RecEx7.pdf (https://canvas.umd.umich.edu/files/39301418/download?downloa

Chem 134 W20 **Rec. Ex. 7** **Name**: _____

Lab sec: _____

Due beginning of recitation March 10 or 11. Answers and work must be shown on **these sheets** wi **stapled** together (or double-sided).

Refer to the Ch. 6A lecture slides or the textbook to answer the following questions.

1. **Energy** is defined as the *capacity to do* _____. **Work** is a _____ acting o

   _____. The SI unit of energy is the joule. **1 J** = _____

3. **Thermal** energy is a form of (*potential, kinetic*) _____ energy. It is the en

   _____

4. **Potential** energy is _____ energy… it has potential to do _____

5. Briefly explain **chemical potential energy**.

6. Systems with high potential energy:

   _____

   _____

7. A piece of warm metal is dropped into a beaker of cold water. Briefly explain the process that oc equation in the diagram.

8. a) How does the energy unit **calorie** compare to the **joule**? A calorie is more than (insert a num