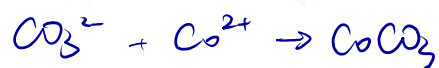
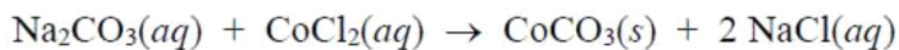
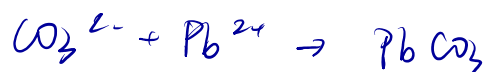
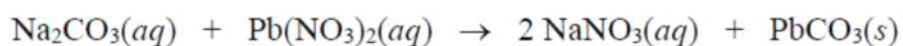
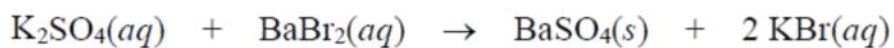
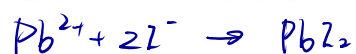
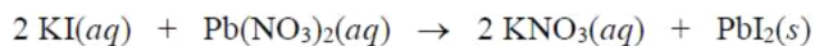
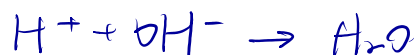
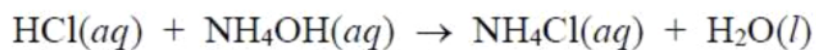


Common rules for solubility of salts

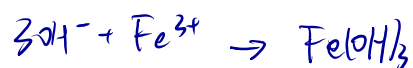
Ions	Solubility of salts	Solubility Exceptions
sodium (Na), potassium (K) and ammonium (NH_4^+)	all soluble	none
nitrates (NO_3^-)	all soluble	none
chlorides (Cl^-) and iodides (I^-)	most soluble	silver (Ag^+), lead (Pb^{2+}), mercury (Hg_2^{2+})
sulfates (SO_4^{2-})	most soluble	Ag^+ , Pb^{2+} , calcium Ca^{2+} , strontium (Sr^{2+}) and barium (Ba^{2+})
carbonates (CO_3^{2-})	most insoluble	Group 1A, NH_4^+ soluble
hydroxide (OH^-)	most insoluble	Group 1A, NH_4^+ soluble

Practice:





No. Ionic Equation



Write down the ionic equations of the following reactions:

1. Zn with sulfuric acid



2. Copper (II) oxide with hydrochloric acid



3. Sodium with water



4. Sodium oxide with water



5. Calcium carbonate with sulfuric acid

