Molecule	Lewis	electron	molecular	bond angle	Polar or not?
	structure	domain	geometry		
		Geometry			
Example: CH ₄	H- C-H	Tetrahedral	tetrahedral	109.5	not
CCl ₄	10 - c - cil	tetrohedral	tetrate drad	(%,5.	not
H ₂ O	4,0°4	testanhedral	V-shape	l=4.ŝ	polar
NH ₃	HAH	tetrahedral	Arigonal pyramodral	107'	polar
CO ₂					
	5= C= 5	linear	linear	18.	han polon
SO ₂					
	0=2	trigonal planar	V-shape	119°	polan
HCN	H-C =N	linear	linear	1%.	Nonpolar
N ₂	(N = N)	1.neer		<i>J</i> .	Nonpolan
OH-	[10-H]	lihear			
SO42-	10-10-01	tetra hudrav	l tetraheurol	195	

CO ₃ 2-	[0 c = 0]	trigonal planer	trigonal planar	7 .°	
BeCl ₂	CI-Be-Ell	lihean	linear	1841	honpolan
BF₃ 3+ V	市岛王	trigonall planar	trigonal planar	lσ.	nonpolan
СО	(C=01	linear			polom
HCOO- (2+4+)+) =)	[H-C- <u>0</u>]	trigonal	trigonal planar	12.	

Predict $\underline{\text{electron domain Geometry}}$, $\underline{\text{molecular geometry}}$ and $\underline{\text{bond angle}}$.

2) N_2H_2 $IV/_{L=6}$ $H-\bar{N}=\bar{N}-H$
ED: trigonal planar molecular: V-shape
4) NO2- Seld No = 18 9 [0 = N - 0,] - ED: trigoned planar holecular: N-shape
4

5) H ₂ S 6-1=8	6) CCl ₄
H-SH ED: tetrihedral moleculan: V-shape 104.5°	(CI) ED tetrahedral (CI-C) EII molecular tetrahedral (CII 10) is
7) 03 ED: trigonal planar planar molecular V-shape	8) NO+ S+6-1=10 [IN=0] +
9) PO43- EP: tetrahedral [] 3- halecular: tetah edral [] 0/]	10) PO3 3- 13 [19