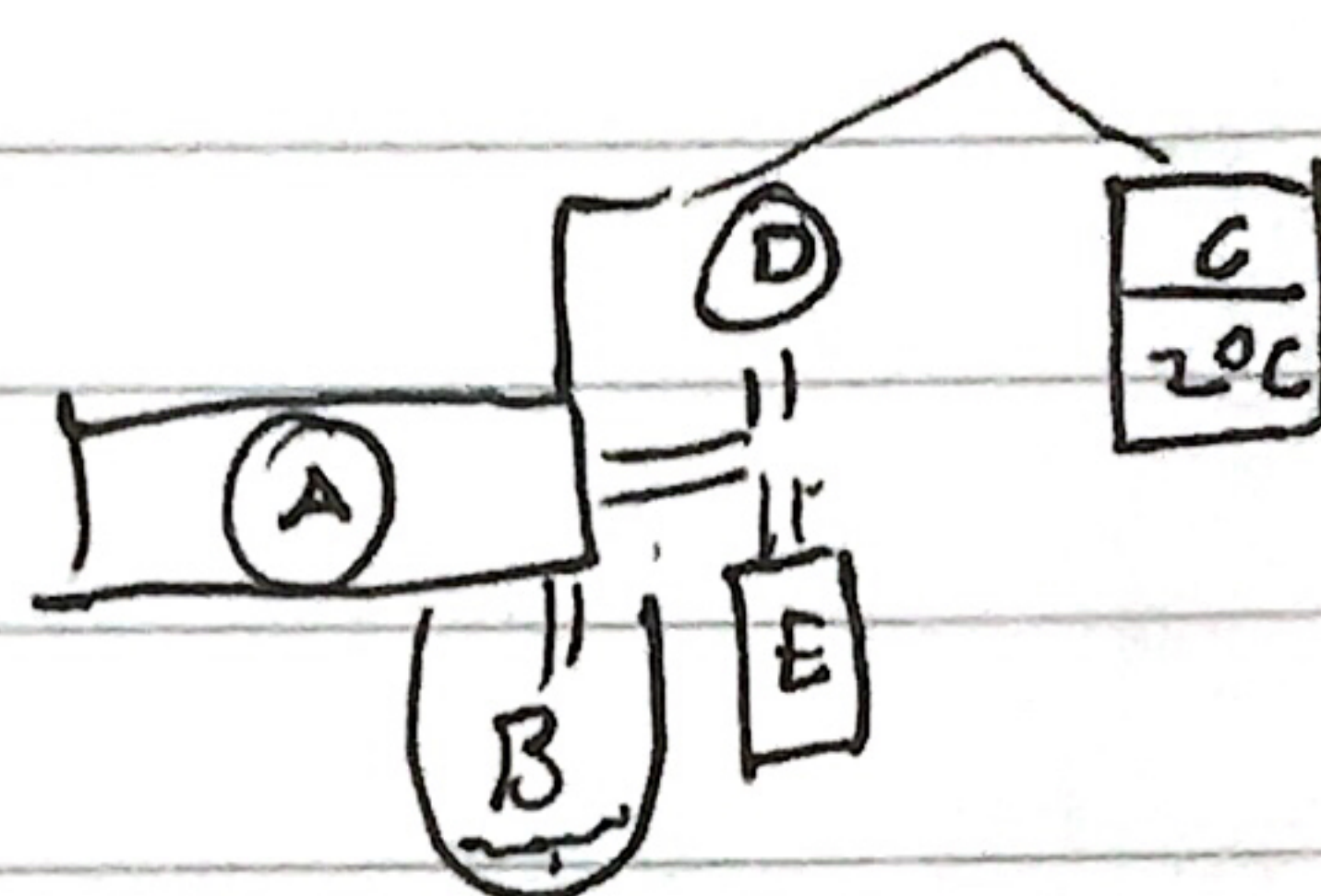


lab

Crush CO<sub>2</sub> dry ice in a plastic open bag.  
Tighten the golden knob. Place CO<sub>2</sub> in canister and close the apparatus. After closing the valve immediately begin taking measurements



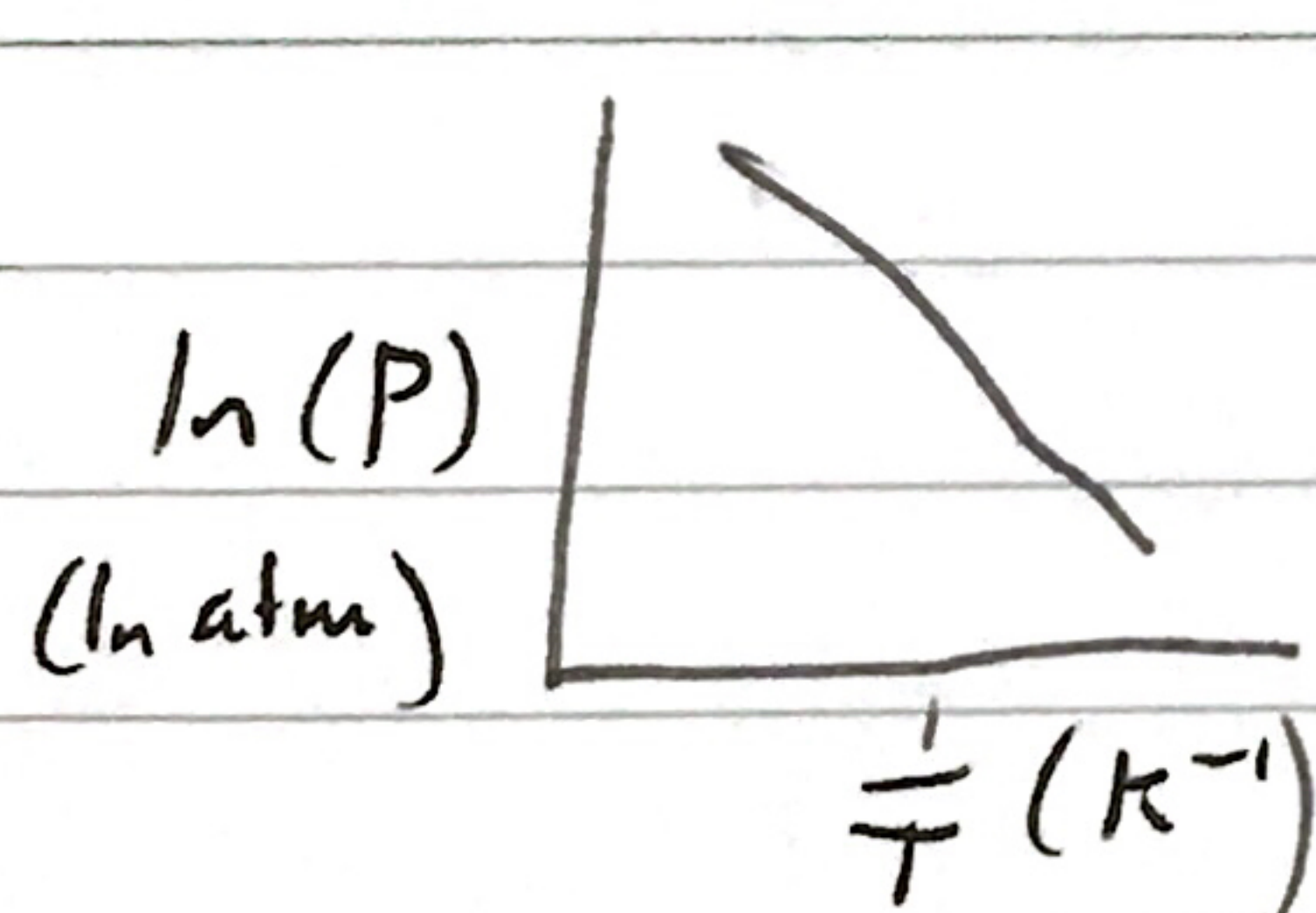
A is the regulator  
B canister  
C temperature reader  
D pressure reader  
E safety valve

chemical potential is equal for two phases at a phase boundary  
 $\mu_s = \mu_l = \mu_g$  @ triple point  
 $G_s = G_l = G_g$

Closed system has no

$$\ln\left(\frac{P_2}{P_1}\right) = -\frac{\Delta H_{\text{sub}}}{R} \left( \frac{1}{T_2} - \frac{1}{T_1} \right)$$

we are calculating  $\Delta H_{\text{sub}}$



$$\Rightarrow \ln(P) = -\Delta H_{\text{sub}} / RT$$

$$m = \frac{-\Delta H_{\text{sub}}}{R}$$

- always add 1 atm, uncertainty of psi.
- ② classius. Klapayran ① phase diagram / put P/T triple point compare to lit value  
No Pressure Temp data only graph.

webbook.nist.gov

$\frac{\Delta H}{R}$  (max) can

error bars

Continued on Page

Read and Understood By

Signed

Date

Signed

Date