

Making fuming nitric acid <https://youtu.be/QmCdrDLyNXQ>

I True or false? Circle the correct answer and justify if it's false

T/F 1 Fuming nitric acid can light common lab gloves on fire.

True.

T/F 2 To make the fuming nitric acid, the speaker uses hydrochloric acid.

False – The speaker uses *sulfuric* acid.

T/F 3 To speed up the reaction, the speaker cools down the reactants.

False – The reaction is sped up by *heating* the reactants, according to the ARRHENIUS equation.

T/F 4 The orange color comes from nitrogen dioxide gas.

True – Excess heat decomposes the nitric acid to form $\text{NO}_{2(g)}$.

T/F 5 The process used to cool down the gas to a liquid form is called a crystallization.

False – It's a *distillation*, a common process in chemical reactions that form gases, and is a means of purification.

T/F 6 The speaker determines the concentration of his product by measuring its density.

True – Most physical properties of mixtures depend on the relative proportions of each compound.

T/F 7 Using gloves in a lab is something one must always do without thinking.

False – Every safety measure is well thought of beforehand.

II Translate into English

◇ Oxydant : oxidizer

◇ Fusée : rocket

◇ Faire bouillir : to boil

◇ Colonne réfrigérante : condenser column

◇ Feuille d'aluminium : aluminium foil

◇ Sous vide : under vacuum

◇ Curieusement (expression idomatique) : oddly enough