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## Iron(II)

Reaction	Baes and Mesmer, 1976	Nordstrom et al., 1990	Hummel et al., 2002	Lemire et al., 2013	Brown and Ekberg, 2016
$Fe^{2+} + H_2O \rightleftharpoons FeOH^+ + H^+$	$-9.3 \pm 0.5$	-9.5	-9.5	$-9.1 \pm 0.4$	$-9.43 \pm 0.10$
$Fe^{2+} + 2 H_2O \rightleftharpoons Fe(OH)_2 + 2 H^+$	$-20.5 \pm 1.0$				$-20.52 \pm 0.08$
$Fe^{2+} + 3 H_2O \rightleftharpoons Fe(OH)_3^- + 3 H^+$	$-29.4 \pm 1.2$				$-32.68 \pm 0.15$
$Fe(OH)_2(s) +2 H^+ \rightleftharpoons Fe^{2+} + 2 H_2O$					$12.27 \pm 0.88$
$\frac{1}{3} \text{Fe}_3 \text{O}_{4 \text{ (s)}} + 2 \text{ H}^+ + \frac{1}{3} \text{H}_2 \rightleftharpoons \text{Fe}^{2+} + \frac{4}{3} \text{H}_2 \text{O}$				12.5	$11.77 \pm 0.22$

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