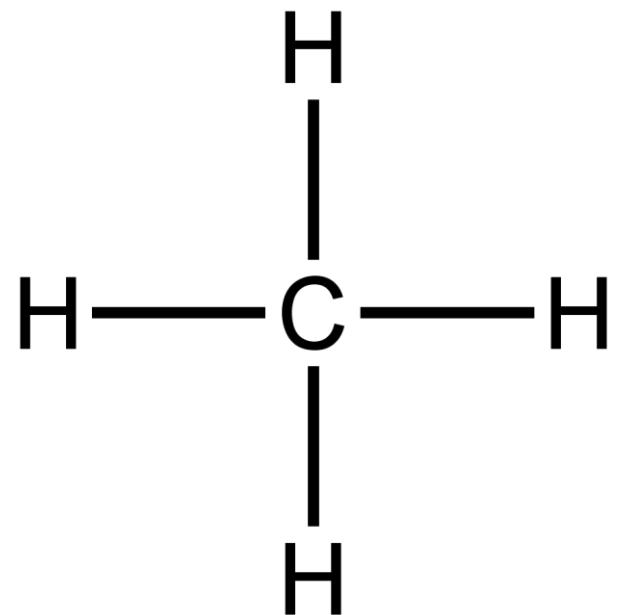


LC3 : Structure spatiale des molécules

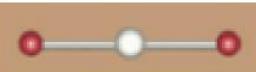
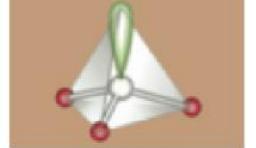
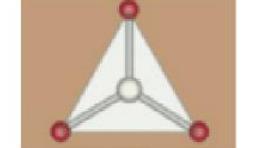
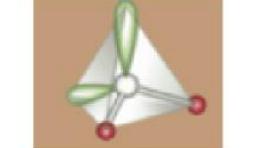
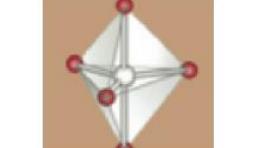
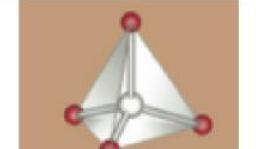
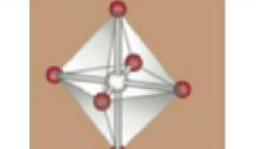
Mathieu Markovitch

Le méthane

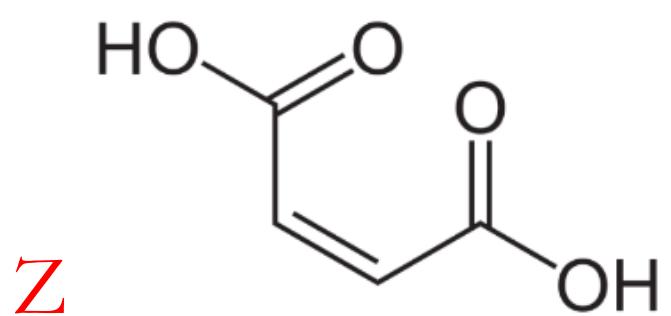


https://phet.colorado.edu/sims/html/molecule-shapes/latest/molecule-shapes_fr.html

Théorie VSEPR

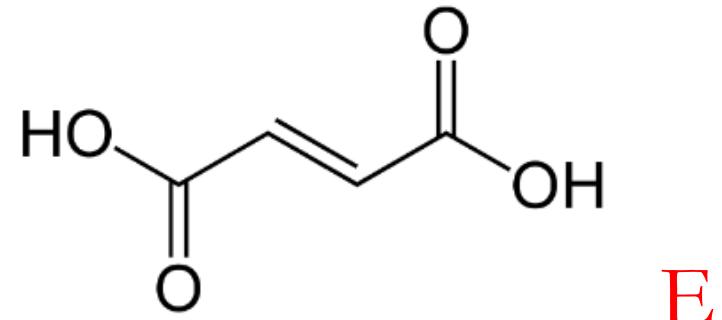
Type de molécule	Forme Idéale	Exemple	Type de molécule	Forme idéale	Exemple
AX_2	 Linéaire	CO_2	AX_3E	 Pyramide à base triangulaire	NH_3
AX_3	 Pyramide trigonale	BF_3	AX_2E_2	 Coudée	H_2O
AX_2E	 Coudée	H_2O	AX_5	 Bipyramide trigonale	PF_5
AX_4	 Tétraédrique	CH_4	AX_6	 Octaédrique	SF_6

Diastéréoisomères



$T_{\text{fus,tab}} = 131 \text{ } ^\circ\text{C}$

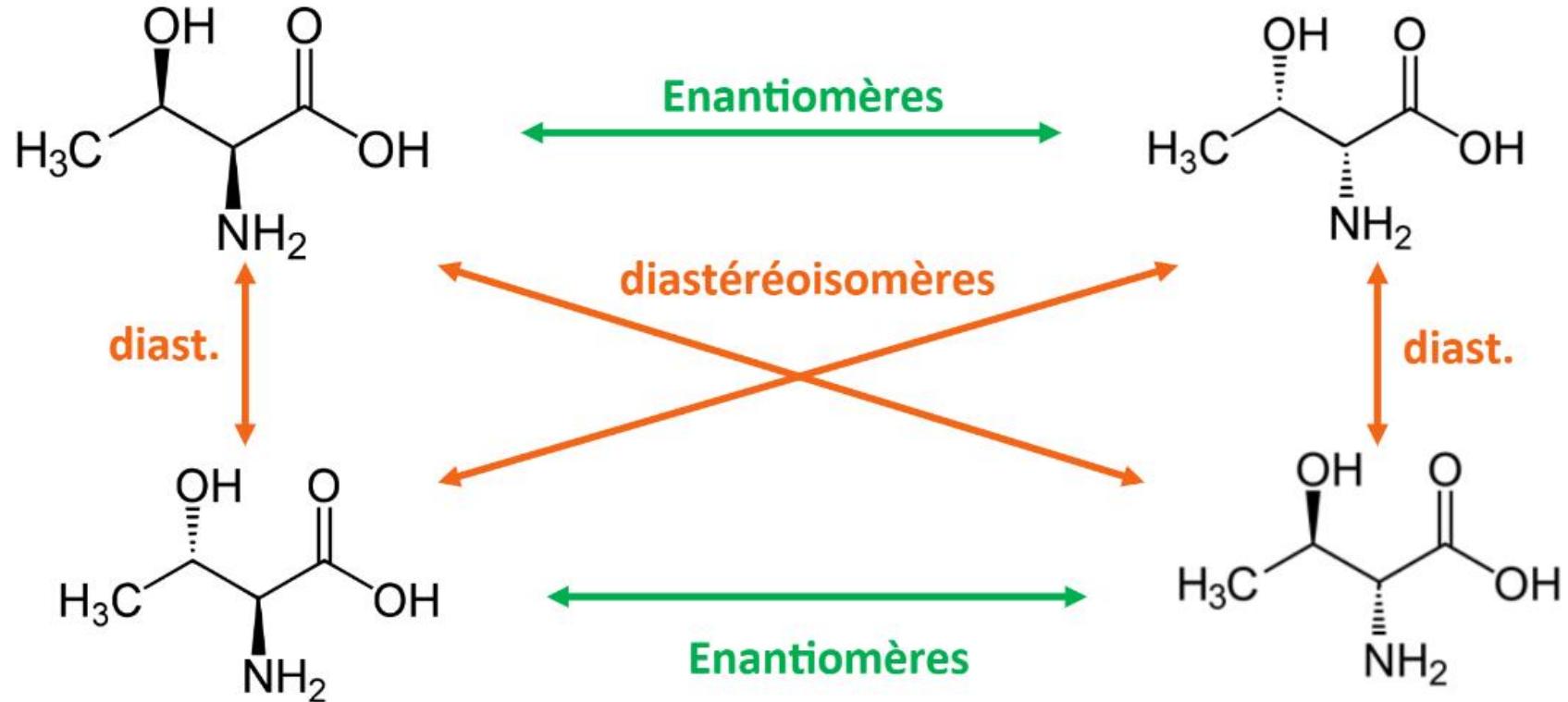
Acide maléique



$T_{\text{fus,tab}} = 287 \text{ } ^\circ\text{C}$

Acide fumarique

Deux carbones asymétriques : la thréonine



Isomérie des molécules

Soient deux molécules de même formule brute

2 isomères

identiques

Formules semi-développées identiques

Formules semi-développées différentes

2 stéréoisomères

2 Isomères de constitution

Simple rotation

Rupture d'une liaison

Stéréoisomères
de conformation

Stéréoisomères
de configuration

Images dans un miroir

Sinon

2 énantiomères

2 diastéréoisomères