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Tellurium(VI)

Tellurate structures in solution are best written as $TeO_2(OH)_4^{2-}$, $TeO(OH)_5^-$ and $Te(OH)_6$. Other notations can be found in the literature.

Reaction	Baes and Mesmer, 1976	Filella and May, 2019[a]
$TeO_2(OH)_4^{2-} + H^+ \rightleftharpoons TeO(OH)_5^-$		10.83
$TeO(OH)_5^- + H^+ \rightleftharpoons Te(OH)_6$	7.68	7.696
$TeO_2(OH)_4^{2-} + 2\ H^+ \rightleftharpoons Te(OH)_6$	18.68	
$TeO_3(OH)_3^{3-} + 3\ H^+ \rightleftharpoons Te(OH)_6$	34.3	
$2\ Te(OH)_6 \rightleftharpoons Te_2O(OH)_{11}^- + H^+$		–6.929

[a]The number of significant figures are retained to minimise propagation of round-off errors; they should not be taken to indicate the relative uncertainty of the values, which is always at least one order of magnitude less than indicated.

C.F. Baes and R.E. Mesmer, The Hydrolysis of Cations. Wiley, New York, 1976.

M. Filella and P.M. May, The aqueous chemistry of tellurium: critically-selected equilibrium constants for the low-molecular-weight inorganic species. Environ. Chem. 16, 289–295 (2019). doi:10.1071/EN19017