$Fe(OH)_3 + 3HNO_3 = Fe(NO_3)_3 + 3H_2O$

19.58 g | 0.8 M | g = 38.70

106.87 | 600 mL | 241.88

0.183 / 0.162 | 0.480 | 0.160

0.023 R

$3Ca(OH)_2 + 2H_3PO_4 = Ca_3(PO_4)_2 + 6H_2O$

g = 14.34 | g = 12.64 | 20.00 g

74.09 | 98.00 | 310.18

0.1935 | 0.129 | 0.0645

$$3\overset{+1}{Na}OH + H_3\overset{-3}{P}O_4 = \overset{+1}{Na}_3\overset{-3}{P}O_4 + 3H_2O$$

$$\overset{+2}{Ca}(OH)_2 + 2H\overset{-1}{N}O_3 = \overset{+2}{Ca}(\overset{-1}{N}O_3)_2 + 2H_2O$$

$$\overset{+3}{Fe}(OH)_3 + 3H\overset{-1}{N}O_3 = \overset{+3}{Fe}(\overset{-1}{N}O_3)_3 + 3H_2O$$

$$\overset{+3}{Fe}(OH)_3 + H_3\overset{-3}{P}O_4 = Fe\overset{-3}{P}O_4 + 3H_2O$$

$$3\overset{+2}{Ca}(OH)_2 + 2H_3\overset{-3}{P}O_4 = Ca_3(PO_4)_2 + 6H_2O$$