

# Maharashtra State Board Class 12 Chemistry Solutions Chapter 13

## Amines

### 1. Choose the most correct option.

Question i.

The hybridization of nitrogen in primary amine is .....

- a. sp
- b. sp<sup>2</sup>
- c. sp<sup>3</sup>
- d. sp<sup>3</sup>d

Answer:

- c. sp<sup>3</sup>

Question ii.

Isobutylamine is an example of .....

- a. 2° amine
- b. 3° amine
- c. 1° amine
- d. quaternary ammonium salt.

Answer:

- a. 2° amine

Question iii.

Which one of the following compounds has the highest boiling point?

- a. n-Butylamine
- b. sec-Butylamine
- c. isobutylamine
- d. tert-Butylamine

Answer:

- a. n-Butylamine

Question iv.

Which of the following has the highest basic strength?

- a. Trimethylamine
- b. Methylamine
- c. Ammonia
- d. Dimethylamine

Answer:

- d. Dimethylamine

Question v.

Which type of amine does produce N<sub>2</sub> when treated with HNO<sub>2</sub>?

- a. Primary amine
- b. Secondary amine
- c. Tertiary amine
- d. Both primary and secondary amines

Answer:

- a. Primary amine

Question vi.

Carbylamine test is given by

- a. Primary amine
- b. Secondary amine
- c. Tertiary amine
- d. Both secondary and tertiary amines

Answer:

- a. Primary amine

Question vii.

Which one of the following compounds does not react with acetyl chloride?

- a. CH<sub>3</sub>-CH<sub>2</sub>-NH<sub>2</sub>
- b. (CH<sub>3</sub>-CH<sub>2</sub>)<sub>2</sub>NH
- c. (CH<sub>3</sub>-CH<sub>2</sub>)<sub>3</sub>N
- d. C<sub>6</sub>H<sub>5</sub>-NH<sub>2</sub>

Answer:

Answer:

c.  $(\text{CH}_3 - \text{CH}_2)_3\text{N}$ 

Question viii.

Which of the following compounds will dissolve in aqueous NaOH after undergoing reaction with Hinsberg reagent?

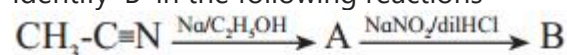
- a. Ethylamine
- b. Triethylamine
- c. Trimethylamine
- d. Diethylamine

Answer:

a. Ethyl amine

Question ix.

Identify 'B' in the following reactions



- a.  $\text{CH}_3\text{-CH}_2\text{-NH}_2$
- b.  $\text{CH}_3\text{-CH}_2\text{-NO}_2$
- c.  $\text{CH}_3\text{-CH}_2\text{N}_2^+\text{Cl}^-$
- d.  $\text{CH}_3\text{-CH}_2\text{-OH}$

Answer:

d.  $\text{CH}_3\text{-CH}_2\text{-OH}$ 

Question x.

Which one of the following compounds contains azo linkage?

- a. Hydrazine
- b. p-Hydroxyazobenzene
- c. N-Nitrosodiethylamine
- d. Ethylenediamine

Answer:

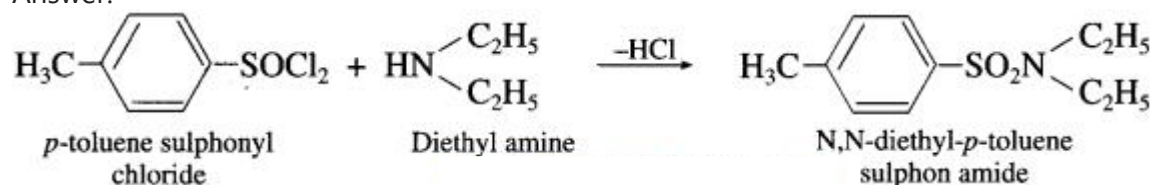
b. p-Hydroxyazobenzene

## 2. Answer in one sentence.

Question i.

Write reaction of p-toluenesulphonyl chloride with diethylamine.

Answer:



Question ii.

How many moles of methylbromide are required to convert ethanamine to N, N-dimethyl ethanamine?

Answer:

2 moles of methylbromide are required to convert ethanamine to N, N-dimethyl ethanamine.

Question iii.

Which amide does produce ethanamine by Hofmann bromamide degradation reaction?

Answer:

Propanamide ( $\text{CH}_3 - \text{CH}_2 - \text{CONH}_2$ ) produces ethanamine by Hofmann bromamide degradation reaction.

Question iv.

Write the order of basicity of aliphatic alkylamine in gaseous phase.

Answer:

The order of basicity of aliphatic alkyl amines in the gaseous follows the order : tertiary amine > secondary amine > primary amine >  $\text{NH}_3$ .

Question v.

Why are primary aliphatic amines stronger bases than ammonia?

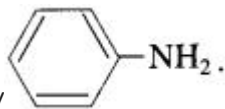
Answer:

The alkyl group tends to increase the electron density on the nitrogen atom. As a result, amines can donate the lone pair of electrons on nitrogen more easily than ammonia. Hence, aliphatic amines are stronger bases than ammonia.

Question vi.

Predict the product of the following reaction. Nitrobenzene Sn/Conc. HCl?

Answer:



The product is aniline/

Question vii.

Write the IUPAC name of benzylamine.

Answer:

The IUPAC name is Phenylmethanamine.

Question viii.

Arrange the following amines in an increasing order of boiling points. n-propylamine, ethylmethyl amine, trimethylamine.

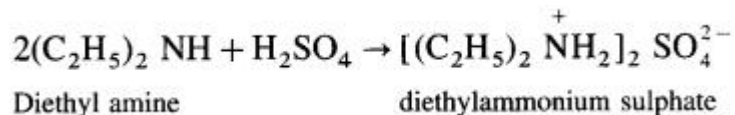
Answer:

Amines in an increasing order of boiling points : trimethyl amine, ethyl methyl amine, n-propyl amine

Question ix.

Write the balanced chemical equations for the action of dil  $\text{H}_2\text{SO}_4$  on diethylamine.

Answer:



Question x.

Arrange the following amines in the increasing order of their  $pK_b$  values. Aniline, Cyclohexylamine, 4-Nitroaniline

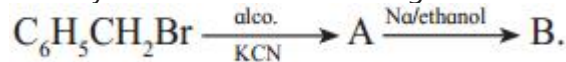
Answer:

Cyclohexyl amine (pK<sub>A</sub> 3.34), aniline (pK<sub>A</sub> 9.13) 4-nitroaniline (pK<sub>A</sub> 12.99)

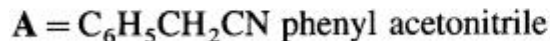
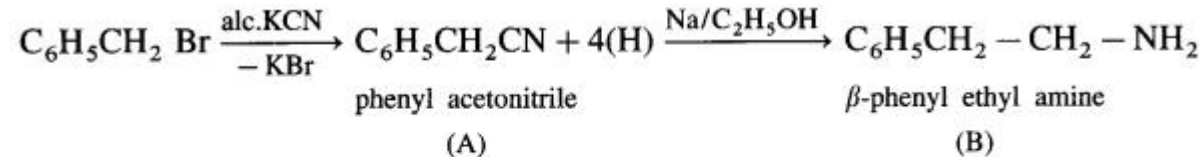
### 3. Answer the following

Question i.

Identify A and B in the following reactions.



Answer:

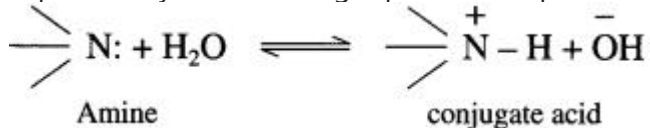


Question ii.

Explain the basic nature of amines with suitable example.

Answer:

The basic strength of amines is expressed in terms of  $K_b$  or  $pK_b$  value. According to Lowry-Bronsted theory the basic nature of amines is explained by the following equilibrium equation.

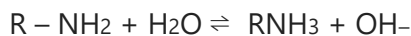


In this equilibrium amine accepts  $H^+$ , hence an amine is a Lowry-Bronsted base.

According to Lewis theory, the species which donates a pair of electrons is called a base.

The nitrogen atom in amides has a lone pair of electrons, which can be donated to suitable acceptor like proton  $H^+$ .

The aqueous solutions of amines are basic in nature due to release of free  $\text{OH}^-$  ions in solutions. Hence amines are Lewis bases. There exists an equilibrium in their aqueous solutions as follows :



Since  $\text{OH}^-$  is a stronger base, equilibrium shifts towards left-hand side giving less concentration of  $\text{OH}^-$ .

Here,  $K_b$  value is smaller and  $pK_b$  value is larger.

Hence amines are weak bases.

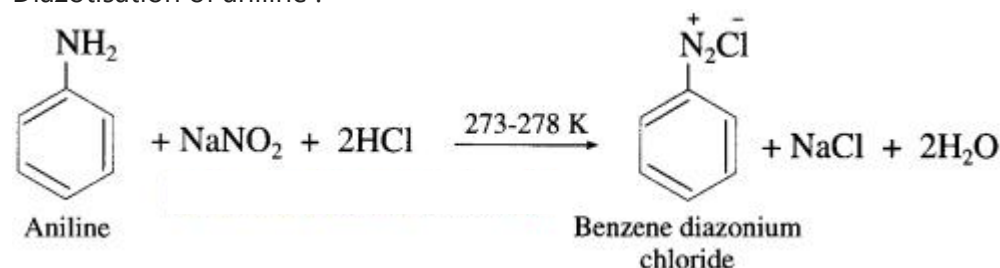
Question iii.

What is diazotisation? Write diazotisation reaction of aniline.

Answer:

Aryl amines react with nitrous acid in cold condition (273 – 278 K) forms arene diazonium salts. The conversion of primary aromatic amine into diazonium salts is called diazotisation.

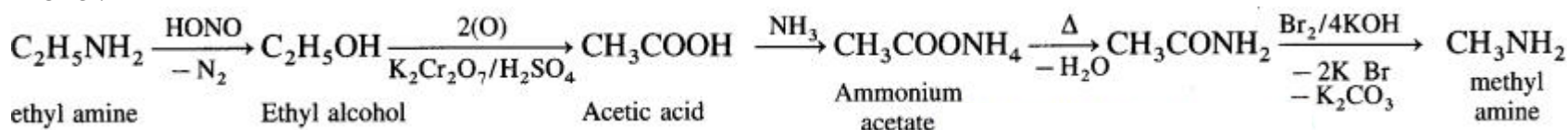
Diazotisation of aniline :



Question iv.

Write reaction to convert acetic acid into methylamine.

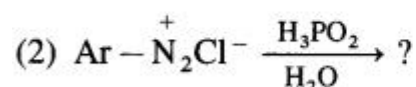
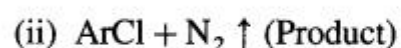
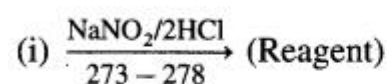
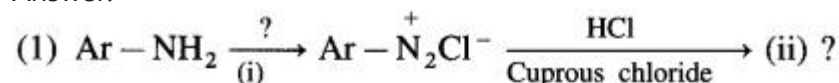
Answer:



Question v.

Write a short note on coupling reactions.

Answer:



Arene

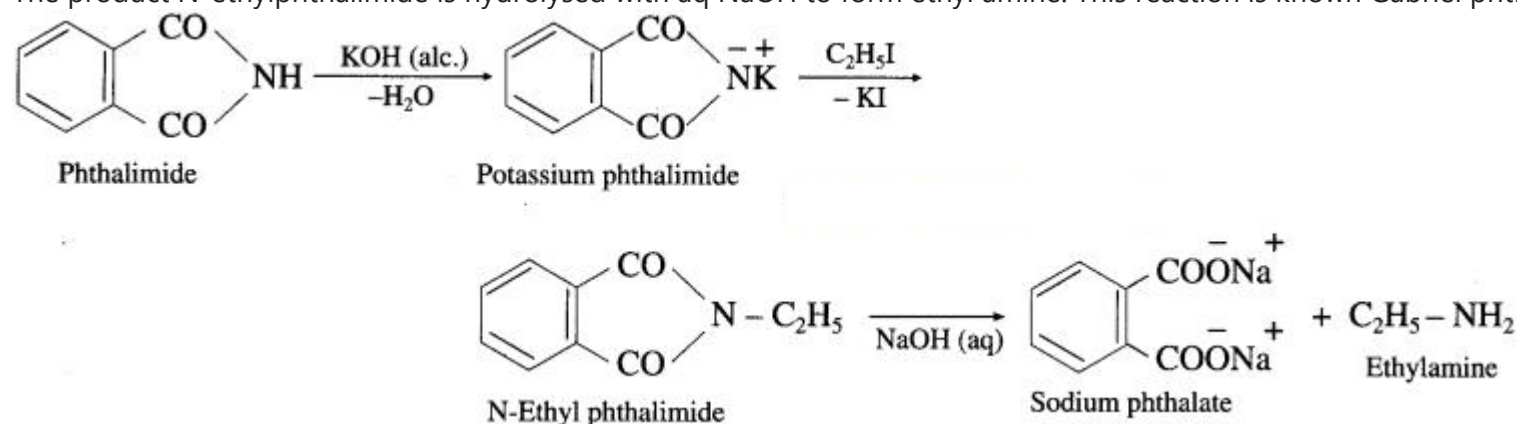
Reactions involving retention of diazo group : (Coupling reactions) :

Question vi.

Explain Gabriel phthalimide synthesis.

Answer:

Phthalimide is reacted with alcoholic KOH to form potassium phthalimide. Further potassium phthalimide is treated with an ethyl iodide. The product N-ethylphthalimide is hydrolysed with aq NaOH to form ethyl amine. This reaction is known Gabriel phthalimide synthesis.



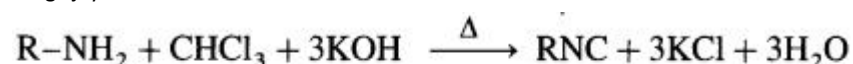
Question vii.

Explain carbylamine reaction with suitable examples.

Answer:

Aliphatic or aromatic primary amines on heating with chloroform and alcoholic potassium hydroxide solution form carbyl amines or isocyanides with extremely unpleasant smell. This reaction is a test for primary amines.

Secondary and tertiary amines do not give this test.



Primary amine (alc.) Alkyl isocyanide



Ethylamine (alc.) Ethyl isocyanide

Question viii.

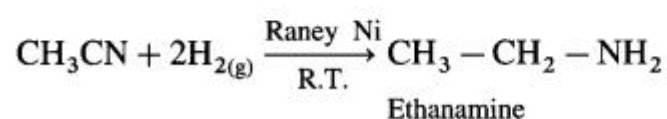
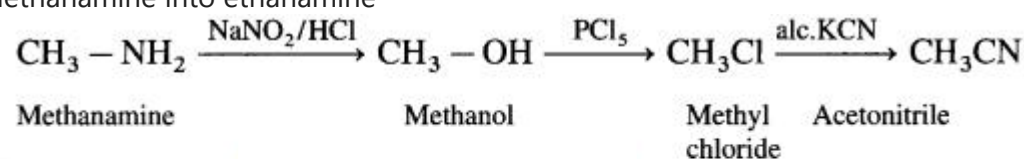
Write reaction to convert

(i) methanamine into ethanamine

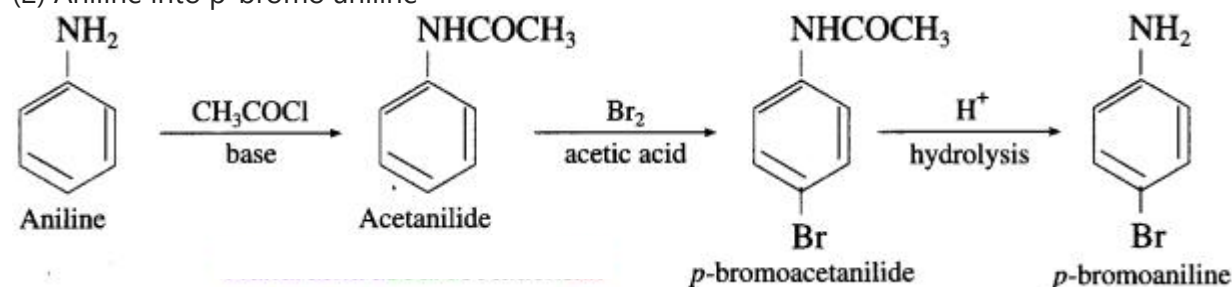
(ii) Aniline into p-bromoaniline.

Answer:

(1) Methanamine into ethanamine



(2) Aniline into p-bromo aniline



Question ix.

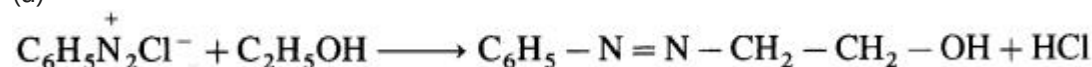
Complete the following reactions :

a.  $\text{C}_6\text{H}_5\text{N}_2^+\text{Cl}^- + \text{C}_2\text{H}_5\text{OH} \rightarrow$

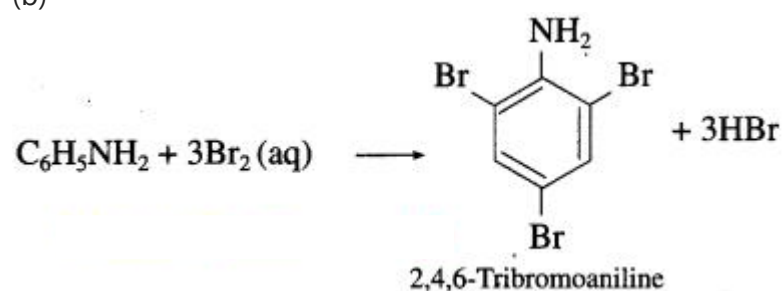
b.  $\text{C}_6\text{H}_5\text{NH}_2 + \text{Br}_2(\text{aq}) \rightarrow ?$

Answer:

(a)



(b)



Question x.

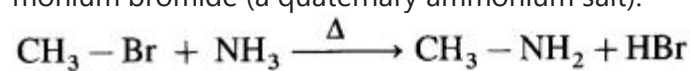
Explain Ammonolysis of alkyl halides.

Answer:

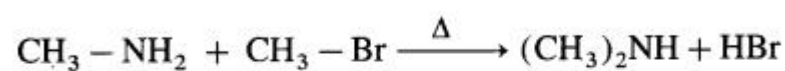
When an alkyl halide is heated with alcoholic ammonia in a sealed tube under pressure at 373 K, a mixture of primary, secondary, tertiary amines and a quaternary ammonium salt is obtained. In this reaction, breaking of C – X bond by ammonia is called ammonolysis of alkyl halides. The reaction is also known as alkylation. For example, when methyl bromide is heated with alcoholic ammonia at 373 K, it gives a mixture of methylamine (a primary amine), dimethylamine (a secondary amine), trimethylamine (a tertiary amine) and tetramethylam-



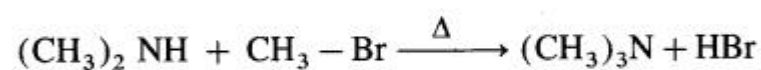
monium bromide (a quaternary ammonium salt).



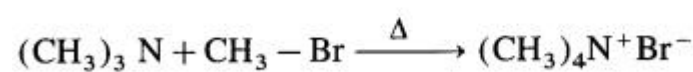
methylbromide alc. , methylamine



methyl amine dimethyl amine



dimethyl amine trimethyl amine



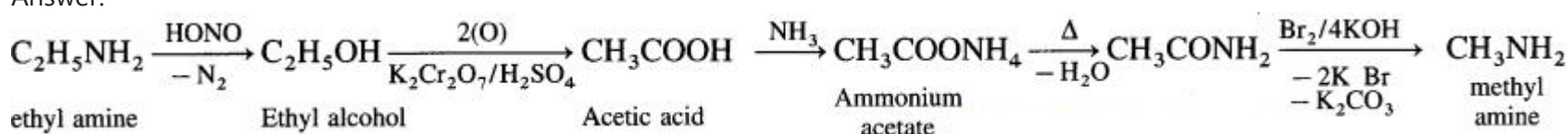
trimethyl amine tetramethyl ammonium bromide  
(quaternary ammonium salt)

The order of reactivity of alkyl halides with ammonia is  $\text{R} - \text{I} > \text{R} - \text{Br} > \text{R} - \text{Cl}$ .

Question xi.

Write reaction to convert ethylamine into methylamine.

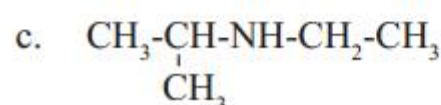
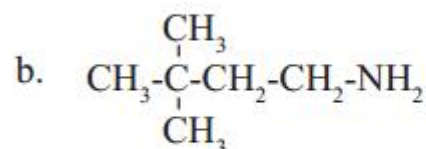
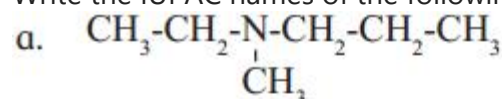
Answer:



4. Answer the following.

Question i.

Write the IUPAC names of the following amines :



Answer:

Compound	IUPAC name
(1) $\text{CH}_3 - \text{CH}_2 - \underset{\text{CH}_3}{\text{N}} - \text{CH}_2 - \text{CH}_2 - \text{CH}_3$	N-Ethyl, N-methyl propan-1-amine
(2) $\text{CH}_3 - \underset{\text{CH}_3}{\overset{\text{CH}_3}{\text{C}}} - \text{CH}_2 - \text{CH}_2 - \text{NH}_2$	3, 3-Dimethylbutan-1-amine
(3) $\text{CH}_3 - \underset{\text{CH}_3}{\text{CH}} - \text{NH} - \text{CH}_2 - \text{CH}_3$	N-Ethyl propan-2-amine

Question ii.

What are amines? How are they classified?

Answer:

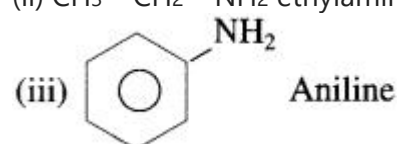
Amines are classified on the basis of the number of hydrogen atoms of ammonia that are replaced by alkyl group. Amines are classified as primary (1°), secondary (2°) and tertiary (3°).

(1) Primary amines (1° amines) : The amines in which only one hydrogen atom of ammonia is replaced by an alkyl group or aryl group are called primary (1°) amines.

Examples :

(i)  $\text{CH}_3 - \text{NH}_2$  methylamine

(ii)  $\text{CH}_3 - \text{CH}_2 - \text{NH}_2$  ethylamine

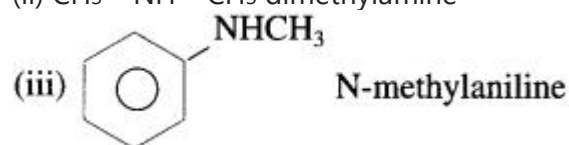


(2) Secondary amines ( $2^\circ$  amines) : The amines in which two hydrogen atoms of ammonia are replaced by two, same or different alkyl or aryl groups are called secondary ( $2^\circ$ ) amines.

Examples :

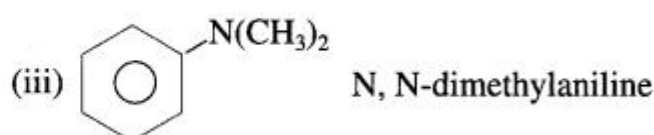
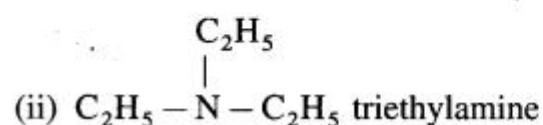
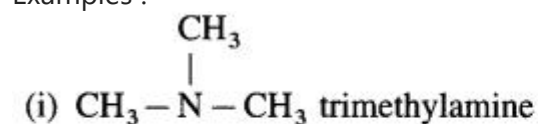
(i)  $\text{C}_2\text{H}_5 - \text{NH} - \text{CH}_3$  ethylmethanamine

(ii)  $\text{CH}_3 - \text{NH} - \text{CH}_3$  dimethanamine



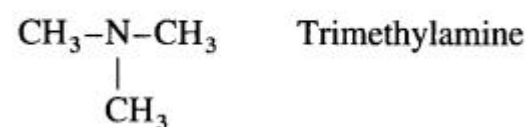
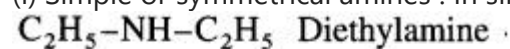
(3) Tertiary amines ( $3^\circ$  amines) : The amines in which all the three hydrogen atoms of ammonia are replaced by three same or different alkyl or aryl groups are called tertiary ( $3^\circ$ ) amines.

Examples :

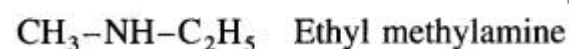
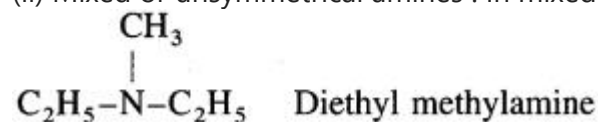


Secondary and tertiary amines are further classified as (1) Simple or symmetrical amines (2) Mixed or unsymmetrical amines.

(i) Simple or symmetrical amines : In simple amines same alkyl groups are attached to the nitrogen e.g.



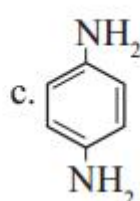
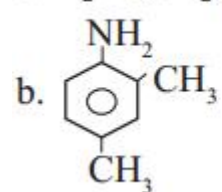
(ii) Mixed or unsymmetrical amines : In mixed amines different alkyl groups are attached to the nitrogen.



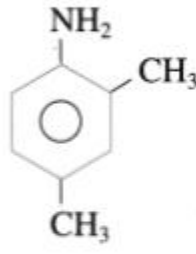

Question iii.

Write IUPAC names of the following amines.

a.  $\text{H}_2\text{N} - (\text{CH}_2)_6 - \text{NH}_2$



Answer:

Compound	IUPAC name
(1) $\text{H}_2\text{N} - (\text{CH}_2)_6 - \text{NH}_2$	Hexan-1, 6-diamine
(2) 	2, 4-Dimethylaniline
(3) 	Benzene-1, 4-diamine

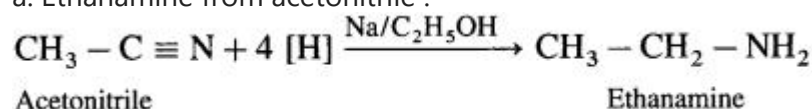
Question iv.

Write reactions to prepare ethanamine from

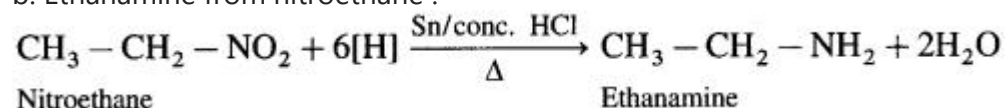
- Acetonitrile
- Nitroethane
- Propionamide

Answer:

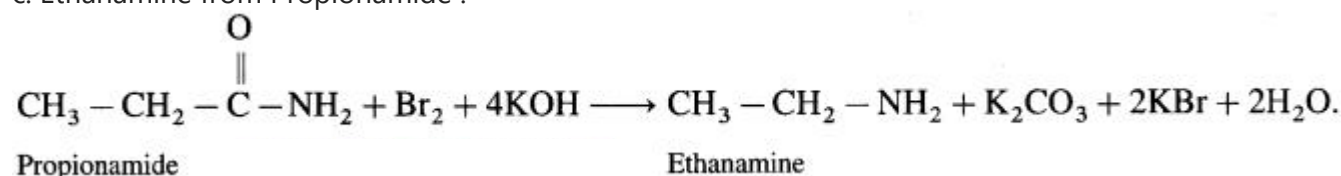
a. Ethanamine from acetonitrile :



b. Ethanamine from nitroethane :



c. Ethanamine from Propionamide :



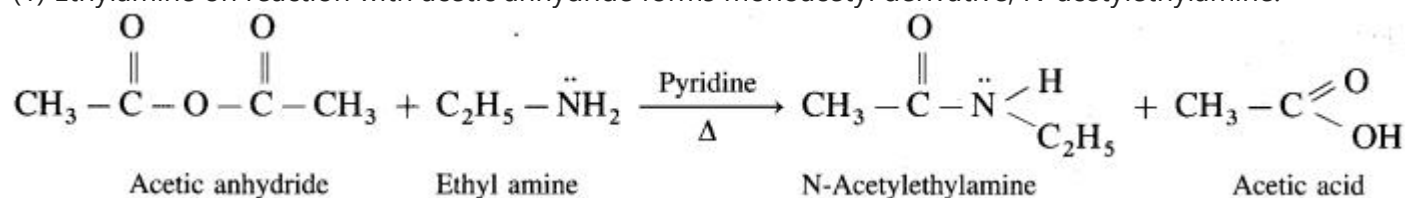
Question v.

What is the action of acetic anhydride on ethylamine, diethylamine and triethylamine?

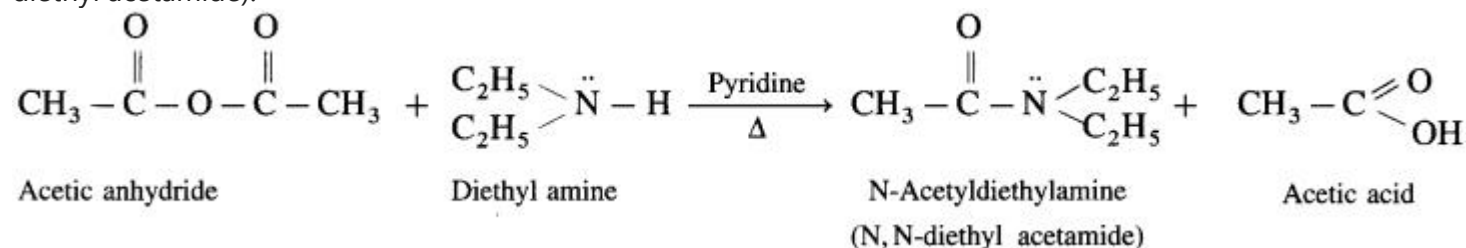
Answer:

Acetylation of amines : The reaction in which the H atom attached to nitrogen in amine is replaced by acetyl group  $(\text{CH}_3 - \overset{\text{O}}{\parallel} \text{C} - )$  is called acetylation of amines.

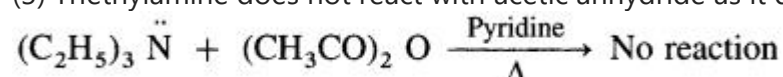
(1) Ethylamine on reaction with acetic anhydride forms monoacetyl derivative, N-acetylethylamine.



(2) Diethylamine (a secondary amine) on reaction with acetic anhydride forms a monoacetyl derivative, N-acetyldiethyl amine (or N,N-diethyl acetamide).



(3) Triethylamine does not react with acetic anhydride as it does not have any H atom attached nitrogen atom of amine



Question vii.

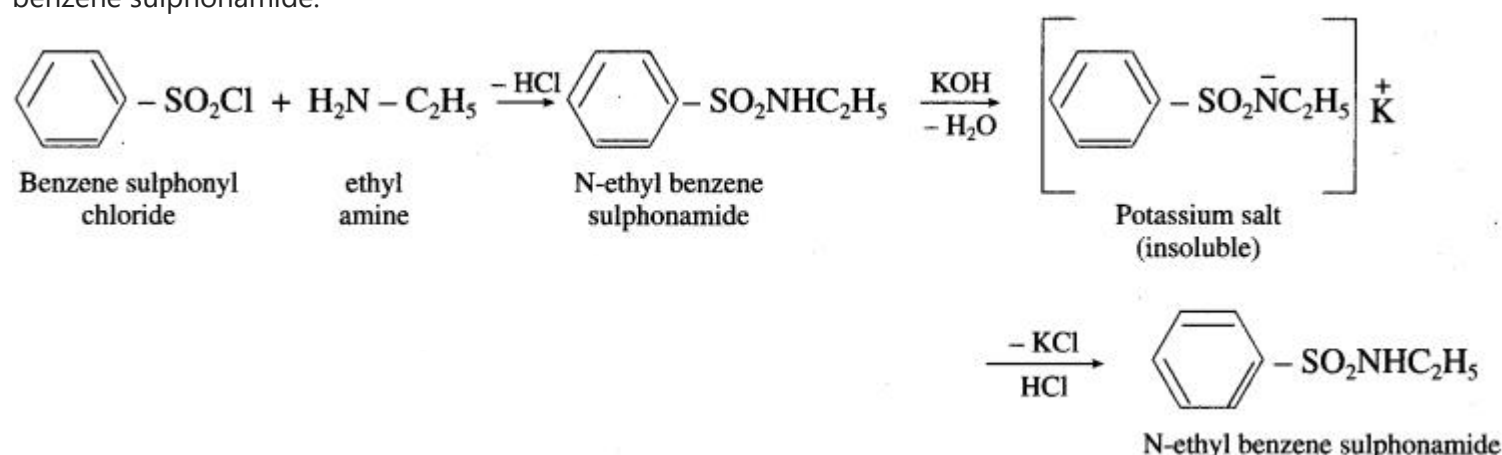
Distinguish between ethylamine, diethylamine and triethylamine by using Hinsberg's reagent?



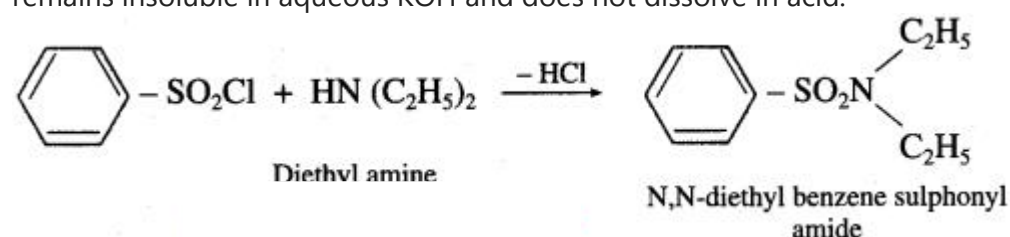
Answer:

This reaction is useful for the distinction of primary, secondary and tertiary amines.

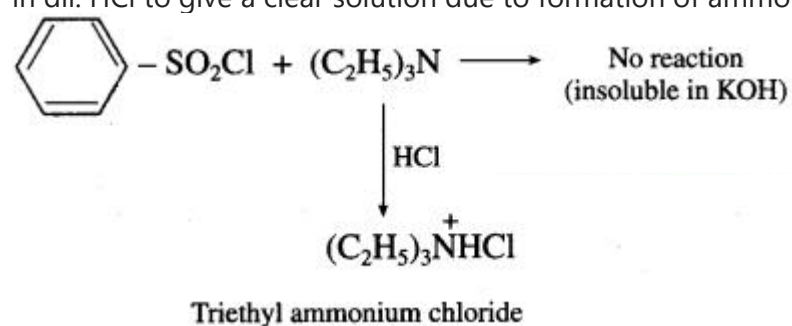
(i) Primary amine (like ethyl amine) is treated with Hinsberg's reagent (benzene sulphonyl chloride) forms N-alkyl benzene sulphonamide which dissolve in aqueous KOH solution to form a clear solution of potassium salt and upon acidification gives insoluble N-alkyl benzene sulphonamide.



(ii) Secondary amine like diethyl amine is treated with benzene sulphonyl chloride forms N,N-diethyl benzene sulphonyl amide remains insoluble in aqueous KOH and does not dissolve in acid.



(iii) Tertiary amine like triethyl amine does not react with benzene sulphonyl chloride and remains insoluble in KOH, however it dissolves in dil. HCl to give a clear solution due to formation of ammonium salt.



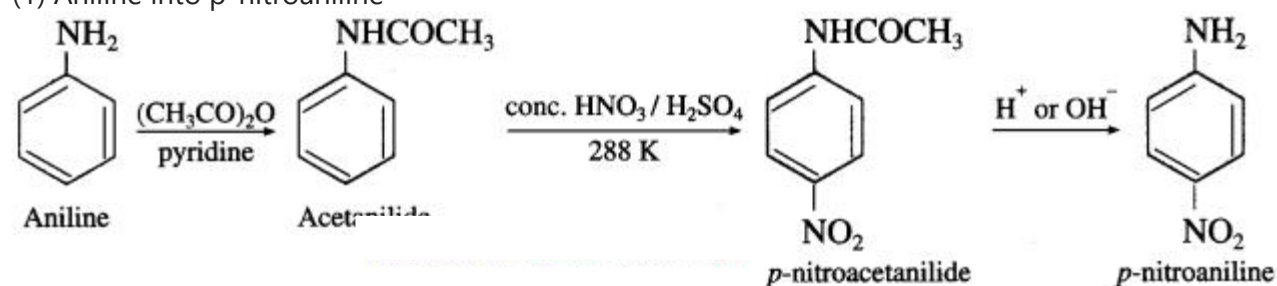
Question viii.

Write reactions to bring about the following conversions :

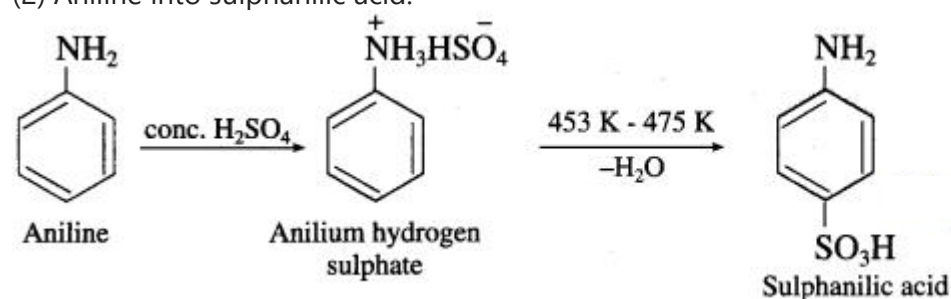
- Aniline into p-nitroaniline
- Aniline into sulphanilic acid?

Answer:

(1) Aniline into p-nitroaniline



(2) Aniline into sulphanilic acid.



Activity :

- Prepare a chart of azodyes, colours and its application.
- Prepare a list of names and structures of N-containing ingredients of diet.

**12th Chemistry Digest Chapter 13 Amines Intext Questions and Answers**

Use your brain power! (Textbook Page No 282)



Question 1.

Draw possible structures of all the isomers of  $C_4H_{11}N$ . Write their common as well as IUPAC names.

Answer:

Compound	Common name	IUPAC name	Primary/Secondary/ Tertiary amines
(1) $CH_3-CH_2-CH_2-CH_2-NH_2$	1-Aminobutane	Butan-1-amine	Primary
(2) $CH_3-CH_2-\underset{\substack{  \\ NH_2}}{CH}-CH_3$	2-Aminobutane	Butan-2-amine	Primary
(3) $CH_3-\underset{\substack{  \\ CH_3}}{CH}-CH_2-NH_2$	1-Amino-2-methyl propane	2-Methyl propan-1-amine	Primary
(4) $CH_3-\underset{\substack{  \\ CH_3}}{\underset{\substack{  \\ CH_3}}{C}}-NH_2$	2-Amino-2-methyl propane	2-Methylpropan-2-amine	Primary
(5) $CH_3-CH_2-CH_2-NH-CH_3$	N-Methylaminopropane	N-Methylpropan-1-amine	Secondary
(6) $CH_3CHNHCH_3$	N-Methyl-2-aminopropane	N-Methylethanamine	Secondary
(7) $CH_3CH_2NHC_2H_5$	N-Ethyl aminoethane	N-Ethylethanamine	Secondary
(8) $CH_3CH_2N\begin{matrix} \nearrow CH_3 \\ \searrow CH_3 \end{matrix}$	N, N-Dimethylaminoethane	N,N-Dimethylethanamine	Tertiary

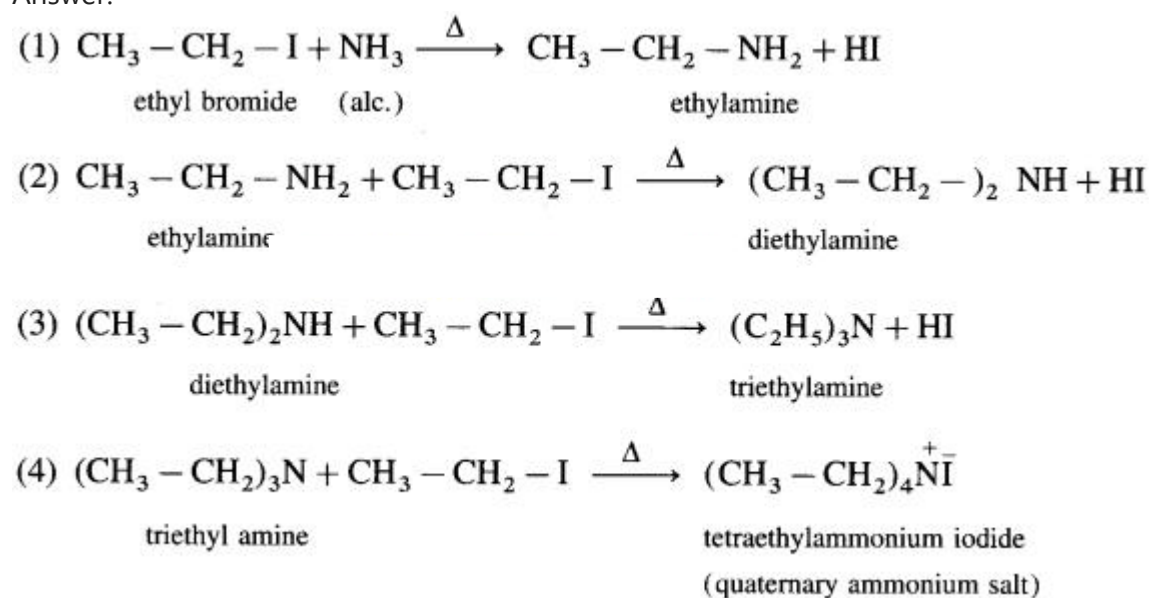
Use your brain power! (Textbook Page No 283)

Question 1.

Write chemical equations for

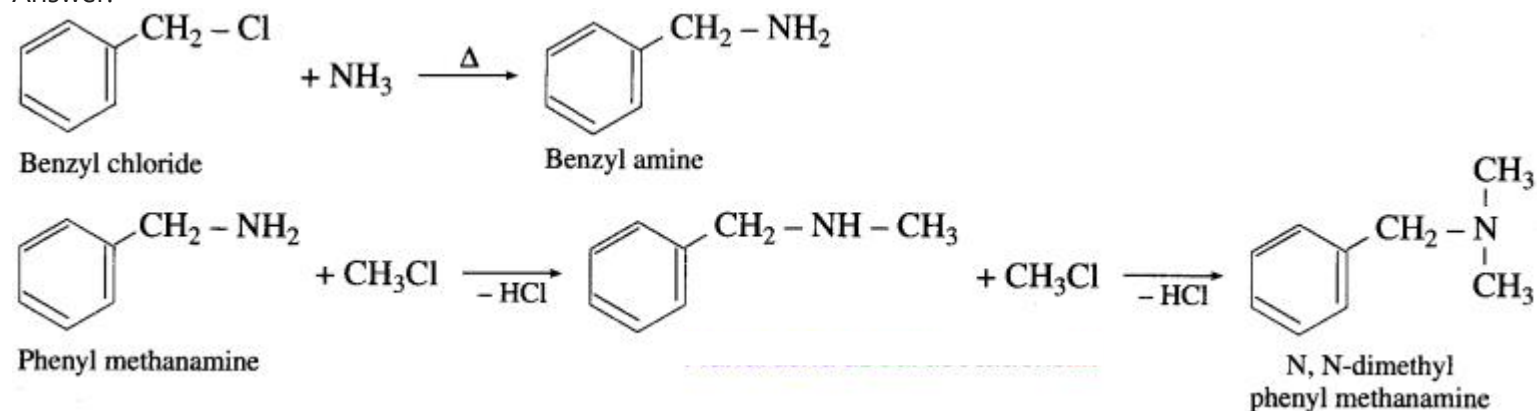
(i) reaction of alc. NH with  $C_2H_5I$ .

Answer:



(ii) Ammonolysis of benzyl chloride followed by the reaction with 2 moles of  $CH_3I$ .

Answer:



(2) Ammonolysis of alkyl halides is not suitable method to prepare primary amines.

Answer:



In the laboratory, ammonolysis of alkyl halides is not a suitable method to prepare primary amines as it gives a mixture of primary, secondary, tertiary amines and quaternary ammonium salts. (Refer to the reaction in answer to Question 16). The separation of primary amine becomes difficult.

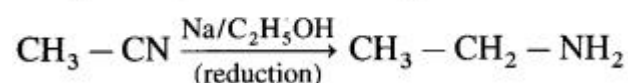
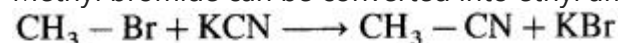
### Problem 13.1 : (Textbook Page No 285)

Question 1.

Write reaction to convert methyl bromide into ethyl amine? Also, comment on the number of carbon atoms in the starting compound and the product.

Solution :

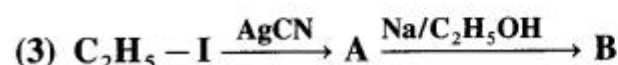
Methyl bromide can be converted into ethyl amine in two stage reaction sequence as shown below.



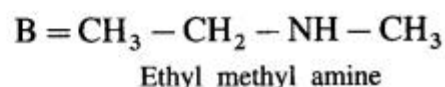
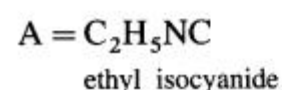
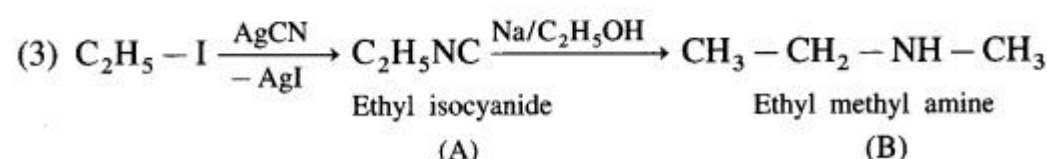
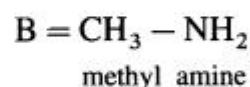
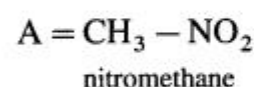
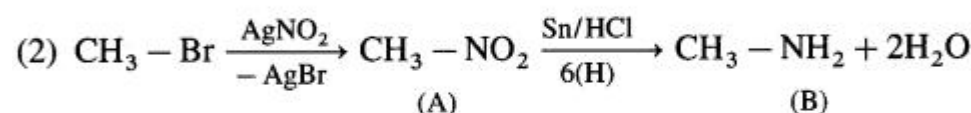
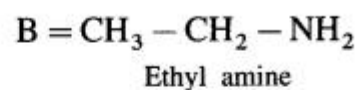
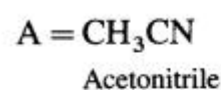
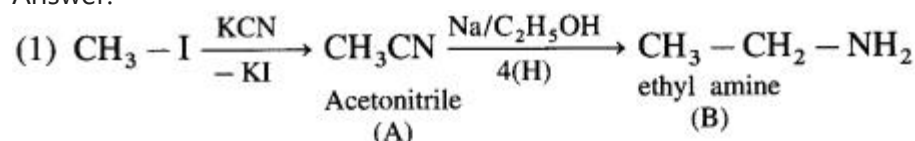
The starting compound methyl bromide contains one carbon atom while the product ethylamine contains two carbon atoms. A reaction in which number of carbons increases involves a step up reaction. The overall conversion of methyl bromide into ethyl amine is a step up conversion.

### Use your brain power! (Textbook Page No 285)

Identify 'A' and 'B' in the following conversions.



Answer:



### Use your brain power! (Textbook Page No 286)

Question 1.

Write the chemical equations for the following conversions :

(1) Methyl chloride to ethylamine.

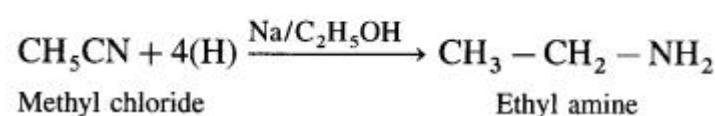
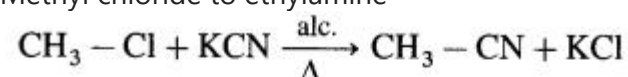
(2) Benzamide to aniline.

(3) 1, 4-Dichlorobutane to hexane-1, 6-diamine.

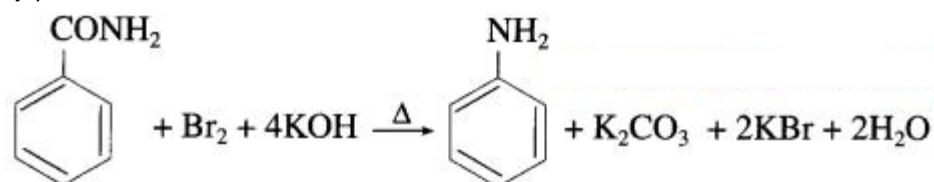
(4) Benzamide to benzylamine.

Answer:

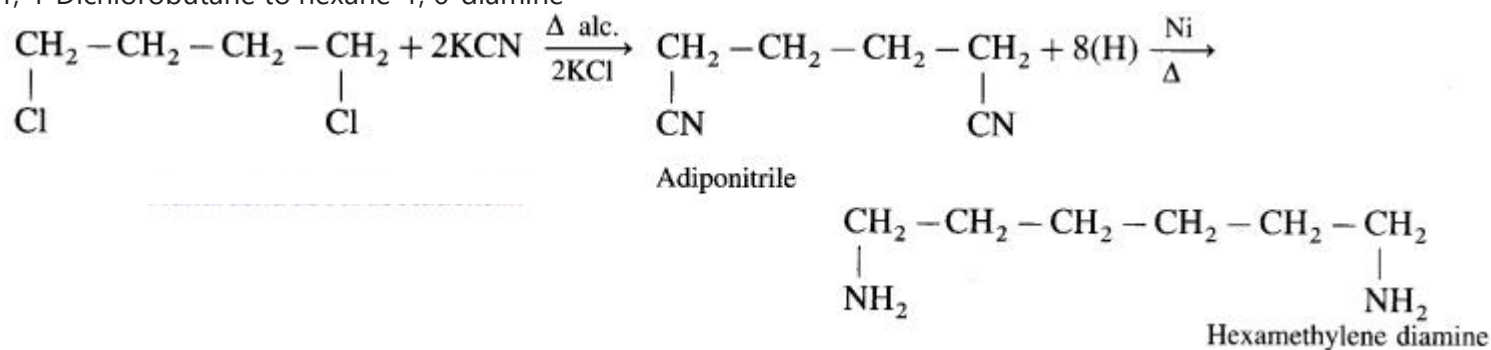
(1) Methyl chloride to ethylamine



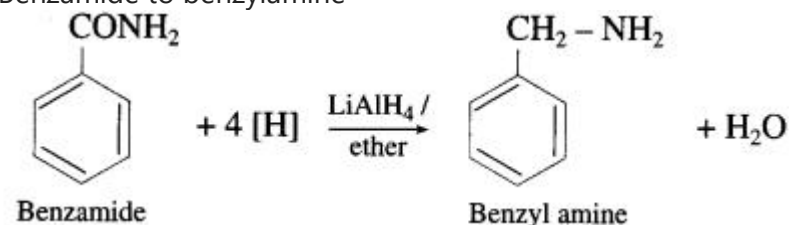
(2) Benzamide to aniline



(3) 1, 4-Dichlorobutane to hexane-1, 6-diamine



(4) Benzamide to benzylamine



Use your brain power! (Textbook Page No 287)

Question 1.

Arrange the following :

- (1) In decreasing order of the boiling point  $\text{C}_2\text{H}_5 - \text{OH}$ ,  $\text{C}_2\text{H}_5 - \text{NH}_2$ ,  $(\text{CH}_3)_2\text{NH}$   
 (2) In increasing order of solubility in water:  $\text{C}_2\text{H}_5 - \text{NH}_2$ ,  $\text{C}_3\text{H}_7 - \text{NH}_2$ ,  $\text{C}_6\text{H}_5 - \text{NH}_2$

Answer:

- (1) Decreasing order of the boiling point :  $\text{C}_2\text{H}_5 - \text{OH}$ ,  $\text{C}_2\text{H}_5 - \text{NH}_2$ ,  $(\text{CH}_3)_2\text{NH}$   
 (2) Increasing order of solubility in water :  $\text{C}_6\text{H}_5\text{NH}_2$ ,  $\text{C}_3\text{H}_7 - \text{NH}_2$ ,  $\text{C}_2\text{H}_5 - \text{NH}_2$

Use your brain power! (Textbook Page No 288)

Question 1.

Refer to  $\text{pK}_b$  values and answer which compound from the following pairs is the stronger base?

- (1)  $\text{CH}_3 - \text{NH}_2$  and  $(\text{CH}_3)_2\text{NH}$   
 (2)  $(\text{C}_2\text{H}_5)_2\text{NH}$  and  $(\text{C}_2\text{H}_5)_3\text{N}$   
 (3)  $\text{NH}_3$  and  $(\text{CH}_3)_2\text{CH} - \text{NH}_2$

Answer:

- (1)  $\text{CH}_3 - \text{NH}_2$  and  $(\text{CH}_3)_2\text{NH}$   
 $(\text{CH}_3)_2\text{NH}$  is a stronger base

- (2)  $(\text{C}_2\text{H}_5)_2\text{NH}$  and  $(\text{C}_2\text{H}_5)_3\text{N}$   
 $(\text{C}_2\text{H}_5)_2\text{NH}$  is a stronger base

- (3)  $\text{NH}_3$  and  $(\text{CH}_3)_2\text{CHNH}_2$   
 $(\text{CH}_3)_2\text{CHNH}_2$  is a stronger base

Use your brain power! (Textbook Page No 290)

Question 1.

Arrange the following amines in decreasing order of their basic strength :

$\text{NH}_3$ ,  $\text{CH}_3 - \text{NH}_2$ ,  $(\text{CH}_3)_2\text{NH}$ ,  $\text{C}_6\text{H}_5\text{NH}_2$

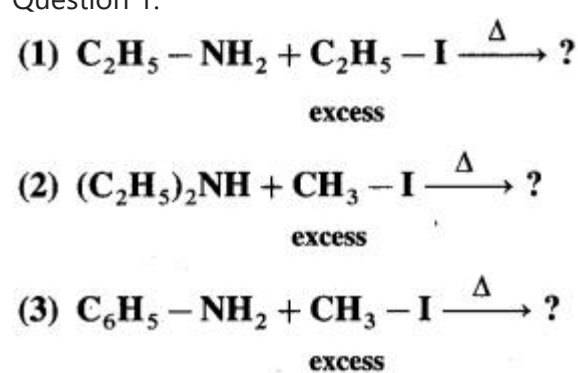
Answer:

Decreasing order of basic strength :  
 $(\text{CH}_3)_2\text{NH}$ ,  $\text{CH}_3 - \text{NH}_2$ ,  $\text{NH}_3$ ,  $\text{C}_6\text{H}_5\text{NH}_2$

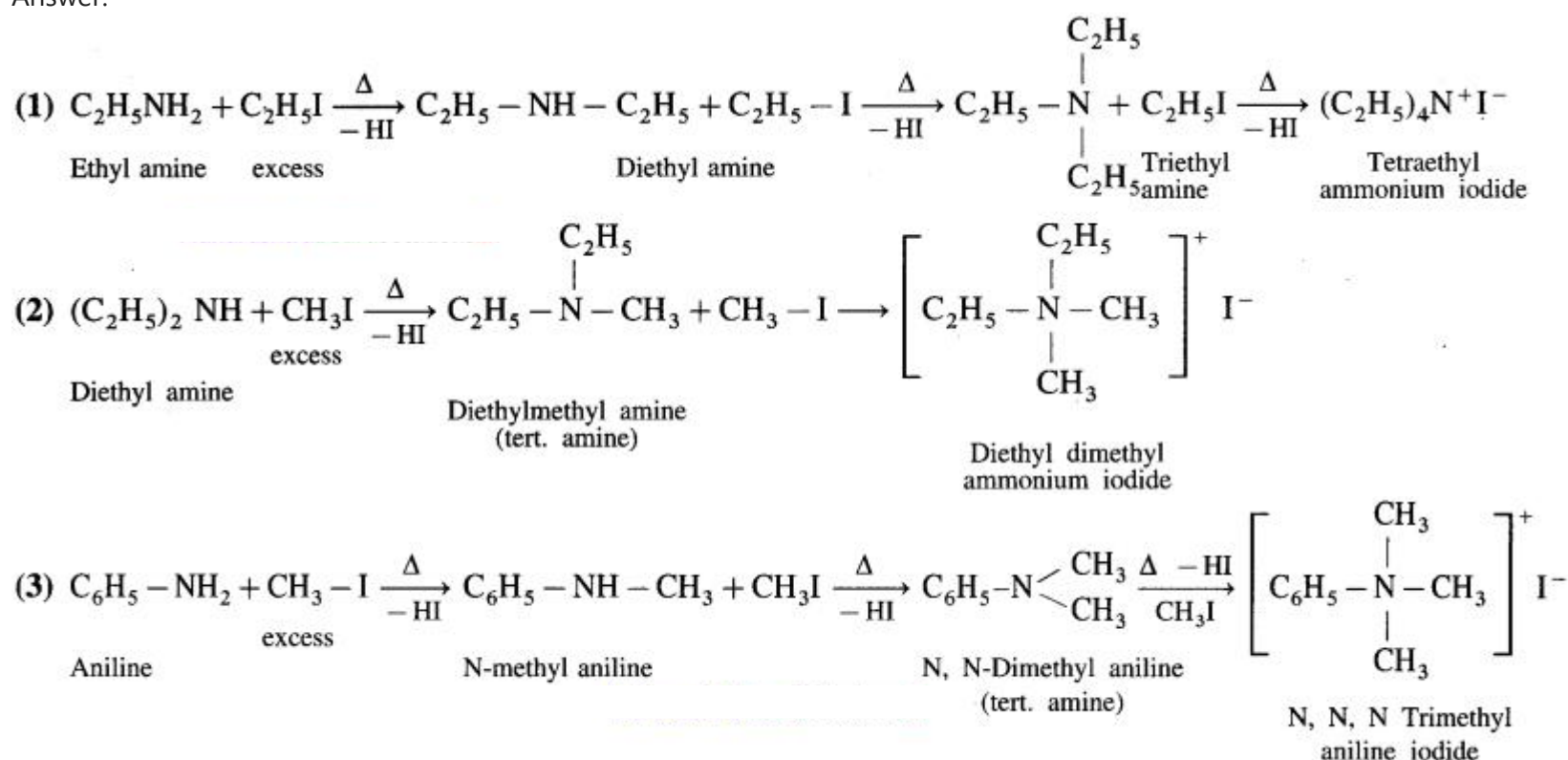
Use your brain power! (Textbook Page No 291)



Question 1.



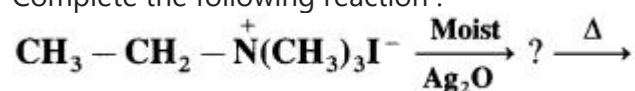
Answer:



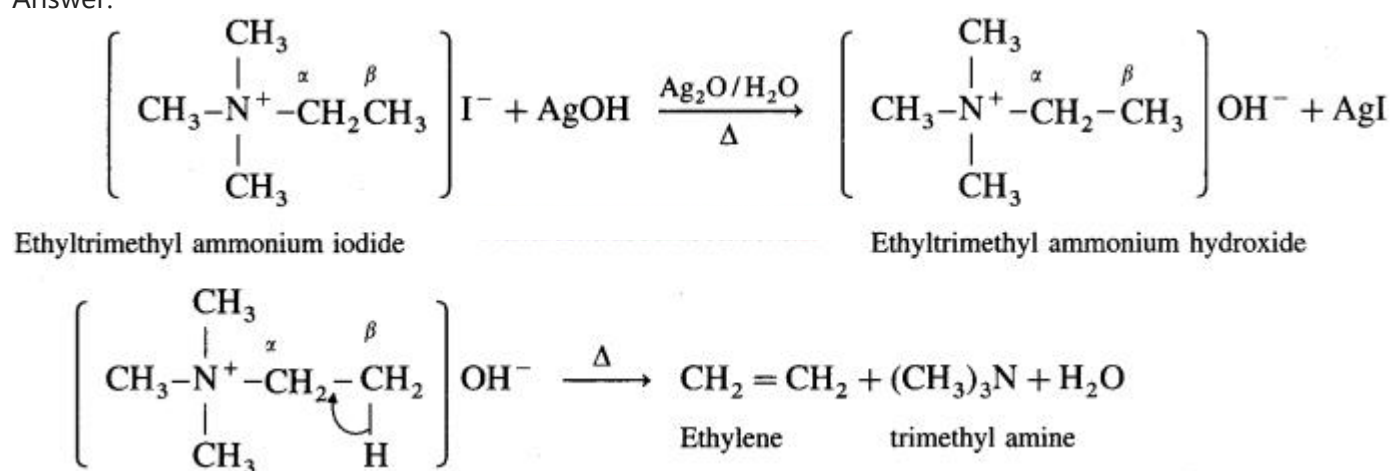
Use your brain power! (Textbook Page No 291)

Question 1.

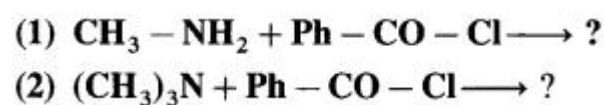
Complete the following reaction :



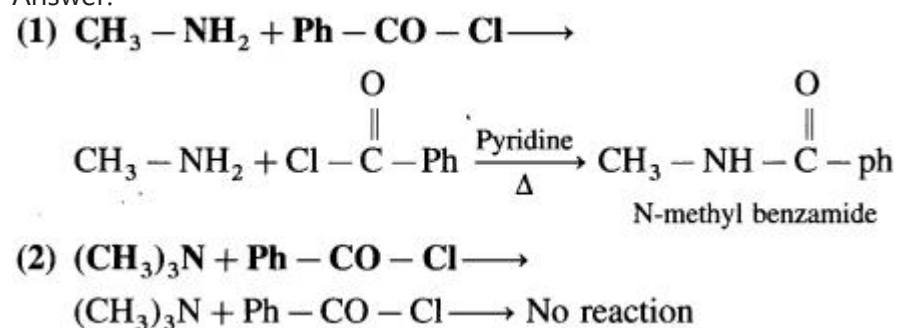
Answer:



Use your brain power! (Textbook Page No 292)



Answer:

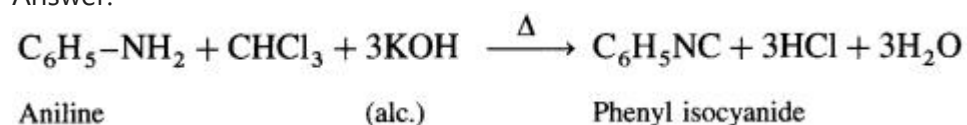


Use your brain power! (Textbook Page No 292)

Question 1.

Write the carbylamine reaction by using aniline as starting material.

Answer:



Can you tell? (Textbook Page No 292)

(1) What is the formula of nitrous acid ?

(2) Can nitrous acid be stored in bottle ?

Answer:

(1) Formula of nitrous acid :  $\text{H}-\text{O}-\text{N}=\text{O}$ 

(2) Nitrous acid cannot be stored in bottle.

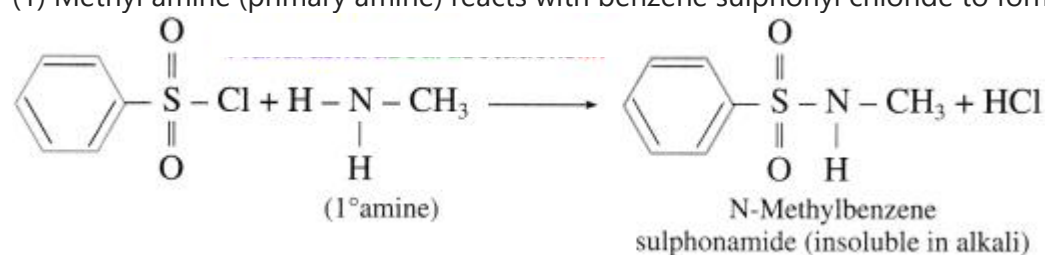
Use your brain power! (Textbook Page No 294)

Question 1.

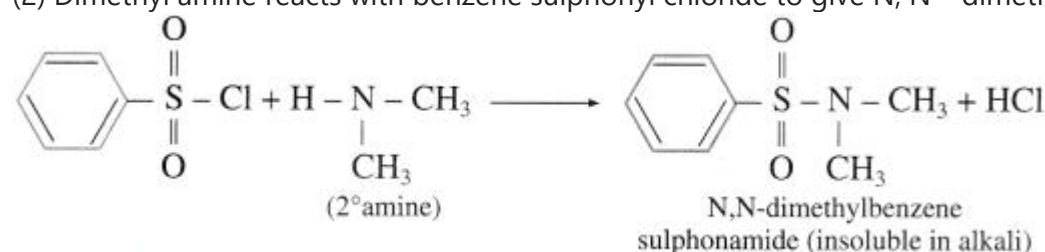
How will you distinguish between methyl amine, dimethylamine and trimethylamine by Hinsberg's test?

Answer:

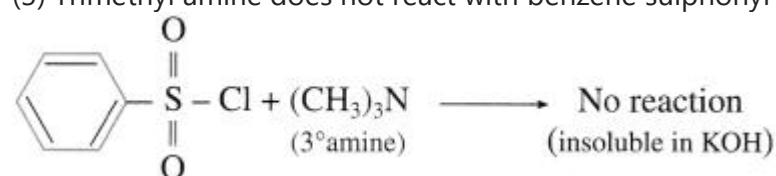
(1) Methyl amine (primary amine) reacts with benzene sulphonyl chloride to form N-methylbenzene sulphonamide



(2) Dimethyl amine reacts with benzene sulphonyl chloride to give N, N – dimethylbenzene sulphonamide.



(3) Trimethyl amine does not react with benzene sulphonyl chloride and remains insoluble in KOH



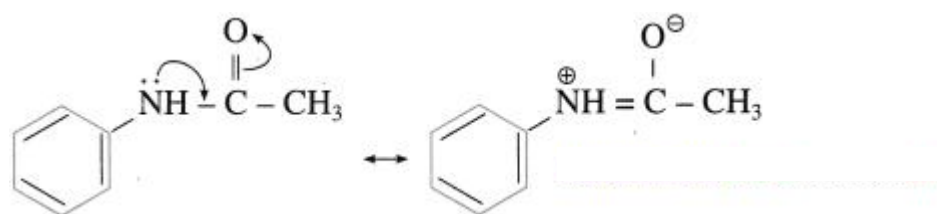
Problem 13.1 : (Textbook Page No 295)

Question 1.

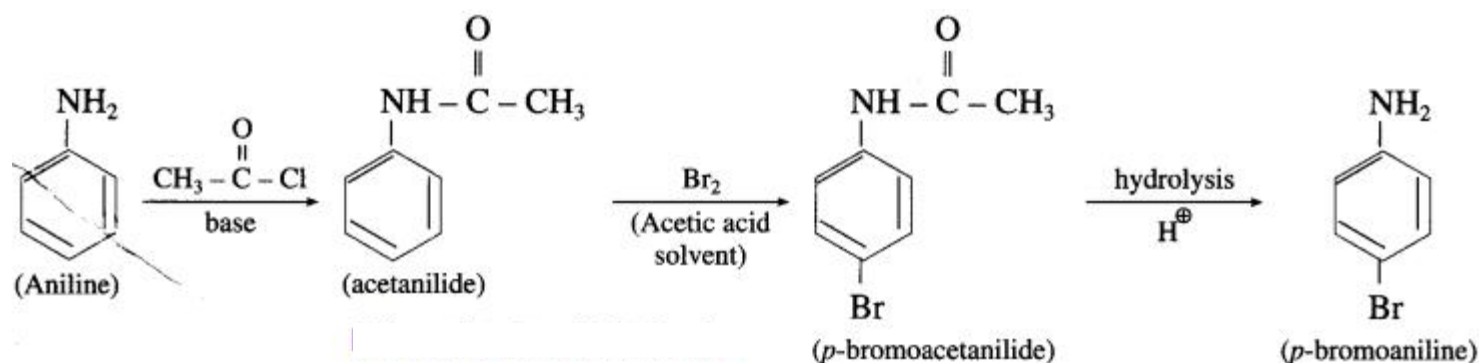
Write the scheme for preparation of p-bromoaniline from aniline. Justify your answer.

Solution :

$\text{NH}_2$  – group in aniline is highly ring activating and *o* – / *p* – directing due to involvement of the lone pair of electrons on 'N' in resonance with the ring. As a result, on reaction with  $\text{Br}_2$  it gives 2,4,6-tribromoaniline. To get a monobromo product, it is necessary to decrease the ring activating effect of  $-\text{NH}_2$  group. This is done by acetylation of aniline. The lone pair of 'N' in acetanilide is also involved in resonance in the acetyl group. To that extent, ring activation decreases.



Hence, acetanilide on bromination gives a monobromo product p-bromoacetanilide. After monobromination the original  $-\text{NH}_2$  group is regenerated. The protection of  $-\text{NH}_2$  group in the form of acetyl group is removed by acid catalyzed hydrolysis to get p-bromoaniline, as shown in the following scheme.



Use your brain power! (Textbook Page No 296)

Question 1.

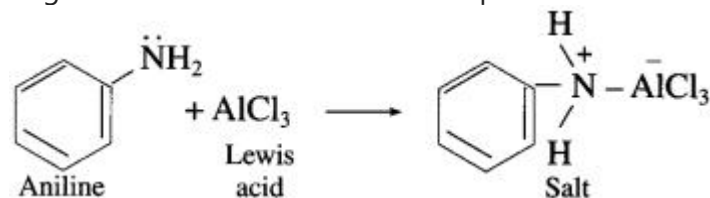
(1) Can aniline react with a Lewis acid?

(2) Why aniline does not undergo Friedel – Craft’s reaction using aluminium chloride?

Answer:

(1) Aniline reacts with a Lewis acid, forms salt.

(2) Aniline does not undergo Friedel-Craft’s reaction (alkylation and acetylation) due to salt formation with aluminium chloride (Lewis acid), which is used as catalyst. Due to this, nitrogen of aniline acquires +ve charge and hence acts as strong deactivating effect on the ring and makes it difficult for electrophilic attack.

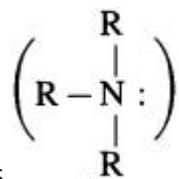


Can you tell? (Textbook Page No 294)

(1) Do tertiary amines have 'H' bonded to 'N'?

(2) Why do tertiary amines not react with benzene sulphonyl chloride?

Answer:



(1) Tertiary amines do not have 'H' bonded to 'N'.

(2) Tertiary amine does not undergo reaction with benzene sulphonyl chloride as it does not have any H atom attached to nitrogen atom of amine.