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Cadmium

Reaction	Baes and Mesmer, 1976	Powell et al., 2011	Brown and Ekberg, 2016
$Cd^{2+} + H_2O \rightleftharpoons CdOH^+ + H^+$	−10.08	–9.80 ± 0.10	−9.81 ± 0.10
$Cd^{2+} + 2\ H_2O \rightleftharpoons Cd(OH)_2 + 2\ H^+$	–20.35	–20.19 ± 0.13	−20.6 ± 0.4
$Cd^{2+} + 3\ H_2O \rightleftharpoons Cd(OH)_3^- + 3\ H^+$	<–33.3	–33.5 ± 0.5	−33.5 ± 0.5
$Cd^{2+} + 4\ H_2O \rightleftharpoons Cd(OH)_4^{2-} + 4\ H^+$	–47.35	–47.28 ± 0.15	−47.25 ± 0.15
$2\ Cd^{2+} + H_2O \rightleftharpoons Cd_2OH^{3+} + H^+$	–9.390	–8.73 ± 0.01	−8.74 ± 0.10
$4\ Cd^{2+} + 4\ H_2O \rightleftharpoons Cd_4(OH)_4^{4+} + H^+$	–32.85		

$Cd(OH)_2(s) \rightleftharpoons Cd^{2+} + 2\ OH^-$		–14.28 ± 0.12	
$Cd(OH)_2(s) + 2\ H^+ \rightleftharpoons Cd^{2+} + 2\ H_2O$	13.65	13.72 ± 0.12	13.71 ± 0.12

C.F. Baes and R.E. Mesmer, The Hydrolysis of Cations. Wiley, New York, 1976.

P.L. Brown and C. Ekberg, Hydrolysis of Metal Ions. Wiley, 2016, pp. 730–738.

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