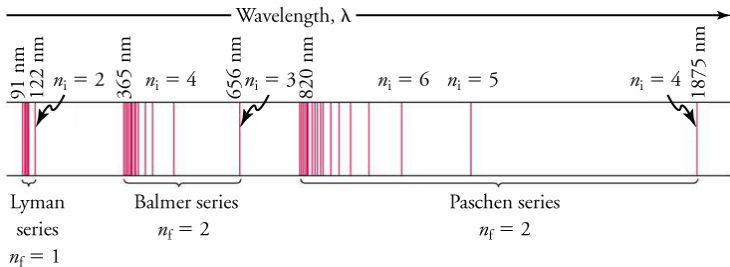
22.1 The Hydrogen Spectrum

Three Series:

- **Lyman series:** Ultraviolet (electrons drop to $n = 1$)
- **Balmer series:** Visible light (electrons drop to $n = 2$)
- **Paschen series:** Infrared (electrons drop to $n = 3$)
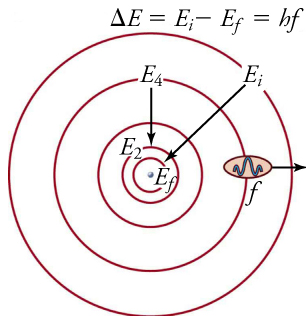
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The Pattern

The wavelengths follow precise mathematical relationships.
But why?

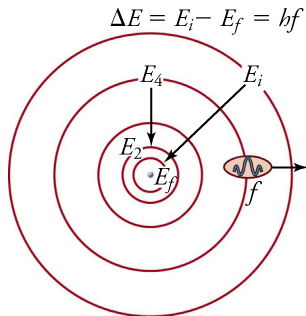# 22.1 Bohr's Quantum Atom (1913)

$$\Delta E = E_i - E_f = hf$$

E_4

E_i

E_2

E_f

f

Bohr's Radical Ideas

1. Only certain orbits allowed (quantized!)
2. Electrons don't radiate while in orbit
3. Energy emitted/absorbed when electron jumps

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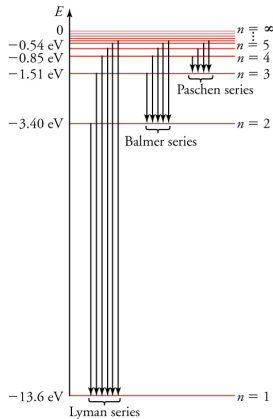### Bohr's Radical Ideas

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Energy Levels for Hydrogen

$$E_n = -\frac{13.6 \text{ eV}}{n^2} \quad (n = 1, 2, 3, \ldots)$$

22.1 Energy Level Diagram

Reading the Diagram

- Arrow down: photon emitted
- Arrow up: photon absorbed
- Longer arrow: higher energy photon

Attempt: Electron Energy Transition

Try this on your own (3 min, silent):
A hydrogen atom absorbs a photon. The electron jumps from the ground state ($n = 1$) to the third energy level ($n = 3$).
Given:

- Initial state: $n_i = 1$
- Final state: $n_f = 3$
- $E_n = -\frac{13.6 \text{ eV}}{n^2}$

Find: How much energy must the photon have?

Work individually. Show your GUESS steps.

Compare: Energy Transition Approach

Turn and talk (2 min):

1. What equation did you use first?
2. Did you get a positive or negative result?
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Name wheel: One pair share your approach (not your answer).

Reveal: Solution

Self-correct in a different color:
U - Unknown: $\Delta E =$?

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S - Substitute:

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$$E_f = -\frac{13.6}{3^2} = -1.51 \text{ eV}$$

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$$\boxed{\Delta E = 12.1 \text{ eV}}$$

Check: Positive means absorption. Photon must provide energy!

22.1 Photon Energy and Wavelength

Energy-Frequency Relationship

$$E = hf = \frac{hc}{\lambda}$$

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Energy-Frequency Relationship

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Where:

- $h = 6.626 \times 10^{-34}$ J·s (Planck's constant)
- $c = 3.00 \times 10^8$ m/s (speed of light)
- $f =$ frequency (Hz)
- $\lambda =$ wavelength (m)

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The Connection

Higher energy transition $\rightarrow$ shorter wavelength photon

22.1 The Rydberg Formula

Wavelength of Emitted Light

$$\frac{1}{\lambda} = R \left(\frac{1}{n_f^2} - \frac{1}{n_i^2} \right)$$

Where:

- $R = 1.097 \times 10^7$ m^{-1} (Rydberg constant)
- $n_i =$ initial quantum number (higher)
- $n_f =$ final quantum number (lower)

22.1 The Rydberg Formula

Wavelength of Emitted Light

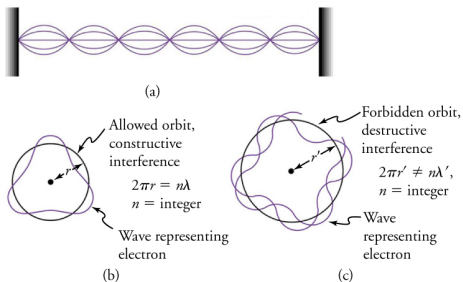$$\frac{1}{\lambda} = R \left(\frac{1}{n_f^2} - \frac{1}{n_i^2} \right)$$

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Historical Note

Rydberg discovered this formula empirically in 1888.
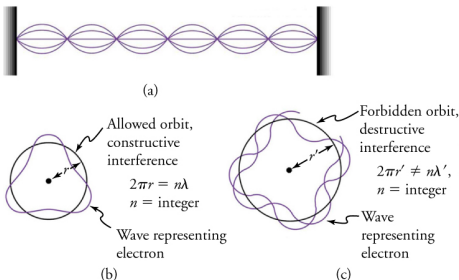Bohr explained why it works in 1913!

22.1 De Broglie's Matter Waves

(a)

Allowed orbit,
constructive
interference

$2\pi r = n\lambda$
n = integer

Wave representing
electron

(b)

Forbidden orbit,
destructive
interference

$2\pi r' \neq n\lambda'$,
n = integer

Wave
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(c)

De Broglie (1924)

If light can be particles,
can particles be waves?

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Yes!

$$\lambda = \frac{h}{p}$$

Only certain wavelengths fit in orbit
$\rightarrow$ quantization!

22.1 Heisenberg Uncertainty Principle

The Fundamental Limit

$$\Delta x \Delta p \geq \frac{h}{4\pi}$$

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What It Means

You cannot simultaneously know exact position AND exact momentum.
This is not a measurement problem - it's reality!

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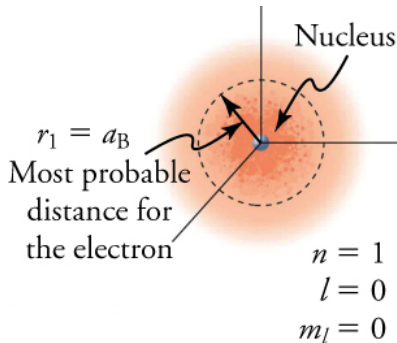$$\Delta x \Delta p \geq \frac{h}{4\pi}$$

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Implication for Atoms

Electrons don't have well-defined orbits.
They exist as probability clouds.

22.1 The Quantum Model

Nucleus

$r_1 = a_B$
Most probable
distance for
the electron

$n = 1$
$l = 0$
$m_l = 0$

Electron Cloud

- Each dot: one position measurement
- Darker region: higher probability
- No defined orbit - only probability distribution

Learning Objectives: 22.2

By the end of this section, you will be able to:

- Describe the structure and forces present within the nucleus

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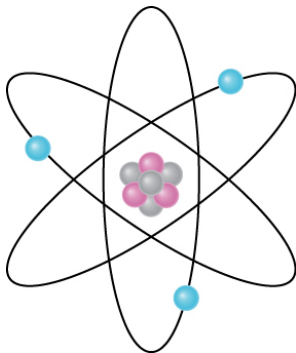- Write nuclear equations associated with radioactive decay
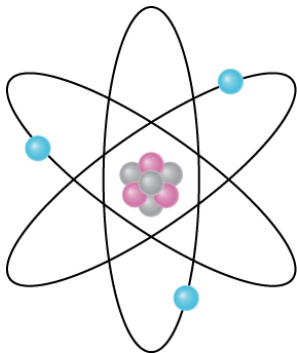
Nucleons:

- Protons: $+1e$ charge
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- Mass ≈ 1 u (atomic mass unit)

22.2 The Nucleus

Nucleons:

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Notation:

$$\,^{A}_{Z}X_{N}$$

- $Z =$ atomic number (protons)
- $A =$ mass number (protons + neutrons)
- $N = A - Z$ (neutrons)

22.2 Isotopes

Same Element, Different Neutrons

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- ^{3}H: 1 proton, 2 neutrons (tritium - radioactive)

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Chemistry vs. Physics

Chemistry: Isotopes behave nearly identically (same electrons)
Physics: Isotopes have very different nuclear stability

22.2 The Strong Nuclear Force

The Problem

Protons repel via Coulomb force.
Why doesn't nucleus explode?

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Protons repel via Coulomb force.
Why doesn't nucleus explode?

The Solution

Strong nuclear force:

- Attractive between nucleons
- Much stronger than EM force
- Very short range ($\sim 10^{-15}$ m)

○ Proton

○ Neutron

22.2 Discovery of Radioactivity (1896)

Becquerel's Accident

- Uranium ore placed on wrapped photographic plate
- Plate darkened - even in complete darkness!
- Uranium emitting invisible, penetrating rays

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Apparent violation of energy conservation!

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Apparent violation of energy conservation!

The Resolution

Einstein's $E = mc^2$ explains it:
Mass converts to energy in nuclear reactions

22.2 Three Types of Nuclear Radiation

Alpha (α) Radiation

2 protons + 2 neutrons (helium nucleus)
Charge: +2e, Penetration: cm in air, stopped by paper

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Charge: −e (or +e), Penetration: m in air, stopped by aluminum

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Beta (β) Radiation

Electron (or positron) from nucleus
Charge: $-e$ (or $+e$), Penetration: m in air, stopped by aluminum

Gamma (γ) Radiation

High-energy photon
Charge: 0, Penetration: km in air, reduced by lead
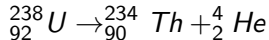
22.2 Alpha Decay
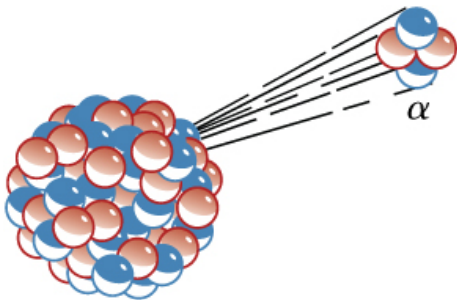
Nuclear Equation
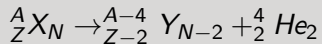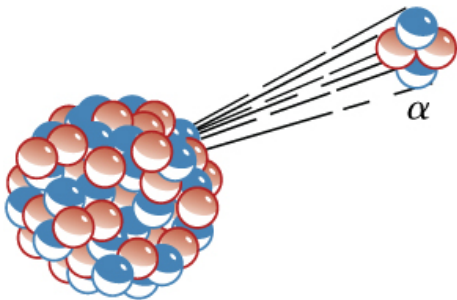
$$^{A}_{Z}X_{N} \rightarrow\, ^{A-4}_{Z-2}Y_{N-2} + \,^{4}_{2}He_{2}$$

22.2 Alpha Decay

Nuclear Equation

$$^{A}_{Z}X_N \rightarrow\, ^{A-4}_{Z-2}Y_{N-2} +\, ^{4}_{2}He_2$$

Example: Uranium-238

$$^{238}_{92}U \rightarrow\, ^{234}_{90}Th +\, ^{4}_{2}He$$

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Nuclear Equation

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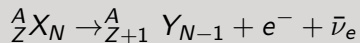$$^{238}_{92} U \rightarrow ^{234}_{90} Th + ^4_2 He$$
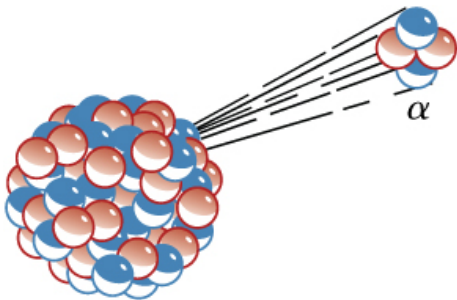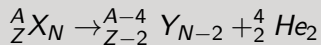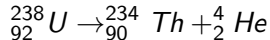
Conservation:

- Mass number: $238 = 234 + 4$ ✓
- Charge: $92 = 90 + 2$ ✓

22.2 Beta Decay
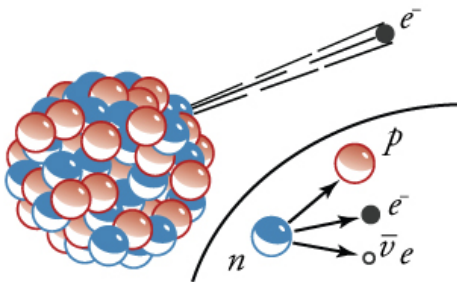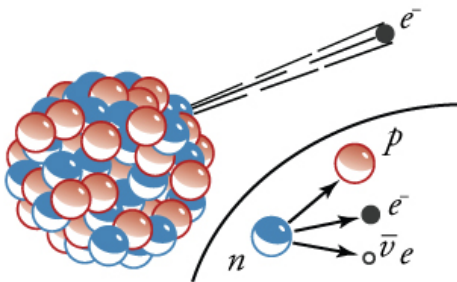
β^- Decay (most common)

Neutron $\rightarrow$ proton + electron + antineutrino

$$^A_Z X_N \rightarrow\, ^A_{Z+1} Y_{N-1} + e^- + \bar{\nu}_e$$

22.2 Beta Decay

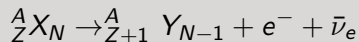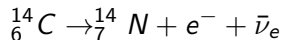## β^- Decay (most common)

Neutron $\rightarrow$ proton + electron + antineutrino

$$\,_Z^A X_N \rightarrow\,_{Z+1}^A Y_{N-1} + e^- + \bar{\nu}_e$$

Example: Carbon-14

$$\,_6^{14} C \rightarrow\,_7^{14} N + e^- + \bar{\nu}_e$$

22.2 Gamma Decay

Excited Nucleus

Nucleus drops from excited state to ground state

$$_{Z}^{A}X^* \rightarrow _{Z}^{A}X + \gamma$$

γ

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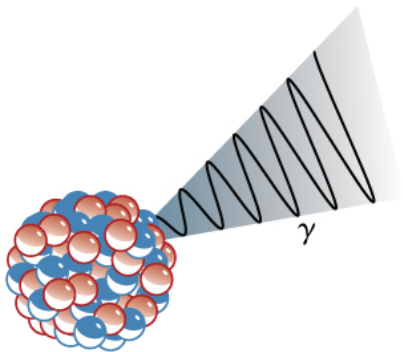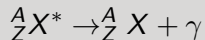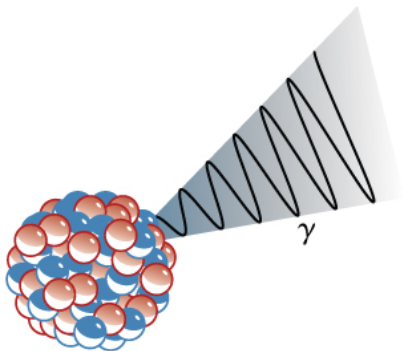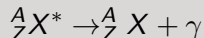
Key Points:

- A and Z unchanged
- Pure energy release
- Often follows α or β decay
- MeV energies (vs. eV for atoms)

Attempt: Write Nuclear Equation

Try this on your own (3 min, silent):
Plutonium-239 undergoes alpha decay.
Given:

- Parent nucleus: $^{239}_{94}Pu$
- Type: alpha decay
- Periodic table for atomic numbers

Find:

1. Complete nuclear equation
2. Identity of daughter nucleus

Remember to conserve mass number and charge!

Compare: Nuclear Equation Strategy

Turn and talk (2 min):

1. What did you subtract from A and Z?
2. How did you identify the daughter element?
3. Did you check conservation?

Compare: Nuclear Equation Strategy
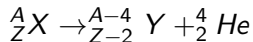
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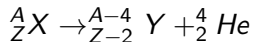# Reveal: Alpha Decay Solution

Self-correct in a different color:
General form:

$$_{Z}^{A}X \rightarrow _{Z-2}^{A-4}Y + _{2}^{4}He$$

Self-correct in a different color:
General form:

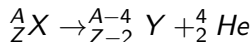$$_{Z}^{A}X \rightarrow _{Z-2}^{A-4} Y + _{2}^{4} He$$

Apply to Pu-239:

$$A_{daughter} = 239 - 4 = 235$$

$$Z_{daughter} = 94 - 2 = 92$$

Reveal: Alpha Decay Solution

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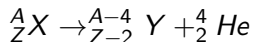
Apply to Pu-239:

$$A_{daughter} = 239 - 4 = 235$$

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Identify element: $Z = 92$ is Uranium (U)

Reveal: Alpha Decay Solution

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General form:
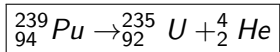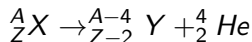
$$^A_Z X \rightarrow ^{A-4}_{Z-2} Y + ^4_2 He$$

Apply to Pu-239:

$$A_{daughter} = 239 - 4 = 235$$

$$Z_{daughter} = 94 - 2 = 92$$

Identify element: $Z = 92$ is Uranium (U)

$$\boxed{^{239}_{94} Pu \rightarrow ^{235}_{92} U + ^4_2 He}$$

Reveal: Alpha Decay Solution

Self-correct in a different color:
General form:

$$\ _{Z}^{A}X \rightarrow\ _{Z-2}^{A-4}Y +\ _{2}^{4}He$$

Apply to Pu-239:

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Identify element: $Z = 92$ is Uranium (U)

$$\boxed{\ _{94}^{239}Pu \rightarrow\ _{92}^{235}U +\ _{2}^{4}He}$$
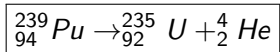
Check: $239 = 235 + 4 \checkmark$, $94 = 92 + 2 \checkmark$

Learning Objectives: 22.3

By the end of this section, you will be able to:
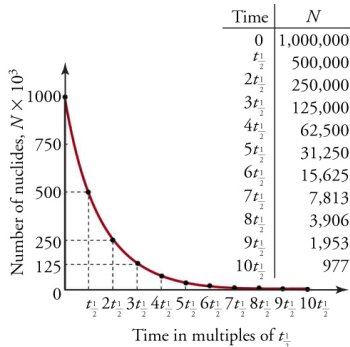
- Explain radioactive half-life and its role in radiometric dating

By the end of this section, you will be able to:

- Explain radioactive half-life and its role in radiometric dating
- Calculate radioactive half-life and solve problems associated with radiometric dating

22.3 What is Half-Life?

Time	N
0	1,000,000
$t_{\frac{1}{2}}$	500,000
$2t_{\frac{1}{2}}$	250,000
$3t_{\frac{1}{2}}$	125,000
$4t_{\frac{1}{2}}$	62,500
$5t_{\frac{1}{2}}$	31,250
$6t_{\frac{1}{2}}$	15,625
$7t_{\frac{1}{2}}$	7,813
$8t_{\frac{1}{2}}$	3,906
$9t_{\frac{1}{2}}$	1,953
$10t_{\frac{1}{2}}$	977

Definition

Half-life ($t_{1/2}$): Time for half the nuclei to decay

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$9t_{\frac{1}{2}}$	1,953
$10t_{\frac{1}{2}}$	977

Time in multiples of $t_{\frac{1}{2}}$
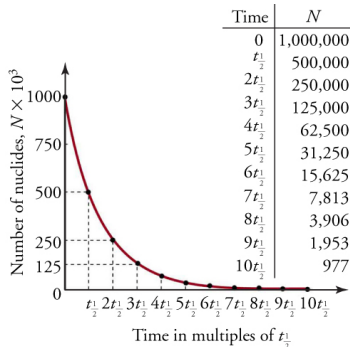
Definition

Half-life ($t_{1/2}$): Time for half the nuclei to decay

- After $t_{1/2}$: $N \to N/2$
- After $2t_{1/2}$: $N \to N/4$
- After $3t_{1/2}$: $N \to N/8$

22.3 Exponential Decay Law

Number of Nuclei vs. Time

$$N(t) = N_0 e^{-\lambda t}$$

Where:

- $N_0 =$ initial number of nuclei
- $N(t) =$ number remaining at time t
- $\lambda =$ decay constant
- $e = 2.71828\ldots$

22.3 Exponential Decay Law

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Where:

- $N_0 =$ initial number of nuclei
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- $\lambda =$ decay constant
- $e = 2.71828\ldots$

Relationship to Half-Life

$$\lambda = \frac{\ln(2)}{t_{1/2}} \approx \frac{0.693}{t_{1/2}}$$

22.3 Activity (Rate of Decay)

Definition

Activity R: Number of decays per unit time

$$R = \frac{\Delta N}{\Delta t} = \lambda N$$

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Units:

- Becquerel (Bq): 1 decay/second (SI unit)
- Curie (Ci): 3.7×10^{10} decays/second (traditional)

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- Becquerel (Bq): 1 decay/second (SI unit)
- Curie (Ci): 3.7×10^{10} decays/second (traditional)

Physical Meaning

More radioactive material $\rightarrow$ higher activity
As sample decays $\rightarrow$ activity decreases

22.3 Carbon-14 Dating

The Method

1. Cosmic rays create ^{14}C in atmosphere
2. Living organisms maintain constant ^{14}C ratio
3. After death: no new ^{14}C absorbed
4. ^{14}C decays with $t_{1/2} = 5,730$ years
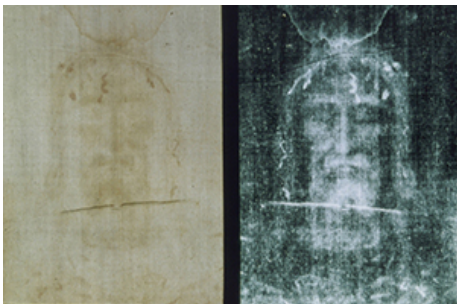5. Measure remaining ^{14}C to find age

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5. Measure remaining ^{14}C to find age

Range

Effective for ages: 100 to 50,000 years
(about 10 half-lives maximum)
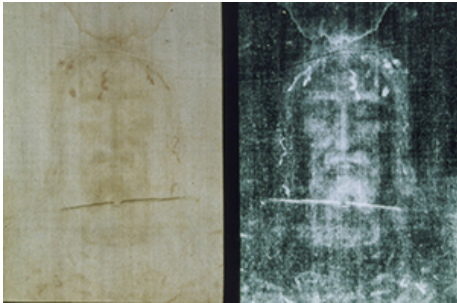
The Claim: Burial shroud of Jesus (33 CE)

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The Test (1988):

- Three independent labs
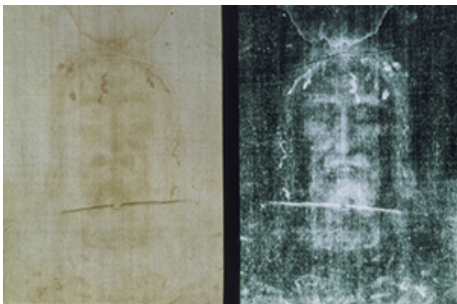- Found 92% of living ^{14}C
- Calculate age...

The Claim: Burial shroud of Jesus (33 CE)

The Test (1988):

- Three independent labs
- Found 92% of living ^{14}C
- Calculate age...

The Result: Dated to 1320 ± 60 CE

Medieval, not ancient!

Attempt: Half-Life Calculation

Try this on your own (3 min, silent):

A radioactive sample has a half-life of 10 days. You start with 80 g of material.

Given:

- $N_0 = 80$ g
- $t_{1/2} = 10$ days
- Time elapsed: $t = 30$ days

Find: How much material remains after 30 days?

Think: How many half-lives have passed?

Turn and talk (2 min):

1. How many half-lives occurred?
2. Did you use repeated halving or a formula?
3. What's your final answer?

Compare: Half-Life Strategy

Turn and talk (2 min):

1. How many half-lives occurred?
2. Did you use repeated halving or a formula?
3. What's your final answer?

Name wheel: Share your approach.

Self-correct in a different color:
Method 1 - Counting Half-Lives:
Number of half-lives: $n = \frac{t}{t_{1/2}} = \frac{30}{10} = 3$

Reveal: Half-Life Solution

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Method 1 - Counting Half-Lives:
Number of half-lives: $n = \frac{t}{t_{1/2}} = \frac{30}{10} = 3$

After each half-life:

- $t = 10$ days: $80 \nabla \cdot 2 = 40$ g
- $t = 20$ days: $40 \nabla \cdot 2 = 20$ g
- $t = 30$ days: $20 \nabla \cdot 2 = 10$ g

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Method 2 - Formula:

$$N = N_0 \left(\frac{1}{2}\right)^n = 80 \left(\frac{1}{2}\right)^3 = 80 \times \frac{1}{8} = 10 \text{ g}$$

Reveal: Half-Life Solution

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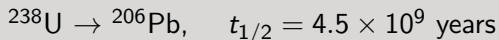$$\boxed{N = 10 \text{ g}}$$

22.3 Other Radiometric Dating

Uranium-Lead Dating

$^{238}U \rightarrow {}^{206}Pb, \quad t_{1/2} = 4.5 \times 10^9$ years
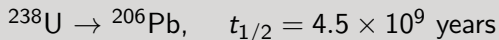
Used for ancient rocks (oldest Earth rocks: 3.5 billion years)

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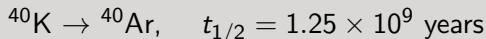
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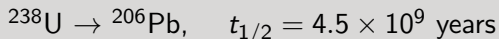
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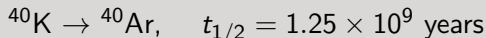
Used for volcanic rocks, human fossils

The Power

Different isotopes cover different time scales:

Years, millennia, millions of years, billions of years

Learning Objectives: 22.4

By the end of this section, you will be able to:

- Explain nuclear fission

Learning Objectives: 22.4
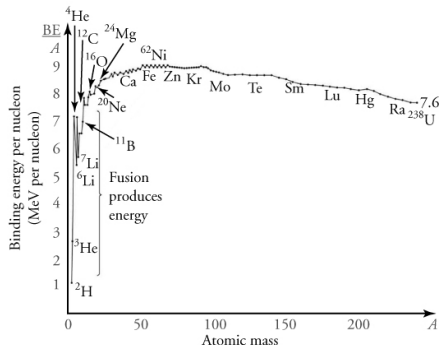
By the end of this section, you will be able to:

- Explain nuclear fission
- Explain nuclear fusion

By the end of this section, you will be able to:

- Explain nuclear fission

- Explain nuclear fusion

- Describe how fission and fusion work in weapons and power generation
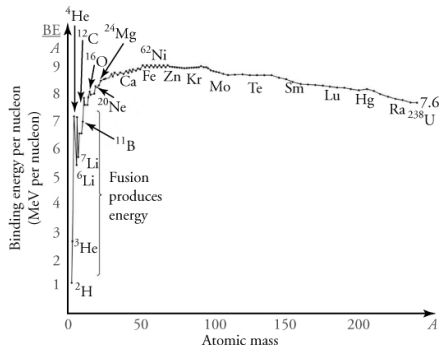
22.4 Nuclear Binding Energy

The Key Insight

Iron-56 has highest binding energy per nucleon
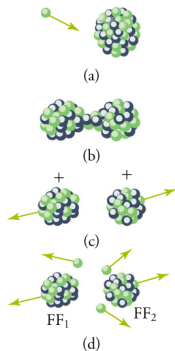$\rightarrow$ Most stable nucleus

22.4 Nuclear Binding Energy

The Key Insight

Iron-56 has highest binding energy per nucleon
$\rightarrow$ Most stable nucleus

- Heavy nuclei (right of Fe): release energy by **fission**
- Light nuclei (left of Fe): release energy by **fusion**

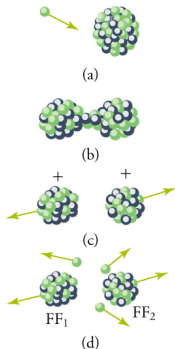# 22.4 Nuclear Fission

(a)

(b)

(c)

(d)

FF_1 FF_2

The Process

1. Neutron strikes heavy nucleus
2. Nucleus elongates
3. EM repulsion overcomes strong force
4. Nucleus splits
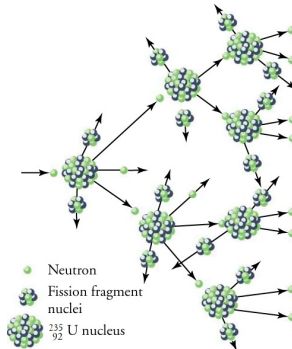5. Releases energy + more neutrons

22.4 Nuclear Fission

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FF_1 FF_2

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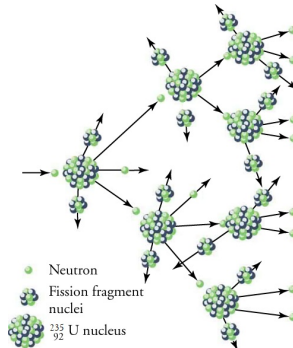**Typical energy:** 200 MeV per fission (vs. 1 eV per chemical reaction)

22.4 Chain Reaction

Neutron
Fission fragment nuclei
$^{235}_{92}$U nucleus

Self-Sustaining Fission

Each fission releases 2-3 neutrons
$\rightarrow$ Those neutrons cause more fissions
$\rightarrow$ Exponential growth

22.4 Chain Reaction

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Critical Mass

22.4 Nuclear Fission Reactor
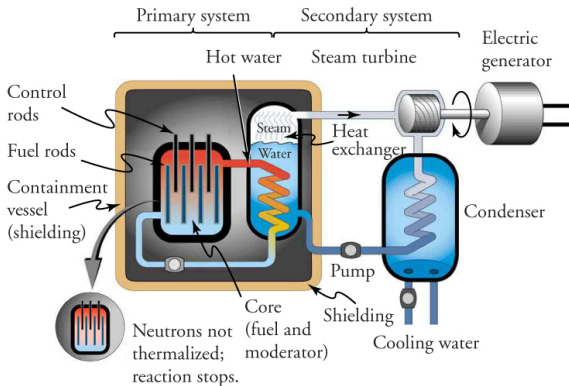
Key Components:

- Fuel rods: enriched ^{235}U
- Moderator: water (slows neutrons)
- Control rods: absorb excess neutrons
- Heat exchanger: produces steam $\rightarrow$ turbine $\rightarrow$ electricity

Einstein's Equation

$$E = mc^2$$

22.4 Mass-Energy Conversion

Einstein's Equation

$$E = mc^2$$

In fission:

- Products have less mass than reactants
- Missing mass $\rightarrow$ converted to energy
- $\Delta m \approx 0.1\%$ of original mass
- Still liberates enormous energy

22.4 Mass-Energy Conversion

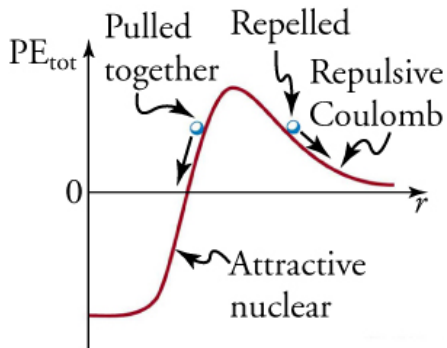### Einstein's Equation

$$E = mc^2$$

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Example

1 kg of ^{235}U fully fissioned:
$E = 8.2 \times 10^{13}$ J $\approx$ 14,000 barrels of oil!
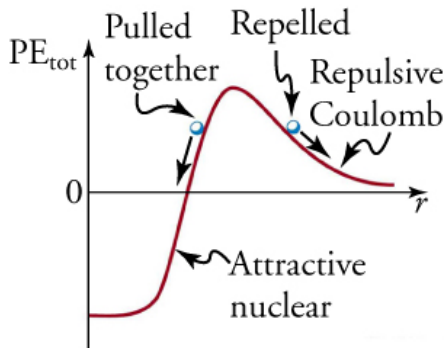
22.4 Nuclear Fusion

The Process

Combine light nuclei
$\rightarrow$ Overcome Coulomb repulsion
$\rightarrow$ Strong force binds them
$\rightarrow$ Release energy

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Requirements:

- Extreme temperature ($\sim 10^7$ K)
- Extreme pressure
- $\rightarrow$ Found in star cores!

22.4 The Proton-Proton Cycle
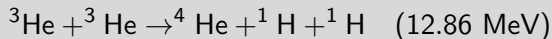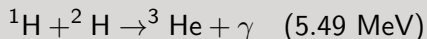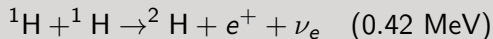
How the Sun Fuses Hydrogen

$$^1\text{H} + {}^1\text{H} \rightarrow {}^2\text{H} + e^+ + \nu_e \quad (0.42 \text{ MeV})$$

$$^1\text{H} + {}^2\text{H} \rightarrow {}^3\text{He} + \gamma \quad (5.49 \text{ MeV})$$

$$^3\text{He} + {}^3\text{He} \rightarrow {}^4\text{He} + {}^1\text{H} + {}^1\text{H} \quad (12.86 \text{ MeV})$$

22.4 The Proton-Proton Cycle

How the Sun Fuses Hydrogen

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Net result:

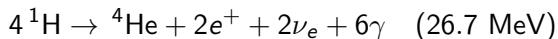$$4\,^1\text{H} \rightarrow\ ^4\text{He} + 2e^+ + 2\nu_e + 6\gamma \quad (26.7 \text{ MeV})$$

22.4 The Proton-Proton Cycle

How the Sun Fuses Hydrogen

$$^1\text{H} +^1 \text{H} \rightarrow^2 \text{H} + e^+ + \nu_e \quad (0.42 \text{ MeV})$$
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Net result:

$$4\,^1\text{H} \rightarrow\,^4\text{He} + 2e^+ + 2\nu_e + 6\gamma \quad (26.7 \text{ MeV})$$
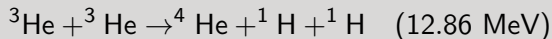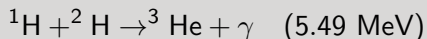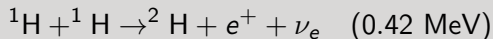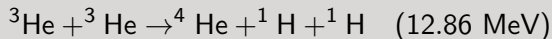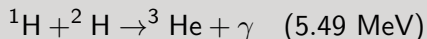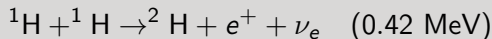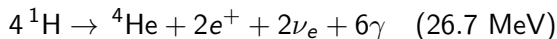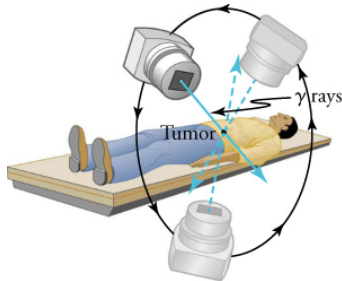
The Wonder

This energy took 32,000 years to reach Sun's surface
Then 8 minutes to reach Earth!

22.4 Fusion Energy Potential

The Promise

- Fuel: deuterium from seawater (virtually unlimited)
- No chain reaction $\rightarrow$ inherently safe
- No long-lived radioactive waste
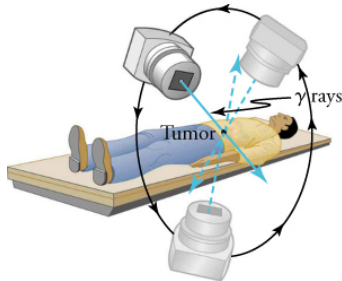- 4x more energy per kg than fission

22.4 Fusion Energy Potential

The Promise

- Fuel: deuterium from seawater (virtually unlimited)
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- No long-lived radioactive waste
- 4x more energy per kg than fission

The Challenge

Confine 100 million degrees plasma long enough for net energy gain

22.4 Fission vs. Fusion Summary

Fission	Fusion
Split heavy nuclei	**Combine** light nuclei
Fuel: ^{235}U, ^{239}Pu (rare)	**Fuel:** H isotopes (abundant)
Products: Radioactive waste	**Products:** Helium (inert)
Status: Mature technology	**Status:** Experimental
Use: Power plants, weapons	**Use:** Stars, H-bombs, (future power?)

Summary: Chapter 22 (Sections 1-4)

22.1 Atomic Structure

Rutherford $\rightarrow$ Bohr $\rightarrow$ Quantum model
Discrete energy levels explain emission spectra

22.2 Nuclear Forces and Radioactivity

Strong force holds nucleus together
Alpha, beta, gamma decay restore stability

22.3 Half-Life

Exponential decay enables radiometric dating
Carbon-14, U-238 date different timescales

22.4 Fission and Fusion

Binding energy curve: both release enormous energy
Fission mature, fusion promising

Key Equations: Chapter 22

Atomic Structure

$$E_n = -\frac{13.6 \text{ eV}}{n^2}, \quad \Delta E = E_f - E_i, \quad E = hf = \frac{hc}{\lambda}$$

$$\frac{1}{\lambda} = R\left(\frac{1}{n_f^2} - \frac{1}{n_i^2}\right), \quad \Delta x \Delta p \geq \frac{h}{4\pi}$$

Nuclear Physics

$$E = mc^2$$

Radioactive Decay

$$N(t) = N_0 e^{-\lambda t}, \quad \lambda = \frac{0.693}{t_{1/2}}, \quad R = \lambda N$$

Nuclear Notation

$$^{A}_{Z}X_N \quad (A = Z + N)$$

Complete the assigned problems

posted on the LMS

Temporary page!

LaTeX was unable to guess the total number of pages correctly. was some unprocessed data that should have been added to the this extra page has been added to receive it.

If you rerun the document (without altering it) this surplus page away, because LaTeX now knows how many pages to expect for document.