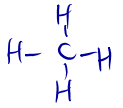
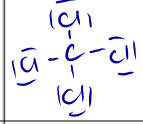
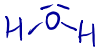
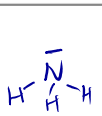
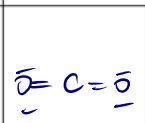
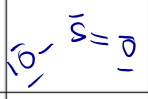
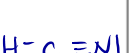
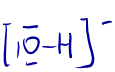
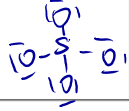
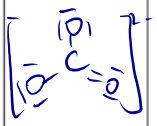
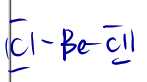
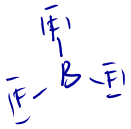
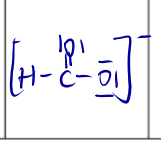
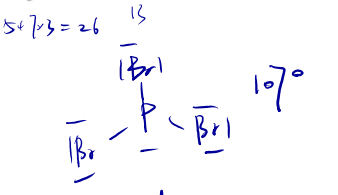
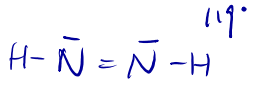
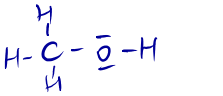
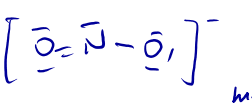
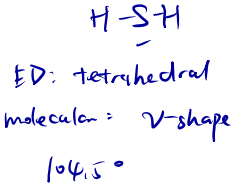
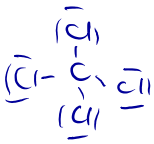
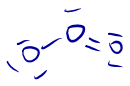
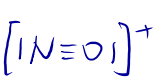
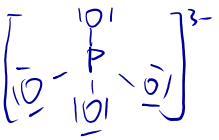
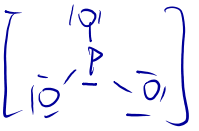
Molecule	Lewis structure	electron domain Geometry	molecular geometry	bond angle	Polar or not?
Example: CH_4		tetrahedral	tetrahedral	109.5°	not
CCl_4		tetrahedral	tetrahedral	109.5°	not
H_2O		tetrahedral	V-shape	104.5	polar
NH_3		tetrahedral	trigonal pyramidal	107°	polar
CO_2		linear	linear	180°	nonpolar
SO_2		trigonal planar	V-shape	119°	polar
HCN		linear	linear	180°	Nonpolar
N_2		linear	linear	180°	nonpolar
OH^-		linear	linear	180°	
SO_4^{2-} $32/2=16$		tetrahedral	tetrahedral	109.5°	

CO_3^{2-}		trigonal planar	trigonal planar	120°	
$BeCl_2$		linear	linear	180°	nonpolar
BF_3 3+21		trigonal planar	trigonal planar	120°	nonpolar
CO		linear	linear	180°	polar
$HCOO^-$ 12+4+1+1 = 9		trigonal planar	trigonal planar	120°	polar

Predict electron domain Geometry, molecular geometry and bond angle.

1) PBr_3  ED: tetrahedral molecular: trigonal pyramidal	2) N_2H_2  ED: trigonal planar molecular: V-shape
3) CH_3OH  ED: tetrahedral molecular: tetrahedral	4) NO_2^-  ED: trigonal planar molecular: V-shape

5) H_2S 6+2=8  ED: tetrahedral molecular: V-shape 104.5°	6) CCl_4  ED: tetrahedral molecular: tetrahedral 109.5°
7) O_3  ED: trigonal planar molecular: V-shape 119°	8) NO^+ 5+6−1=10  linear 180°
9) PO_4^{3-} ED: tetrahedral molecular: tetrahedral 109.5° 	10) PO_3^{3-} 5+18+3=26, 13  ED: tetrahedral molecular: trigonal pyramidal 107°