

## Eco-friendly Jet Engine

- Uses ethanol as fuel
- ethanol can be converted into ethylene then oligomerization to form hydrocarbons.
- This fuel can be ignited in a Ram Jet Engine



- air
- ethanol based fuel
- ignition
- fuel nozzle

Signature: *Bizz*

Date: 10/7/19

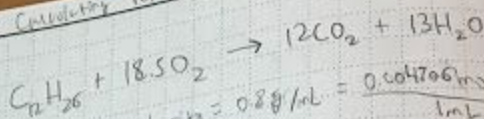
Team Members:

Witness: Dante Ciricola

Date: 10/7/19

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$$C_{12}H_{26} \text{ mm} = 170 \text{ g/mol}$$

$$\text{density} = 0.8 \text{ g/mL} = \frac{0.04706 \text{ mol } C_{12}H_{26}}{1 \text{ mL}}$$

$$0.08706 \text{ mol } O_2$$

$$0.0565 \text{ mol } CO_2$$

$$0.0612 \text{ mol } H_2O$$

$$PV = nRT$$

$$V = \frac{nRT}{P}$$

$$V = \frac{0.08706 \times 8.314 \times 273 \text{ K}}{100 \text{ kPa}} = 1.98 \text{ L}$$

$$1.98 \text{ L} \times 1000 = 1980 \text{ mL} = 1980 \text{ cm}^3 O_2 \times \frac{100 \text{ cm}^3}{21.02} = 9428 \text{ cm}^3$$

$$\frac{12 CO_2}{18.5 O_2} \times 1980 \text{ cm}^3 O_2 = 1284 \text{ cm}^3 CO_2$$

$$\frac{13 H_2O}{18.5 O_2} \times 1980 \text{ cm}^3 O_2 = 1391.4 \text{ cm}^3 H_2O$$

$$1 \text{ cm}^3 C_{12}H_{26} \rightarrow \frac{9428 \text{ cm}^3 \text{ air}}{8.56 \text{ cm}^2} = \frac{1100 \text{ cm}}{5} \times \frac{1 \text{ m}}{160934} \times \frac{3600 \text{ s}}{1 \text{ hr}} = 24.6 \text{ mph}$$

↑  
area of  
cylinder

$$\frac{24.6 \text{ mph}}{1 \text{ mL/5}}$$

$$22.4 \text{ L} = 1 \text{ mol}$$

$$1.95 \text{ L} = 0.08706 \text{ mol}$$

$$VO_2 = 1.9$$

$$VCO_2 = 1$$

$$VH_2O = 1$$

$$VC_{12}H_{26} = 1$$

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$$22.4 \text{ L} \cdot \text{mol}^{-1}$$

$$1.95 \text{ L} \cdot 0.08206$$

$$V_{O_2} = 1.95 \text{ L}$$

$$V_{CO_2} = 1.27 \text{ L}$$

$$V_{H_2O} = 1.37 \text{ L}$$

$$V_{C_2H_6} = 0.001 \text{ L}$$

Kelvin Pressure 37°C x 3

Per 1 mL

$$\frac{1.95 \text{ L} \cdot 0.08206}{1.01} = \frac{1.95}{0.08206} = 37.6773$$

$$Q = n \Delta T$$

$$37.6773 - (0.001 \text{ kg}) (2000 \text{ J/kg}) (\Delta T)$$

$$2,400 = \Delta T \cdot n \cdot K$$

$$P = \frac{101325 \text{ Pa}}{1.01} = 100321.78 \text{ Pa}$$

$$0.376773 \text{ L} \cdot \text{mol}^{-1} \cdot \text{atm}^{-1} \cdot \text{K}^{-1} = 0.376773 \text{ L} \cdot \text{mol}^{-1} \cdot \text{atm}^{-1} \cdot \text{K}^{-1}$$

$$m = m_a + m_{\text{combustion}}$$

$$127.42 - 53.23 =$$

$$P_{\text{air}} = \frac{m_{\text{air}}}{V_{\text{air}}} \Rightarrow m_{\text{air}} = P_{\text{air}} \times V_{\text{air}}$$

$$243.83 - 21.87 = 221.96$$

$$m_{\text{air}} = 1.225 \frac{\text{kg}}{\text{m}^3} \times 5.4 \times 10^{-4} \text{ m}^3$$

$$m_{\text{air}} = 6.6 \times 10^{-4} \text{ kg}$$

$$m_{O_2} = 0.00279 \text{ kg}$$

$$m_{CO_2} = 0.00249 \text{ kg}$$

$$m_{H_2O} = 0.00110 \text{ kg}$$

$$m_{C_2H_6} = 0.0008 \text{ kg}$$

$$m_{\text{total}} = 0.0078 \text{ kg}$$

$$P_{\text{total}} = \frac{n_{O_2} RT}{V} + \frac{n_{CO_2} RT}{V}$$

$$P_{\text{total}} = 248 \text{ psi}$$

or

560 psi using

$$P = nRT$$

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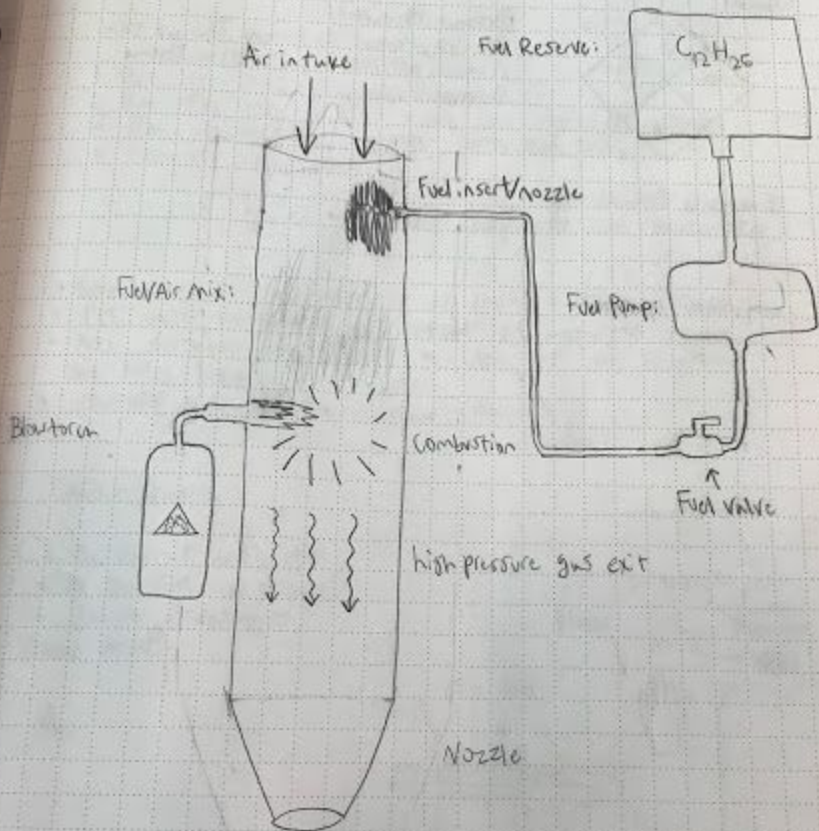
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## Test Plan

## Procedure

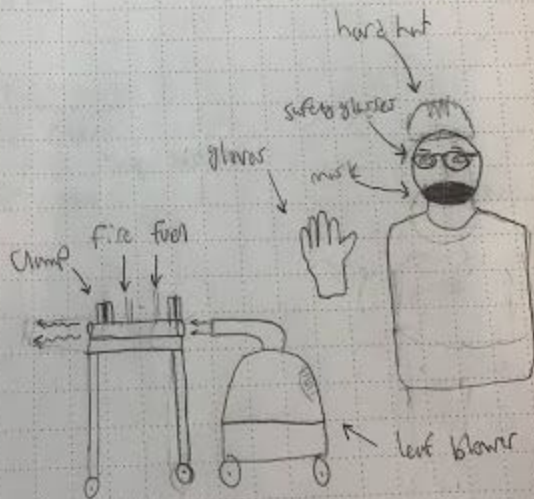
1. Run. wir. Provider. In correct position
2. Run. Air. for 5 Sec. making sure mixture is correct
3. Run. Ignition for 20 Sec. with fuel coming
4. Turn off air systems

## Safety:

- Engine must be held down by at least 2 clamps
- PPE must be used
- Area surrounding engine must be clear of all flammable and living entities
- Shutoff valve will be in place

## Central Question:

Can the current configuration ignite the fuel and produce the gasses necessary to produce thrust?

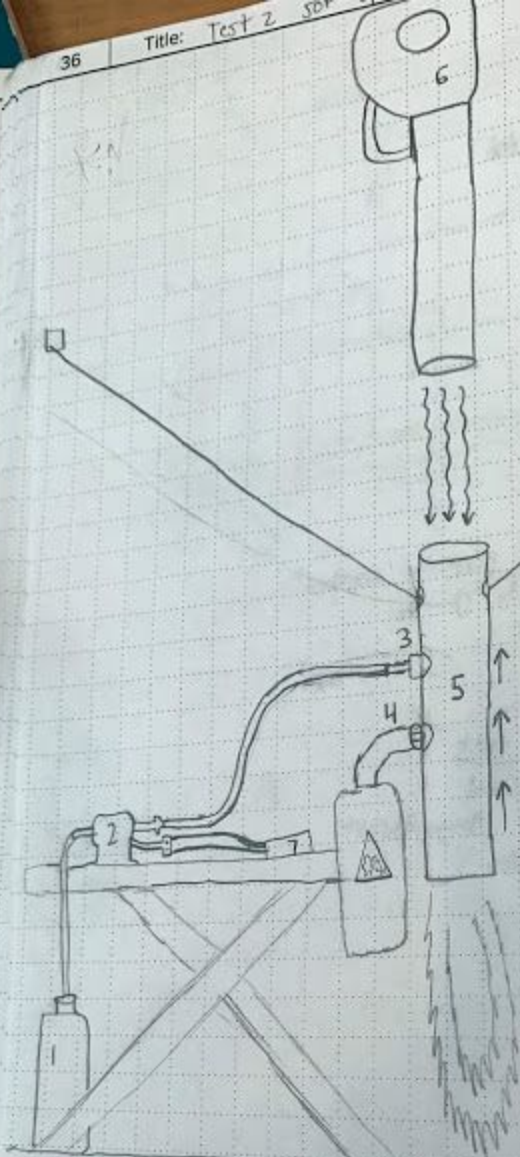


Signature:

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1. Jet engine
2. Jet engine
3. Fuel pump
4. Leaf blower
5. From

- 1: Pure ethanol
- 2: Fuel Pump
- 3: Nozzle

4. Torque
5. Jet engine
6. Leaf Blower

7. Battery

Signature:

Date:

Witness:

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Team Members:

Signature

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