

Physics 2610H: Assignment IV

Jeremy Favro

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Problem 1. Use the tables of $R(r)$, $\Theta(\theta)$, and $\Phi(\phi)$ to describe the mathematical dependence on the r , θ , and ϕ variables of the $n = 3$, $\ell = 2$, $m_\ell = 2$ energy eigenstate of the hydrogen atom. Use computer software of your choice (Excel, Matlab, *etc.*) to plot the shape of the radial probability density, $P(r)$, of this state against r . Include units! Denote on your plot the most probable radius for an electron in this state.

Solution 1.

$$R(r) = \frac{1}{(3a_0)^{3/2}} \frac{2\sqrt{2}r^2}{27\sqrt{5}a_0^2} e^{-r/3a_0}$$
$$\Theta(\theta) = \sqrt{\frac{15}{32\pi}} \sin^2 \theta$$
$$\Phi(\phi) = e^{2i\phi}$$

Problem 2. Show that the degeneracy of the n -th state of the hydrogen atom *if* the electron had no spin (so course material only up to Chapter 7) would be n^2 . That is, prove the following equality and in a sentence or two state why this equation answers this question:

$$\sum_{\ell=0}^{n-1} (2\ell + 1) = n^2$$

Solution 2. Note the sum is

$$\sum_{\ell=0}^{n-1} (2\ell + 1) = 1 + (2 + 1) + \cdots + (2(n - 1) + 1)$$

which is the sum of odd integers up to $n - 1$. The sum we are interested in can be split into two parts,

$$\sum_{\ell=0}^{n-1} (2\ell) + \sum_{\ell=0}^{n-1} 1$$

The second part of this sum, $\sum_{\ell=0}^{n-1} 1$, is the sum of 1 $n - 1 + 1$ times, the $+1$ arising due to the beginning index being at 0 (Fencepost problem! That took me a minute). This reduces to n . $\sum_{\ell=0}^{n-1} (2\ell)$ is simply twice the sum of the integers up to $n - 1$ which is equal to

$$(n - 1)(n)$$

so the entire sum is

$$(n - 1)(n) + (n) = n^2.$$

This equation gives the degeneracy of the n -th state of the hydrogen atom as ℓ runs from 0 to $n - 1$ and m_ℓ runs from $-\ell$ to $+\ell$ which means m_ℓ could be any value from 0 to $\pm(n - 1)$. Because degeneracy (without spin) is effectively a measure of the number of possible values of m_ℓ the degeneracy can be expressed as the sum given in the problem statement.

Problem 3. For a hydrogen atom in the $1s$ state

- (a) What is the most probable value of r from a measurement of r ?
- (b) What is the average value of r ?
- (c) Where in 3-dimensional space is the electron most likely to be within a small volume of fixed size dV ?

Solution 3.

- (a) Here we are looking for the maximum of the probability,

$$R^2(r)r^2 = \left(\frac{2e^{-r/a_0}}{a_0^{3/2}} \right)^2 r^2 = \frac{4r^2 e^{-2r/a_0}}{a_0^3} = P(r)$$

This can be found by solving

$$\frac{dP}{dr} = \frac{8r(r - a_0)e^{-2r/a_0}}{a_0^4} = 0$$

Which gives solutions of $r = 0, a_0, \infty$. 0 and ∞ are non-physical so the most probable value is $r = a_0$.

- (b)

$$\int_0^\infty R^2(r)r^3 dr = \int_0^\infty \left(\frac{2e^{-r/a_0}}{a_0^{3/2}} \right)^2 r^3 dr = \frac{4}{a_0^3} \int_0^\infty r^3 e^{-2r/a_0} dr$$

Which, using the integral table given on the formula sheet, is

$$\frac{4}{a_0^3} \frac{3!}{\left(\frac{2}{a_0}\right)^4} = \frac{3a_0}{2}$$

- (c)

$$dP(r, \theta, \phi) = |\psi(r, \theta, \phi)|^2 dV = |\psi(r, \theta, \phi)|^2 r^2 \sin \theta d\theta d\phi = \left(\frac{2e^{-r/a_0}}{a_0^{3/2}} \sqrt{\frac{1}{4\pi}} \right)^2 r^2 \sin \theta d\theta d\phi$$

TODO: this isn't done

Problem 4. Evaluate, in electron volts, the energies of the three lowest energy levels of the hydrogen atom. Then calculate the frequencies in Hz, and the wavelengths in Å of all the photons that can be emitted by the atom in transitions between these levels. In what range of the electromagnetic spectrum are these photons?

Solution 4.

Problem 5. Consider the probability of finding the electron in the hydrogen atom somewhere inside a cone of $\theta = 23.5^\circ$ ("arctic polar region").

- (a) If the electron were equally likely to be found anywhere in space, what would be the probability of finding the electron in the arctic polar region?
- (b) Suppose the electron is in the state $n = 2, \ell = 1, m_\ell = 0$ (i.e. a $2p_z$ orbital); recalculate the probability of finding the electron in the arctic polar region.

Solution 5. (a)

$$\begin{aligned} &= \int_{0^\circ}^{23.5^\circ} \int_{0^\circ}^{360^\circ} \int_0^R r^2 \sin \theta dr d\phi d\theta \\ &= \int_{0^\circ}^{23.5^\circ} \int_{0^\circ}^{360^\circ} \frac{R^3 \sin \theta}{3} d\phi d\theta \\ &= \int_{0^\circ}^{23.5^\circ} 120R^3 \sin \theta d\theta \\ &= -\frac{2\pi}{3} R^3 \cos \left(\frac{47\pi}{360} \right) + \frac{2\pi}{3} R^3 \cos(0) \end{aligned}$$

TODO: infinite?

(b)

$$\begin{aligned} &= \iiint \left(\frac{1}{(2a_0)^{3/2}} \frac{r}{\sqrt{3}a_0} e^{-r/2a_0} \sqrt{\frac{3}{4\pi}} \cos \theta \right)^2 r^2 \sin \theta dV \\ &= \frac{1}{16a_0^5} \iint r^4 e^{-r/a_0} \cos^2 \theta \sin \theta dr d\theta \\ &= \frac{3}{2} \int \cos^2 \theta \sin \theta d\theta \\ &= -\frac{\cos^3(\theta)}{2} \Big|_0^{47\pi/360} \approx 0.114 \end{aligned}$$

TODO: Check this

Problem 6. An electron (with spin) in a hydrogen atom is in the $4f_{5/2}$ state. Find the values of the quantum numbers n , ℓ and j . What is the magnitude of the electron's total angular momentum? What are the possible values for the z -component of the electron's total angular momentum?

Solution 6.

Problem 7. Neglecting any effects of spin (i.e. if both particles are spinless), how many quantum states are possible for a two-particle system with orbital angular momentum quantum numbers $\ell_1 = 2$ and $\ell_2 = 1$? Write down all these states in what is called the *uncoupled* form: $(\ell_1 m_{\ell_1} \ell_2 m_{\ell_2})$. Then find all possible values of the total orbital angular momentum quantum number ℓ_T and the total orbital magnetic quantum number m_{ℓ_T} . Write down all these states in the *coupled* form: $(\ell_1 \ell_2 \ell_T m_{\ell_T})$.

Solution 7.

Problem 8. Two spinless particles, assumed initially to be distinguishable from each other, occupy two different states ($n = 2$ and $n' = 3$) in a 1-D infinite well of length L . The *product wave function* for the case with particle 1 in state $n = 2$ and particle 2 in state $n' = 3$ is

$$\psi(x_1, x_2) = A \sin\left(\frac{2\pi x_1}{L}\right) \sin\left(\frac{3\pi x_2}{L}\right), \quad A = \frac{2}{L}$$

- (a) Show that with this normalization constant this product wave function is correctly normalized.
- (b) If the two particles are instead *identical* how should this product wave function be adapted to form two realistic wave functions? Given that the normalization constant for each of these wave functions is $L = \sqrt{2}/L$ in each case what is the probability of both particles being found in the left half of the well (i.e. between 0 and $L/2$).

Solution 8.