

Cite this: DOI: 10.1039/xxxxxxxxxx

Molecular electrometer and binding of cations to phospholipid bilayers[†]

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Received Date
Accepted Date

DOI: 10.1039/xxxxxxxxxx

www.rsc.org/journalname

Despite the vast amount of experimental and theoretical studies on the binding affinity of cations into phospholipid bilayers, especially the biologically relevant Na⁺ and Ca²⁺ ions, there is no consensus in the literature. In this paper, we show that the ion binding affinity can be directly compared between simulations and experiments by using the choline headgroup order parameters according to the 'molecular electrometer' concept [Seelig *et al.*, *Biochemistry*, 1987, **26**, 7535]. Our findings strongly support the view that Na⁺ and other monovalent ions (except Li⁺) do not specifically bind to phosphatidylcholine lipid bilayers with sub-molar concentrations, in contrast to Ca²⁺ and other multivalent ions. Especially the Na⁺ binding affinity is overestimated by several molecular dynamics simulation models, leading to an artificially positively charged lipid bilayer and exaggerated structural effects in the headgroups. Qualitatively correct headgroup order parameter response is observed with Ca²⁺ binding in all the tested models, however, none of them has a sufficient quantitative accuracy to interpret the Ca²⁺:lipid stoichiometry or the induced atomistic resolution structural changes. This work has been done as a fully open collaboration, using nmrlipids.blogspot.fi as a main communication platform; all the scientific contributions were made publicly on this blog.

1 Introduction

Due to its high physiological importance — nerve cell signalling being the prime example — interaction of cations with phospholipid membranes has been widely studied via theory, simulations, and experiments. The relative ion binding affinities are gener-

ally agreed to follow the Hofmeister series^{1–9}, however, consensus on the quantitative affinities is currently lacking. Until 1990, the consensus (documented in two extensive reviews^{2,3}) was that while multivalent cations interact significantly with phospholipid bilayers, for monovalent cations (with the exception of Li⁺) the interactions are weak. This conclusion has since been strengthened by further studies showing that bilayer properties remain unaltered upon the addition of sub-molar concentrations of monovalent salt^{4,10,11}. Since 2000, however, another view has emerged, suggesting much stronger interactions between phospholipids and monovalent cations, and strong Na⁺ binding in particular^{6–9,12–18}.

The pre-2000 view has the experimental support that (in contrast to the significant effects caused by any multivalent cations) sub-molar concentrations of NaCl have a negligible effect on phospholipid infrared spectra⁴, area per molecule¹⁰, dipole potential¹⁹, lateral diffusion¹¹, and choline head group order parameters²⁰; in addition, the water sorption isotherm of a NaCl–phospholipid system is highly similar to that of a pure NaCl solution — indicating that the ion–lipid interaction is very weak⁴.

The post-2000 'strong binding' view rests on experimental and above all simulation findings. At sub-molar NaCl concentrations, the rotational and translational dynamics of membrane-

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[†] Electronic Supplementary Information (ESI) available: 5 figures, detailed technical discussion and simulation details. See DOI: 10.1039/b000000x/

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embedded fluorescent probes decrease^{7,9,12}, and atomic force microscopy (AFM) experiments show changes in bilayer hardness^{14–18}; in atomistic molecular dynamics (MD) simulations, phospholipid bilayers consistently bind Na⁺, although the binding strength depends on the model used^{12,13,21–26}.

Some observables have been interpreted in favour of both views. For example, as the effect of monovalent ions (except Li⁺) on the phase transition temperature is tiny (compared to the effect of multivalent ions), it was initially interpreted as an indication that only multivalent ions and Li⁺ specifically bind to phospholipid bilayers²; however, such a small effect in calorimetric measurements was later interpreted to indicate that also Na⁺ binds^{8,12}. Similarly, the lack of significant positive electrophoretic mobility of phosphatidylcholine (PC) vesicles in the presence of NaCl (again in contrast to multivalent ions and Li⁺) suggested weak binding of Na⁺^{1,8,14,15,27}; however, these data have also been explained by a countering effect of the Cl[−] ions^{22,28}. To reduce the area per lipid in scattering experiments, molar concentrations of NaCl are required¹⁰, which indicates weak ion–lipid interaction; in MD simulations, however, already orders of magnitude lower concentrations result in Na⁺ binding and clear reduction of area per lipid^{12,23}. Finally, in noninvasive NMR experiments, lipid lateral diffusion is unaltered by NaCl¹¹; however, it is reduced in simulations upon Na⁺ binding, which supports interpreting the reduced lateral diffusion of fluorescent probes^{7,9,12} as favouring the post-2000 view.

In this paper we set out to solve the apparent contradictions between the pre-2000 and post-2000 views. To this end we employ the ‘molecular electrometer’ concept, according to which the changes in the order parameters of the α and β carbons in the phospholipid head group (see Fig. 1) can be used to measure the ion affinity to PC lipid bilayer^{20,29?–31}. As order parameters can be accurately measured in experiments and directly compared to simulations³², employing the molecular electrometer as a function of cation concentration allows the comparison of binding affinity between simulations and experiments. In addition to demonstrating the usefulness of this general concept, we show that the response of order parameters to penetrating cations is qualitatively correct in MD simulations, but that in several models the affinity of Na⁺ for PC bilayers is grossly overestimated. Moreover, we show that the accuracy of lipid–Ca²⁺ interactions in current models is not enough for atomistic resolution interpretation of NMR experiments.

This work has been done as an Open Collaboration at nmrlipids.blogspot.fi; all the related files¹²⁹ and almost all the simulation data (<https://zenodo.org/collection/user-nmrlipids>) are openly available.

2 Results and Discussion

2.1 Background: Molecular electrometer in experiments

The molecular electrometer concept is based on the experimental observation that binding of any charged objects (e.g. ions, peptides, anesthetics, amphiphiles) on a PC bilayer interface induces systematic changes in the choline β and α segment order parameters^{20,29–31,33?–38}. Thus, these changes can be used to



Fig. 1 Chemical structure of 1-palmitoyl-2-oleoylphosphatidylcholine (POPC), and the definition of γ , β , α , g_1 , g_2 and g_3 segments.

determine binding affinities of the charged objects. The molecular electrometer was originally devised for cations^{20,29?}, but further experimental quantification with various positively and negatively charged molecules showed that the choline order parameters S_{CH}^{α} and S_{CH}^{β} in general vary linearly with small amount of bound charge per lipid^{29–31,33–38}. The empirically observed linear relation can be written as³⁹

$$S_{CH}^i(X^{\pm}) = S_{CH}^i(0) + \frac{4m_i}{3\chi} X^{\pm}, \quad (1)$$

where $S_{CH}^i(0)$ is the order parameter in the absence of bound charges, m_i is an empirical constant depending on the valency and position of bound charge, X^{\pm} is the amount of the bound charge per lipid, i refers to either α or β , and the value of quadrupole coupling constant is $\chi \approx 167$ kHz. The change in order parameters with respect to a bilayer without bound charges then becomes

$$\Delta S_{CH}^i = S_{CH}^i(X^{\pm}) - S_{CH}^i(0) = \frac{4m_i}{3\chi} X^{\pm}. \quad (2)$$

For Ca²⁺ binding to POPC bilayer (in the presence of 100 mM NaCl), combination of atomic absorption spectra and ²H NMR experiments gave $m_{\alpha} = -20.5$ and $m_{\beta} = -10.0$ ²⁹.

The absolute values of order parameters increase for β and decrease for α segment with bound positive charge and *vice versa* for negative charge^{20,29–31,33,38?}. However, as the β carbon order parameter is negative while α carbon order parameter is positive^{40–42}, we can conclude that both ΔS_{CH}^{β} and ΔS_{CH}^{α} decrease with bound positive charge and increase with bound negative charge. Consequently, values of m_i are negative for bound positive charges and *vice versa*. This can be rationalised by electrostatically induced changes in choline P–N dipole tilt^{30,31,44}, which is also seen in simulations^{23,24,45,46}. This is in line with order parameter decrease related to the P–N vector tilting more parallel to membrane plane seen with decreasing hydration levels⁴³.

The quantification of ΔS_{CH}^{β} and ΔS_{CH}^{α} with different cations have revealed that $\Delta S_{CH}^{\beta}/\Delta S_{CH}^{\alpha} \approx 0.5$ for a wide range of different cations (aqueous cations, cationic peptides, cationic anes-

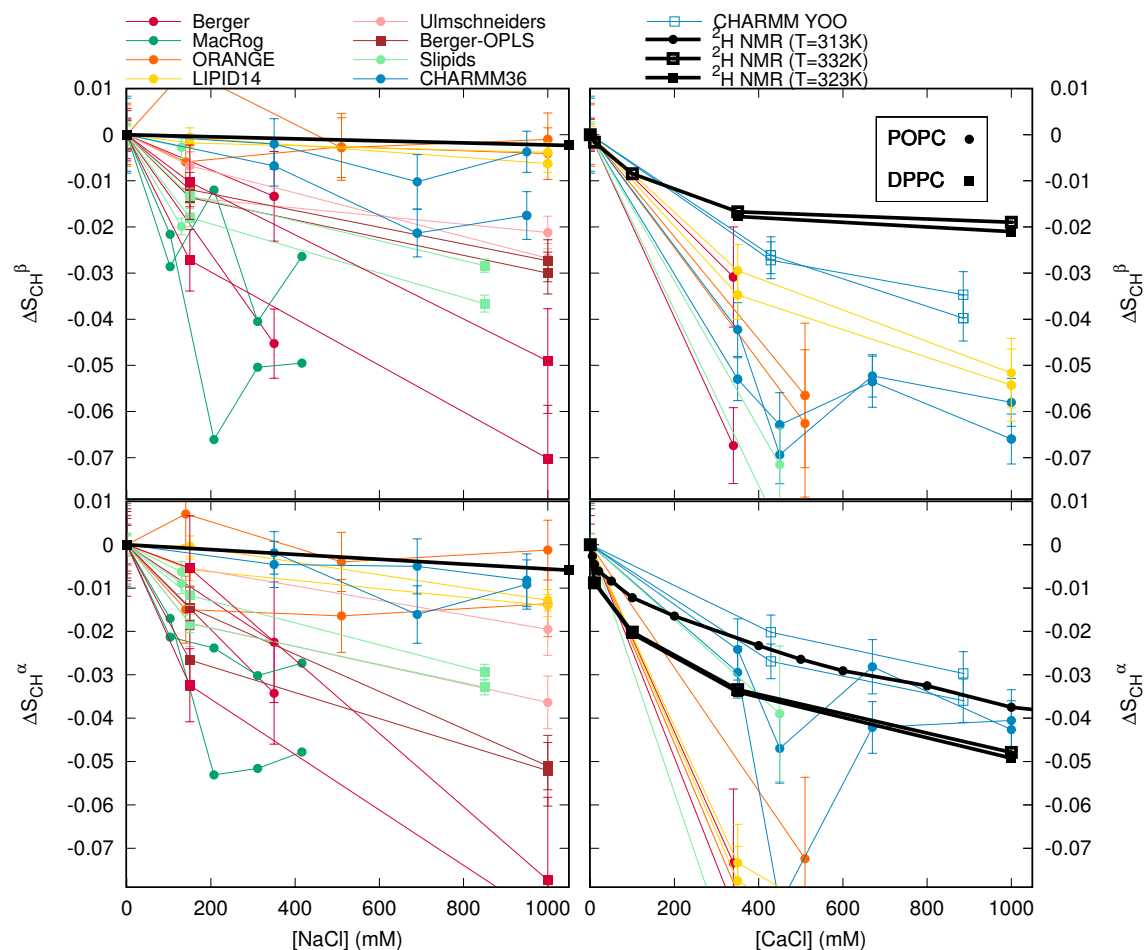


Fig. 2 The order parameter changes for β and α segments as a function of NaCl (left column) and CaCl_2 (right column) concentration, from simulations and experiments²⁰ (POPC with CaCl_2 from²⁹). The signs of the experimental order parameters, taken from experiments without ions^{40–42}, can be assumed to be unchanged with concentrations represented here^{29,32}. It should be noted that none of the models used here reproduces the order parameters within experimental error for pure PC bilayer without ions, indicating structural inaccuracies with varying severity in all models⁴³. Note that the relatively large decrease in CHARMM36 with 450 mM CaCl_2 arise from more equilibrated binding affinity due to long simulation times, see ESI[†].

thetics)^{36,38}. More specifically, the relation $\Delta S_{\text{CH}}^{\beta} = 0.43\Delta S_{\text{CH}}^{\alpha}$ was found for a DPPC bilayer with various CaCl_2 concentrations²⁰.

2.2 Molecular electrometer concept in MD simulations

The headgroup order parameter changes as a function of ion concentration in solution from ^2H NMR experiments are shown in Fig. 2 for DPPC and POPC bilayers^{20,29}. Only minor changes in order parameters are seen as a function of NaCl in solution, while the effect of CaCl_2 is an order of magnitude larger. Thus, according to the molecular electrometer concept, monovalent Na^+ ions have negligible affinity for PC lipid bilayers at concentrations up to 1 M, while binding of Ca^{2+} ions at the same concentration is significant^{20,29}.

Figure 2 also reports order parameter changes calculated from MD simulations of DPPC and POPC lipid bilayers as a function of NaCl or CaCl_2 initial concentrations in solution (for details of the simulated systems see Tables 1, 2 and ESI[†]). Note that none of these MD models reproduced within experimental uncertainty the order parameters for a pure PC bilayer without ions (Figure 2 in Ref. 43), indicating structural inaccuracies of varying severity in

all models⁴³. However, the experimentally observed headgroup order parameter increase with dehydration was qualitatively reproduced by all the models⁴³, and similarly here the presence of cations leads to the decrease of S_{CH}^{β} and S_{CH}^{α} (Fig. 2), in qualitative agreement with experiments. The changes are, however, overestimated by most models. According to the electrometer concept this indicates overbinding of cations in most MD simulation models.

While electrometer concept is well established in experiments (see previous section), it is not *a priori* clear that it works in simulations. The overestimated order parameter decrease could, in principle, arise also from the oversensitivity of choline headgroups on cation binding, instead of overbinding. Here we analyse the relation between cation binding and choline order parameter decrease in simulations in order to evaluate the usability of the electrometer concept in MD simulations.

According to the molecular electrometer concept, order parameter changes are linearly proportional to the amount of bound cations in bilayer (Eq. (2)). Figure 3 shows the changes in order parameter as a function of bound charge in MD simula-

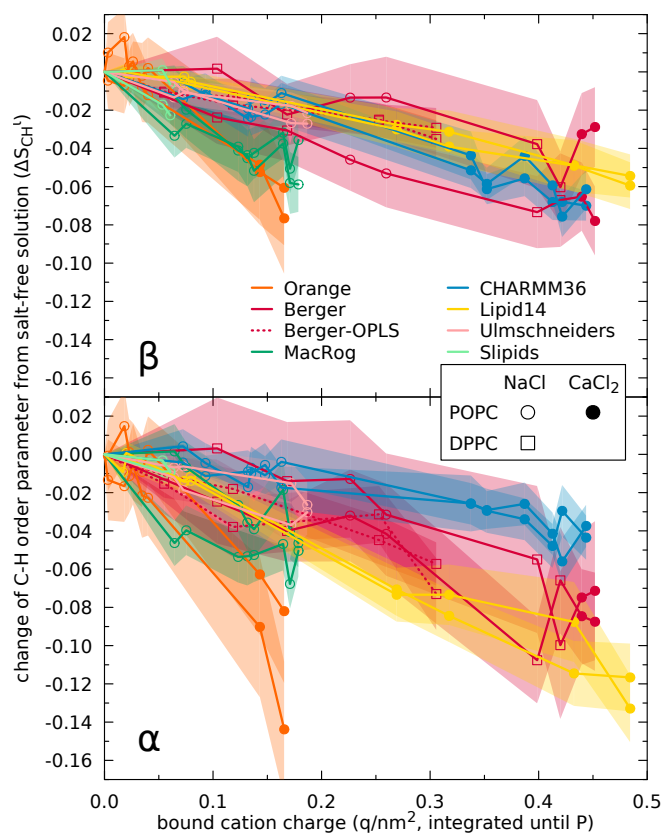


Fig. 3 Change of order parameters (from salt-free solution) of the β and α segments, ΔS_{CH}^{β} and ΔS_{CH}^{α} , shown as a function of bound cation charge. Eight MD simulation models compared. The order parameters as well as the bound charge calculated separately for each leaflet; cations residing between the bilayer centre and the density maximum of Phosphorus considered bound; error bars show standard error of mean over lipids.

tions (see ESI[†] for the definition of bound ions); in keeping with the molecular electrometer, roughly linear correlation between bound charge and order parameter change is found in all models. Note that quantitative comparison of the proportionality constants (i.e. slopes in Fig. 3) between different models and experimental slopes ($m_{\alpha} = -20.5$ and $m_{\beta} = -10.0$ for Ca^{2+} binding in DPPC bilayer in the presence of 100mM NaCl in Eq. 1²⁹) is not straightforward since the simulation slopes depend on the definition used for bound ions (see ESI[†]).

The comparison of order parameter changes in response to bound charge is more straightforward for systems with charged amphiphiles fully associated in bilayer, as the amount of bound charge is then explicitly known in both simulations and experiments. Such comparison between previously published simulation data⁴⁷ and experiments^{31,48} could not rule out overestimation of order parameter response to bound cations (i.e., slopes m_{β} and m_{α}) in a Berger-based model (ESI[†]). This might, in principle, explain the overestimated order parameter response of Berger model to CaCl_2 , but not to NaCl (see discussion in ESI[†]). Since simulation data with charged amphiphiles from other models is not available, the extended comparison with different models is left for further studies.

Figure 3 shows that the decrease on order parameter clearly

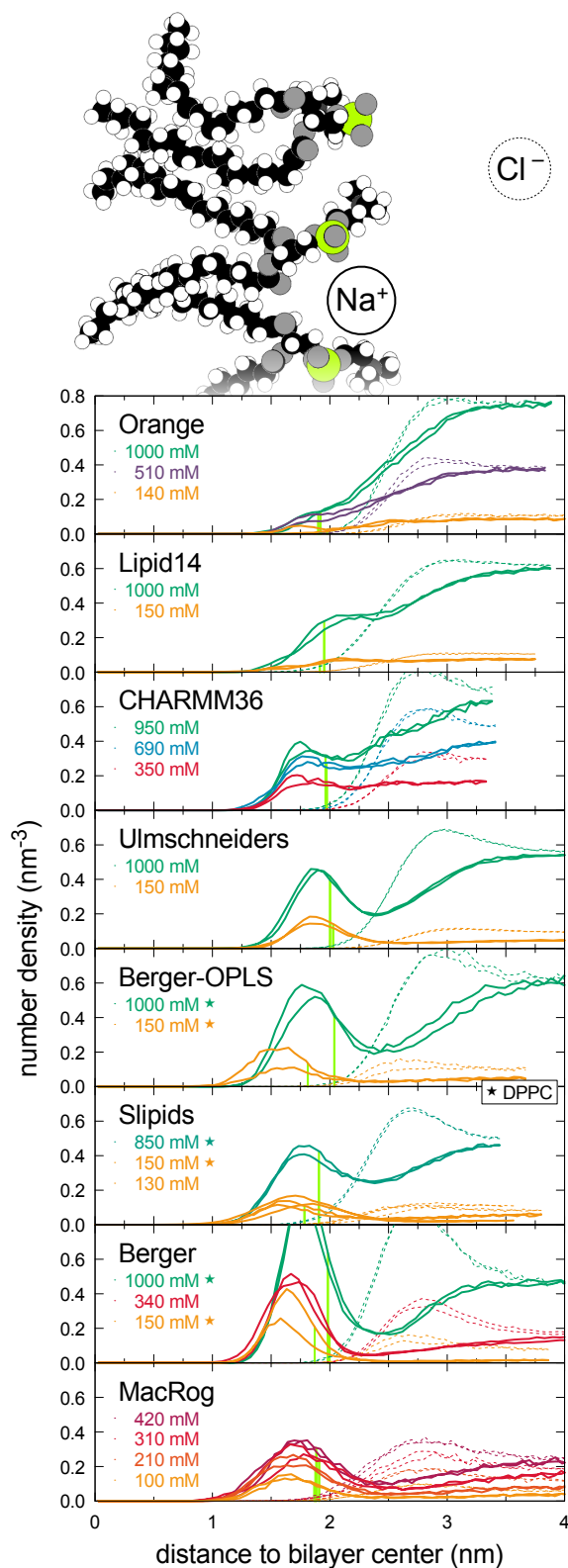


Fig. 4 Na^+ (solid line) and Cl^- (dashed) distributions along the lipid bilayer normal from MD simulations at several NaCl concentrations. The eight MD models are ordered according to their strength of order parameter change in response to NaCl (Fig. 2) from the weakest (top panel) to the strongest (bottom). The light green vertical lines indicate the locations of the Phosphorus maxima, used to define bound cations in Fig. 3.

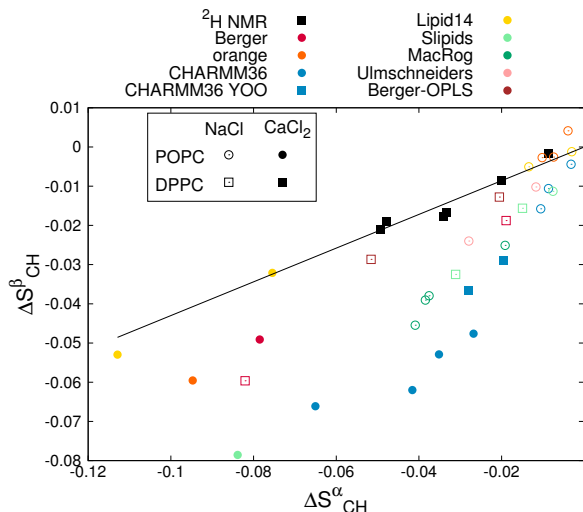


Fig. 5 Relation between $\Delta S_{\text{CH}}^{\beta}$ and $\Delta S_{\text{CH}}^{\alpha}$ from experiments²⁰ and different simulation models. Solid line is $\Delta S_{\text{CH}}^{\beta} = 0.43\Delta S_{\text{CH}}^{\alpha}$ determined for DPPC bilayer from ^2H NMR experiment with various CaCl_2 concentrations²⁰.

correlates with the amount of bound cations also in simulations. This is also evident from Fig. 4, which shows the Na^+ density profiles of the MD models ordered according to the order parameter change (reported in Fig. 2) from the smallest (top) to the largest (bottom). The Na^+ density peaks are larger for models with larger changes in order parameters, in line with the observed correlation between cation binding and order parameter decrease in Fig. 3.

Figure 5 compares the relation between $\Delta S_{\text{CH}}^{\beta}$ and $\Delta S_{\text{CH}}^{\alpha}$ in experiments²⁰ and different simulation models. Only Lipid14 gives $\Delta S_{\text{CH}}^{\beta}/\Delta S_{\text{CH}}^{\alpha}$ ratio in agreement with the experimental ratio. In all the other models the α order parameter decrease with bound cations is underestimated with respect to β order parameter decrease.

In conclusion, the clear correlation between bound cations and order parameter decrease is observed in all the tested simulation models. Consequently, the electrometer concept can be used to compare the cation binding affinity between experiments and simulations. However, we find that the quantitative response of α and β segment order parameters to bound cations in simulations do not generally agree with the experiments. The $\Delta S_{\text{CH}}^{\beta}/\Delta S_{\text{CH}}^{\alpha}$ ratio agrees with experiments only in Lipid14 model (Fig. 5). Thus, the observed overestimation of the order parameter changes with cation concentrations may, in principle, arise from overbinding of ions or from too sensitive lipid headgroup response on bound cation (see also discussion in ESI[†]). A careful analysis with current lipid models is performed in the next section.

2.3 Cation binding in different simulation models

The order parameter changes (Fig. 2) and density distributions (Fig. 4) demonstrate significantly different Na^+ binding affinities in different simulation models. The best agreement with experiments (lowest $\Delta S_{\text{CH}}^{\alpha}$ and $\Delta S_{\text{CH}}^{\beta}$) is observed for those models (Orange, CHARMM36, and Lipid14; see Fig. 2) that also predict the

lowest Na^+ densities in the membrane proximity (Fig. 4). In all the other tested models, the choline order parameter responses to NaCl are clearly overestimated (Fig. 2), and the strength of the overestimation is clearly linked to the strength of the Na^+ binding affinity (compare Figs. 2 and 4); this leads us to conclude that sodium binding affinity is overestimated in all these models.

In the best three models, the order parameter changes with NaCl are small (< 0.02), so with the achieved statistical accuracy we cannot conclude which of the three has the most realistic Na^+ binding affinity, especially at physiological NaCl concentrations ($\sim 150\text{mM}$) relevant for most applications. The overestimated binding in the other models raise questions on the quality of the predictions from these models when NaCl is present. Especially interactions between charged molecules and lipid bilayer might be significantly affected by the strong Na^+ binding, as it makes the bilayer effectively positively charged.

Significant Ca^{2+} binding affinity to a phosphatidylcholine bilayer at sub-molar concentrations is agreed in the literature^{2,3,20,29}, however, several details are yet under discussion. Simulations suggest that Ca^{2+} bind to lipid carbonyl oxygens with coordination number of 4.2¹³, while interpretation of NMR and scattering experiments suggest that one Ca^{2+} interacts mainly with choline groups^{102–104} of two phospholipid molecules²⁹. A simulation model correctly reproducing the order parameter changes would resolve the discussion by giving atomistic resolution interpretation for the experiments.

As a function of CaCl_2 concentration, all but one (CHARMM36 with recent ion model by Yoo et al.⁷³), model overestimate the order parameter decrease (Fig. 2). According to the molecular electrometer, this indicates overestimated Ca^{2+} binding. This is the most likely scenario for the models where changes in both order parameters were overestimated, however, in the case of CaCl_2 we cannot exclude the possibility that the headgroup response is oversensitive to bound cations (see ESI[†]). In CHARMM36 with ion model by Yoo et al.⁷³, ΔS_{CH} is overestimated for β but underestimated for α , in line with Fig. 5 where $\Delta S_{\text{CH}}^{\beta}/\Delta S_{\text{CH}}^{\alpha}$ ratio in CHARMM36 is larger than in experiments. Since we do not know if $\Delta S_{\text{CH}}^{\beta}$ or $\Delta S_{\text{CH}}^{\alpha}$ is more realistic in CHARMM36, we cannot conclude if Ca^{2+} binding is too strong or weak in this simulation model. This could be resolved by comparing CHARMM36 model to the experimental data with known amount of bound charge (e.g., experiments with amphiphilic cations^{31,48}), however, such simulation data are not currently available.

The ion density distributions with CaCl_2 in Fig. 6 show significant Ca^{2+} binding in all models, however, some differences occur in details. The Berger model predicts deeper penetration depth (density maxima close to $\pm 1.8\text{ nm}$) compared to other models (density maxima close to $\pm 2\text{ nm}$). The latter value is probably more realistic since ^1H NMR and neutron scattering data indicate that Ca^{2+} interacts mainly with the choline group^{2,102–104}. In CHARMM36, almost all Ca^{2+} ions present in simulation bind in bilayer indicating strongest binding affinity among the tested models. The difference is not as clear in Fig. 2 because α carbon order parameters are the least sensitive to bound charge in CHARMM36 (Fig. 3).

The origin of inaccuracies in lipid–ion interactions and binding

Table 1 List of simulations performed in this work. The ion concentrations are calculated as $[\text{ion}] = (N_{\text{ion}} \times [\text{water}]) / N_{\text{w}}$, where $[\text{water}] = 55.5\text{M}$. These correspond the concentrations reported in the experiments by Akutsu et al.²⁰. The lipid force fields are named as in our previous work⁴³.

Force field (lipid, ion)	lipid	[Ion] mM	^a N _l	^b N _w	^c N _{Na}	^d N _{Ca}	^e N _{Cl}	^f T (K)	^g t _{sim} (ns)	^h t _{anal} (ns)	ⁱ Files
Berger-POPC-07 ⁴⁹	POPC	0	128	7290	0	0	0	298	270	240	50
Berger-POPC-07 ⁴⁹ , ffgmx ⁵¹	POPC	340 (NaCl)	128	7202	44	0	44	298	110	50	52
Berger-POPC-07 ⁴⁹ , ffgmx ⁵¹	POPC	340 (CaCl ₂)	128	7157	0	44	88	298	108	58	53
Berger-DPPC-97 ⁵⁴	DPPC	0	72	2880	0	0	0	323	60	50	55
Berger-DPPC-97 ⁵⁴ , ffgmx ⁵¹	DPPC	150 (NaCl)	72	2880	8	0	8	323	120	60	56
Berger-DPPC-97 ⁵⁴ , ffgmx ⁵¹	DPPC	1000 (NaCl)	72	2778	51	0	51	323	120	60	57
BergerOPLS-DPPC-06 ⁵⁸	DPPC	0	72	2880	0	0	0	323	120	60	59
BergerOPLS-DPPC-06 ⁵⁸ , OPLS ⁶⁰	DPPC	150 (NaCl)	72	2880	8	0	8	323	120	60	61
BergerOPLS-DPPC-06 ⁵⁸ , OPLS ⁶⁰	DPPC	1000 (NaCl)	72	2778	51	0	51	323	120	60	62
CHARMM36 ⁶³	POPC	0	72	2242	0	0	0	303	30	20	64
CHARMM36 ⁶³ , CHARMM36 ⁶⁵	POPC	350 (NaCl)	72	2085	13	0	13	303	80	60	66
CHARMM36 ⁶³ , CHARMM36 ⁶⁵	POPC	690 (NaCl)	72	2085	26	0	26	303	73	60	67
CHARMM36 ⁶³ , CHARMM36 ⁶⁵	POPC	950 (NaCl)	72	2168	37	0	37	303	80	60	68
CHARMM36 ⁶³ , CHARMM36	POPC	350 (CaCl ₂)	128	6400	0	35	70	303	200	100	69
CHARMM36 ⁶³ , CHARMM36	POPC	450 (CaCl ₂)	200	9000	0	73	146	310	2000	100	70
CHARMM36 ⁶³ , CHARMM36	POPC	670 (CaCl ₂)	128	6400	0	67	134	303	200	120	71
CHARMM36 ⁶³ , CHARMM36	POPC	1000 (CaCl ₂)	128	6400	0	100	200	303	200	100	72
CHARMM36 ⁶³ , Yoo ⁷³	DPPC	0 (CaCl ₂)	128	8000	0	0	0	323	170	150	-
CHARMM36 ⁶³ , Yoo ⁷³	DPPC	430 (CaCl ₂)	128	7760	0	60	120	323	200	170	-
CHARMM36 ⁶³ , Yoo ⁷³	DPPC	886 (CaCl ₂)	128	7520	0	120	240	323	200	170	-
MacRog ⁷⁴	POPC	0	288	14400	0	0	0	310	90	40	75
MacRog ⁷⁴ , OPLS ⁶⁰	POPC	100 (NaCl)	288	14554	27	0	27	310	90	50	76
MacRog ⁷⁴ , OPLS ⁶⁰	POPC	210 (NaCl)	288	14500	54	0	54	310	90	50	76
MacRog ⁷⁴ , OPLS ⁶⁰	POPC	310 (NaCl)	288	14446	81	0	81	310	90	50	76
MacRog ⁷⁴ , OPLS ⁶⁰	POPC	420 (NaCl)	288	14392	108	0	108	310	90	50	76

^a The number of lipid molecules

^b The number of water molecules

^c The number of Na⁺ molecules

^d The number of Ca²⁺ molecules

^e The number of Cl⁻ molecules

^f Simulation temperature

^g The total simulation time

^h Time frames used in the analysis

ⁱ Reference for simulation files

Table 2 List of simulations performed in this work. The ion concentrations are calculated as $[\text{ion}] = (N_{\text{ion}} \times [\text{water}]) / N_w$, where $[\text{water}] = 55.5 \text{ M}$. These correspond the concentrations reported in the experiments by Akutsu et al.²⁰. The lipid force fields are named as in our previous work⁴³.

Force field (lipid, ion)	lipid	[Ion] mM	$^a N_l$	$^b N_w$	$^c N_{Na}$	$^d N_{Ca}$	$^e N_{Cl}$	$f T$ (K)	$^g t_{\text{sim}}$ (ns)	$^h t_{\text{anal}}$ (ns)	i Files
Orange, OPLS ⁶⁰	POPC	0	72	2880	0	0	0	298	60	50	77
Orange, OPLS ⁶⁰	POPC	140 (NaCl)	72	2866	7	0	7	298	120	60	78
Orange, OPLS ⁶⁰	POPC	510 (NaCl)	72	2802	26	0	26	298	120	100	79
Orange, OPLS ⁶⁰	POPC	1000 (NaCl)	72	2780	50	0	50	298	120	80	80
Orange, OPLS	POPC	510 (CaCl ₂)	72	2802	0	26	52	298	120	60	81
Slipids ⁸²	DPPC	0	128	3840	0	0	0	323	150	100	83
Slipids ⁸² , AMBER ^{84,85}	DPPC	150 (NaCl)	600	18000	49	0	49	323	100	40	-
Slipids ⁸² , AMBER ^{84,85}	DPPC	850 (NaCl)	128	3726	57	0	57	323	105	100	86
Slipids ⁸⁷	POPC	0	128	5120	0	0	0	303	200	150	88
Slipids ⁸⁷ , AMBER ⁸⁹	POPC	130 (NaCl)	200	9000	21	0	21	310	105	100	90
Slipids ⁸⁷ , AMBER ⁶⁰	POPC	450 (CaCl)	200	9000	0	73	146	310	2000	100	91
Lipid14 ⁹² , AMBER ⁶⁰	POPC	0	128	5120	0	0	0	298	205	200	93
Lipid14 ⁹² , AMBER ⁶⁰	POPC	150 (NaCl)	128	5120	12	0	12	298	205	200	94
Lipid14 ⁹² , AMBER ⁶⁰	POPC	1000 (NaCl)	128	5120	77	0	77	298	205	200	95
Lipid14 ⁹² , AMBER ⁶⁰	POPC	350 (CaCl ₂)	128	6400	0	35	70	298	200	100	96
Lipid14 ⁹² , AMBER ⁶⁰	POPC	1000 (CaCl ₂)	128	6400	0	100	200	298	200	100	97
Umschneiders ⁹⁸ , OPLS ⁶⁰	POPC	0	128	5120	0	0	0	298.15	205	200	99
Umschneiders ⁹⁸ , OPLS ⁶⁰	POPC	150 (NaCl)	128	5120	12	0	12	298.15	205	200	100
Umschneiders ⁹⁸ , OPLS ⁶⁰	POPC	1000 (NaCl)	128	5120	77	0	77	298.15	205	200	101

^a The number of lipid molecules

^b The number of water molecules

^c The number of Na⁺ molecules

^d The number of Ca²⁺ molecules

^e The number of Cl molecules

^f Simulation temperature

^g The total simulation time

^h Time frames used in the analysis

ⁱ Reference for simulation files

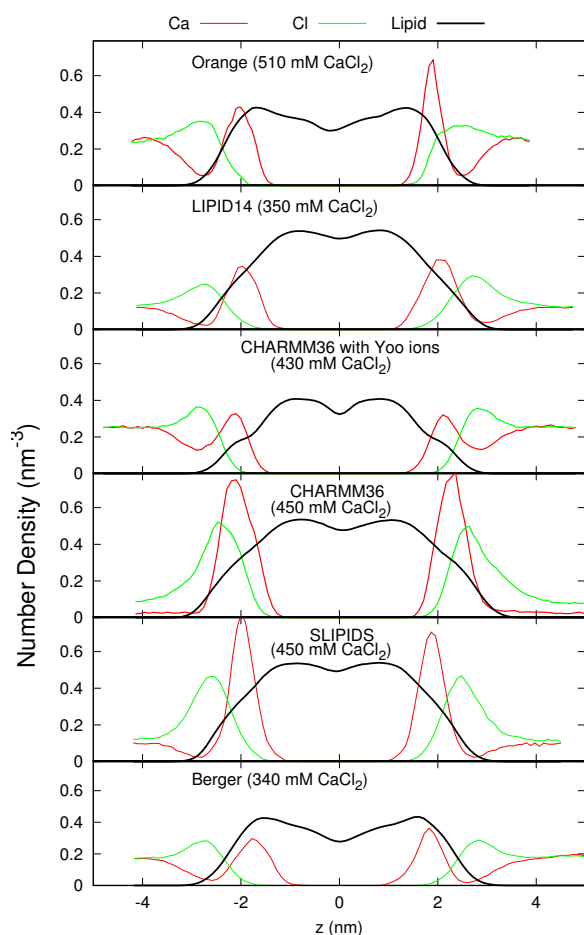


Fig. 6 Atom number density profiles along the membrane normal coordinate z for lipids, Ca^{2+} and Cl^- ions from simulations with different force fields. The profiles only with smallest available CaCl_2 concentration are shown for clarity. Figure including all the available concentrations is shown in ESI[†]. The lipid densities are scaled with 100 (united atom) or 200 (all atom model) to make them visible with the used y-axis scale. The Cl^- density is scaled with 2 to equalise charge density of ions.

affinities in different models is far from clear. Potential candidates could be, for example, discrepancies in the ion models^{105–107}, incomplete treatment of electronic polarizability¹⁰⁸, or inaccuracies in the lipid headgroup description⁴³. Cordomi et al.²⁴ showed that the Na^+ binding affinity decreases when ion radius increases in the model, however, also the models with the largest radius show significant binding in DPPC bilayer simulated with OPLS-AA force field¹⁰⁹. In our results, the Slipids model gives essentially similar binding affinity with ion parameters from Refs. 89 and 84,85. Further, the compensation of missing electronic polarizability by scaling ion charge^{108,110} reduced Na^+ binding in Berger, BergerOPLS and Slipids models, but not enough to be in agreement with experiments (ESI[†]). The charge-scaled Ca^{2+} model¹¹¹ slightly reduced binding in CHARMM36, but did not have significant influence on binding in Slipids (ESI[†]). Significant reduction of Ca^{2+} binding was observed with ion model by Yoo et al.⁷³, however, the CHARMM36 lipid model must be further analysed to fully interpret the results.

On the other hand, also the lipid models may have significant

influence on ion binding behaviour. For example, the same ion model and non-bonded parameters are used in the Orange and BergerOPLS⁵⁸ simulations, but while Na^+ ion binding affinity appears realistic in the Orange model, it is significantly overestimated in the BergerOPLS (Fig. 4). However, realistic Na^+ binding does not directly relate to realistic Ca^{2+} binding (see Orange, Lipid14 and CHARMM36 in Fig. 2) or realistic choline order parameter response to bound charge (see Orange and CHARMM36 in Fig. 5). It should be also noted that the low binding affinity of Na^+ in CHARMM36 model is due to the additional repulsion added between sodium ions and lipid oxygens (NBFI⁶⁵ (ESI[†]). Altogether, our results indicate that probably both, lipid and ion force field parameters, need improvement to correctly predict the cation binding affinity, and the associated structural changes.

3 Conclusions

As suggested by the molecular electrometer concept^{20,29–31}, the decrease in order parameters of α and β carbons in the PC head group of lipids bilayers is related to cation binding in all tested simulation models (Fig. 3), despite of known inaccuracies in the actual atomistic resolution structures⁴³. Hence the molecular electrometer concept allows a direct comparison of Na^+ binding affinity between simulations and noninvasive NMR experiments. The comparison reveals that most models overestimate Na^+ binding; only Orange, Lipid14, and CHARMM36 predict realistic binding affinity. None of the tested models has the required accuracy to interpret the Ca^{2+} :lipid stoichiometry or induced structural changes with atomistic resolution.

In general, our results support the pre-2000 view that at sub-molar concentrations, in contrast to Ca^{2+} and other multivalent ions^{1–4,10,11,19,20,27,29}, Na^+ and other monovalent ions (except Li^+) do not specifically bind to phospholipid bilayers. Concerning the interpretation of existing experimental data, our work supports Cevc's view² that the observed small shift in phase transition temperature is not indicative of Na^+ binding. Further, our findings are in line with the noninvasive NMR spectroscopy work of Filippov et al.¹¹ that proved the results of Refs. 7,9,12 to be explainable by direct interactions between Na^+ ions and fluorescent probes. Finally, as spectroscopic methods are in general more sensitive to atomistic details in fluid-like environment than AFM, our work indirectly suggests that the ion binding reported from AFM experiments on fluid-like lipid bilayer systems^{14–18} might be confounded with other physical features of the system. Concerning contradictions in MD simulation results, we reinterpret strong Na^+ binding as an artefact of several simulation models, e.g., the Berger model used in Refs. 12,13.

The artificial specific Na^+ binding in simulations may lead to doubtful results, since it effectively leads to positively charged phosphatidylcholine (PC) lipid bilayers even at physiological NaCl concentration. Such a PC bilayer has distinctly different interactions with charged objects compared to a (more realistic) model without specific Na^+ binding. Furthermore, the overestimation of Na^+ binding affinity may extend also to other positively charged objects, say, membrane protein segments. This would affect lipid–protein interactions and could explain, for example, contradicting results on electrostatic interactions between charged protein seg-

ments and lipid bilayer^{112,113}. In conclusion, more careful studies and model development on lipid bilayer-charged object interactions are called for to make molecular dynamics simulations directly usable in a physiologically relevant electrolytic environment.

This work has been done as a fully open collaboration, using nmrlipids.blogspot.fi as the communication platform. All the scientific contributions have been communicated publicly through this blog or GitHub repository¹²⁹. All the related content and data is available at Ref. 129.

Acknowledgements: AC and VSO wish to thank the Research Computing Service at UEA for access to the High Performance Computing Cluster. VSO acknowledges the Engineering and Physical Sciences Research Council in the UK for financial support (EP/L001322/1). OHSO acknowledges Tiago Ferreira for very useful discussions, the Emil Aaltonen foundation for financial support, Aalto Science-IT project and CSC-IT Center for Science for computational resources. MSM acknowledges financial support from the Volkswagen Foundation (86110). M.G. acknowledges financial support from Finnish Center of International Mobility (Fellowship TM-9363). J. Melcr acknowledges computational resources provided by the CESNET LM2015042 and the CERIT Scientific Cloud LM2015085 projects under the program "Projects of Large Research, Development, and Innovations Infrastructure" LM acknowledges funding from the Institut National de la Sante et de la Recherche Medicale (INSERM).

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