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Roll no →

B10 - 211

Assignment - 1

Solution 1

An e^- microscope magnify a cell ~~as~~ 10,000 folds (10^4)
So when viewing a eukaryotic cell i.e $50 \mu\text{m}$ diameter
magnified cells have diameter $50 \times 10^4 \mu\text{m}$

$$= 500 \text{ mm}$$

$$= 19.685 \text{ inches}$$

Solution 2.

~~Area of sphere~~

$$\text{Volume of sphere} = \frac{4}{3} \pi r^3 = \frac{4}{3} \pi \left(\frac{d}{2}\right)^3$$

$$\text{Diameter of muscle cell} = 50 \mu\text{m}$$

$$\text{radius of muscle cell} = 25 \mu\text{m}$$

$$\text{Vol. of muscle cells} = \frac{4}{3} \pi (25 \times 10^{-6})^3$$

$$= 20.9 \times 10^3 \times 10^{-18} \pi$$

$$= 20.9 \times 10^{-15} \pi \text{ m}^3$$

$$\text{Diameter of Actin molecule} = 3.6 \text{ nm}$$

$$\text{radius of Actin molecule} = 1.8 \text{ nm}$$

$$\begin{aligned}\text{Volume of Actin molecule} &= \frac{4}{3} \pi (1.8 \times 10^{-9})^3 \\ &= \frac{4}{3} \times 1.8^3 \times 10^{-27} \pi \\ &= 7.776 \times 10^{-27} \pi \text{ m}^3\end{aligned}$$

$$\begin{aligned}\text{No. of molecules muscle cell can hold} &= \frac{\text{vol. of muscle cell}}{\text{vol. of Actin molecule}} \\ &= \frac{20.9 \times 10^{-15} \pi \text{ m}^3}{7.776 \times 10^{-27} \pi \text{ m}^3}\end{aligned}$$

$$\approx 2.68 \times 10^{12} \text{ molecules}$$

2.68×10^{12} (approx.) molecules can be held by muscle cell in absence of any other organelle

Solution 3

$$\text{mass} = 3.0 \times 10^9 \text{ Da}$$

$$\text{M.W of nucleotide pair} = 660 \text{ Da}$$

$$\text{length of DNA molecule} = 0.34 \text{ nm}$$

(i)

~~No. of pairs~~

$$\text{No. of nucleotide pairs} = \frac{3.0 \times 10^9}{660} = 4.5 \times 10^6 \text{ bp}$$

$$\begin{aligned}\text{length of E. coli DNA molecule} &= 4.5 \times 10^6 \times 0.34 \times 10^{-9} \\ &= 1.53 \times 10^{-3} \text{ m}\end{aligned}$$

(i) Dimensions of typical E. coli cell
→ Rod-shaped

About $2\mu\text{m}$ long having $0.25-1.0\mu\text{m}$ in diameter

molecular weight = $3.1 \times 10^9 \text{ g/mol}$

M. w of single pair = $0.66 \times 10^3 \text{ g/mol}$

$$\begin{aligned}\text{Amount of nucleotide pairs in DNA} &= \frac{\text{M. w of DNA}}{\text{M. w of single pair}} \\ &= \frac{3.1 \times 10^9 \text{ g/mol}}{0.66 \times 10^3 \text{ g/mol}}\end{aligned}$$

$$\text{length per pair} = 0.34 \text{ nm/pair} \quad = 4.7 \times 10^6 \text{ bp}$$

$$= 4.7 \times 10^6 \times 0.34 \text{ nm}$$

$$= 1.6 \text{ mm}$$

$$\approx 2 \text{ mm}$$

(iii)

$$\frac{1}{3} \text{ cod are codons} = \frac{4.7 \times 10^6}{3} = 1.6 \times 10^6 \text{ codons}$$

avg. proteins in E. coli made of 360 AA

$$\text{no. of proteins} = \frac{1.6 \times 10^6 \times (1 \text{ AA/codon})}{360 \text{ AA/protein}}$$

$$= 4444.44 \text{ proteins}$$

Solution 4

yeast cell spherical nucleus diameter = $2 \mu\text{m}$

nucleus house contains 1.2×10^7 bp of genome

8 histones make cylindrical core where 150 bp wrapped

50 bp space exist b/w 2 nucleosomes

let us take the no. of nucleosomes to be n
the eqⁿ we get is,

$$150n + (n-1)50 = 1.2 \times 10^7$$

$$150n + 50n - 50 = 1.2 \times 10^7$$

$$200n - \underbrace{50}_{\text{negligible}} = 1.2 \times 10^7$$

$$n = \frac{1.2 \times 10^7}{200} \approx 6 \times 10^4 \text{ nucleosomes}$$

we get 6×10^4 no. of nucleosomes

Solution 5

Yeast cell enclosed by lipid bilayer

$$SA = 80 \times 10^{-6} \text{ m}^2$$

$\frac{1}{2}$ av. covered by proteins so $40 \times 10^{-6} \text{ m}^2$

new SA left = $40 \times 10^{-6} \text{ m}^2$

$$\text{No. of lipid molecules we get} = \frac{2 \times 40 \times 10^{-6} \text{ m}^2}{0.25 \times 10^{-9} \text{ m}^2}$$

$$= \frac{80 \times 10^{-6} \text{ m}^2}{25 \times 10^{-11}}$$

$$= 3.2 \times 10^5 \text{ m}^2$$

Solution-6 Concentratⁿ of glucose in E. coli = 1 mM

$$(i) M = \frac{n}{V} = 1 \text{ mM} = \frac{1 \times 10^{-3}}{1 \text{ L}} \times \frac{180}{1}$$

$$= 0.18 \text{ g/L}$$

$$= 0.18 \text{ mg/mL}$$

(ii) For no. of glucose molecules, as we know

$$M = \frac{n}{V}$$

~~$n = M \times V$~~ ~~for molecules~~

~~$n = 1 \times 180$~~

~~$n = 1 \times 10^{-3} \text{ mole}$~~

~~no. of molecules = $1 \times (6.023 \times 10^{23})$~~

$$n = M \times V$$

$$= 1 \times 10^{-15} \times 10^{-3}$$

$$n = 10^{-18} \text{ moles}$$

$$\begin{aligned} \text{no. of molecules} &= 10^{-18} \times (6.023 \times 10^{23}) \\ &= 6.023 \times 10^5 \text{ molecules} \end{aligned}$$

Solution 7 -

225 μ M of barium cycloheximide

$$M = \frac{n}{V} = 225 \mu\text{M} = \frac{225 \times 10^{-6}}{1\text{L}} \times 12327$$

$$= 2.773 \text{ g/L}$$

$$= 2.773 \text{ mg/mL}$$

$$= 2.773 \mu\text{g}/\mu\text{L}$$