

Effective and Complete Decomposition of Per- and Polyfluoroalkyl¹ Substances and Byproducts in Incineration

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Executive Summary¹

The study, led by Dr. Feng ‘Frank’ Xiao from the University of Missouri, investigates various PFAS compounds under different incineration conditions to understand their stability and decomposition mechanisms better. This research is timely as it addresses the persistent challenge posed by per- and polyfluoroalkyl substances (PFAS) in municipal solid wastes, particularly in groundwater and soil impacted by industrial activities and waste disposal practices.²

The report details the conditions under which complete degradation of PFAS occurs,³ emphasizing the temperature-time profiles that achieve high degradation efficiencies. For example, higher temperatures (up to 1000 °C) significantly expedite the degradation process, with all studied PFAS achieving nearly total decomposition within minutes, if not seconds. Findings of this kind are crucial for the development of more effective and efficient waste treatment technologies that can operate at optimal temperatures, thus minimizing the environmental impact of PFAS. The study provides a scientific basis for these operational improvements and also offers a benchmark for regulatory standards concerning the thermal treatment of PFAS.

Moreover, we focused on the post-treatment of products of incomplete degradation (PIPs) of PFAS, in particular perfluoroalkanes and perfluoroalkenes. We demonstrated that near-complete degradation ($\geq 99\%$) of unsaturated perfluorocarbons (PFCs) like C₇F₁₄ and C₈F₁₆ is achievable at temperatures above 300°C, forming organofluorine compounds as the dominant pathway, whereas saturated PFCs require temperatures exceeding 500°C for similar results.⁴ The use of granular activated carbon significantly enhances the degradation efficiency, particularly at lower temperatures, by influencing the pathways and increasing the yield of fluorine. By integrating these findings into their operations, facilities can achieve higher degradation efficiencies, thereby reducing the long-term environmental footprint of PFAS disposal.

Based on the results of this project, incineration appears to be a promising approach for PFAS waste management. However, it is important to note that existing incineration systems were not specifically designed to handle PFAS-containing waste. Our previous studies have shown that certain PFAS compounds can become volatile at moderate temperatures. Therefore, suboptimal incinerator operation may lead to the release of PFAS from the system. Sustainability and efficiency depend on addressing PIPs, potentially by implementing emission control systems to capture fugitive emissions, though such measures would require a thorough cost-benefit analysis and appropriate regulatory support.⁵

In conclusion, the report highlights the potential applications of its findings in improving solid waste management practices through the adoption of scientifically backed thermal treatment strategies. By providing a comprehensive analysis of PFAS degradation under various conditions, it equips stakeholders with the knowledge to implement more effective waste treatment solutions aimed at limiting the environmental circulation of PFAS, thereby contributing to environmental protection and public health. The report advocates for continued research into PFAS treatment technologies, underscoring the importance of innovation in tackling one of the most persistent and challenging pollutants in modern waste management.⁶

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INTRODUCTION ¹

The U.S. Environmental Protection Agency (EPA) has regulated a few per- and polyfluoroalkyl substances (PFAS), including perfluorooctanoic acid (PFOA) and perfluorooctane sulfonic acid (PFOS) in under the Safe Drinking Water Act.¹ Elevated levels (several tens of µg/L) of PFAS have been found in groundwater at a number of locations in the United States.²⁻¹⁰ Groundwater contamination of PFAS is directly associated with surface soil contamination.^{2-6, 10} The major contamination sources include fire-training sites using aqueous film-forming foams (AFFFs)⁴ and municipal solid waste (MSW) landfills.¹¹⁻¹³ Concentrations of PFOA/PFOS as high as 87,140 ng/L have been measured in landfill leachates.¹¹⁻¹³ PFAS can move off-site from contamination sources (landfills, AFFF-applied sites, or biosolids) and migrate to aquifers via leaching and surface water via runoff. PFAS-laden MSW pose considerable challenges to established waste disposal and recycling practices.

Thermal treatment is a promising approach for PFAS management in solid waste. The studies of PI Xiao and his group are the earliest attempts to investigate PFAS thermal stability and degradation mechanisms.¹⁴⁻²⁷ The temperature–time profile has a significant impact on PFAS degradation. The thermal mineralization rate of PFOA and PFOS can reach to 87–92 mol% at 700–800 °C. At a lower temperature (500 °C), a brief treatment for 7 min resulted in 45–55% degradation of all perfluoroalkyl sulfonates (PFSAs) in soil, while a treatment time of 15 min increased the extent of degradation to 94–98%.²⁸ A small portion of PFOS, however, continued to decompose at a much slower rate.²⁸ The mass loss followed first-order kinetics. The apparent thermal decomposition half-lives of PFSAs in soil were 28.7–51.8 min and 2.4–3.1 min, respectively, at 400 °C and 500 °C.²⁸ Highly porous materials can alter the thermal decomposition pathway of perfluoroalkyl carboxylates (PFCAs) (e.g., PFOA) and GenX and substantially accelerate their decomposition at low and moderate temperatures (<400 °C).¹⁵ In this research, we hypothesize that half-lives of PFAS, including PFOA and PFOS, in MSW may be in the range of several seconds to one minute at a typical incineration temperature (e.g., ~900 °C). This hypothesis will be tested in this research.

In addition to our studies, Strathmann at the Colorado School of Mines and co-workers developed a hydrothermal treatment system,²⁹⁻³⁰ in which water at a high temperature and pressure is used to catalyze hydrolytic reactions of PFAS. Crownover et al. examined the decontamination of soils containing perfluoroalkyl substances using thermal desorption³¹ by heating the contaminated soil to a specific temperature to vaporize the contaminants. The goal of thermal desorption, however, is simply to separate PFAS from soil, and subsequent treatment of vaporized PFAS molecules is still needed. Gerhard at the University of Western Ontario and Major at Geosyntec Consultants investigated thermal degradation of PFAS in soil by smoldering combustion.³²⁻³³ Dastgheib et al. investigated PFAS degradation on spent resins by means of thermogravimetric analysis to simulate the incineration process.³⁴ Wang et al. suggested direct F elimination from PFCAs during thermal treatments.³⁵ Blotevogel et al. found that the primary decomposition mechanism of PFOA is C-COOH cleavage, whereas α-elimination predominates in the breakdown of hexafluoropropylene oxide dimer acid (HFPO-DA).³⁶ Using density functional theory, Adi and Altarawneh explored various potential thermal decomposition pathways for HFPO-DA.³⁷ Kennedy and colleagues conducted a series of experiments via thermogravimetric analysis coupled with FTIR (TGA-FTIR) for the simultaneous and continuous real-time analysis of the volatile products of incomplete combustion (PIPs) of PFOA and PFOS (the acid form).³⁸⁻⁴¹ Hughey et al. employed sampling bags to capture gaseous PIPs of K-PFOS, which were

subsequently analyzed by FTIR.⁴² Granular activated carbon,¹⁵ pure platinum (Pt),³⁵ alumina (Al_2O_3),⁴³ and alkali and alkaline-earth metal additives^{30, 43-46} have shown to improve the thermal degradation of PFAS while reducing the production of PIDs. 1

Despite the insights gained from previous studies, some fundamental issues have not been properly addressed. We are not aware of systematic studies on PFAS degradation and mineralization in MSW incineration processes. The PFAS incineration mechanisms and thermokinetics are unclear, and the effects of the combustion temperature–time profile, the thermokinetic control, and the identity of decomposition products remain largely elusive. It is imperative to obtain a comprehensive understanding of thermal PFAS decomposition mechanisms (pathways, products, and thermodynamics) to facilitate the industrialization of incineration processes. 2

Furthermore, post-treatment may be needed for further treatment of PIDs. In one of our recent studies, ultra-short-chain perfluoroalkanes of significant concern,⁴⁷ such as perfluoromethane and perfluoroethane, were not observed in the exhaust gas during thermal treatment of PFOA and PFOS.²⁸ Perfluoroalkenes, in particular *perfluoroheptene*, are main PIDs identified in pyrolysis processes of PFOA, PFOS, cationic/zwitterionic PFAS, and AFFF samples at moderate temperatures. We further found that *perfluoroheptene* thermally decomposed at temperatures as low as 200 °C to shorter-chain homologues following a radical chain-scission mechanism.¹⁴ An afterburner temperature of 890 °C appears to be effective to decompose fugitive PFAS and volatile degradation products.¹⁴ The exhaust gas will be passed through a column filled with CaO⁴⁴ or kaolinite²⁸ to remove HF generated from both tube furnaces. 3

The goal of this project is to develop effective and complete destruction technologies for PFAS in incineration and fuel blending operations. We have determined the incineration conditions (temperature–time profiles, air flow, PFAS loading, and air–fuel ratio) to achieve a destruction and removal efficiency (DRE) of 99–99.999% for PFAS in municipal solid waste (MSW); optimized the incineration conditions to achieve maximum mineralization of PFAS; identified incineration pathways and end-products of PFAS; and established a post-treatment train to remove and degrade volatile incineration products. We also tested other (sensor) approaches for measuring PIDs (co-PI Zhang at New Jersey Institute of Technology). 4

RESULTS AND DISCUSSION ¹

Chapter 1. Thermal Destruction of PFAS at Incineration Temperatures ²

1.1. Incineration of PFCAs ³

This experiment investigated the incineration of a series of perfluorocarboxylic acids (PFCAs) – PFBA (C4), PFPeA (C5), PFHpA (C6), PFHxA (C7), PFOA (C8), PFNA (C9), PFDA (C10), and PFUnDA (C11) – different incineration temperatures over varying durations of time (Figure 1). The observed degradation percentages were plotted against time to track the progression of the thermal degradation process. ⁴

From the onset of heating, a discernible trend was evident across all PFCAs at 600 °C. At the 2-min mark, degradation rates ranged from 15.3% for PFHpA to 30.62% for PFBA. PFBA, with a shorter chain length, displayed the highest degradation, emphasizing the impact of chain length on PFCA stability. PFBA, which is much more volatile than its homologs²³ included in this research, also showed the highest degradation rate (Figure 1). By the 5-min mark, degradation rates had increased significantly, ranging from 36% for PFUnDA to 51.3% for PFHpA. The relative order of degradation rates had shifted, suggesting different rates of reaction for different PFCAs as the process progressed. At 10 min, a majority of PFCAs exhibited degradation rates above 70%. The least degraded PFCA at this point was PFNA, with 71.4% degradation. By the 15-min mark, nearly all PFCAs nearly complete (>99%) degradation. The exception was PFUnDA, which was 85%. The PFCAs exhibited more than 95% degradation, highlighting their vulnerability to prolonged heat exposure. Finally, at the 20 and 30-min marks, almost all PFCAs reached near or total degradation, displaying the potency of extended heat exposure in driving complete PFCA degradation. ⁵

When the temperature was raised to 700 °C, a markedly quicker degradation process was noted. By the 2-min mark, degradation levels were generally higher than the corresponding values at 600 °C. For instance, PFBA demonstrated 60.5% degradation at 2 min, compared to 30.62% at the same time interval at 600 °C. It's clear from these results that the increase in temperature had a significant impact on the speed and extent of PFCA degradation. At 800 °C, the rate of degradation became even more rapid. By just 2 min, degradation percentages were already at around 80–90% for most PFCAs, with some compounds such as PFHpA and PFOA demonstrating nearly complete degradation ($\geq 99\%$). It appears that 800 °C is a critical temperature threshold where almost complete thermal degradation of PFCAs can be achieved in less than 5 min. The trend of accelerated degradation with increased temperature continued as we observed the data for 900 °C (Figure 1). Almost all PFCAs exhibited near-total or total degradation at the 2-min mark. After this point, degradation rates appeared to stabilize, showing no significant change up to 30 min. This suggests that by 900 °C, nearly all PFCAs had been degraded almost completely within just a few min. Finally, at 1000 °C, it was observed that all PFCAs achieved complete thermal degradation within the first 2 min. In summary, the data suggests a clear temperature-dependent trend in the thermal degradation of PFCAs. As temperature increases from 600 °C to 1000 °C, there is a significantly faster degradation process. At the higher temperatures of 800 °C, 900 °C, and 1000 °C, the PFCAs undergo near to total degradation within a matter of min, with 1000 °C ensuring complete degradation in just 2 min. This detailed investigation provides valuable insights into the thermal stability of PFCAs, potentially guiding future research and management strategies related to these environmentally persistent compounds. ⁶

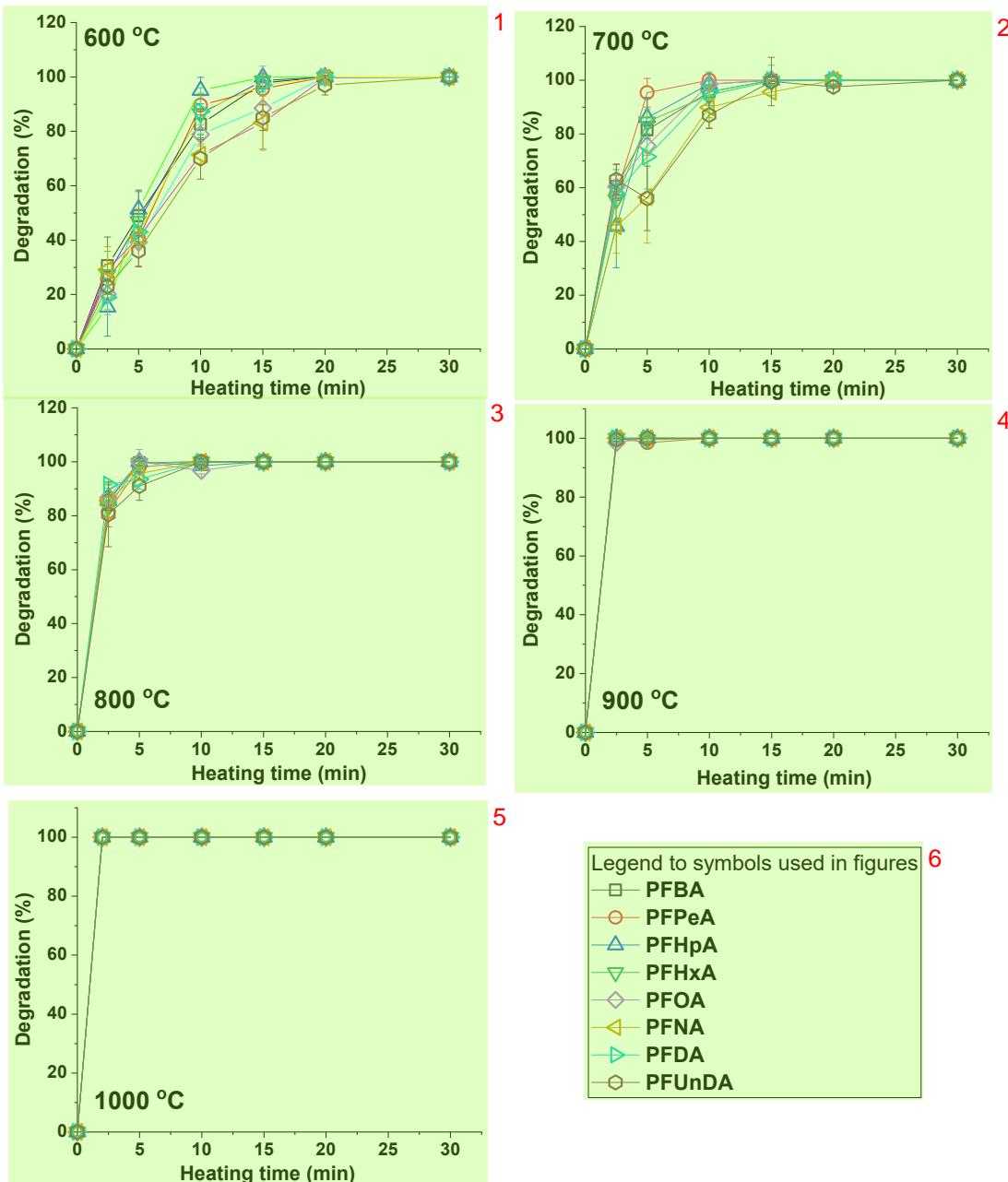


Figure 1. Degradation of PFCAs during incineration at different times. 7

1.2. Incineration of PFSAs, PFAES, and FTS 8

The data in Figure 2 summarize the degradation rates of PFSAs, PFBS (C4), PFHxS (C6), and PFOS (C8), across a temperature range of 600 °C to 1000 °C during incineration, measured over time. At 600 °C, the degradation rates were noticeably slower than at higher temperatures. Initially, the degradation percentages after 2 min for PFBS, PFOS, and PFHxS were 5.6%, 11%, and 10.2% respectively. As time progressed, the degradation rates increased gradually, with PFOS showing a notably higher rate of 70.9% after 20 min. However, even after 30 min, none of the compounds had reached complete degradation ($\geq 99.99\%$ or non-detectable residuals), which

signifies the high thermal stability of these compounds at this temperature. At 700 °C, all compounds showed a much higher degradation rate. In just 2 min, PFBS, PFOS, and PFHxS degradation percentages were 25.3%, 15.6%, and 35.6%, respectively, showing a substantial increase compared to the results at 600 °C. At the 10-min mark, PFHxS had already nearly reached complete degradation at 98.5%. 1

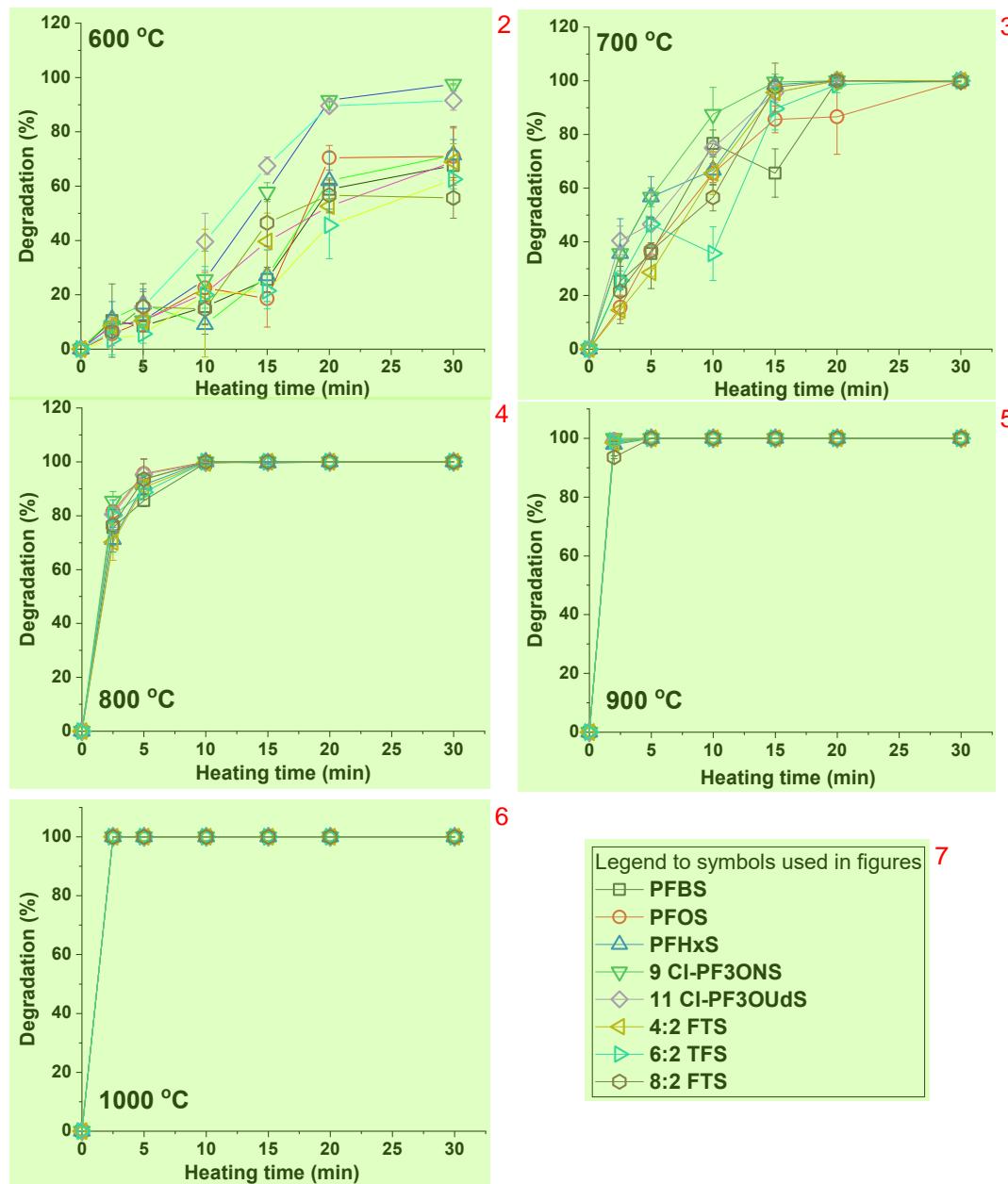


Figure 2. Degradation of PFSAs, PFAES, and FTS during incineration at different times. 8

By the 20-min mark, all compounds had achieved complete thermal degradation during incineration. When the temperature was raised to 800 °C, the degradation rate increased significantly for all compounds. By 2 min, degradation rates of 75.6%, 81.5%, and 71.1% for PFBS, PFOS, and PFHxS respectively were observed. Complete degradation ($\geq 99.99\%$) was reached for 9

all compounds by the 15-min mark. At 900 °C and 1000 °C, the compounds degraded rapidly, reaching near-complete ($\geq 99\%$) or complete ($\geq 99.99\%$) degradation within the first 2 min. This suggests an extreme increase in the thermal degradation rate at these temperatures. Overall, it can be concluded that the thermal degradation rate of PFBS, PFOS, and PFHxS increases drastically with incineration temperature. At lower temperatures (600 °C), these PFCs exhibit high thermal stability, whereas, at higher temperatures (900 °C-1000 °C), they degrade within 2 min. These insights are critical for the safe handling and disposal of PFSAs such as PFOS that are much more difficult to degrade than PFOA by non-thermal approaches.¹

The incineration of two homologs, 9 Cl-PF3ONS (also known as 6:2 Cl-PFESA or F53-B) and 11 Cl-PF3OUDS, was examined over a range of temperatures, from 600 to 1000 °C, at varying time intervals. At 600 °C, it is apparent that both homologs demonstrate a steady increase in degradation over time. 9 Cl-PF3ONS reached substantial degradation (97.5% or 90–99%) after 57.8 min, whereas 11 Cl-PF3OUDS achieved a slightly lesser degradation rate of 91.5% at the same time. However, with additional heating, 11 Cl-PF3OUDS caught up, reaching nearly complete degradation ($\geq 99\%$) at the end of the experiment. At 700 °C, the rate of thermal degradation accelerated for both homologs, particularly notable for 9 Cl-PF3ONS, which achieved complete ($\geq 99.99\%$) degradation after 87.5 min. 11 Cl-PF3OUDS exhibited slower degradation, reaching nearly complete degradation only at the end of the experiment. When heated at 800 °C, both homologs reached complete degradation at a much quicker rate. It's noteworthy that the degradation rate for 11 Cl-PF3OUDS, which had lagged slightly at lower temperatures, caught up and mirrored that of 9 Cl-PF3ONS at this higher temperature. The thermal degradation rates of both homologs at 900 °C and 1000 °C were almost identical. Near-complete degradation was observed almost instantly at 900 °C, with complete degradation observed by the 2nd time-point for both homologs. This was similarly observed at 1000 °C. Overall, these results indicate a clear temperature-dependent trend in the thermal degradation of these two homologs. With increasing temperature, the time required for complete degradation decreased significantly. This is a characteristic behavior of thermally labile compounds, which decompose more readily with increasing temperature. Furthermore, despite their structural differences, both homologs displayed very similar thermal degradation profiles, suggesting comparable thermal stabilities and potentially similar decomposition mechanisms. These findings not only enhance our understanding of the thermal properties of these compounds but also provide valuable insights for their safe handling and disposal. The knowledge that complete degradation can be achieved at a certain temperature and timeframe can guide strategies for their incineration treatment, ensuring that potential environmental and health risks are minimized.²

Figure 2 also presents the incineration degradation data of a group of fluorotelomer sulfonic acids (4:2 FTS, 6:2 FTS, 8:2 FTS). The degradation data suggest that temperature significantly impacts the degradation rates of all studied FTS. The overall trend is similar to that of PFSAs. At the lowest temperature of 600 °C, there is a clear and gradual increase in degradation with longer heating times. For most compounds, the degradation percentage reached a maximum of about 70–100% after 20 to 30 min of heating. However, when the temperature was increased to 700 °C, a noticeable shift was observed. FTS exhibited significantly higher degradation rates at shorter times. They reached near-complete ($>99\%$) or complete ($\geq 99.99\%$) degradation within 15 to 30 min. Further increases in temperature to 800 and 900 °C only continued this trend, with many compounds reaching near-complete or complete degradation within as little as 5 to 10 min of heating. At 1000 °C, complete degradation was reached almost instantly for all compounds, with virtually no differences observed between the different heating times.³

1.3. Incineration of PFECAs and FOSAAs ¹

The data in Figure 3 present a comprehensive incineration degradation analysis of PFECAs (PFMPA, HFPO-DA, ADONA) and two FOSSAs (NMeFOSAA and NEtFOSAA) at different heating temperatures (600, 700, 800, 900, and 1000 °C). In general, it can be observed that the degradation rate of these compounds increases with the temperature. At 600 °C, none of the compounds reach 100% degradation within the provided timeframe. While ADONA and HFPO-DA exhibited complete degradation ($\geq 99.99\%$ to non-detectable residues), PFMPA lags slightly with the highest recorded degradation at 86.7%. NEtFOSAA and NMeFOSAA display a somewhat intermediate behavior, achieving 100% and 99.99% degradation respectively. At 700 °C, all substances except PFMPA reach 100% degradation, PFMPA still shows slower degradation reaching a maximum of 100%. At this stage, the reason for PFMPA's slower degradation remains unclear.

ADONA and HFPO-DA show a slight increase in degradation rate at this temperature, achieving 100% degradation faster than at 600 °C (Figure 3). At 800 °C, all substances reach 100% degradation. At 900 and 1000 °C, all substances consistently reach 100% degradation within 2 min. Interestingly, at 900 °C, NEtFOSAA shows a slightly slower degradation at 99.5% compared to other compounds. However, this difference is negligible and could be within the error margin of the experiment. It is also worth noting the variation in degradation patterns between these substances. While PFMPA exhibits the slowest degradation rate across all temperatures, ADONA and HFPO-DA show relatively fast degradation. NEtFOSAA and NMeFOSAA exhibit intermediate degradation rates (Figure 3).

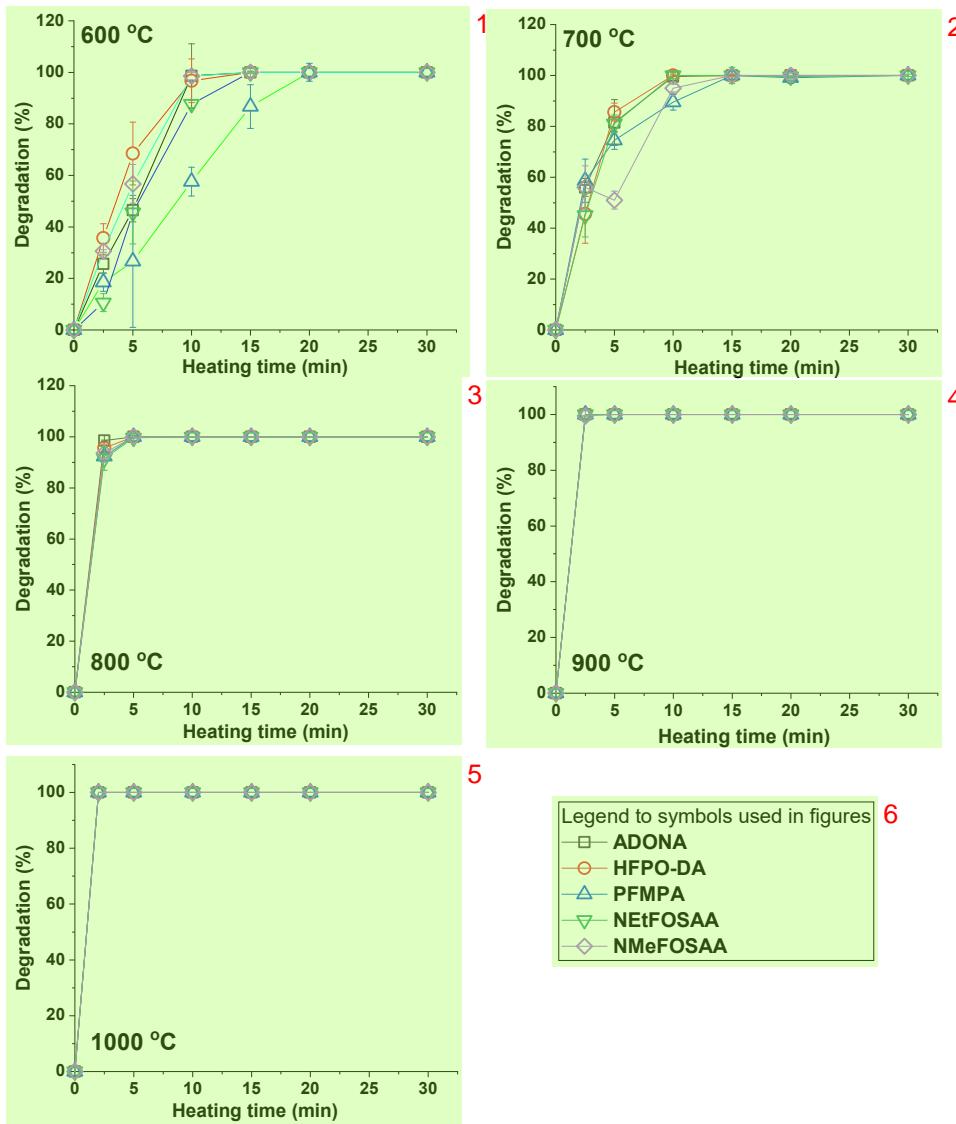


Figure 3. Degradation of PFECAs and FOSAAs during incineration at different times. 7

1.4. Minimum incineration time required to achieve 99%, 99.9%, 99.99%, and 8 99.999% degradation of PFAS

By heating these substances to temperatures of 600, 700, 800, 900, and 1000 °C, and measuring the minimum time in minutes needed to achieve 99% (2-log), 99.9%, 99.99%, and 99.999% degradation, we have garnered a deep understanding of the robustness and thermal stability of these compounds. The results are presented in Tables 1 and 2. The degradation of PFAS is significantly influenced by the heating temperature. For instance, at a lower temperature of 600 °C, the time required to attain 99% degradation for PFCAs ranged from 15 to 30 minutes. However, when the temperature was escalated to 1000 °C, the degradation time dramatically reduced to less than 2.5 minutes, indicating a faster degradation rate at higher temperatures. This pattern of accelerated degradation with increased temperature was observed across all studied PFAS. However, there was variability among different PFAS groups. Some substances, such as PFSAs,

demonstrated a slower degradation rate at 600 °C, requiring more than 30 minutes to achieve a 99% ¹ degradation rate. Yet, a significant reduction in degradation time was noted at 1000 °C, wherein degradation was achieved in less than 2.5 minutes. In the case of FTS, a noticeable resistance to thermal degradation was exhibited at lower temperatures with times greater than 30 minutes needed for 99% degradation at 600 °C. However, at elevated temperatures, especially at 1000 °C, this resistance diminished, and rapid degradation was accomplished in less than 2.5 minutes. PFECAs and FOSSAs showed relatively less resistance to thermal degradation across all temperatures. At 600 °C, both PFECAs and FOSSAs achieved 99% degradation within 15 minutes, and this time was further reduced to less than 2.5 minutes at 1000 °C. The data distinctly highlights the thermal dependency of PFAS degradation, reaffirming the crucial role of high temperatures in achieving rapid and comprehensive degradation. While there was a notable difference in the degradation rate among different PFAS, the data establish that their thermal degradation is possible and can be significantly expedited at higher temperatures. These findings are integral in shaping effective waste management strategies and providing robust solutions for PFAS disposal or remediation.

Table 1. Minimum incineration time (min) to achieve 99% (2-log), 99.9% (3-log), 99.99% (4-log), and 99.999% (5-log) degradation of ¹ PFCAs, PFAs, and PFESAs.

Temperature (°C)	PFCA 99%	PFCA 99.9%	PFCA 99.99%	PFCA 99.999%	PFSA 99%	PFSA 99.9%	PFSA 99.99%	PFSA 99.999%	PFESA 99%	PFESA 99.9%	PFESA 99.99%	PFESA 99.999%
600	15	20	20	30	>30	>30	>30	>30	>30	>30	>30	>30
700	10	15	20	30	20	20	20	30	15	20	20	30
800	5	5	10	15	5-10	10	15	20	5-10	10	15	15
900	<2	<5	5	5	2	5	5	5	2	<5	5	5
1000	<2	<2	<2	<2	<2	<2	<2	<2	<2	<2	<2	<2

Table 2. Minimum incineration time (min) to achieve 99% (2-log), 99.9% (3-log), 99.99% (4-log), and 99.999% (5-log) degradation of ³ FTS, PFECAs, and FOSSAs.

Temperature (°C)	FTS 99%	FTS 99.9%	FTS 99.99%	FTS 99.999%	PFECAs 99%	PFECAs 99.9%	PFECAs 99.99%	PFECAs 99.999%	FOSSAs 99%	FOSSAs 99.9%	FOSSAs 99.99%	FOSSAs 99.999%
600	>30	>30	>30	>30	10-15	15	15	15	15	15	15	15
700	15-20	20	20	30	10	10	15	15	10	10	15	15
800	5-10	10	15	15	2	2	2-5	2-5	5	10	10	10
900	2	5	5	5	<2	<2	<2	<2	<5	<5	<5	<5
1000	<2	<2	<2	<2	<2	<2	<2	<2	<2	<2	<2	<2

Notes: ">" means that the actual time required is more than the value given. "<" means that the actual time required is less than the value ⁵ given. "X-Y" means the actual time required falls within this range.

Chapter 2. Thermal Degradation of PIDs.¹

2.1. Thermal Degradation and Mineralization.²

Temperature shifts have demonstrated a distinct effect on PFAS degradation efficiency,¹⁴⁻³
^{16, 20} however, information on the temperature effects on PFC thermal degradation remains limited. As depicted in Figure 4a, near-complete degradation of C₇F₁₄ was achieved at temperatures above 300 °C, with the yield of fluorine (F) increasing modestly and peaking at 24.0 mol% at 400 °C, before slightly decreasing at higher temperatures. Similarly, C₈F₁₆ showed trends of complete degradation at 300 °C and above, with a significant increase in F yield at elevated temperatures, as shown in Figure 4b. In contrast, saturated PFCs (C₇F₁₄ and C₈F₁₆) exhibited complete degradation only at temperatures exceeding 500 °C. The generation of organofluorine compounds is the dominant pathway during thermal decomposition, with a notably low yield of inorganic F (<10 mol%) across the temperature range studied (Figures 4a and 4b).

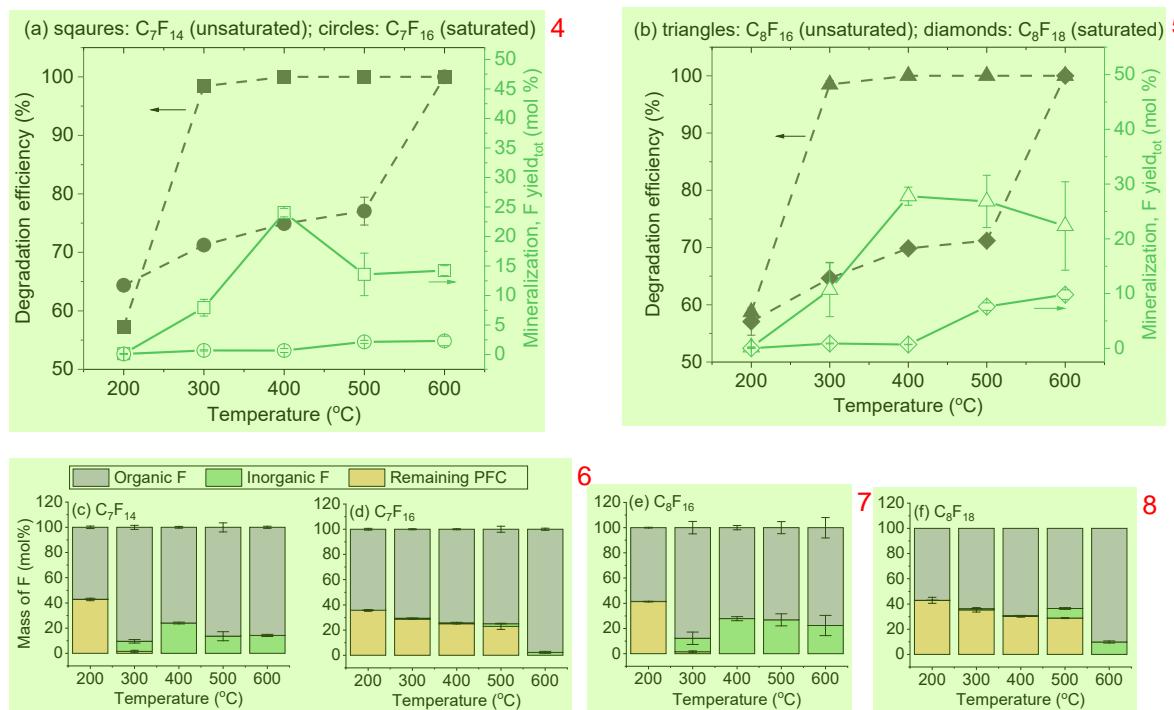


Figure 4. Degradation and mineralization of four PFCs heated in air at different temperatures (200, 300, 400, 500, and 600 °C) for 60 min. (a): C₇ PFCs. (b): C₈ PFCs. (c)–(f): Mass flows of organic fluorinated degradation products, inorganic F, and remaining PFC.

The organofluorine products of PFCs at a given temperature were estimated based on the difference between the initial mass of PFC added to the reactor and the mass of inorganic F at this temperature. As shown in Figure 4c, the mass flows of organic F from C₇F₁₄ started at 53.6 mol% at 200 °C and reached a peak of 90.5 mol% at 300 °C before slightly dropping to 75.9 mol% at 400 °C and stabilizing around 85.0 mol% at 500 and 600 °C. This result suggests that as C₇F₁₄ decomposes, a significant portion of the resulting products are organic F compounds.

In the case of C₈F₁₆, the organic F content followed a similar trend, starting at 58.6 mol% at 200 °C and rising to 87.8 mol% by 300 °C. At higher temperatures, the organic F slightly

declined and stabilized approximately at 75 mol% (Figure 4e). For saturated PFCs, the majority of F remained in the undegraded PFCs, as depicted in Figures 4d and 4f.

Further supporting these observations, Khan et al. explored the thermal mineralization mechanisms of perfluorooctanesulfonic acid (PFOS), identifying that at 1000 K, PFOS preferentially forms a-sultone, which swiftly decomposes into perfluorooctanal and SO₂ with a rapid half-life of 0.2 seconds. These insights, highlighting a highly efficient breakdown process, emphasize the potential of tailored thermal treatments to significantly enhance the degradation efficiency of PFAS compounds under relatively modest temperatures, achieving up to approximately 40% mineralization between 300–500 °C.⁴⁸

The presence of GAC significantly enhanced the degradation efficiency of C₇F₁₄, achieving 100% at 200 °C (Figure 5a), reflecting a 42.7% degradation enhancement. Additionally, at 500 °C, the F yield for C₇F₁₄ increased to 32 mol%, representing an 8% increase in comparison with the no-GAC scenario. Similarly, as highlighted in Figure 5b, for C₈F₁₆, the presence of GAC resulted in a degradation efficiency of 99.6% at 200 °C, marking a 40.8% enhancement, and the F yield rose to 41.8 mol%, indicating a 14.1% increase at 500 °C.

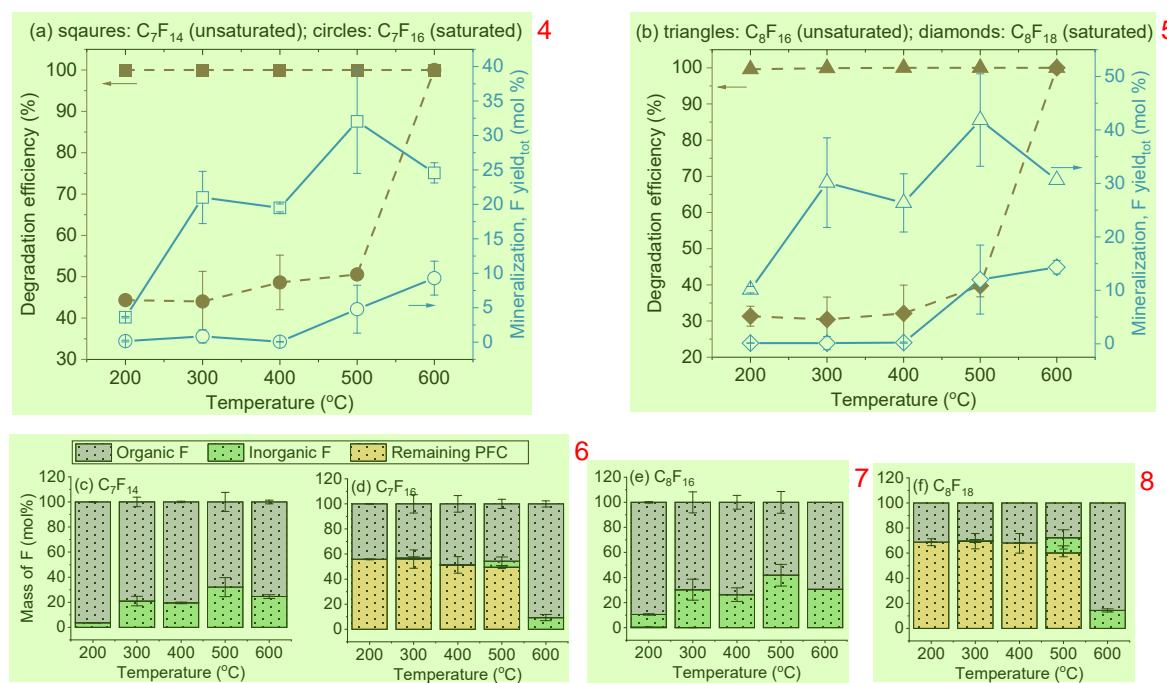


Figure 5. Degradation and mineralization of four PFCs heated in air at different temperatures (200, 300, 400, 500, and 600 °C) for 60 min in the presence of GAC. (a): C₇ PFCs. (b): C₈ PFCs. (c)–(f): Mass flows of organic fluorinated degradation products, inorganic F, and remaining PFC after heating at these temperatures.

On the other hand, the effect of GAC on the thermal degradation of saturated PFCs is not evident. In fact, in some cases, GAC appears to have an inhibitory effect on the degradation of saturated PFCs, leading to a around 20% drop in degradation efficiency compared to scenarios without GAC (Figures 4 and 5).

The mass flow of organic F from C₇F₁₄ began at a high of 96.4 mol% at 200 °C in the presence of GAC, then gradually declined to 68.0 mol% by 500 °C (Figure 5c). C₈F₁₆ exhibited a similar trend, though it had a higher proportion of inorganic F compared to C₇F₁₄ (Figure 5e).

For saturated PFCs, with GAC, the mass percentage of organic F was low across 200 to 1 500 °C, approximately 45 mol% for C₇F₁₆ and 30 mol% for C₈F₁₈, due to the poor degradation of these two saturated compounds. At 600 °C, significant degradation occurs, increasing the organic F mass percentage to 90.7 mol% for C₇F₁₆ and 85.7 mol% for C₈F₁₈.

As shown in Figures 4 and 5, we observed a decline in the yield of F at temperatures exceeding 400 °C in certain cases. The precise reason behind this trend is currently unclear. A plausible explanation could be that the consumption of F radicals by intermediate products during the degradation process contributes to the observed decline in yield. 2

In summary, all four PFCs effectively degrade at specific temperatures—above 200 °C for perfluoroalkenes and above 600 °C for perfluoroalkanes. The release of F generally increases with temperature, although the specific yield percentages for saturated PFCs remain low. Saturated PFCs, such as perfluoroalkanes, show significantly greater thermal stability compared to perfluoroalkyl carboxylic acids (PFCAs),^{15, 20, 49} which start to degrade between 150 °C and 300 °C. Conversely, unsaturated PFCs, like perfluoroalkenes, exhibit similar thermal stability to PFCAs. 3

2.2. Organofluorine Byproducts of PFCs. 4

Figure 2.1 (Appendices) depicts the abundance of selected byproducts formed during the 5 degradation of a mixture of four tested PFCs, both with and without the presence of GAC. The thermal treatment duration (10, 60, 120 min), as shown in Figure 2.1a, did not significantly affect the overall abundance of potential byproducts. A notable finding is the distinct difference in byproduct abundance between scenarios 2.1 with GAC and without GAC. The presence of GAC notably enhanced the yield of linear and cyclic organofluorine byproducts throughout the process, which is attributed to GAC's promotion of PFC degradation.

Moreover, as shown in Figure 2.1b (Appendices), there was a decrease in the abundance 6 of all detectable byproducts as the temperature increases from 200 °C to 600 °C. These products became non-detectable at temperatures exceeding 500 °C, indicating the decomposition of these organofluorine byproducts. Our previous investigation of the thermal decomposition products from aqueous film forming foam (AFFF), PFOA, PFOS, and other PFAS revealed a comparable pattern.¹⁸⁻²⁰ These compounds exhibited peak byproducts diversity and intensity at intermediate temperatures (200 °C to 500 °C) subsequently diminishing to trace levels at temperatures above 890 °C.

It is worth mentioning that the degradation products analyzed in Figure 2.1 (Appendices) 7 were only those with a GC retention time greater than 8.7 min. In accordance with our GC-MS center's policy, we conducted MS analysis on products that have a retention time exceeding that of n-hexane (8.4 min). This procedure was to ensure accurate differentiation and analysis of compounds of interest. To gain further insight into potential byproducts formed at a wider GC retention time range, we used two approaches. The first approach involved the analysis of selected samples at 300 °C with a wider GC retention time range. Perfluoro-1-butene (C₄F₈; retention time 7.40 min) emerged as the dominant decomposition product of PFCs under both oxidative and inert atmospheres. Interestingly, experiments with GAC did not exhibit a corresponding peak for this compound, suggesting GAC's ability to change the degradation pathways of PFCs and/or inhibit the generation of this compound. The generation of C₄F₈ has also been reported as a fluorinated volatile byproduct of AFFF and individual PFAS, such as PFOS and perfluorovaleric acid (C5) decomposition in prior studies.¹⁸⁻²⁰

The second approach involved the analysis of the thermal degradation products of PFCs 8 using TD-Pyr-GC-MS, where the compound was directly heated, and the resultant species were

examined through GC–MS. For both PFCs, the presence of byproducts increased significantly as temperatures rose from 50 to 500 °C, indicating enhanced thermal degradation. At lower temperatures (50–200 °C), byproducts with relatively low molecular weights were predominantly observed, suggesting initial breakdown processes.²⁴ As the temperature reached 200–300°C and above, a greater variety of linear and cyclic degradation products with varying complexities appeared. Notably, by 500°C, almost no peaks were detectable for either compound, implying near-total degradation.

2.3. Bond Energy Distribution. ²

To further understand the thermal degradation behaviors observed in our experiments with PFCs, we analyzed bond dissociation free energies (BDFEs) among the investigated compounds. BDFEs reflect the energy required to break a specific chemical bond, with a lower BDFE indicating a weaker bond that is more susceptible to cleavage during thermal treatment.²⁰ C–F bonds and double C–C bonds are generally very strong due to the high electronegativity of F and the presence of two shared electron pairs in the double bond, respectively (Figure 2.2; Appendices). As shown in Figure 2.2, saturated PFCs (C_7F_{16} and C_8F_{18}), due to their symmetrical structure with a repeated pattern of single C–C bonds, exhibit less variation in their BDFE values (Figure 2.2). All their BDFE values fall within a relatively narrow range of 71.8 kcal/mol to 75.9 kcal/mol. This consistency reflects the uniform distribution of bond strengths across the molecule in saturated PFCs.

Unsaturated PFCs (C_7F_{14} and C_8F_{16}) possess the strongest bonds (C2–C3) among the investigated molecules, with a BDFE of 91.2 kcal/mol, likely reinforced by the adjacent double bond (C1=C2) as depicted in Figure 2.2. Conversely, these compounds also exhibit the weakest bonds (C3–C4) among all C–C bonds involved, with BDFE values of approximately 66.1 kcal/mol. This discrepancy highlights that the single bond between the third and fourth C atoms is particularly susceptible to cleavage (Figure 2.2), contributing to the lower thermal stability observed in these unsaturated PFCs compared to their saturated counterparts (Figures 4 and 5).

Thermal Decomposition Pathways of Perfluoroalkenes. Based on our qualitative and quantitative analysis of byproducts, we propose possible thermal degradation pathways of perfluoroalkenes as shown in Figure 2.3 (Appendices).

Two primary transformation routes are possible: chain scission and cyclization. Chain scission occurs through a series of steps involving hemolytic cleavage of C–C bonds and F atom transfer, resulting in the generation of shorter-chain perfluoroalkenes and perfluoroalkanes (e.g., perfluoro-1-butene; C_4F_8). Following the bond cleavage, the radical perfluoroalkyl chains can also undergo intramolecular free radical cyclization to generate perfluorocycloalkenes. This process is energetically favorable since it reduces the overall energy of the system by eliminating the radicals and forming more stable compounds. The cyclization process yields different perfluorocycloalkenes, such as perfluorocyclopentane (C_5F_{10}) and perfluorocyclohexane (C_6F_{12}). Similar cyclic perfluorocarbons have been identified as gaseous PFAS byproducts in other studies investigating the degradation of perfluorinated compounds.⁵⁰⁻⁵¹

Intermediates and radical fluorocarbon chains can undergo reaction pathways during the heating process and subsequent cooling period, potentially influencing the final product distribution. However, our analysis suggests that the bulk of the initial mass of perfluoroalkenes follows pathways through C3–C4 carbon cleavage, as this bond poses lower BDFE than others. Following this initial cleavage, the molecules proceed along a path that results in the formation of

perfluoro-1-butene and perfluorocyclohexane, considering the high abundance of these ¹ compounds in thermally treated samples.

2.4. Post-treatment of PIDs. ²

In our previous studies, we have found that GAC enhanced the mineralization of PFOA ³ and its homologs,^{15, 20} suggesting that the thermal degradation pathways of PFAS can be altered. As C₇F₁₄ and other PFCs included in these studies are PIDs of PFOA, we examined whether GAC can also enhance the mineralization of these PFCs. In addition to GAC, we also included a few commercial catalysts such as platinum on alumina (Pt/Al₂O₃) and palladium on alumina (Pd/Al₂O₃).

As shown in Figure 2.4 (Appendices), while all amendments have shown a positive impact ⁴ on inorganic fluoride yield, the extent of this enhancement varied considerably. The highest total fluoride yield was observed by introducing hydrogen zeolite Y to the system, which yielded 24.8 mol%. GAC amendments (F200, Norit 400, and HD 400) also boosted the F yields of C₈F₁₆ upon heating, ranging from 15.5 mol% (F200) to 18.3 mol% (HD 4000). Other commercial catalysts such as a silica-alumina catalyst support, Pd/Al₂O₃, and Pt/Al₂O₃ only slightly increased the yield of F, ranging from 3.7 mol% to 3.9 mol%. Among the other amendments, including molecular sieve, activated alumina, biochar, and iron filings showed no sign of increasing the yield of F (Figure 2.4; Appendices).

The thermal surface reactions of PFAS with these additives are potentially complicated and ⁵ are likely influenced by the physicochemical properties of the surfaces involved and the specific environmental conditions. Biswas and Wong have significantly advanced our understanding of PFOA's thermal mineralization mechanisms through *ab initio* molecular dynamics simulations.⁵² Their research elucidates that PFOA does not degrade on pristine γ -Al₂O₃ (100) surfaces even under high temperatures. However, introducing an oxygen vacancy dramatically alters this behavior, facilitating rapid defluorination and bond cleavage, thus offering a potential pathway to enhance the effectiveness of thermal treatments in environmental remediation of such persistent pollutants.⁵²

Among the investigated amendments, GAC offers several distinct advantages. Firstly, ⁶ GAC is a well-established material with extensive applications in environmental remediation, particularly for the treatment of PFAS in water.⁵³⁻⁵⁶ Its effectiveness in adsorbing PFAS from aqueous environments is a well-documented approach for mitigating PFAS contamination. Secondly, GAC possesses a relatively low cost and a readily scalable production process compared to some commercial catalysts. This economic feasibility is crucial for the widespread applications. GAC outperforms other additives (Figure 2.4; Appendices) in promoting the thermal mineralization of PFCs primarily due to its high surface area and strong adsorptive capacity. These properties enable GAC to more effectively capture and break down PFCs, especially at lower temperatures where other additives may not be as efficient. Furthermore, GAC's structure allows for the adsorption of intermediate products that might otherwise inhibit further degradation, enhancing the overall efficiency of the mineralization process. This capability is particularly evident in Figures 2.4 and 2.5 (Appendices), where the use of GAC not only boosted the F yields significantly across various temperatures but also facilitated the rapid degradation of both saturated and unsaturated PFCs, underscoring its superior performance in environmental remediation applications.

Next, we conducted further detailed investigations into the effects of GAC on the thermal ⁷ mineralization of PFCs with varying structures. Figure 2.5 shows that GAC enhances the total F

yield during the thermal degradation of both saturated and unsaturated PFCs, with a more notable effect on unsaturated PFCs. Specifically, GAC significantly increased F production in unsaturated PFCs, achieving up to 32.2 mol% for C₇F₁₄ and 42.5 mol% for C₈F₁₆ at 500 °C. This enhancement was evident even at lower temperatures, such as 200 °C. Interestingly, similar effects of GAC on increasing F yields at low temperatures were previously observed for PFCAs, including PFOA).¹⁵ The findings from this study (Figure 2.5; Appendices) suggest that the positive impact of GAC on thermal PFCA degradation may be linked to its role in enhancing the mineralization of their fluorinated precursors (i.e., unsaturated PFCs).

For the saturated PFCs, C₇F₁₆ and C₈F₁₈, the data from Figure 2.5 also shows a considerable increase in F yield with the presence of GAC at 500 °C and 600 °C. This enhancement indicates that GAC not only facilitates the breakdown of unsaturated PFCs but also significantly improves the thermal mineralization of their saturated counterparts at elevated temperatures.

Note that the *total* F yield (Eqs 3 and 4) generally aligns with the *relative* F yield, due to the efficient degradation of perfluoroalkenes during thermal treatment, resulting in negligible residual mass ($M_{PFC,r} \approx 0$).

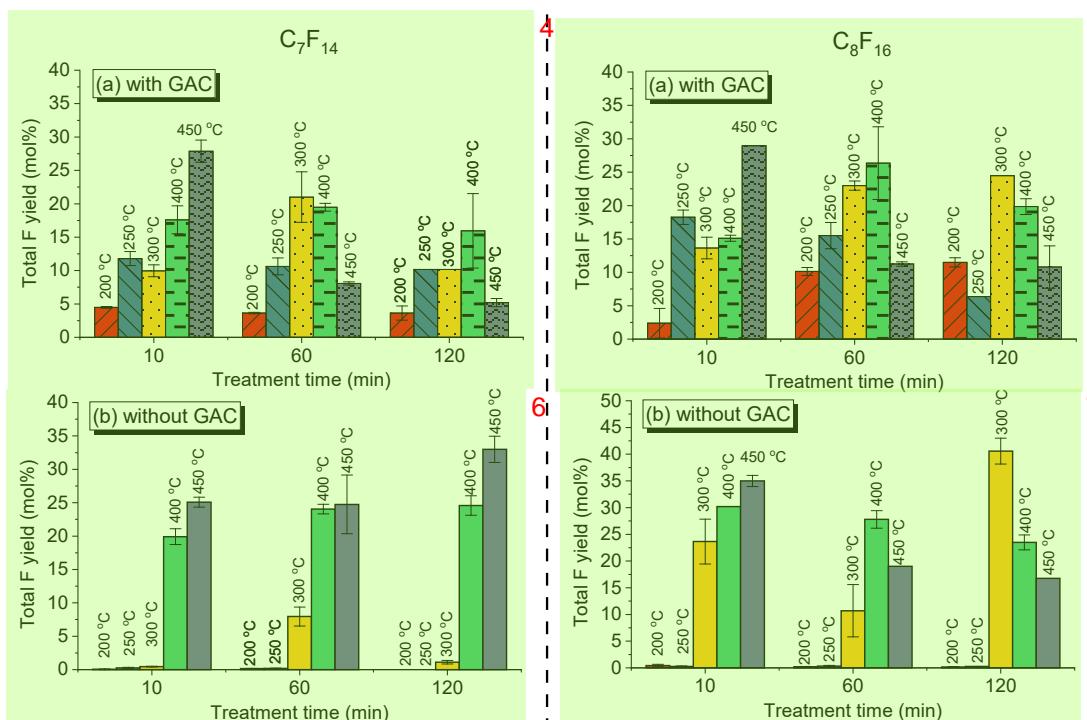


Figure 6. Thermal mineralization of PFCs heated in selective temperatures and durations.

Figure 6 shows the time-dependent effects of GAC on the mineralization of these PFCs (Figure 6). At 200 °C, the yield of F from C₇F₁₄ after 10-min heating was much higher (4.5 mol%) compared to the absence of GAC (<0.1 mol%). This trend continued at 60 min, where the yield of F with GAC was 3.6 mol%, compared to 0.2 mol% without GAC. At higher temperatures (250 °C and 300 °C), the data shows a similar pattern where GAC presence enhanced the mineralization substantially. For instance, at 300 °C, the yield of F with GAC was 10.0 mol% after 10 min and peaked at 21.0 mol% at 60 min, whereas without GAC, the yield of F was 8.0 mol% after 60 min

heating at 300 °C. The influence of GAC on the yield of F from C₇F₁₄ diminished at 400 °C, at 1 which the yields of F with GAC were comparable with those without the presence of GAC (Figure 6). Interestingly, at 500 °C and 600 °C, the catalytic effect of GAC was once again evident, increasing the yield of F from C₇F₁₄ as high as 57% (Figure 2.7; Appendices).

Similar to C₇F₁₄, at 200 °C, the presence of GAC showed a substantial effect on the yield 2 of C₈F₁₆ (Figure 6), significantly enhancing it across all time points. For example, with GAC, the yield of F from C₈F₁₆ was 2.4 mol% after 10-min treatment and increased to 11.5 mol% at 120 min. In contrast, without GAC, the maximum yield of F was ~0.4 mol% over the same period. At 250 °C, the effect of GAC continues to be prominent. The yield of F peaked at 18.2 mol% after 10-min thermal treatment at 250 °C in the presence of GAC. In contrast, without GAC, the yield of F was never above 0.5 mol%. The effect of GAC diminished at 400 °C but became significant again at 500 °C and 600 °C. For example, the yield of F from C₈F₁₆ rose to 48.0 mol%, compared to just 23.5 mol% without GAC after 60-min heating at 500 °C.

2.5. Effects of Atmosphere.³

We also compared the effects of atmosphere on the thermal mineralization of these PFCs. 4 As shown, air favored the mineralization of both unsaturated and saturated PFCs at low temperatures (200 and 300 °C).²⁴

In both atmospheres, the addition of GAC boosted the mineralization of both PFCs at 5 temperatures up to 300 °C. However, at 400 °C, there was a notable surge in the mineralization (~60 mol% of F) of both unsaturated PFCs in N₂ without GAC, compared to the mineralization conditions in air. We repeated multiple times on this datapoint, and the observation is consistent. This difference could be attributed to the consumption of F radicals by oxygen or alternative mineralization pathways of these compounds in the air.²⁴

Chapter 3. Nondestructive detection of PIDs of PFAS⁶

While GC-MS/(MS) can provide important information regarding the PIDs from PFAS, 7 some small PIDs may not be easy to ionize in MS. In this regard, we explored a nondestructive, real-time analytical method (TGA-FTIR) to analyze PIDs of PFAS.

3.1. Reference FTIR Spectra of PFCs, HFCs, and Other Potential Thermal 8 Degradation Products of PFAS.

We begin with a brief overview of the FTIR spectra of reference compounds gathered from 9 the U.S. National Institute of Standards and Technology database^{42, 57-64} and a recent study by the U.S. Pacific Northwest National Laboratory.⁵⁹ The spectrum^{35, 57} of CF₂O exhibits a prominent absorption near 1935 cm⁻¹, characteristic of the C=O stretching vibration, with additional peaks at ~1232 and ~1028 cm⁻¹ representing the stretching and bending vibrations of C–F (Figure 3.1a; Appendices). The spectrum of SiF₄ shows a sharp and intense peak at ~1030 cm⁻¹,^{58, 65} attributable to the stretching vibration of the strong Si–F bond (Figure 3.1a; Appendices).

Characteristic peaks at ~1250 and 1114 cm⁻¹ were observed in both PFCs (Figures 3.1b and 10 3.1d; Appendices)^{59, 61-64} and HFCs (Figure 3.1c; Appendices).⁵⁹⁻⁶⁰ Compared to short-chain PFCs, long-chain PFCs and HFCs exhibit two characteristic peaks at 1220–1206 and 1155–1145 cm⁻¹ (Figure 3.1c). The spectrum of CF₄ is noted for a single peak at 1250 cm⁻¹.

For perfluoroalkenes, the peaks at 1789 cm⁻¹ and 1720 cm⁻¹ represent the C=C stretch of a 11 double bond in perfluoro-1-alkenes (e.g., perfluorohept-1-ene) and perfluorohept-2-ene,

respectively. Cyclooctafluorobutane (C_4F_8) displays a dominant peak at 962 cm^{-1} (Figure 3.1b; 1 Appendices).

3.2. Thermal Degradation Products of PFOA Investigated by TGA-FTIR. 2

Detailed investigations focused on the thermal decomposition of PFOA and K-PFOS 3 across a range of temperatures. Panels (a) and (b) in Figure 3.2 present the FTIR spectra of PFOA when heated up to $700\text{ }^\circ\text{C}$ in N_2 and air, respectively. In N_2 , the spectrum for PFOA at $150\text{ }^\circ\text{C}$ is very similar to the reference spectrum,⁶⁶ suggesting little to no decomposition at this temperature. The results align with our previous observation¹⁵ that the thermal degradation of PFOA is either thermodynamically or kinetically unfavorable at $150\text{ }^\circ\text{C}$. Above $150\text{ }^\circ\text{C}$, characteristic peaks of PFOA associated with the O–H stretch at 3600 cm^{-1} , the C=O stretch at 1817 cm^{-1} , and CF_3 vibrations around 1250 cm^{-1} began to decline in intensity, pointing to initiation of thermal degradation (Figure 3.2a; Appendices).

Simultaneously, new peaks emerged at ~ 1205 and $\sim 1145\text{ cm}^{-1}$ (Figure 3.2; Appendices). 4 Weber et al. recently identified perfluorooctanoyl fluoride ($C_8F_{16}O$) from PFOS in thermal processes, based on peaks observed at approximately 1207 cm^{-1} and 1145 cm^{-1} (both C–F stretching) and at $\sim 1879\text{ cm}^{-1}$ (stretching vibration of the ketone C=O bond).³⁹⁻⁴⁰ The absorbance at $\sim 1879\text{ cm}^{-1}$ falls slightly outside the typical range for ketone C=O bonds (1775 – 1650 cm^{-1}).⁶⁷

In comparison, the thermal degradation of PFOA in the presence of air began at $150\text{ }^\circ\text{C}$, as 5 evidenced by the diminishing peak ($\sim 1819\text{ cm}^{-1}$) associated with the C=O bond in PFOA's carboxyl group (Figure 3.2b). The emergence of new peak(s) at $\sim 1723\text{ cm}^{-1}$ ($\nu C=C$) could signify the formation of perfluoro-2-alkene, such as perfluorohept-2-ene (C_7F_{14}). FTIR peaks of HF and CO_2 were not observed when PFOA was heated alone in N_2 or air (Figure 3.2; Appendices). As shown below, the yield of F from PFOA heated alone is quite low. Carboxyl radicals, if released from PFOA via the decarboxylation pathway, may recombine with some nonpolar degradation products to form short-chain homologs; we previously demonstrated the formation of short-chain homologs from PFOA upon heating.^{15, 20} This might explain the absence of CO_2 peaks in Figure 3.2 of Appendices. Collectively, these findings demonstrate that the degradation of vaporized PFOA begins at approximately $150\text{ }^\circ\text{C}$, accompanied by an increasing presence of PIDs.

3.3. Thermal Degradation Products of K-PFOS Investigated by TGA-FTIR. 6

Figure 3.3 presents the FTIR spectra of K-PFOS when heated up to $700\text{ }^\circ\text{C}$ in air. In a N_2 7 atmosphere, the stability of K-PFOS was evident up to approximately $400\text{ }^\circ\text{C}$, as peaks for the S=O stretch in sulfonate at 1347 cm^{-1} and the C–F stretch at 1154 and 1250 cm^{-1} for CF_2 and CF_3 groups, respectively, maintained their presence.²⁵ As the temperature increased above $500\text{ }^\circ\text{C}$, the characteristic peaks of K-PFOS began to fade and peaks (1500 and 1000 cm^{-1}) corresponding to potential PIDs began to emerge.

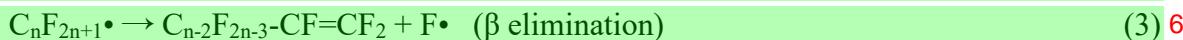
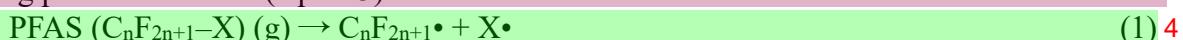
The decomposition pattern of K-PFOS in air differs significantly from that in N_2 , with 8 some new peaks being only detected under oxidative conditions. The peaks at ~ 1950 , 1028 , 978 , 774 cm^{-1} represent CF_2O , which was observed only when K-PFOS was heated at $450\text{ }^\circ\text{C}$ and $500\text{ }^\circ\text{C}$ in air. It was not detected in PFCAs, C8-N-TAmP-FASA (see below), or when K-PFOS was heated in N_2 or in air at $>500\text{ }^\circ\text{C}$ (Figure 3.3; Appendices). These findings suggest that CF_2O formation is species-dependent, temperature-dependent and sensitive to the presence of oxygen.

Furthermore, the medium peak at 1430 cm^{-1} (C–O–H bending)⁶⁷ confirms the emergence 9 of oxygen-containing PIDs in an oxidative environment, akin to the observations with PFOA. Additionally, peaks between 4000 and 3500 cm^{-1} corresponding to HF or $(HF)_n$ are particularly

noteworthy. HF is known to self-associate in the vapor phase, forming $(HF)_n$ through strong hydrogen bonds.⁶⁸⁻⁶⁹ These clusters can show a range of interactions, leading to multiple peaks in the FTIR spectrum in that region.⁶⁸⁻⁶⁹ The appearance of $(HF)_n$ highlights the influence of oxygen in promoting the decomposition and mineralization of K-PFOS.

To enhance the visibility of subtle changes observed in the spectra (Figure 3.3a; Appendices), we normalized the absorbance (Figures 3.3b and 3.3c; Appendices). We then focused on the wavenumber ranges of 3412–3250 cm^{-1} (Figure 3d) and 1600–1300 cm^{-1} (Figure 3.e), highlighting the peaks⁶⁷ corresponding to O–H (3330 cm^{-1}) and C–O–H (1425 cm^{-1}). The peak intensities remained relatively low at temperatures up to 500 °C, but exhibited a significant increase at 550 °C and peaked at 650 °C (Figure 3.g; Appendices), suggesting the thermal degradation of K-PFOS within this temperature range. A previous study revealed that K-PFOS begins to melt at approximately 280 °C, and the molecules subsequently vaporize above this temperature, reaching maximum vaporization rates around 430 °C.²³ The observed changes in peak intensities (Figures 3.d–3.g) align with the thermal phase transition stages of K-PFOS,²³ indicating that the formation of PIDs was mainly due to the decomposition of vapor phase of K-PFOS.

The thermal degradation mechanisms of PFAS, such as PFOA and PFOS, are not yet fully understood. Only a few studies have developed mechanisms based on experimental evidence. Wang et al. suggested direct F elimination from PFCAs during thermal treatments.³⁵ Blotevogel et al. found that the primary decomposition mechanism of PFOA is C–COOH cleavage, whereas α -elimination predominates in the breakdown of hexafluoropropylene oxide dimer acid (HFPO-DA).³⁶ Using density functional theory, Adi and Altarawneh explored various potential thermal decomposition pathways for HFPO-DA.³⁷ Additionally, Alinezhad et al.²⁰ and Sasi et al.¹⁵ have investigated the thermal degradation mechanisms of PFCAs, HFPO-DA, and the long-chain homologs of HFPO-DA. The DFT computations⁷⁰ have illustrated that direct C–C β bond fission is thermodynamically (162 kJ/mol) preferred over fission of the terminal C–C bond forming CF_2 (232 kJ/mol) or elimination of an F atom (245 kJ/mol) producing perfluorinated alkenes. The findings from the present work, along with those of these previous studies,^{15, 20, 35-36, 70} collectively indicate that the initial homolytic cleavage or HF elimination reactions in parent PFAS leads to the formation of perfluoroalkyl radicals (eq 1), which then transform into organofluorine PIDs, including perfluoroalkenes (eqs 2–3).^{14-16, 18, 24}



This mechanism (eqs 1–3) is supported by the observed formation of perfluoropropene ($\text{CF}_3(\text{FC}=\text{CF}_2)$) from perfluorobutanoic acid ($\text{CF}_3(\text{CF}_2)_2\text{COOH}$), perfluoro-1-butene ($\text{C}_2\text{F}_5(\text{FC}=\text{CF}_2)$) from perfluoropentanoic acid ($\text{CF}_3(\text{CF}_2)_3\text{COOH}$), perfluoro(propyl vinyl ether) ($\text{C}_3\text{F}_7\text{O}(\text{FC}=\text{CF}_2)$) from HPFO-DA ($\text{C}_3\text{F}_7\text{OFCCF}_3\text{COOH}$), and perfluorohept-1-ene ($\text{C}_5\text{F}_{11}(\text{FC}=\text{CF}_2)$) from both PFOA and PFOS.^{18, 20, 71} In an oxygen-rich environment, the formed perfluoroalkyl radicals react with O_2 to form oxygen-containing products as observed in this study.

CONCLUSIONS

Comprehensive investigations into the thermal degradation of PFAS, particularly PFCAs and PFSAs, across a range of temperatures (600 °C to 1000 °C), demonstrate a clear temperature-dependent efficacy in their incineration. At higher temperatures, nearly complete to total

degradation is consistently achieved within minimal time frames, underscoring the critical role of temperature in optimizing PFAS disposal. This is particularly evident for shorter-chain PFCAs which show significant degradation even at lower temperatures compared to the more robust PFSAs and PFAS, which require higher temperatures to break down effectively. Overall, the results align with recent studies⁷²⁻⁷³ on the high-temperature destruction of PFAS under incineration conditions, supporting incineration as a promising approach for managing PFAS-containing waste materials. However, the variability observed in full-scale incinerators⁷⁴⁻⁷⁸ suggests that while incineration is a mature technology for conventional solid waste, it remains in its early stages for PFAS and PFAS-impacted materials. This gap highlights the need for tailored operational guidelines and standardized emission control strategies. Our findings on GAC suggest it may serve as an effective post-treatment option to capture fugitive emissions and PFAS intermediates (PIPs). Additionally, co-injection of powdered activated carbon with solid waste may help mitigate emissions, though further research is needed to confirm its effectiveness.

The experiments also highlight the distinctive behaviors of different PFAS classes under thermal stress. While all PFAS types show increased degradation rates with temperature, the extent and speed of degradation vary. This variability suggests that incineration strategies might need to be tailored to the specific PFAS being targeted, to maximize efficiency and minimize residual environmental impact. Understanding these nuances allows for more informed decisions in the design of waste treatment facilities and could lead to the development of more effective PFAS remediation technologies.

A clear trend shows that incineration temperatures (800 °C to 1000 °C) drastically reduce the time required for near-total to total degradation of all PFAS types, emphasizing the effectiveness of high-temperature incineration processes.

PFAS, PFCAs, and PFESAs degrade more rapidly than FTS, PFECAs, and FOSSAs,⁴ suggesting differences in chemical stability and degradation dynamics. For instance, achieving over 99.99% degradation at 800 °C happens within 5 to 10 min for PFAS but takes 15 min or more for FTS.

The data suggest optimal operational temperatures for incinerators should be at least 800 °C⁵ to ensure efficient PFAS degradation within practical time frames, potentially leading to significant energy and cost savings in waste management processes.

Facilities may need to adjust incineration temperatures depending on the specific PFAS compounds in the waste to ensure efficient degradation, considering that some compounds like PFECAs and FOSSAs respond differently at varying temperatures.

Furthermore, the findings offer substantial insights into the practical applications of thermal treatments for environmental management. The ability to achieve high degradation rates within a few minutes at temperatures of 800°C and above could significantly enhance the operational efficiencies of waste treatment processes. This could lead to reduced operational costs and lower environmental footprints of such facilities. Moreover, these results provide a valuable dataset that can guide future research into PFAS decomposition, facilitating the development of more robust models for predicting the behavior of these compounds under various conditions.

This study offers new insights into the thermal degradation mechanisms of PFCs and highlights the role of GAC in enhancing fluorine release. It is the first to comprehensively examine the thermal stability, degradation byproducts, and pathways for both unsaturated and saturated long-chain PFCs, which are significant as industrial chemicals and potential degradation products of PFAS. Understanding these mechanisms is crucial for their degradation and for predicting the fate of PFCs in thermal environments.

Key findings demonstrate a C–C scission mechanism along the perfluorinated backbone,¹ clarifying the thermal breakdown process. This research is also one of the first to investigate the catalytic thermal degradation of PFCs, indicating that GAC significantly enhances the thermal mineralization of PFCs, outperforming most commercial catalysts.

Additional results include:²

- Unsaturated PFCs (C_nF_{2n}) showed nearly complete degradation at temperatures³ above 300 °C.
- Saturated PFCs (C_nF_{2n+2}) required up to 600 °C for complete degradation, due to stronger C–F bonds and greater molecular stability.
- Byproducts from thermal degradation primarily included shorter-chain saturated and unsaturated PFCs, with variations in temperature significantly affecting the variety and abundance of these byproducts.
- Atmospheric conditions impacted outcomes; air presence boosted mineralization at lower temperatures, while an inert nitrogen atmosphere was more effective above 400 °C possibly due to reduced competition from oxygen radicals.

Lastly, the research opens avenues for exploring continuous improvement in thermal treatment technologies. By understanding the specific conditions under which PFAS decompose most efficiently, efforts can be directed towards optimizing incinerators and perhaps even developing new technologies that can achieve similar results at lower temperatures or with enhanced environmental safety. This could ultimately contribute to the global efforts in reducing the persistence of harmful pollutants in the environment, aligning with broader environmental protection and sustainability goals.⁴

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