Solutions and Their Properties

In chemistry, a solution is a homogeneous mixture composed of two or more substances.
In such a mixture, a solute is a substance dissolved in another substance, known as a solvent.

The concentration of a solute in a solution is a measure of how much solute is dissolved in the solvent.
Common units of concentration include molarity, molality, and mass percentage.

For example, a 1 molar (1 M) solution of sodium chloride contains 1 mole of NaCl in 1 liter of solution.

Colligative properties such as vapor pressure lowering, boiling point elevation, freezing point depression,
and osmotic pressure depend on the number of solute particles in a solution and not on the type of solute.

Understanding these properties helps in determining molecular masses and in studying properties of solutions in various contexts.