

12_11_2014

week 46: Economy

2.5

Suppose that coal-fired power plant burns lignite coal. Based on the energy content of lignite – brown coal (Table 2.1), estimate the CO₂ emissions in kg/GJ and compare your result with emission data in Table 2.2.

Assume that all the C in coal goes into CO₂, with only negligible amount of creating CO.

Lignite: <46% of C (coal) \approx 40%

Energy content in lignite: 5500-8300 Btu/lb = (12.8 - 19.3) MJ/kg

1 btu = 1055 J

1 lb = 0.45 kg

1 Btu/lb = 2.33 kJ/kg

C \rightarrow CO₂ (reaction in mol equivalent)

1 kmol C \rightarrow 1 kmol CO₂

(12 kg/kmol) C \rightarrow (44 kg/kmol) CO₂

0.08 kg C \rightarrow 0.02 kg CO₂

M(C) = 12 g/mol

M(O) = 16 g/mol

M(CO₂) = 44 g/mol

CO₂ emission:

$0.4 \cdot 1/12.8 \text{ (kg/MJ)} = 0.4 \cdot 78.1 \text{ kg/GJ} = 31.2 \text{ kg/GJ}$

$0.4 \cdot 1/19.3 \text{ (kg/MJ)} = 0.4 \cdot 51.8 \text{ kg/GJ} = 20.7 \text{ kg/GJ}$

12 g/mol 31.2 kg/GJ

44g/mol....x kg/GJ

$x_{\max} = (44\text{g/mol} / 12 \text{ g/mol}) \cdot 31.2 \text{ kg/GJ} = 114.4 \text{ kg/GJ (CO}_2\text{)}$

$x_{\min} = 20.7 \text{ kg/GJ}$

$x(\text{average}) = 95.1 \text{ kg/GJ (+/- 5 kg/GJ)}$

Table 2.2 : Emission from lignite power plant: 101 kg/GJ of CO₂

2.8

World proven natural gas reserves $1.9 \times 10^{14} \text{ m}^3$. 1000 kg occupies a volume of 48700 ft^3 at atmospheric pressure, how does the amount trapped in hydrates (estimated as $6.4 \times 10^{12} \text{ tons} = 64 \times 10^{14} \text{ m}^3$) compared to the proven reserves?

$$1 \text{ ft} = 0.0283 \text{ m}^3$$

$$48700 \text{ ft}^3 / 1000 \text{ kg} = 1378 \text{ m}^3 / 1000 \text{ kg}$$

World gas reserves in tones:

$$1.9 \times 10^{14} \text{ m}^3 \text{ t} / 1378 \text{ m}^3 = 1.38 \times 10^{11} \text{ t}$$

2.11

Consider a power-plant having a fixed electrical power output 1000 MW. Show that if the efficiency of the plant were to increase from 33% to 50%, the amount of rejected heat (waste) per MW generated is halved.

$$\eta = \text{efficiency} = (\text{power in} - \text{waste}) / (\text{power in})$$

$$\eta = \text{Power out} / \text{Power in}; \text{Power in} \Rightarrow 0.33 = 1000 \text{ MW} / x \rightarrow 3030.3 \text{ MW}$$

Solution:

if $\eta = 33\%$ waste 67%

if $\eta = 50\%$ waste 50%

$$0.67 \text{ waste} / 0.33 \text{ generated} \approx 2 \text{ MW waste} / 1 \text{ MW generated}$$

$$0.5 \text{ waste} / 0.5 \text{ generated} = 1 \text{ MW waste} / 1 \text{ MW generated}$$

2.14

If your view of “fracking” (hydraulic fracturing) is that it is too risky to be pursued, look up some sources that support this view and, in a one page description, see if you can find any flaws in the arguments. Do the same if your view happens to be that fracking should be pursued.

Hydraulic fracturing

PRO

3x less price of natural gas

cheaper than wind, solar energy

CONS

leaching of liquid and chemicals -> drinking-water contamination

air pollution -> global warming

structural stability of that shale bedrock -> man-made earthquakes

costs of cleaning pollution

amount of water gone

Facts:

underground rivers that provide our drinking water - 20 to 100 m The gas-producing shale rock formations - 1000 to 1800 m that the well you drill to pump the water and released gas- sealed properly. (Density of CO₂ increases with depth)