Name: _____ Per: ____ Date: _____

Balancing Act III – Predicting Products in Equations

Look at the reactants and determine the type of reaction (S, D, SR, DR, C). Using the type of reaction, predict the products. Write the correct formula of the product(s) and then balance the entire equation.

_____ a. _____ $C_6H_{6(L)}$ + _____ $O_{2(g)}$ →

_____ b. _____ $Zn_{(s)}$ + _____ $H_3PO_{4(aq)}$ →

_____ c. _____ $PB(ClO_3)_{2(aq)}$ + _____ $KI_{(aq)}$ →

_____ d. _____ $BaCO_{3(s)}$ →

_____ e. _____ $Br_{2(aq)}$ + _____$FeI_{3(aq)}$ →

_____ f. _____ $BaCl_{2(aq)}$ + _____ $H_2SO_{4(aq)}$ →

_____ g. _____ C_4H_{10} + _____ $O_{2(g)}$ →

_____ h. _____ $Ca_{(s)}$ + _____ $O_{2(g)}$ →

_____ i. _____ $HgO_{(s)}$ →

_____ j. _____ $Li_2SO_{4(aq)}$ + _____ $BaCl_{2(aq)}$ →

_____ k. _____ $F_{2(g)}$ + _____ $KCl_{(aq)}$ →

_____ l. _____ $HC_2H_3O_{2(aq)}$ + _____$LiOH_{(aq)}$ →

_____ m. _____ $Ni_{(s)}$ + _____ $FeSO_{4(aq)}$ →

_____ n. _____ $Sr_{(s)}$ + _____$N_{2(g)}$ →

_____ o. _____ $C_4H_{8(L)}$ + _____ $O_{2(g)}$ →

_____ p. _____ $CoBr_{2(s)}$ →

_____ q. _____ $H_3PO_{4(aq)}$ + _____$Al(OH)_{3(aq)}$ →

_____ r. _____ $CH_3OH_{(L)}$ + _____ $O_{2(g)}$ →

_____ s. _____ $Br_{2(aq)}$ + _____ $CuI_{2(aq)}$ →

_____ t. _____ $Pb(NO_3)_{2(aq)}$ + _____ $NaCl$ →

_____ u. _____ $Zn_{(s)}$ + _____ $CuCl_{2(aq)}$ →

_____ v. _____ $AlCl_{3(aq)}$ + _____ $Pb(NO_3)_{2(aq)}$ →