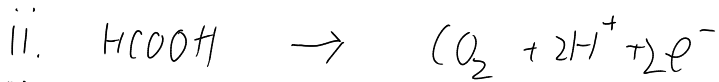
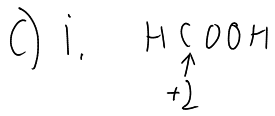
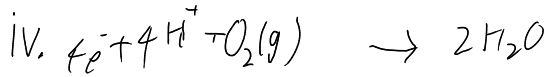


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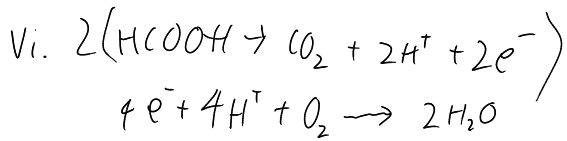


iii. minus pol

Elektronerna vandrar ifrån minuspolen till pluspolen



v. H^+ vandrar ifrån anoden till katoden



vii. $E_{\text{cell}} = E_{\text{red}}(\text{katod}) - E_{\text{red}}(\text{anod})$

$$E_{\text{cell}} = 1.23 - 0.2$$

$$E_{\text{cell}} = 1.03 \text{ V}$$

viii. $M(\text{HCOOH}) = 1 + 12 + 2 \cdot 16 + 1 = 46 \text{ g/mol}$

$$n = \frac{m}{M}$$

$$\frac{0.2}{46} = 0.00217 \text{ mol}$$

$$n(\text{e}^-) = 2 \cdot n(\text{HCOOH})$$

$$n(\text{e}^-) = 0.00434 \text{ mol}$$

$$9.65 \cdot 10^4 \cdot 0.00434 = 418.87 \text{ C} \quad 1 \text{ h} = 3600 \text{ s}$$

$$I = \frac{Q}{t}$$

$$I = \frac{418.87}{3600} \approx 0.11635 \text{ A} = 116.35 \text{ mA}$$