

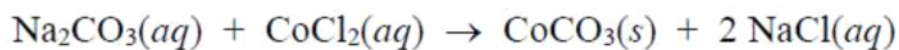
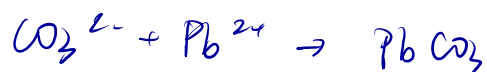
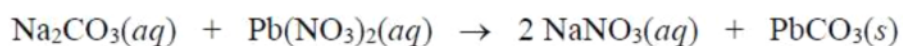
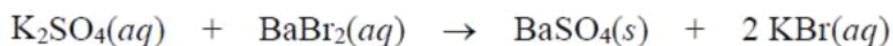
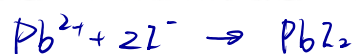
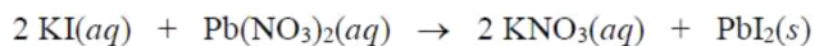
# Printout

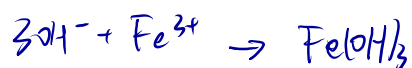
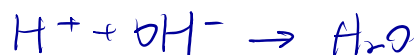
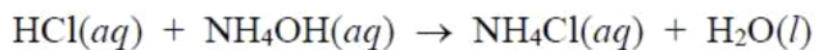
2023年3月26日 13:39

### Common rules for solubility of salts

Ions	Solubility of salts	Solubility Exceptions
sodium (Na), potassium (K) and ammonium ( $\text{NH}_4^+$ )	all soluble	none
nitrates ( $\text{NO}_3^-$ )	all soluble	none
chlorides ( $\text{Cl}^-$ ) and iodides ( $\text{I}^-$ )	most soluble	silver ( $\text{Ag}^+$ ), lead ( $\text{Pb}^{2+}$ ), mercury ( $\text{Hg}_2^{2+}$ )
sulfates ( $\text{SO}_4^{2-}$ )	most soluble	$\text{Ag}^+$ , $\text{Pb}^{2+}$ , calcium $\text{Ca}^{2+}$ , strontium ( $\text{Sr}^{2+}$ ) and barium ( $\text{Ba}^{2+}$ )
carbonates ( $\text{CO}_3^{2-}$ )	most insoluble	Group 1A, $\text{NH}_4^+$ soluble
hydroxide ( $\text{OH}^-$ )	most insoluble	Group 1A, $\text{NH}_4^+$ soluble

### Practice:





**Write down the ionic equations of the following reactions:**

1. Zn with sulfuric acid



2. Copper (II) oxide with hydrochloric acid



3. Sodium with water



4. Sodium oxide with water



5. Calcium carbonate with sulfuric acid

