

CROSS-SECTION MEASSUREMENTS FOR $^{nat}\text{Zn}(n, p)^{64, 67}\text{Cu}$ REACTION

by

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Abstract

For me DEDICATION TO

Acknowledgements

THANK YOU

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DATO

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Chapter 1

Introduction

“Count only the good days.”

— Irene Jensen, Ahus 2017

A short introduction to nuclear medicine

The usage of medical radionuclides can be traced back as long as late 1890s when Henry Becquerel discovered radioactivity when he was studying a sulfate of uranium and Marie Curie together with her husband, Pierre Curie, found activity in the ore where they had extracted uranium from. Since then, many studies on animal and humans have been done which led to development of radionuclides used in therapy and imaging with radiotracers.

Therapeutic application involves using radioisotopes together with a molecule or proteins to e.g. irradiate cancer cells. The advantage of this treatment method is that the biological molecules are being produced for one purpose only, to find a specific area in the body and irradiates it from the inside unlike external beam therapy. This approach is also used in diagnostics. A PET-scan makes an image of the activity in vivo, this imaging technique is commonly used to detect cancer, heart problems and brain disorders with the usage of different molecules.

In this report, medical isotopes are being discussed. The usage of radiopharmaceuticals in both therapy and diagnostics are used more and more in hospitals to treat and detect illnesses. The characteristics of a medical isotope and why those are important aspects is described in the first part of this report. The second part gives an introduction to different production methods of medical isotopes and the final part of this report is the usage of these radiopharmaceuticals.

Chapter 2

Theory

quote

— by,

2.1 Background of Nuclear medicine

According to the National cancer institute [1] there were 1.735.350 people diagnosed with cancer in 2018 in the United States alone. The amount of deaths caused by cancer worldwide is approximate 9.6 million people [2]. If we did not have modern medicine like chemotherapy and internal and external radiation, the number of deaths would be much higher.

From when radioactivity were discovered in the 1890s there were a lot of new findings done in the medicine field, such as x-rays by Roentgen in 1895. Not long after, the biological foundation of nuclear medicine followed [3]. The tracer approach was developed in 1913 by Georg de Hevesy [4] when he was investigating the absorbtion of radioactive led in plants. Not much later, in 1927, Blumgart and Weiss [5] used aqueous solution of radon to study the pulmonary circulation ¹ in a man. Then, a cyclotron was build. Ernest Lawrence build the first cyclotron in Berkeley with the intention of accelerating particles to a high enough energy so they can be used to produce new particles [6]. From a German article written by Winderöe on multiple acceleration on a positive ion Lawrence got the idea that he could make the particle reach a high enough energy by using a magnetic field such that the particle would bend and go around in a circle, and for each round in the circle the particle would increase its energy. It was his brother however, John, who realized that this can be used in medical applications. John was a medical doctor and was interested in the usage of radioisotope to treat cancer

¹Pulmonary circulation is the portion of the circulatory system that carries deoxygenated blood from the right side of the heart to the lungs, where the blood oxygenates and returns to the left side of the heart.

and in 1935 he came to Berkeley to do some testing on mice which had a positive outcome. In 1937 John moved to Berkeley and performed his first treatment on a human using radioisotope, in fact Ernest and John's mother got diagnosed with cancer in her uterus with a few months to live. John and Ernest got her to a special clinic that treated her with radiation and she lived for another 15 years [6].

In the 1950s, positron emitters and their usage in imaging techniques were described by Frank R. Wrenn, Myron L. Good and Philip Handler [7]. This opened up to the possibility to use radioactive isotopes together with a discover and diagnose different diseases. The mainly used radionuclide that was used before the discovery of ^{99m}Tc in 1964 [8] was ^{131}I and was used to study and diagnose thyroid disorders [3]. After Paul Harper and his colleagues used ^{99m}Tc for brain scanning [8], because of its flexibility for labeling, ^{99m}Tc was being used to study various of organs in the body. The gamma property of ^{99m}Tc was proven to be good for imaging, and the fact that it could be produced in a long-lived generator made it attractive for usage in hospitals [8]. One other big development was the mathematics to produce an image of a body based on angular views around the patient [3]. This technique was important for the evolution of nuclear imaging, and are the basis for imaging procedure that is being used today like SPECT, CT, MRI and PET.

Today, nuclear medicine is used for both diagnostics and treatment of a patient. PET and SPECT is two imaging techniques that are most used in diagnostics. ^{18}F is the most used radionuclide in PET scans. Attached to a sugar molecule it will find cancerous cells in the body. In SPECT scans, ^{99m}Tc is often used to diagnose coronary artery diseases and stroke. These two illnesses are the number one killer worldwide and ^{99m}Tc is therefore the most widely used radioisotope. Near 80 % of all procedures are using ^{99m}Tc . In therapy, the radioisotope that is being used depends on the illness to the patient. What type of radionuclide that is being used in both diagnostics and therapy depends on the characteristics of the isotope.

2.2 Characteristics of medical radionuclides

Radionuclides for therapeutic and diagnostic use have three principle factors that affect the ability for it to perform as a medical isotope in medicine [9], biological-, physical-, and chemical properties. The biological and chemical feature involves the stability in a living organism, biological half-life, toxicity, tissue targeting and retention of radioactivity in the tumor [10]. The physical characteristics include physical half-life, energy of the radiation, purity of the radionuclide, type of emission, daughter product(s) and method of production [10].

In this section, a book by Krane [11] is used to describe the general nuclear reactions.

2.2.1 Half-life

The physical half-life, $t_{1/2}$, of a radioactive substance is the time it takes for a given amount to be reduced by half as a consequence of its decay and is described by the formula

$$N(t) = N_0 e^{-\lambda t} \quad (2.1)$$

where N_0 is the amount of initial substance, $N(t)$ is then how much there is left after time, t , and λ is the decay constant. This equation can be solved and the half-life of the decaying quantity is then given as

$$t_{1/2} = \frac{\ln 2}{\lambda} \quad (2.2)$$

When a radioactive pharmaceutical gets injected into a human body, the biological half-life is important. It is the time it takes for a living body to reduce a biological substance by half through its biological processes. Therefore, when a radiopharmaceutical is produced to be used in a patients body, this is something that has to be considered carefully. It should be long enough to do a procedure, but short enough to not damage unnecessary healthy tissue.

In diagnostic, the half-life should only be a few hours. From when the radioisotope is produced it starts to decay, and it is important that the half-life is long enough such that it can be used. If a patient is taking a PET-scan, depending on the radionuclide, a procedure using ^{18}F usually takes 30-60 minutes to execute. Therefore, the radiopharmaceutical have to have a half-life longer than it takes from production of the radioactive substance to the end of the procedure.

For a therapeutic isotope, it should have a half-life of approximately a few days. It should be long enough to deliver the right amount of dose to the area of interest. But the most important factor is the effective half-life, that is both physical and biological half-life within a patients organ[10]. The physical half-life is well known, but the biological half-life is dependent on several things, like what kind of tracer is being used, metabolism, uptake and how the pharmaceutical leaves the body[10]. It depends on the uptake mechanism, type of tumor and method of administration[10] on what kind of tracer and isotope that should be used. If the patient has a low uptake, the physical half-life should be longer such that it will not emit radiation before it has reached the tumor, but not too long such that it sends out radiation on its way out of the body.

Travel time is also a factor when choosing an isotope. Based on the physical half-life, some hospitals can produce some radioisotopes themselves. If they have a cyclotron in the hospital they can produce some isotopes, such as ^{18}F with half-life of 109 minutes, that is used in PET scans. In those cases the half-life can be shorter than if it has to be transported to another part of the county.

2.2.2 Stopping power

When a charged particle penetrates an absorbing medium it will slowly slow down due to energy loss, the parameter that describes this is called stopping power. There are two known cases of stopping power[12], collision, S_{col} , and radiative, S_{rad} , stopping power. S_{col} is a result from charged particles interaction with orbital electron of the material while S_{rad} is when the charged particle interact with the nuclei in the material. The total stopping power is $S_{tot} = S_{col} + S_{rad}$

Electrons can lose energy by Bremsstrahlung and by ionization. At high energies where a particle travels faster than the speed of light in a medium, Cherenkov radiation happens. This is an important phenomena but the energy loss in this process is small[13].

Even though the majority of the energy loss for electron is by collisions, the emission of bremsstrahlung photons is also important. When an electron is traveling close to an atomic nucleus it can decelerate due to the Coulomb field of the nucleus and atomic electrons. When this happens, the electron will transfer energy to a photon that gets emitted. When the kinetic energy of the electron increases the yield of bremsstrahlung increases.

When it comes to photons, the beam intensity $I(x)$ will decrease with

$$I(x) = I_0 e^{-\mu x} \quad (2.3)$$

where μ is the attenuation coefficient and are dependent of the energy $h\nu$ of the photon and the atomic number Z of the absorber[12]. I_0 is the initial intensity of the beam at depth in medium where $x = 0$.

Photons will also interact with a medium through photoelectric effect, Compton scattering and pair production, these are briefly discussed in section 2.1.3 under γ and X-rays.

LET For therapy, the biological effect depends on how the isotopes decay, and how it distributes its energy to the surrounding medium. High linear energy transfer (LET) will concentrate the deposited energy to a small area near its tracks, as a result the biological damage is grater and there will be more damage to the DNA and other cellular structure. LET is closely related to the linear energy loss $\frac{dE}{dx}$ except that LET does not include radiative losses of energy.

As shown in figure 2.1, high LET will do more damage in a smaller area and the therapeutic effect is significantly better than for low LET. α is a nuclei with high LET. Low LET, as figure 2.1 shows, is when the particle distribute little ionization throughout the medium, which results in little to no damage on the cells along the way. In diagnostic, we want to use isotopes that decays with particles who follows low LET and the opposite for therapy.

Since the ionization of a material increases when the energy decreases for particles[12], the radiation will be more significant at the end of its tracks. Figure 2.2 shows how the ionization alters as the particle goes through a medium. The

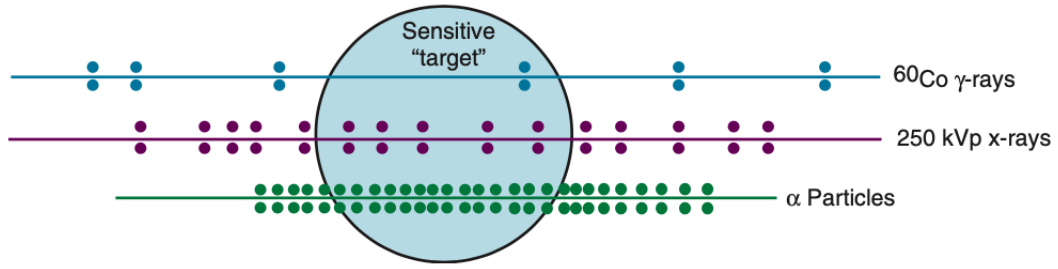


Figure 2.1: An illustration from [14] of low LET vs high LET

peak at the end is the bragg-peak and is correlated with high LET where most of the ionization happens at the end of its range.

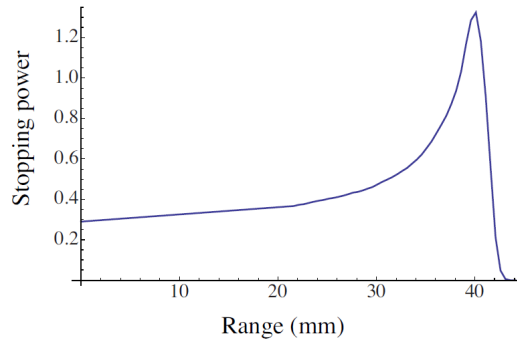


Figure 2.2: Illustration of the bragg curve for charged particles. Picture from [15]

2.3 Decay modes

When a patient have to go trough diagnostic and therapy with radioisotopes, the isotopes that is being used have do undergo decay that gives the wanted radiation. For therapy, decay modes that gives a γ or e^+ is important. In therapy, the wanted decay mode gives off an e^- , α and smetimes γ .

2.3.1 β -decay

Decay by β can happen by β^- , EC (see auger electrons) or β^+ decay. β^- :

$$n \rightarrow p + e^- + \bar{\nu}_e \quad (2.4)$$

β^+ :

$$p \rightarrow n + e^+ + \nu_e \quad (2.5)$$

For therapy, e^- will travel longer in a medium than e^+ (will annihilate fast). **WHY?** e^- will ionize matter on its path and we want to take advantage of that. Therefore, it is β^- decay we are interested in when we treat a patient. β^+ decay are used in diagnostic. If we want to take a picture of the biological activity in a patient, e^+ is the source of interest for usage of PET scans, which I will discuss in section 4.1.

Compared to α particles, β particles travels a lot further and creates less ionization to the surrounding medium on its path. This is important when the biological effect matters. There are however, difference in the range of various of isotopes that undergo β decay. Short range β -particle emitters have a relatively short range, a couple of μm , in tissue that is suitable for treatment of small tumor metastases. The β -emitters with long range in tissue can penetrate with an average of more than 2 mm and can be used to treat i.e. rheumatoid arthritis, which is an autoimmune disorder that often affect joints.

The energy of β particles are not distinct, but rather more spread out in a spectrum.

The spectrum in figure shows how the energy are divided between the electron and the anti neutrino. The electrons of beta particles will interact with the surrounding electrons by dismissing the it from its orbit and producing a ion pair or cause excitation. Since the amount of damage tissue is related to LET and β particles penetrates further in to a medium than α particles, β particles has lower LET than α does. Exposure to β results therefore in less damage.

^{90}Y is one popular long range isotope used in therapy, it has a mean range of 3.900 μm and an average energy of 935 KeV, while a widely used short range beta emitter is ^{131}I . It has a mean range of 910 μm and an average energy of 182 KeV. The energy will vary from almost zero to the maximum energy, and the average energy is approximately one third of the maximum energy. For therapy, the mean energy dependent mean and maximum range in tissue for beta emitters.

2.3.2 Electron capture, Internal conversion and Auger electrons

When low-energy X-ray photons collides with an atom and electron from the inner shell can get ejected. If the nucleus has too many protons, the electron can be captured by the nucleus where the electron and a proton turns into a neutron. This is called electron capture (EC). When this happens, there will be an empty spot in the K shell which one of the electrons in the L shell will try to fill. There will therefore be a vacancy in the L shell. This process releases an energy of $E_K - E_L$ and can be transferred to another electron in the L shell, resulting in ejecting it from the atom. This electron is called auger electron. It is a low energy electron that will have a kinetic energy of $T = E_k - E_{L1} - E_{L2,3}$ [16], where $E_k - E_{L1}$ is the difference in binding energy. It has a short range and will deposit

all of its energy in a highly localized site. The distance the auger electrons travels from $0.5 \mu\text{m}$ to a few nm in biological tissue [17], it follows high LET and will therefor do a lot of damage to the DNA and other important structure in the cell nucleus that are important for the cell to live. Auger electrons are highly effective for killing these structure and breakage of double DNA strands due to its energy deposit in a short distance (see figure ??).

One other way auger electrons are created is due to the effect of internal conversion (IC). This is when an electron get knocked out of the inner shell and one other electron from a higher orbital jumps down to the inner shell. In this process energy gets realised by photons which can knock out an electron from the outer shell, the auger electron.

2.3.3 γ -decay and X-rays

γ and X-rays interaction with matter

Emission of γ -rays are usually a product from the radioactive decay of atomic nuclei. If the nucleus is left in an excited state, it will de-excite by either emit a photon.[13]. γ decay are described as

$${}^A_Z X^* \rightarrow {}^A_Z X + \gamma \quad (2.6)$$

where the excitation energy is transferred to a γ photon plus the recoil energy.

If the notation includes a "m"

$${}^{Am}_Z X \rightarrow {}^A_Z X + \gamma \quad (2.7)$$

in that case, the nucleus is in an isomer state where it is long-lived ($T_{1/2} < 1\mu\text{s}$) in an excited state[13]. The transition will then be described as an isometric transition (IT).

Competing with γ -ray emission, internal conversion (IC) is a way for the nucleus to de-excite. It will electromagnetically interact with a shell electron witch leads to emitting of the orbiting electron[18].

X-rays, as γ -rays, are electromagnetic radiation. The difference is that X-rays has its origin outside of the nucleus. When X-rays and γ penetrates a medium it can create different effects in the body by ionization. Photoelectric effect, Compton effect and pair production.

The photoelectric effect happens when a nucleus absorbs the incoming photon. The energy absorbed will be used to ejected an orbiting electron from the atom. This electron is called a photoelectron[13]. The electron cannot be emitted if the incoming photon energy is lower than the binding energy of the shell to the electron. The energy of the electron emitted is then $E_e = h\nu - E_b$. The

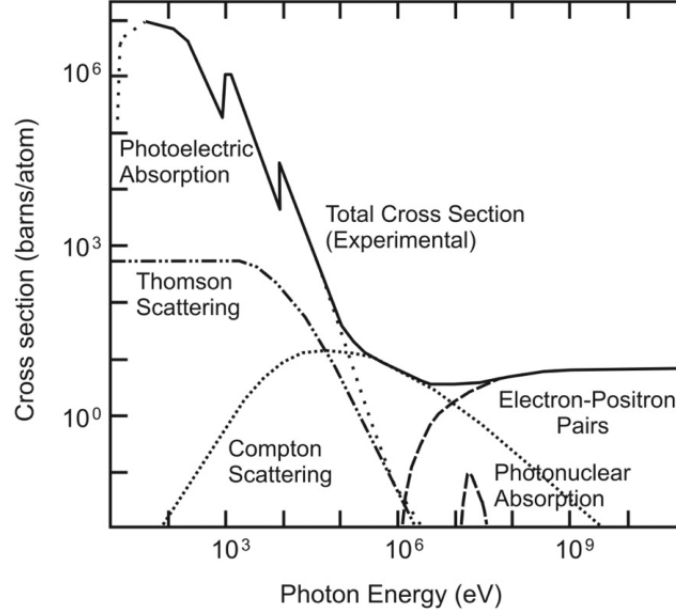


Figure 2.3: Cross section for photoelectric effect, Compton scattering and pair production. Taken from [19]

probability of photoelectric effect happens is measured as the cross section

$$\sigma = \text{constant} \cdot \frac{Z^n}{E^3} \quad (2.8)$$

where n is equal to 4 or 5[20]. The photoelectric effect will increase with the atomic number and by that reason we are using lead ($Z=82$) as shielding for gamma radiation.

Compton effect is scattering of a photon due to an electron. When an incident photon interacts with an atomic or free electron, it will bend from its original path with an angle θ . The electron are then being emitted with an angle of ϕ . The energy is shared between the emitted electron and the photon, where the incident photon will transfer some of its energy to the electron. The final photon energy is

$$h\nu' = \frac{h\nu}{1 + \frac{h\nu}{m_e c^2}(1 - \cos\theta)} \quad (2.9)$$

The atomic cross section is $\sigma_a = Z\sigma_e$ where σ_e is the electronic cross section, assuming a free electron[13].

The last main effect is pair production. The incoming photon disappears, and spontaneously turns into an electron-positron pair. This can only happen if the energy of the photon is greater than the total rest mass of an positron-electron

pair, 1.022 MeV. Since X-rays and γ can penetrate long distances without ionizing the medium, emerges from the body and has a low LET, they are at good use for diagnostic of a patient with different imaging technique. The cross section of pair is production is dependent on Z^2 .

Figure 2.3 shows how the three main interaction of a photon is dependent on energy. For lower energies it is clear that the photoelectric effect is dominant, pair production is more probable to happen in higher energies.

γ decay

When a nucleus is being excited, either by the creation of a compound nucleus, with a direct reaction or if a nucleus decays with one of the decay modes described in 2.3, the nucleus will decay into its groundstate by emitting a γ -ray. This is illustrated in 2.4.

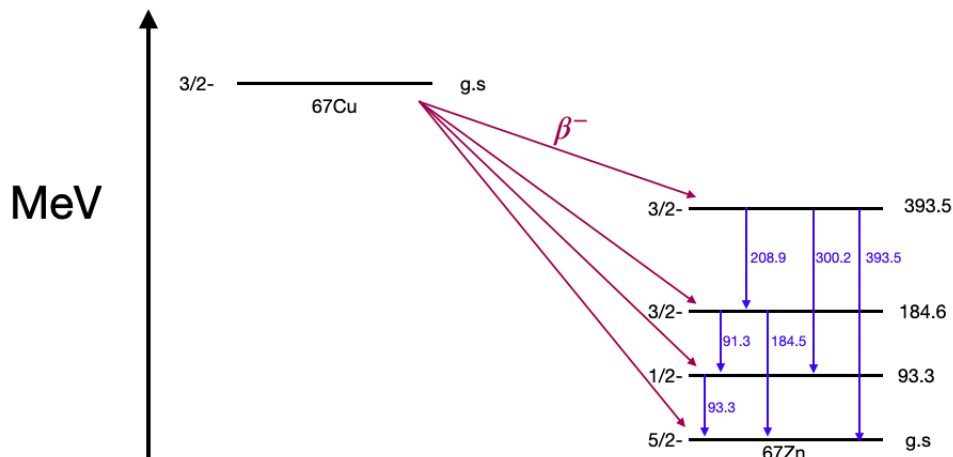


Figure 2.4: An illustration on how ^{67}Cu is β -decaying down to ^{67}Zn .

As 2.4 shows, the γ s that are emitted in ^{67}Zn when the nucleus deexcites can have different energies. This depends on **what? why is 300 MeV more favourable than 393 MeV gamma? - Sunniva?**

It is the γ -rays from ^{67}Zn that we are using to claim that ^{67}Cu were made, see 5.

2.4 Production of medical isotopes

2.4.1 Production rates and cross-sections

Production rate

When there is a nucleon-nucleon interaction in a nuclei there will be change in the tribe of the nucleus and therefore the production of different nuclei can happen. This can happen by exciting a nuclei, exchange of nucleons or fusion. [13] If a nucleus ends up in an excited state after a collision before it decays into a new nuclei, the states are described as

$$a + A \rightarrow C^* \rightarrow b + B \quad (2.10)$$

[13] where A is the target, a is the projectile, b is the light product and B is the new heavy product. If however, the nucleus doesn't get excited and decays into a new product, the general equation for a nuclear reaction is

$$a + A \rightarrow b + B + Q \quad (2.11)$$

where Q is the reaction energy. The nuclear reaction is often written as

$$A(a, b)B \quad (2.12)$$

Equation 6 describes the creation of a compound nucleus, C^* . With low energies and a direct hit on the center of the nucleus gives the highest probability of creating a compound nucleus. When a direct reaction happens with higher energies and often at the edge of the nucleus, equation 7 defines this process[12]. The production of radionuclides can therefore a mixture of these two reactions.

There are many ways as to which radionuclides can be produced. Neutron induced reaction, charged particle induced nuclear reaction, fission, selective separation and gamma induced are the most important ones[12].

The neutron induced nuclear reaction can happen with a thermal neutron, where fission and (n, γ) is the production route, and with fast neutrons, where we can get (n, n') , $(n, 2n')$, .., (n, p) , .., (n, α) , and fission reactions. One advantage of using neutrons as a beam is that it can penetrate the target with quite low energies, in some cases it only needs an energy of 0.0025 eV to penetrate a nucleus[12]. Charge particles however, needs to overcome the coulomb barrier to enter the nucleus.

BILDE HER AV CROSS SECTIONS

Charged particles can give rise to elastic-, inelastic- as well as other reactions. Elastic reaction is where the product before the collision is the same as the product after the collision, a (p, p) reaction. The masses, momentum and the kinetic energy before and after the interaction are identical[12]. Inelastic reaction, (p, p') , happens when energy is transferred to the product and the difference in kinetic energy is saved as excitation energy in the nucleus. Depending on the beam and energy, figure ?? shows that there are different nuclear reaction channels that can happen. With higher energies, the outgoing particle can be heavier or more in number.

2.5 High purity Germanium detector

Most semiconductors are based on the n-p junction, which is a boundry between two different materials. A crystal is being doped on one side with a material that has one excess electron in the outer shell and therefore is called the negative side, or n-type. The other side is doped with a material that has one less electron in its outer shell, leaving a hole. This side is called p-type, since it is more positive. When these two types meet the electrons from the n-type will be attracted to the holes in the p-type and leaves a depletion zone in the middle of the crystal with a negative charge on the p-side and a positive charge on the n-side. This makes it hard for the electrons to travel to the p-side unless there is a voltage applied to the crystal.

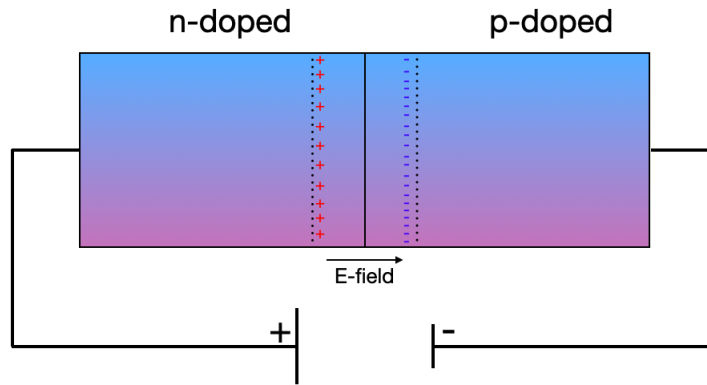


Figure 2.5: n-p junction of a semiconductor

Chapter 3

Cu64,67

QUOTE.

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3.1 what can we do to make the nuclear medicine better?

Nuclear medicine is a growing field within diagnostic and treatment of a patient. The usage of many different medical isotopes is being used to different illnesses, and there is always good to have new isotopes that can be used in nuclear medicine. Nuclear medicine is being produced as a personalized medicine. This indicates that the patients characteristics such as anatomy, physiological and genetic are being evaluated when radioisotopes for diagnostic and therapy are being chosen [26].

To make the planning and execution of a treatment easier and more efficient, researchers are looking at different ways to combine therapy and diagnostics. ^{64}Cu and ^{67}Cu are two theranostics isotopes that are interesting due to their half-life, decay scheme,

Isotope	Half-life	Decay (%)	Usage	Source
^{64}Cu	12.07 hours	$\beta+$	Diagnostic: PET	Cyclotron
^{67}Cu	2.57 days	$\beta-$ (100)	Therapy: Tumor	Cyclotron

-medical perspective: we want to introduce a new theragnostic pairs to use in hospitals - how wonderful Cu64,67 are – properties – papers – better than a lot of the stuff we are already using – motivation for my work — can adjust ratio for 64,67 Cu by tuning the energy of the beam

3.1.1 My motivation

3.1.2 Physics motivation

- cu64,67 are amazin but now, we have not a good way for making them. tell about my way of create them – deuterium breakup (n,-) way – how we are doing it

3.2 Theranostic

Chapter 4

The experiment

In this theses we wanted to study the $^{nat}\text{Zn}(n, x)^{64,67}\text{Cu}$ reaction. In the following chapter, the experiment in itself is being discussed. Sections 4.1.1, 4.1.2 and 4.1.3 describes the 88-Inch cyclotron used int the experiment, the deuterium breakup process and how the targets were stacked together before being irradiated.

4.1 The experimental setup

The experiment was preformed on August 2018 at the Lawrence Berkeley National Laboratory where a deuterium beam with an energy of 16 and 33 MeV was used. The deuterium was used in the purpose to hit a berillium target, creating a neutron beam. Five targets, zinc, zirconium, indium, yttrium and aluminum were then irradiated with the neutrons, for illustration see fig 4.3. For the 16 MeV run, the targets were irradiated for 7 hours and 57 minutes, and 1 hour and 52 minutes for the 33 MeV run. The length of the irradiation is different since the energies are different. A higher energy beam and a longer irradiation time will increase the dose and open up for other production rates. Since the germanium detector that was used to count the foils after irradiation have a couple of nS of deadtime, there will be a lot of information lost if the 33 MeV run was longer than 1.52 hours. The activity would have been much higher and the detector would not had been able to catch all the activity. When 33 MeV is used, we know from section 4.1.2 that the number of neutrons that are being produced in the prosses is two to three times higher than for 16 MeV. This is also a reason for the shorter irradiation time for 33 MeV.

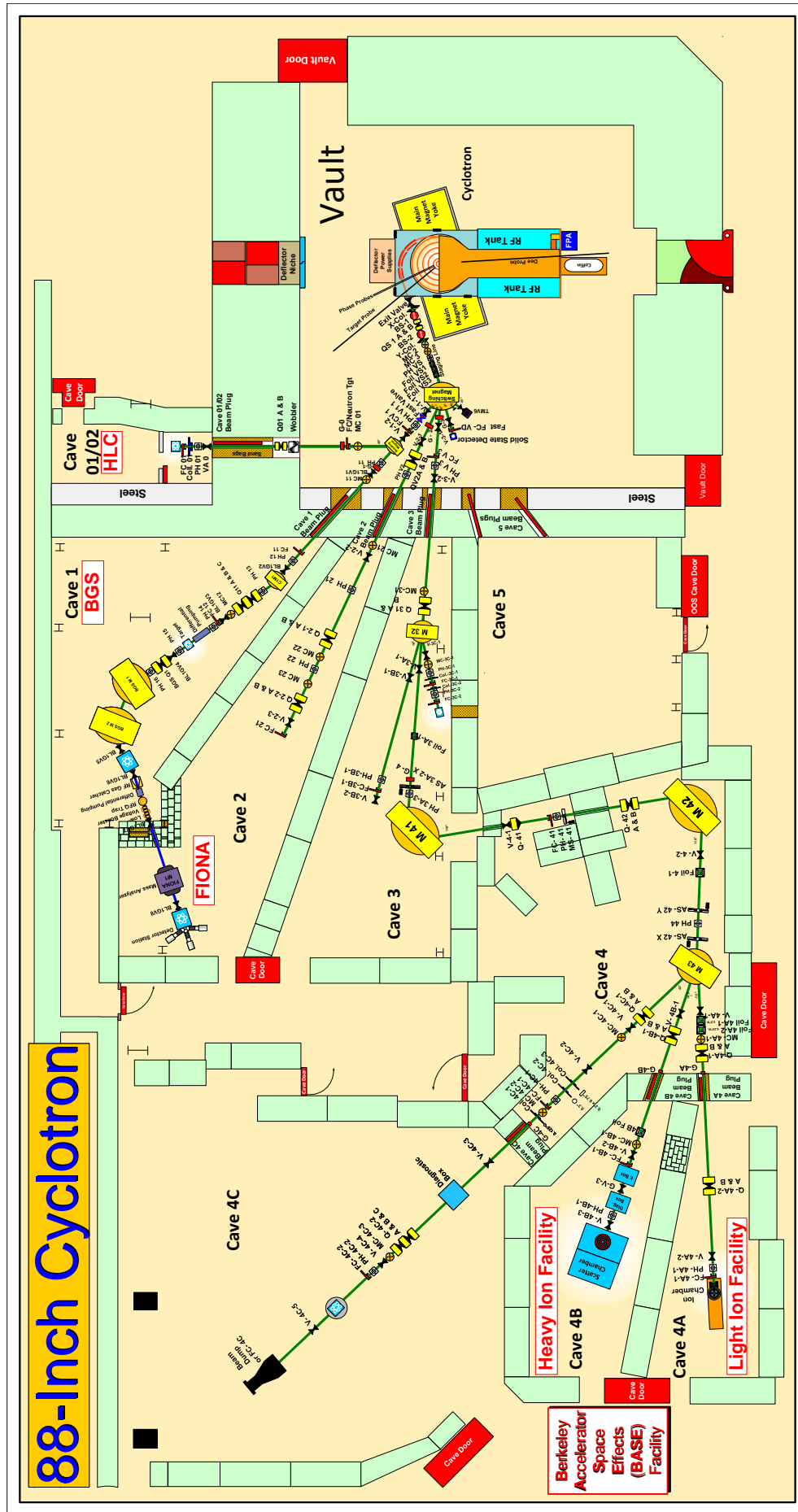


Figure 4.1: Map over the 88-inch cyclotron at Lawrence Berkeley National

The irradiation took place in cave 0, see figure 4.1. In figure 4.1 there is a green line, which is the beam line. For the beam to get to cave 0 it has to go through two bending magnets where the deuterions with the highest and lowest energies will be lost when the beam are being bend to the right beamline. Cave 0 has good shielding and when high energy beam enters the cave there will be more x-rays in the room. Shielding is therefore one of the main reasons to why high energy experiments are being done in cave 0.

The beam had to be tuned such that it didn't have a lot of angular distribution before it hits the targets. The neutrons gets emitted from the beryllium target with an angular distribution and the targets are based on the energy the neutrons hits the target with. The angle should be about zero degrees and the targets are therefore placed 10 cm back from the be disk. The targets will not be activated with the full beam energy and the intensity of the neutrons will decrease with distance-squared, x^2 . reference to where you talk about solid angle

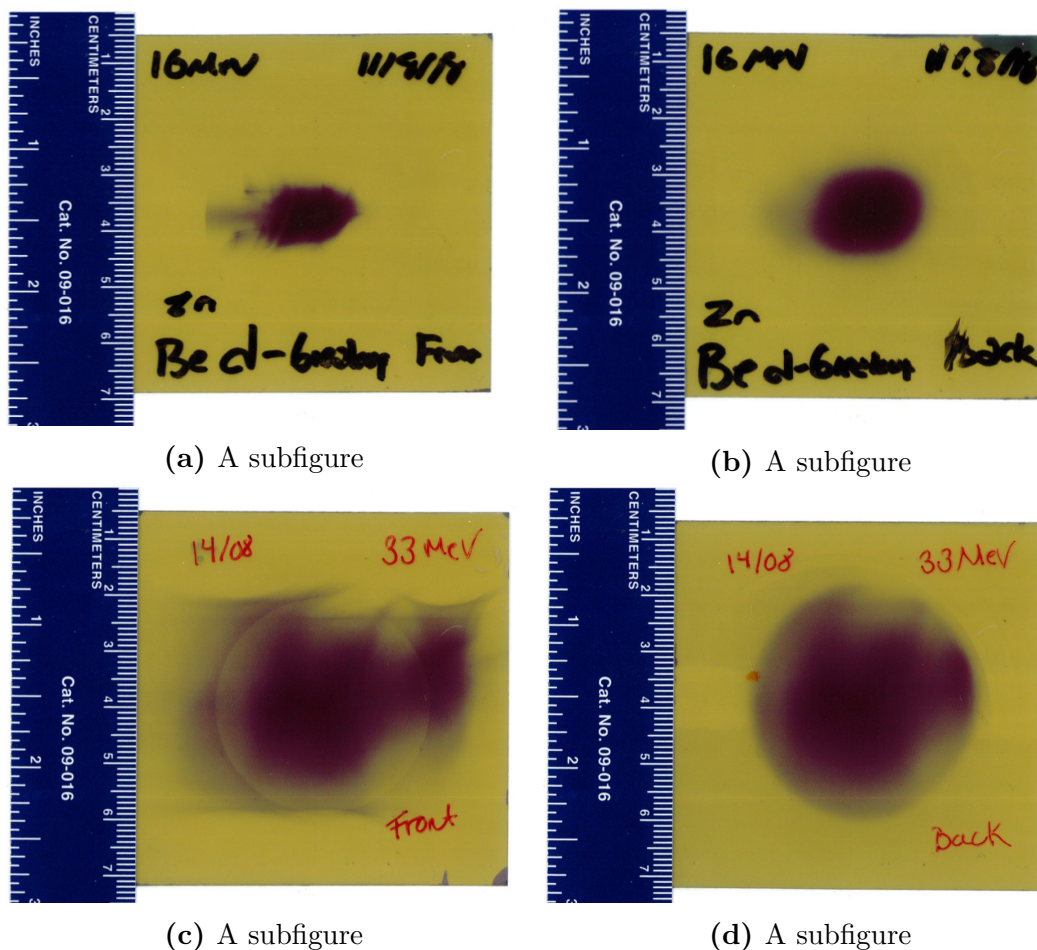


Figure 4.2: Pictures of the beam shape. (a) and (b) shows the beam during the 16 MeV run, and (c) and (d) shows the beam for the 33MeV run.

After the the tuning, the targets were placed in a metallic box, see figure 4.7, and then placed in a tube that were sealed with vacuum. The vacuume were there to decrease the amount of interactions between the air prticles and the charge particle inside the targetholder.

When the irradiation was done, the targets were removed to the counting room where they were placed in a gernanium detector [see section???](#).

4.1.1 The 88-Inch cyclotron

Being one of the first cyclotrons used to produce medical isotopes [6], the Berkeley lab is still used to produce radionuclides that are a part of the research in the medical industry. Lawrence national laboratory is located in the hills, directly over the university of California at Berkeley overlooking San Francisco bay.

The 88-Inch cyclotron is used in research on varies of fields including astro-physics and nuclear structure. The accelerator is a 300-ton, $K = 140$ sector-focused cyclotron with the ability to run with both heavy and light ions. The K-value represents the kinetic energy a particle can reach in the cycotron and can be used to rescale the energy reach for protons to other charge-to-mass ratio $E_k = AK(\frac{Q}{A})^2$. Where A is the change in atomic mass, Q is the amount of energy released in a reaction and E_k is the kinetic energy to a proton. The facility of the cyclotron is shown in figure 4.1, it mainly has five experimental caves where cave 0 is the one used for neutron beams and isotope production. [The cyclotron can accelerate ions up to 70 MeV ...](#)

- k-value (descuss energies for hospital cyclotrons and what energies for Cu64,67), what is it, how does it work - no more than a page or two (look at other theses o se how deep you should go)

4.1.2 Deuterium breakup prosess

Since a neutron beam has to be created, the usage of deuterium to make neutrons was done by thick target deuterium breakup. This prosess has been studied by Meulders [27] and Saltmarsh [28] and it involves splitting up deuterium into a proton and a neutron using a thick target of beryllium.

Deuterium break up can be done on any material but natural beryllium is a dense medium that has the quality as a good conductor of heat, it is stable and it is impossible to produce a radioactive activation product using deuterium. Be has a low Z wich means that the atom dencity is lower and as a result, the breakup yield will increas. There is no excited states in deuterium [29] and because of the low binding energy of 2.22 MeV it will easly split into one proton and one neutron when it hits the beryllium target with energy greater than the bindng energy of the deuterium.

Deuterium breakup reaction can happen in two ways: direct process and excitation of the deuterium. The first one happens at higher energy where a proton

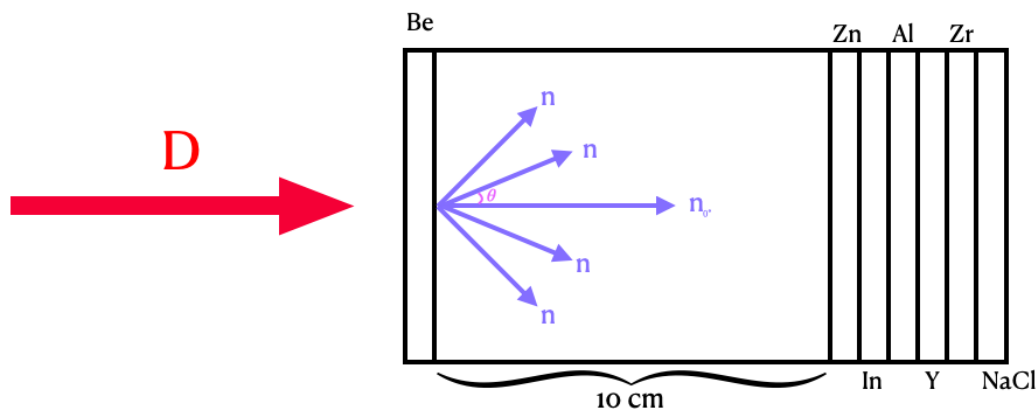


Figure 4.3: Illustration of the deuterium breakup and how the neutron beam is dependent on the angle it gets emitted by.

interacts with an atom or the electron cloud in the Be target and strips off the proton, allowing the neutron to pass through the beryllium disk. At lower energies there is a probability to excite the deuterium. If the deuterium is being excited above the binding energy, there will be a compound nucleus and it will decay by sending out a proton or a neutron. The energy distribution of the neutron beam is visualized in fig 4.5 where the neutron flux increases with materials that have lower atomic number. This is one of the main reasons to why beryllium was used in this experiment.

Fig 4.4 shows how the energy varies with the angle of the neutron that is being emitted in the $\text{Be}(d,n)$ reaction. When the angles of the neutrons get bigger, the neutron flux decreases [27]. Even going from zero to ten degrees decreases the neutron flux by a factor of two. In this experiment the angles of the neutron flux were aimed to be zero degrees since the energy drops with the angle [27]. The assumption that the neutron flux spectra for 33 MeV and 16 MeV as in fig 4.4 is made with the knowledge that the energy will shift down. The expectation of a neutron flux energy of 16 MeV deuterium beam is approximately 6.5 MeV and approximately 16 MeV for the 33 MeV beam. This is shown in fig 4.5.

4.1.3 Stack design

A specific design of the targets were used to make sure that the cross section for each target could be measured for the different energies we irradiated the targets with. Five different targets, Al, In, Y, Zn, and Zr, were made into small packings using kapton tape to tightly seal them (see figure 4.6), **braker dette ogs til hinde contamination siden targetset kan vre sjr og unng spredning av lst radiaktivt stoff.**

The kapton contains polymer carbon, hydrogen and oxygen but the sticky part

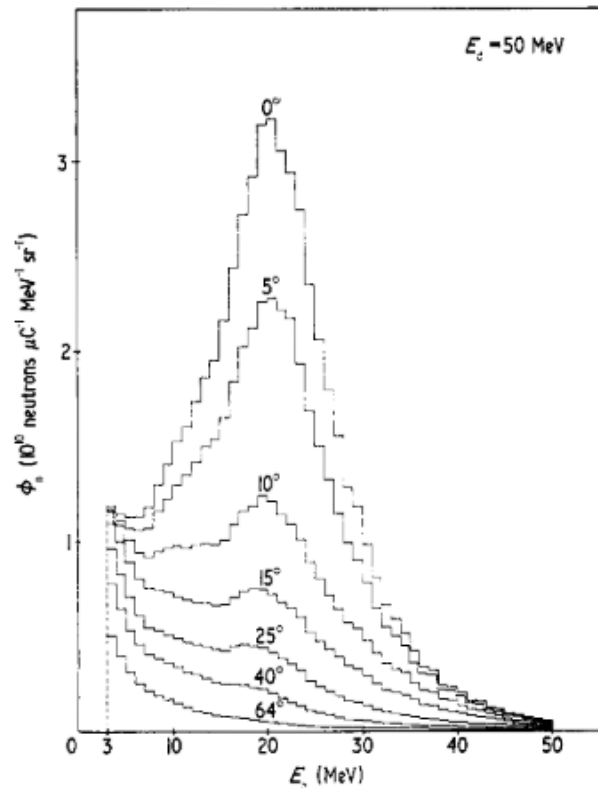


Figure 4.4: angles

of the tape are made from silicone. For each target we measured their weight, thickness and length **Using what?** to calculate the uncertainty of each target. The targets were attached to a plastic frame and were placed in the end of a metallic **box made of?** where we wanted the targets to be 10 cm from the Br disk, the targets were held in place by a spring between them and the disk in front.

The metallic box was then placed in the beamline by inserting it into a vacuum tube. The vacuum was there to prevent the beam to hit air molecules and to make the beamline as straight as possible, such that more of the beam would be able to hit the targets. The radiation of the targets were on for approximately **12 hours?** for both 16 MeV and 33 MeV. [6]

- photos and stuff - monitor foils

4.2 Radiation

- beam monitor to measure the – plot the beam current as a function of time to justify that we can make the math that we do

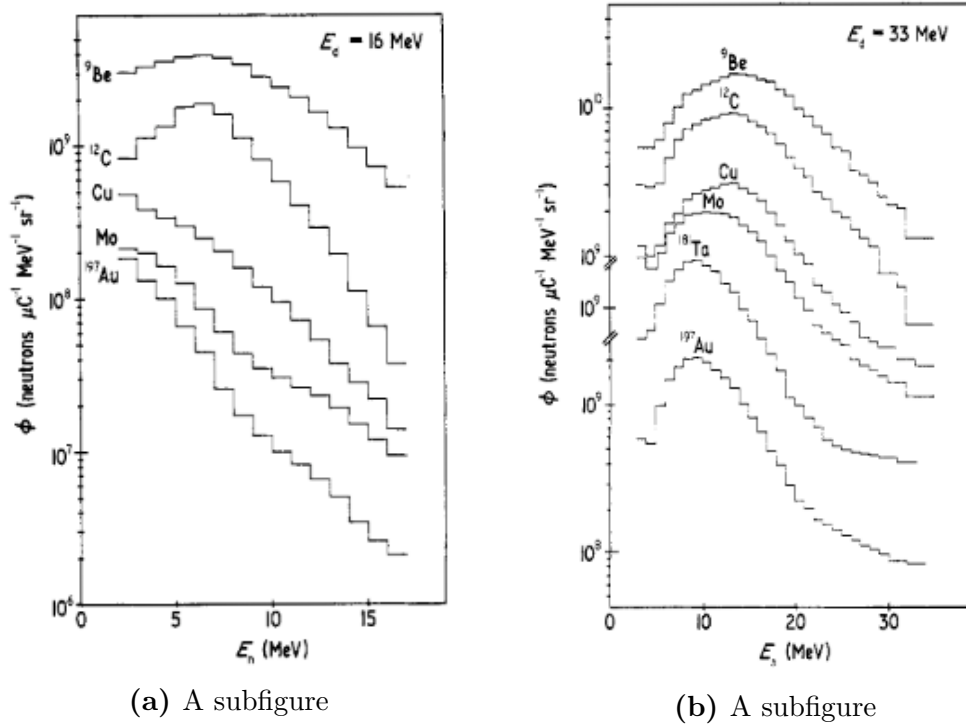


Figure 4.5: Neutron energy spectra produced with 16 MeV in (a) and 33 MeV in (b) that were produced in the (d,n) reaction. Taken from [27]



Figure 4.6: Picture of how the targets were sealed before irradiation

4.2.1 Counting

- after each radiation, we removed the foils to the counting room (how long did this take?) hvor lang ti tok det fr vi begynte telle etter EOB?

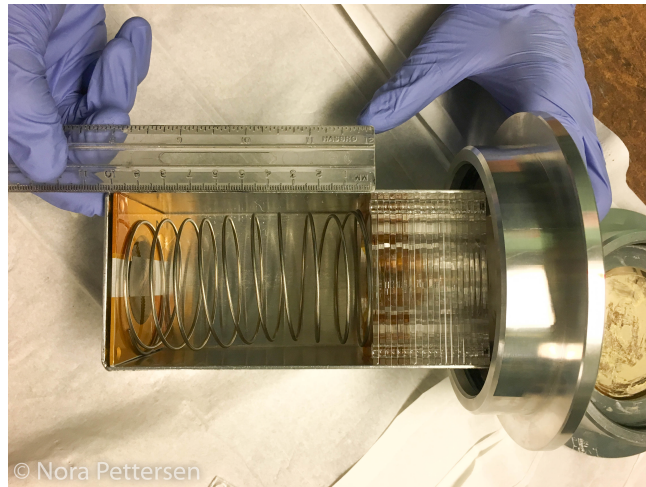


Figure 4.7: Picture of how the targets were placed in the metallic box before irradiation

4.3 Gamma spectroscopy

- Detector, ortec gmx detector – forklare hvor dypt jeg skal g inn i physics (doping, n-p junction)

4.3.1 Gamma spectra

- deadtime, og alt det der

4.3.2 Calibrating

- detector efficiency - curves at different position

Chapter 5

Analyse

“If you ever start thinking too seriously, just remember that we are talking monkeys on an organic spaceship flying through the universe”

— Joe Rogan

5.1 Fitz peak

FitzPeaks Gamma Analysis and Calibration Software is a program that is being used for gamma spectroscopy analysis and calibration. The program uses linear least square routine **DESCRIBED WHERE?** and tries to fit the data to a gaussian function.

- Data to activity (the math). Peak counts to activity - i - Aktivitet to A_EOB - Production physics (Analyse eller Resultater?) - how we calculate cross-section, the work with John to use the monitor data to neutron fluxes that I'm going to use for calculate the cross section for my isotopes

5.1.1 Regression process

χ^2 - distribution is a set of distribution depending on how many sets you have
- non linear regression

5.2 Production physics

Activity is written as

$$A(t) = A_0 e^{-\lambda \Delta t_{decay}} \quad (5.1)$$

And that decays with

$$\Delta D = \int_{t_{live}}^{t1} A_0 e^{-\lambda t} dt = \frac{A_0}{\lambda} (1 - e^{-\lambda \Delta t_{live}}) \quad (5.2)$$

Efficiency is then

$$\epsilon_{eff} = \frac{N_c}{\Delta D I_\gamma} \quad (5.3)$$

where N_c is the number of counts. The efficiency then becomes

$$\epsilon_{eff} = \frac{\lambda N_c}{A_0 I_\gamma (1 - e^{-\lambda \Delta t_{live}}) (e^{-\lambda \Delta t_{decay}})} \quad (5.4)$$

A_0 :

$$A_0 = N_T \cdot (1 - e^{-\lambda t}) \cdot \int \sigma(E) \frac{d\phi}{dE} dE \quad (5.5)$$

Average cross section:

$$\sigma_{avg} = \frac{A_0}{N_T (1 - e^{-\lambda t}) \cdot \phi_{avg}} \quad (5.6)$$

From deuterium to neutron energy:

$$E_{avg} = \frac{\int E \cdot d\phi(E)}{\int d\phi(E)} \quad (5.7)$$

where $d\phi dE$ is the energy dependent neutron flux.

To find the errorbars in the neutron energy a calculation of FWHM of $\frac{d\phi}{dE}$ from the meulders paper[27]. A list of energies and a list of the $d\phi dE$ is being plotted where the maximum value of the neutron flux is collected. **INCERT FIGURE OF MEULDERS** The half of the maximum value is being used to find the intersection between E_n and $d\phi dE$ which gives two energy values, x_1 and x_2 . To find the errorbars δE of the neutron flux these equations are being used.

$$E_{x_1} = E_{avg} - \delta E_{lower} \quad , \quad E_{x_2} = E_{avg} + \delta E_{upper} \quad (5.8)$$

where $\delta E_{lower} \neq \delta E_{upper}$ because $\frac{d\phi}{dE}$ is not symmetric, as shown in fig??

Chapter 6

Results and dscussion

- Cross sections for all the isotopes - experimentall data with the data that has been meassured so fare

6.1 TALYS

- tolking av resultat. Hvilke energier er best i forhold til hva vi ser? - Verdien av dette i fremtiden? - how can we desien target for kunne produsere mer Cu67. hvilke cyclotrons can we use for that?

6.2 ALICE

6.3 EXFOR

Chapter 7

Summary and outlook

7.1 Future work

quote

— by

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