

CML103 - Major (Part-1)

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Q2.

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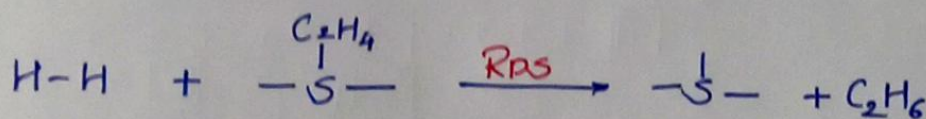
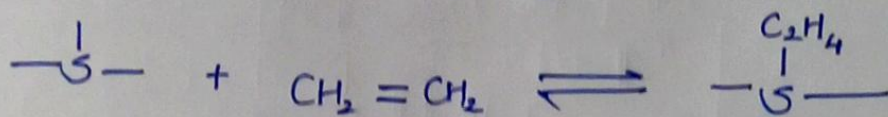
$$V = \frac{K[H_2][C_2H_4]}{(1 + K[C_2H_4])}$$

→ Mechanism of Addition of Hydrogen on Ethylene on Nickel surface

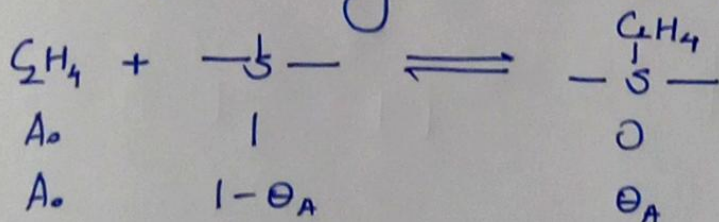
The Given Rate eqⁿ suggest 3 possibilities ;

1. Rate = $K\theta_A\theta_B$; Rate = $K\theta_A[B]$; Rate = $K\theta_B[A]$

∴ As in the question H_2 is gas free Hence (2) Eqⁿ



∴ Now proceeding to calculate θ_A



$$\frac{K_1}{K_{-1}} = \frac{\theta_A}{[C_2H_4][1 - \theta_A]}$$

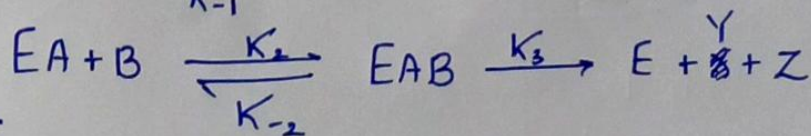
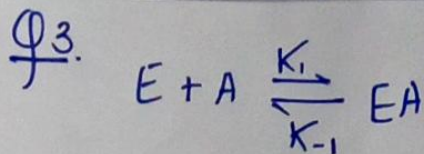
$$\frac{K_1}{K_{-1}} = K = \frac{\theta_A}{[1 - \theta_A][C_2H_4]}$$

$$\theta_A = K[1 - \theta_A][C_2H_4]$$

$$\theta_A = \frac{K[C_2H_4]}{1 + K[C_2H_4]}$$

$$\begin{aligned} \therefore \text{Rate of Rxn} &= K\theta_A[B] = \overset{K_0}{K \times K} [C_2H_4] \times [H_2] \\ &= \frac{K_0[H_2][C_2H_4]}{1 + K[C_2H_4]} \end{aligned}$$

→ Proved the mechanism.



$[E_0] = [E] + [EA] + [EAB]$

$\frac{[E_0]}{[EAB]} = \frac{[E]}{[EAB]} + \frac{[EA]}{[EAB]} + 1 \quad \text{--- (1)}$

$\frac{k_1}{k_{-1}} = \frac{[EA]}{[E][A]} \quad \text{--- (2)}$

$\frac{k_2}{k_{-2}} = \frac{[EAB]}{[EA][B]} \quad \text{--- (3)}$

$\therefore \frac{k_{-2}}{k_2[B]} = \frac{[EA]}{[EAB]} \quad \text{--- (4)}$

\therefore from (1) :-

$\frac{[E_0]}{[EAB]} = \frac{[E]}{[EAB]} + \frac{[EA]}{[EAB]} + 1$

$= \frac{[E]}{[EAB]} + \frac{k_{-2}}{k_2[B]} + 1 \quad \dots \text{from (4)} \quad \text{--- (x)}$

\therefore from (2) & (3) :-

$[E] = \frac{k_{-1}[EA]}{k_1[A]} \quad ; \quad [EA] = \frac{k_{-2}}{k_2} \times \frac{[EAB]}{[B]}$

$[E] = \frac{k_{-1}}{k_1} \times \frac{k_{-2}}{k_2} \times \frac{[EAB]}{[A][B]} \quad \text{--- (5)}$

\therefore Putting (5) in (x) :-

$\frac{[E_0]}{[EAB]} = \frac{k_{-1}k_{-2}}{k_1k_2} \frac{[EAB]}{[A][B][EAB]} + \frac{[EA] \times k_1[A]}{k_2[B]} + \frac{1}{1} \times \frac{k_1k_2[A][B]}{k_1k_2[A][B]}$