

Del 1.

Probability. Nuts.

$$f(\mu, \sigma, x) = \frac{1}{\sqrt{2\pi}\sigma} e^{-\frac{1}{2}\left(\frac{x-\mu}{\sigma}\right)^2}$$

vi skal bare integrere udtrykket med hensyn på  $x$ .

Hvis vi antager at  $a, b$  er  $\mathbb{R}$ .

$$\int_{-\infty}^{\infty} \frac{1}{\sqrt{2\pi}\sigma} e^{-\frac{1}{2}\left(\frac{x-\mu}{\sigma}\right)^2} dx = \frac{(x-\mu)^3}{3\sqrt{2\pi}\sigma^3}$$

Hvis vi tar med  $a$  og  $b$  grænse.

$$\int_a^b \frac{1}{\sqrt{2\pi}\sigma} e^{-\frac{1}{2}\left(\frac{x-\mu}{\sigma}\right)^2} dx$$

Så blir svaret.

$$\frac{-(a-b)(a^2 + a(b-3\mu) + b^2 - 3b\mu + 3\mu^2)}{3\sqrt{2\pi}\sigma^3}$$

— Når sannsynlighetsfordelingen ligger mellom  $a$  og  $b$ .

— Integreret for verdier  $(x-\mu)$  er større enn  $-\sigma$  og mindre enn  $\sigma$ .

$$-0 \leq x - \mu \leq \sigma \quad \sigma \leq x - \mu \leq \sigma$$

Det samme som  $\sigma - 1 \leq x \leq \sigma + 1$

$$\int_{-\infty}^{\infty} \frac{1}{\sqrt{2\pi}\sigma} e^{-\frac{1}{2}\left(\frac{x-\mu}{\sigma}\right)^2} dx = \mu \left| \frac{(x-\mu)^2}{\sqrt{2\pi}\sigma^3} \right| + \mu \quad \text{Care...}$$

Videa...

Vi integrerer x...

Svar skal bli 6.8%

Sett  $\mu = 0$  og  $\sigma = \sqrt{\frac{kT}{m}}$  Det her er Maxwell Boltzmann? Temperatur til gassen?

$$\sigma = \frac{FWHM}{\sqrt{8 \ln 2}}$$

6.8% Vi trenger middelværdi  $\mu$  og standarddeviasjon.

—  $FWHM = \frac{2\sqrt{2 \ln 2} \sigma}{\text{mulig feil}}$  bruk  $\mu = 168 \text{ Å}^2$   
 $\sigma = 610 \text{ Å}^2$   
 blir feil

$$\sigma = \frac{FWHM}{\sqrt{8 \ln 2}}$$

AST 3.2

Derive the Ideal gas law.  $P = n k T$

$$P = \frac{1}{3} \int_0^{\infty} p v n(p) dp$$

$n(p) = n(p)$  er tetthetsdistribusjonen.

$$P = \frac{1}{3} \int_0^{\infty} p v n(p) dp$$

$$P = \frac{1}{3} \int_0^{\infty} \frac{p^2}{m} n(p) dp = \frac{1}{3m} \int_0^{\infty} n \left( \frac{1}{2 \tilde{n} m k T} \right)^{3/2} -$$

$$4 \tilde{n} e^{-p^2 / 2 m k T} \cdot p^4 dp$$

Substitusjonen  $x = \frac{p^2}{2 m k T}$  slik at  $p^2 = 2 m k T x$

$$\text{videre blir } dp = \frac{1}{2} \sqrt{\frac{2 m k T}{x}} dx$$

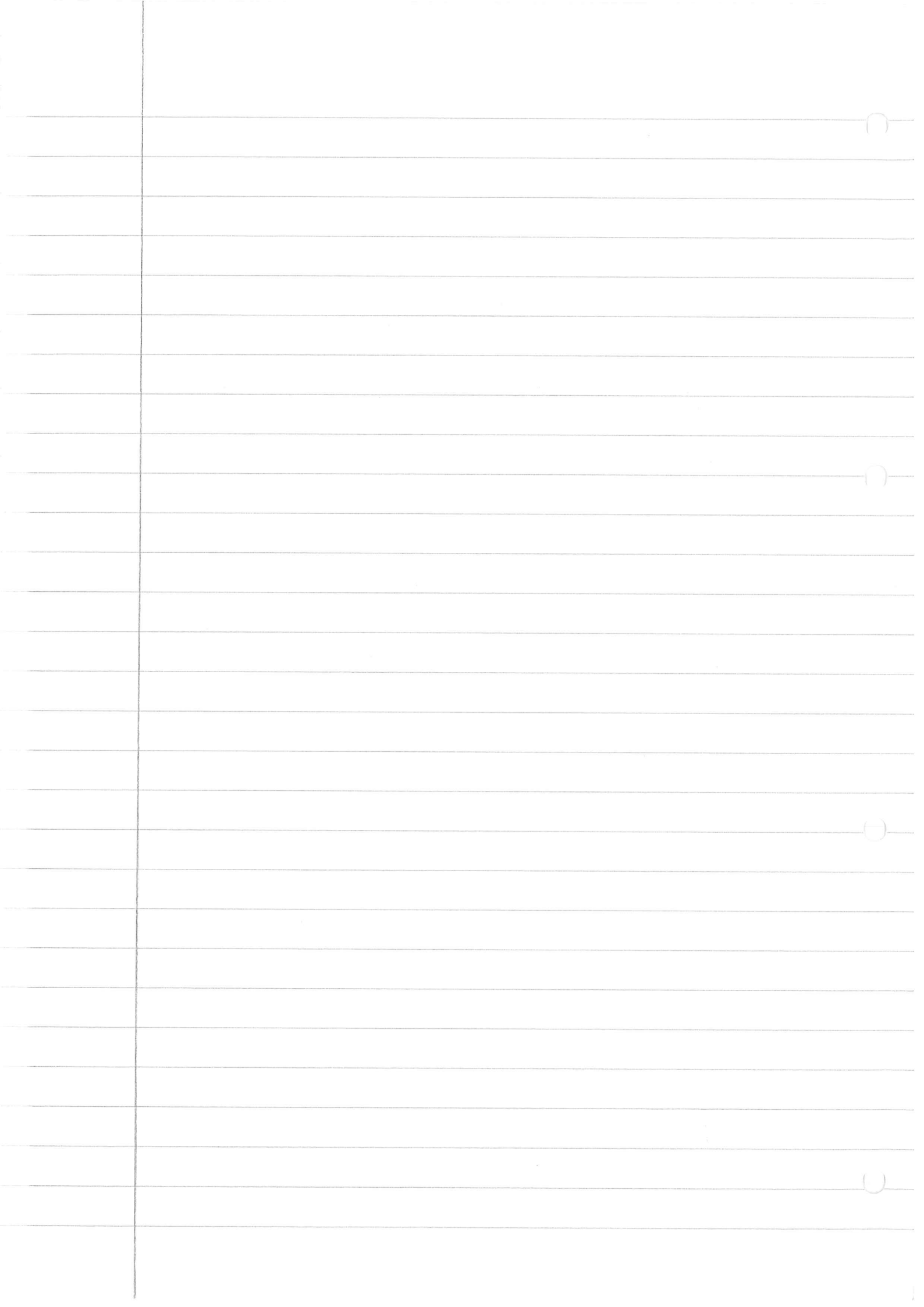
$$\text{Substituerer } P = \frac{4 \tilde{n} m}{3 m} \left( \frac{1}{2 \tilde{n} m k T} \right)^{3/2} \int_0^{\infty} e^{-x (2 m k T x)^2} \frac{1}{2} \sqrt{\frac{2 m k T}{x}} dx$$

$$\text{legger sammen. } P = \frac{2 \tilde{n} m}{3 m} \left( \frac{1}{2 \tilde{n} m k T} \right)^{3/2} (2 m k T) \int_0^{\infty} e^{-x} x^{3/2} dx$$

Vi bruker gamma funksjonen  $\Gamma\left(\frac{5}{2}\right) = \frac{3}{2} \Gamma\left(\frac{3}{2}\right) = \frac{3 \sqrt{\pi}}{4}$

Da får vi

$$P = \frac{\tilde{n} m}{2 m} \left( \frac{1}{2 \tilde{n} m k T} \right)^{3/2} (2 m k T)^{5/2} = \underline{\underline{n k T}}$$



Åst 3,3

Udled den gennemsnitlige energi til  
molekyler i en ideel gas.

$$E = \frac{1}{2} m v^2 = \int_0^{\infty} P(v) \cdot \frac{1}{2} m v^2 dv$$

$$\frac{m}{2} \left( \frac{m}{2\pi kT} \right)^{3/2} 4\pi \int_0^{\infty} e^{-\frac{1}{2} \left( \frac{m v^2}{kT} \right)} \cdot v^2 dv$$

Vi substituerer  $x = \frac{1}{2} \frac{m v^2}{kT}$  således at  $v^2 = \frac{2 x kT}{m}$

Da bliv  $dv = \frac{kT}{v m} dx = \frac{kT \sqrt{m}}{m \sqrt{2 x kT}} dx = \sqrt{\frac{kT}{2 m x}} dx$

Vi sætter ind

$$E = \frac{m}{2} \left( \frac{m}{2\pi kT} \right)^{3/2} 4\pi \int_0^{\infty} e^{-x \left( \frac{2 x kT}{m} \right)^2} \sqrt{\frac{kT}{2 m x}} dx$$

Vi løser integralet.

$$E = \frac{m}{2} \left( \frac{m}{2\pi kT} \right)^{3/2} 4\pi \left( \frac{2 kT}{m} \right)^2 \sqrt{\frac{kT}{2 m}} \cdot \frac{3}{4} \sqrt{\pi}$$

Udtrykket bliver da

$$E = \frac{3}{2} m \left( \frac{m}{2\pi kT} \right)^{3/2} \left( \frac{2\pi kT}{m} \right)^{3/2} \left( \frac{kT}{m} \right)$$

Dette tilsvare  $E = \frac{3}{2} kT$

