

Indium

Reaction	Baes and Mesmer, 1976	NIST46	Brown and Ekberg, 2016
$\text{In}^{3+} + \text{H}_2\text{O} = \text{In}(\text{OH})^{2+} + \text{H}^+$	-4.00	-3.927	-3.96
$\text{In}^{3+} + 2 \text{H}_2\text{O} = \text{In}(\text{OH})_2^+ + 2 \text{H}^+$	-7.82	-7.794	-9.16
$\text{In}^{3+} + 3 \text{H}_2\text{O} = \text{In}(\text{OH})_3 + 3 \text{H}^+$	-12.4	-12.391	
$\text{In}^{3+} + 4 \text{H}_2\text{O} = \text{In}(\text{OH})_4^- + 4 \text{H}^+$	-22.07	-22.088	-22.05
$\text{In}(\text{OH})_3(\text{s}) = \text{In}^{3+} + 3 \text{OH}^-$	-36.92	-36.9	-36.92
$1/2 \text{In}_2\text{O}_3(\text{s}) + 3/2 \text{H}_2\text{O} = \text{In}^{3+} + 3 \text{OH}^-$	-35.3		-35.24

C.F. Baes and R.E. Mesmer, The Hydrolysis of Cations. Wiley, New York, 1976.

P.L. Brown and C. Ekberg, Hydrolysis of Metal Ions. Wiley, 2016, pp. 135-145.

NIST46, NIST Critically Selected Stability Constants of Metal Complexes: Version 8.0. Available at: www.nist.gov/srd/nist46