Molecule	Lewis	electron	molecular	bond angle	Polar or not?
	structure	domain	geometry		
		Geometry			
Example: CH ₄	H- C-H H	Tetrahedral	tetrahedral	109.5	not
CCl ₄	(Q1) (Q - C- (Q1)	tetrohedral	tetrahedrod	(28,5.	not
H ₂ O	H/0-H	testahedral	V-shape	l=4.ĝ	polar
NH ₃	HAH	-tetrahedral	trigonal pyramidrul	107'	polar
CO ₂					
	5= C= 5	linear	linear	18.	hon polon
SO ₂					
	0=2	trigonal planar	V-shape	119°	polar
HCN	H-C =N	linear	linear	130.	Nonpolar
N ₂					
	(N =N)	1.neer	lihear	(6-3)	nonpolan
OH-	[10-H]	lihear	linear	1%	
SO42-	10-4-01	tetra hudrad	l tetrahedral	19.5	

CO ₃ ²⁻	[[p] [p] [] [p] [p] [p] [p] [p] [p] [p	trigonal planer	trigonal planar	1200	
BeCl ₂	cl-Be-cll	likean	lihear	1841	honpolen
BF ₃	市岛主	trigonal planar	trigonal planor	ls• .	nonpolar
СО	(C=01	line	linear	182.	polem
HCOO- (2+4+)+) =)	[H-C- <u>0</u>]	trigonal planar	trigonal planar	احر,	polar

Predict <u>electron domain Geometry</u>, <u>molecular geometry</u> and <u>bond angle</u>.

1) PBr3 5+7-3=26 Br1 Br1 ED: tetrahedral molecules: trigonal pyramidnal	2) N2H2 11/2-6 H-N=N-H ED: trysnel planar molecular: V-shape
3)CH3OH H-C-O-H H Ep: tetrahedral m(ecular: tetra hedral	4) NO2- S+(+ 12 = 18 9 [0= 17 - 0,] - ED: trigoned planar molecular: N-shape

5) H ₂ S 6-12=8	6) CCl ₄
H-SH ED: tetrahedral molecular: V-shape 104,5°	(CI) ED: tetrahedral (CI-C, CI) molecular: tetrahedra (CII 10):5.
7) O3 ED: trigonal planar molecular: V-shape	8) NO+ S+6-1=10 [IN=0] + linear [180.
9) PO43- Ep: tetrohedral [0] 3- molecular: tetah edral [0] 0)	10) PO33 13 [9]