Latex 语法

这里只是一些例子，语法手册请打开链接查看

https://katex.org/docs/supported.html

latex 语法支持

https://katex.org/docs/support_table.html

$$\mathring{\mathrm{A}}$$

$$\begin{matrix} a \\ b \end{matrix}$$

$$\binom{n+1}{k+2}$$

$$y^2 + y^2$$

$$\cancel{5}$$

$$\cfrac{2}{1+\cfrac{2}{1+\cfrac{2}{1}}}$$

$$\frac{a}{b+1}$$

$$\sqrt{y}$$

$$\frac{a-1}{b-1}$$

$$123\ \frac{\mathrm{kJ}}{\mathrm{mol}}$$

$$x_1 + \cdots + x_n$$

$$\$x_1 x_2 \cdots x_n$$

$$\{x \in \mathbb{R} \mid x > 0\}$$

$$\int_{A_1} \int_{A_2} \cdots$$

$$\left\|\begin{matrix} a & b \\ c & d \end{matrix}\right\|$$

$$\text{Æ}$$

$$\sqrt[3]{x}$$

$$\overbrace{a+b+c}^{\text{note}}$$

$$\begin{cases} a & \text{if } b \\ c & \text{if } d \end{cases}$$

$$\underbrace{a+b+c}_{\text{note}}$$

$$x = \begin{cases} a & \text{if } b \\ c & \text{if } d \end{cases}$$

$$a = b + c$$

$$d + e = f$$

$$\stackrel{!}{=}$$

$$\begin{Bmatrix} n \\ k \end{Bmatrix}$$

$$\sum_{i=1}^{n}$$

$$\left[\frac{a}{a+1}\right]$$

$$\overleftarrow{AB}$$

$$10x + 3y = 2$$

$$3x + 13y = 4$$

$$\sum_{i=1}^{n}$$

$$\underbrace{a+b+c}_{\text{note}}$$

$$\sum_{\substack{0<i<m\\0<j<n}}$$

$$\lim_{x}$$

$$\overleftrightarrow{AB}$$

$$\sum_{1\leq i\leq j\leq n} x_{ij}$$

a	b	c
d	e	f
g	h	i

$$\begin{Bmatrix} a & b \\ c & d \end{Bmatrix}$$

$$\begin{vmatrix} a & b \\ c & d \end{vmatrix}$$

$$\mathrm{C_6H_5-CHO}$$

$$\begin{bmatrix} a & b \\ c & d \end{bmatrix}$$

$$\overbrace{AB}$$

$$\widehat{AB}$$

数学公式

$$\left(\sum_{k=1}^{n} a_k b_k\right)^2 \leq \left(\sum_{k=1}^{n} a_k^2\right) \left(\sum_{k=1}^{n} b_k^2\right)$$

$$\frac{1}{\left(\sqrt{\phi\sqrt{5}}-\phi\right) e^{\frac{2}{5}\pi}} = 1+\frac{e^{-2\pi}}{1+\frac{e^{-4\pi}}{1+\frac{e^{-6\pi}}{1+\frac{e^{-8\pi}}{1+\cdots}}}}$$

$$1+\frac{q^2}{(1-q)}+\frac{q^6}{(1-q)(1-q^2)}+\cdots = \prod_{j=0}^{\infty}\frac{1}{(1-q^{5j+2})(1-q^{5j+3})}, \qquad \text{for } |q|<1.$$

$$\forall x \in X, \quad \exists y \leq \epsilon$$

$$\cos(2\theta) = \cos^2\theta - \sin^2\theta$$

$$\lim_{x \to \infty} \exp(-x) = 0$$

$$k_{n+1} = n^2 + k_n^2 - k_{n-1}$$

$$f(n) = n^5 + 4n^2 + 2|_{n=17}$$

$$\sqrt[n]{1 + x + x^2 + x^3 + \cdots + x^n}$$

$$P\left(A = 2 \middle| \frac{A^2}{B} > 4\right)$$

$$\left.\frac{x^3}{3}\right|_0^1$$

$$x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$$

$$\begin{aligned} (a+b)(a^2 - ab + b^2) \\ = a^3 - a^2b + ab^2 + a^2b + ab^2 + b^2 \\ = a^3 + b^3 \end{aligned}$$

化学公式

注意mathjax和katex的语法不同之处

https://mhchem.github.io/MathJax-mhchem/

ce

$$\mathrm{CO_2}$$

$$\mathrm{H^+}$$

$$\beta^+$$

$$\mathrm{[\{(X_2)_3\}_2]^{3+}}$$

$$\mathrm{NH_4{}^+}$$

$$\mathrm{\tfrac{1}{2}\,H_2O}$$

$$\mathrm{^{227}_{90}Th^+}$$

$$\mathrm{H^3HO}$$

$$\mathrm{A \xrightarrow[text below]{text above} B}$$

$$\mathrm{NaOH\,(aq, \infty)}$$

$$\mathrm{ZnS\,(c)}$$

$$\mathrm{ZnS(\sim c\ )}$$

$$\mathrm{A + B}$$

$$\mathrm{A \pm B}$$

$$\mathrm{N_2\,(g) + 3\,H_2\,(g) \longrightarrow 2\,NH_3\,(g)} \qquad \Delta H_\mathrm{f}^\circ = -92.5 kJ$$

$$\mathrm{CO_2 + C \xrightarrow[below]{above} 2\,CO}$$

$$\mathrm{2\,H_2 + O_2 \xrightarrow{\Delta} H_2O}$$

$$\mathrm{2\,NaHCO_3 \xlongequal{heating} Na_2CO_3 + H_2O + CO_2\uparrow}$$

$$\mathrm{Hg^{2+} \xrightarrow{I^-} HgI_2 \xrightarrow{I^-} [Hg^{II}I_4]^{2-}}$$

$$\mathrm{^{40}_{18}Ar + \gamma + \nu_e}$$

$$\mathrm{Fe(CN)_{\frac{6}{2}}}$$

$$\mathrm{cis\text{-}[PtCl_2(NH_3)_2]}$$

$$\mathrm{{SO_4}^{2-} + Ba^{2+} \longrightarrow BaSO_4\downarrow}$$

$$\mathrm{CH_4 + 2\,\left(\,O_2 + \tfrac{79}{21}\,N_2\,\right)}$$

$$\mathrm{x\,Na(NH_4)HPO_4 \xrightarrow{\Delta} (NaPO_3)_x + x\,NH_3\uparrow + x\,H_2O}$$

$$\mathrm{[Pt(\eta^2\text{-}C_2H_4)Cl_3]^-}$$

$$\mathrm{A{-}B{=}C{\equiv}D}$$

$$\mathrm{KCr(SO_4)_2 \cdot 12\,H_2O}$$

$$\mathrm{Li^{\times}_{Li,1-2x}Mg^{\bullet}_{Li,x}\ V^{\prime}_{Li,x}Cl^{\times}_{Cl}}$$

$$\mathrm{CO_2 + 3\,H_2 \rightleftharpoons CH_3OH + H_2O}$$

$$\mathrm{CO_2 + H_2 \rightleftharpoons CO + H_2O}$$

$$\mathrm{CO + 2\,H_2 \rightleftharpoons CH_3OH}$$

$$\mathrm{CH_3OH \rightleftharpoons CH_3OCH_3 + H_2O}$$

$$\mathrm{2\,NaHCO_3 \xlongequal{heating} Na_2CO_3 + H_2O + CO_2\uparrow}$$

$$\mathrm{Zn^{2+} \underset{+2\,H^+}{\overset{+2\,OH^-}{\rightleftharpoons}} \underset{\text{amphoteres Hydroxid}}{Zn(OH)_2\downarrow} \underset{+2\,H^+}{\overset{+2\,OH^-}{\rightleftharpoons}} \underset{\text{Hydroxozikat}}{[Zn(OH)_4]^{2-}}}$$

$$\mathrm{K = \frac{[Hg^{2+}][Hg]}{[{Hg_2}^{2+}]}}$$

$$\mathrm{K = \frac{[Hg^{2+}][Hg]}{[{Hg_2}^{2+}]}}$$

$$\mathrm{Hg^{2+} \xrightarrow{I^-} \underset{\text{red}}{HgI_2} \xrightarrow{I^-} \underset{\text{red}}{[Hg^{II}I_4]^{2-}}}$$

pu

$123\ \mathrm{kJ}$

$123\ \mathrm{mm^2}$

$123\ \mathrm{J\,s}$

$123\ \mathrm{J{\cdot}s}$

$123\ \mathrm{kJ/mol}$

$123\ \mathrm{\frac{kJ}{mol}}$

$1.2 \cdot 10^3\ \mathrm{kJ}$

$1{,}2 \cdot 10^3\ \mathrm{kJ}$

$1.2 \times 10^3\ \mathrm{kJ}$