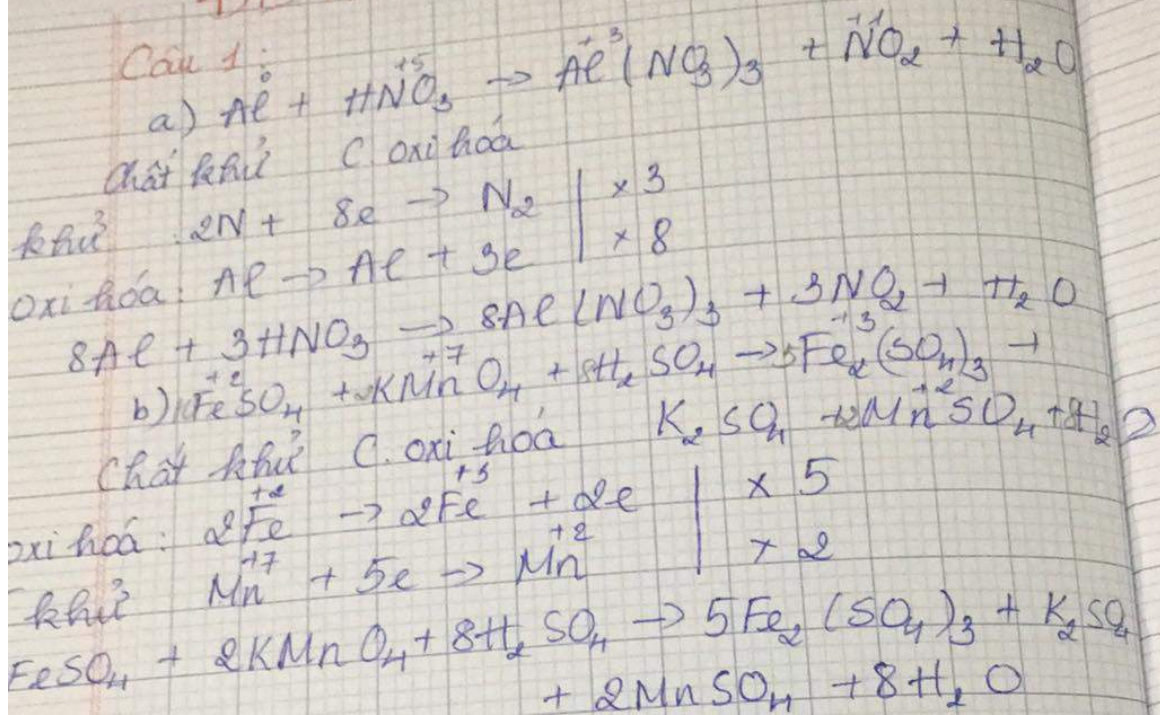
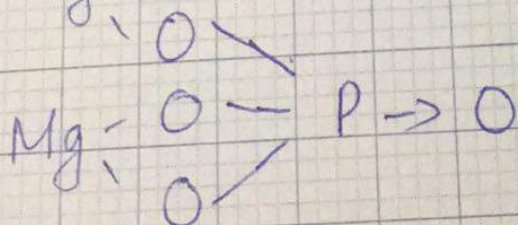
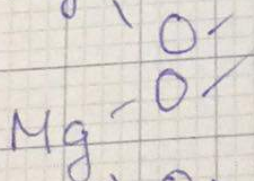
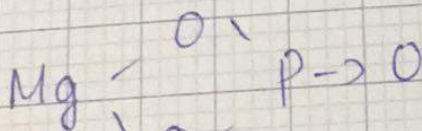
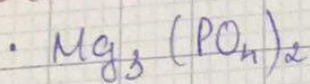


## ĐỀ 4

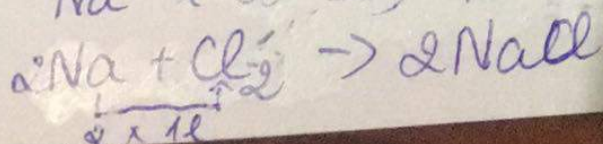
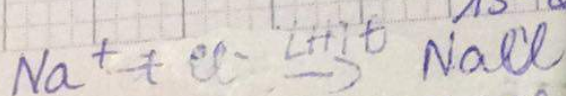
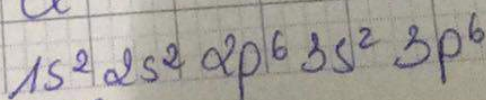
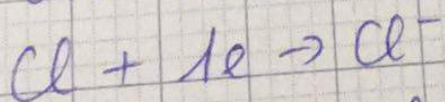
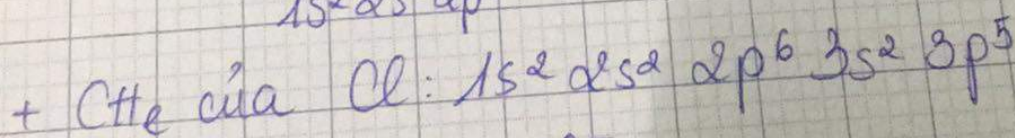
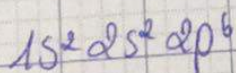
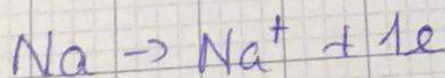
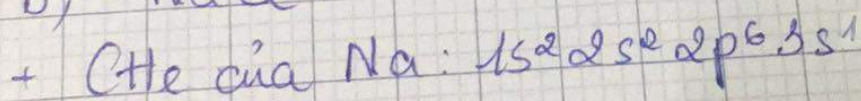
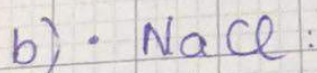
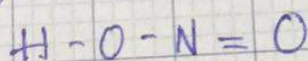
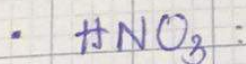
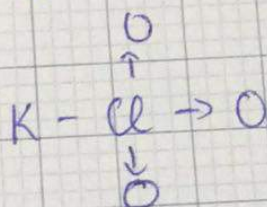
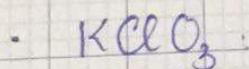
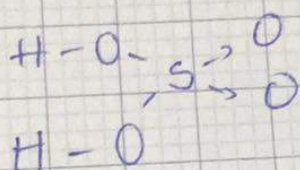
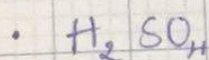
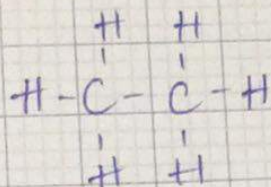
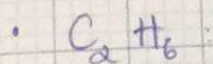
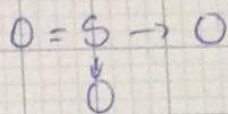
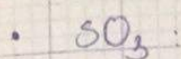
**Câu 1:**



**Câu 2:**



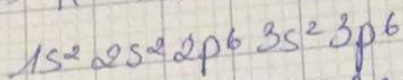
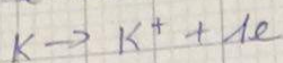




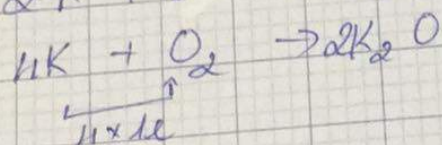
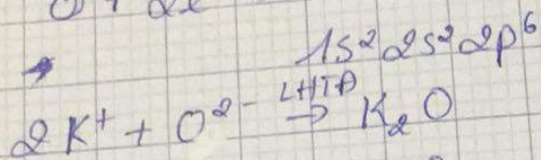
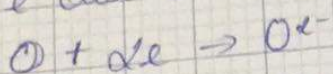


•  $K_2O$ :

+ CHe của K:  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$

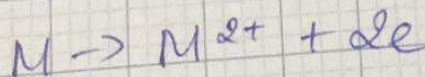
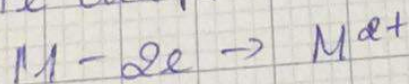


+ CHe của O:  $1s^2 2s^2 2p^4$



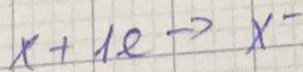
**Câu 3:**

a) • CHe của  $M^{2+}$ :  $1s^2 2s^2 2p^6$



CHe của M:  $1s^2 2s^2 2p^6 3s^2 \Rightarrow M \text{ là Mg}$

• CHe của  $X^-$ :  $1s^2 2s^2 2p^6 3s^2 3p^6$



CHe của X:  $1s^2 2s^2 2p^6 3s^2 3p^5$  ( $Z = 17$ )

$\Rightarrow X \text{ là Cl}$



b) • Vị trí của M:

+ O: 12 ( $Z = 12$ )

+ Chu kỳ: 3 (vì có 3 lớp e)

+ Nhóm IIA (vì có 2e hóa trị lớp ngoài)

• Vị trí của X:

+ O: 17 ( $Z = 17$ )

+ Chu kỳ: 3 (vì có 3 lớp e)

+ Nhóm VII A (vì có 7e hóa trị lớp ngoài)

c)  $\text{Fe} \rightarrow \text{Fe}^{3+} + 3e$

• CHe của Fe:  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$

• CHe của  $\text{Fe}^{3+}$ :  $1s^2 2s^2 2p^6 3s^2 3p^5$

**Câu 4:**

• Trong 1 nhóm A, theo chiều tăng của ĐTHN, tính kim loại tăng dần.

• Trong chu kỳ, theo chiều tăng của ĐTHN, tính kim loại giảm dần.

• Sắp xếp theo chiều giảm dần tính kim loại: K; Ca; Al; S; Cl; F

**Câu 5:**

• Hợp chất khí có dạng  $\text{RH} \Rightarrow \text{R}$  thuộc nhóm VIIA

$\Rightarrow$  Công thức oxit cao nhất:  $\text{R}_2\text{O}_7$

$$\% \text{R} = \frac{2 \cdot M_{\text{R}}}{M_{\text{R}} \cdot 2 + 7 \cdot 16} = 38,80\%$$

Một trang ở một chương lại



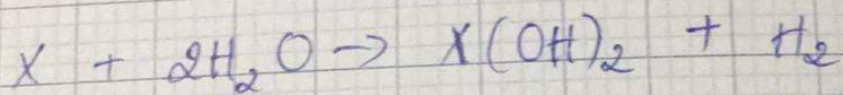
$$\Rightarrow M_R = 35,5\%$$

$\Rightarrow R$  là nguyên tố (lo (II))

Câu 6:

Gọi  $M$  là hỗn hợp hai kim loại

$$n_{H_2} = \frac{V}{22,4} = \frac{6,72}{22,4} = 0,3 \text{ mol}$$



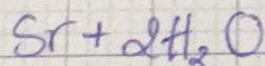
$$0,3 \leftarrow 0,6 \leftarrow 0,3 \leftarrow 0,3$$

$$M_x = \frac{36,2}{0,3} \approx 120,6$$

$$M_A < \bar{M}_x < M_B$$

$\Rightarrow A$  là Sr

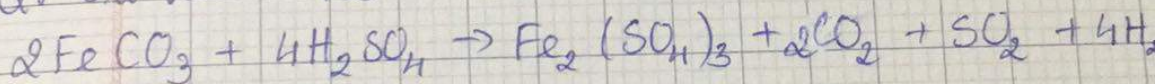
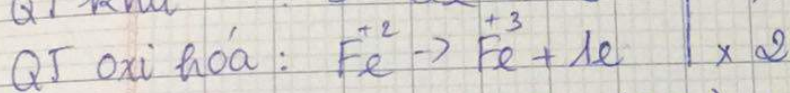
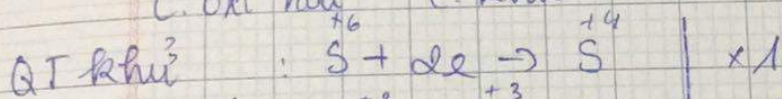
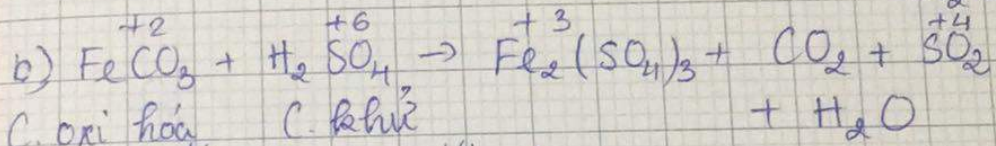
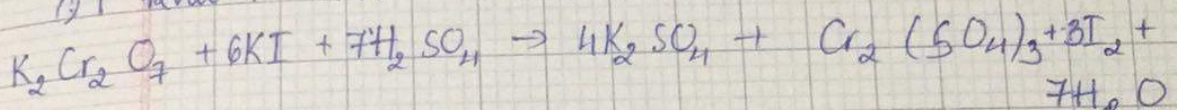
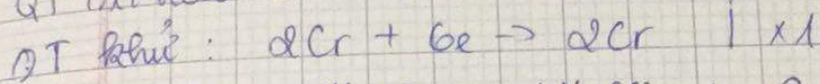
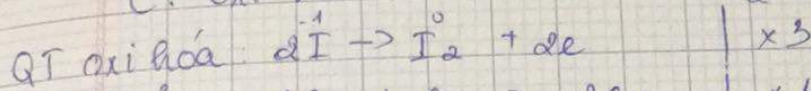
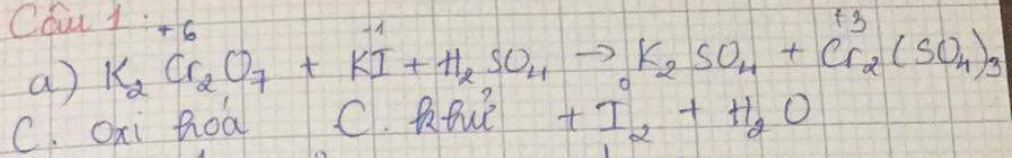
$\{ B$  là Ba



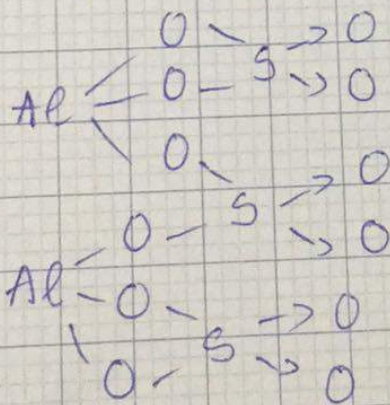
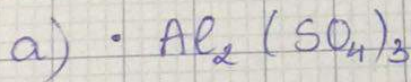


## ĐỀ 5

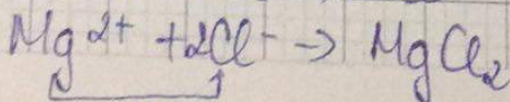
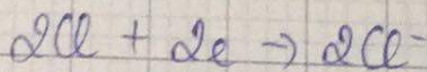
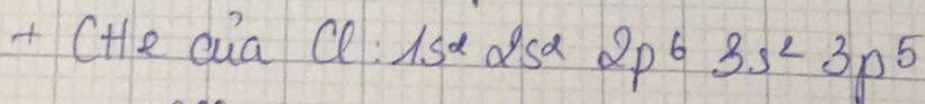
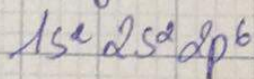
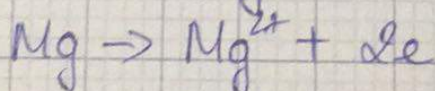
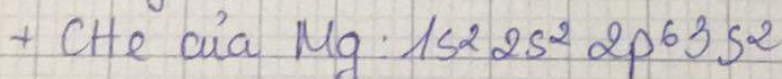
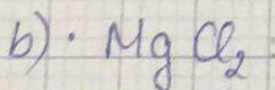
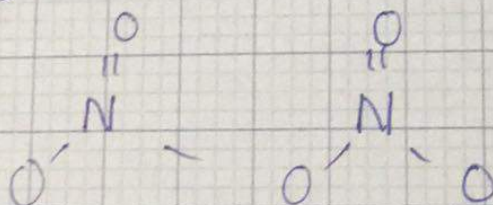
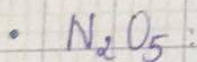
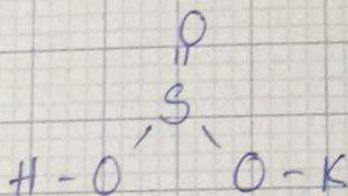
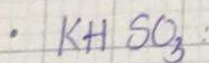
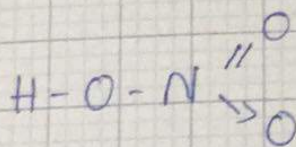
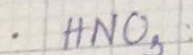
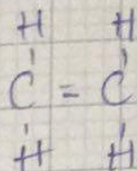
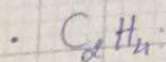
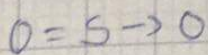
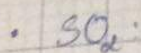
Câu 1:



Câu 2:





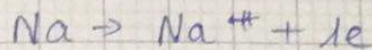


2. Một trang vở một tương lai



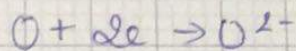
•  $\text{Na}_2\text{O}$ :

+ Cte của Na:  $1s^2 2s^2 2p^6 3s^1$

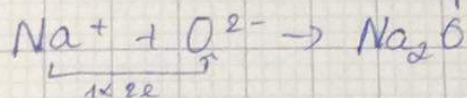


$1s^2 2s^2 2p^6$

+ Cte của O:  $1s^2 2s^2 2p^4$



$1s^2 2s^2 2p^6$



**Câu 3:**

a) • M:  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 \rightarrow$  Kali

• X:  $1s^2 2s^2 2p^4 \rightarrow$  Oxi

b) • Vị trí của M:

+ Ô: 19 ( $Z=19$ )

+ Chu kỳ: 4 (vì có 4 lớp e)

+ Nhóm IA (vì có 1e hóa trị lớp ngoài cùng)

• Vị trí của X:

+ Ô: 8 ( $Z=8$ )

+ Chu kỳ: 2 (vì có 2 lớp e)

+ Nhóm: VIA (vì có 6e hóa trị lớp ngoài cùng)

c)  $\text{Fe} \rightarrow \text{Fe}^{2+} + 2e$

+ Cte của Fe:  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$

+ Cte của  $\text{Fe}^{2+}$ :  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6$

Một trang ở một thương lái



### Câu 4:

- Trong 1 nhóm, theo chiều tăng của ĐTHH, tính phi kim giảm dần
- Trong chu kỳ, theo chiều tăng của ĐTHH, tính phi kim tăng dần
- Sắp xếp theo chiều giảm dần tính phi kim: F; N; Cl, P; Mg; K

### Câu 5:

- Hợp chất khí có dạng  $R_2O_5 \Rightarrow R$  thuộc nhóm VA
- $\Rightarrow$  Công thức oxit cao nhất:  $R_2O_5$

$$\%R = \frac{2M_R}{M_R \cdot 2 + 5 \cdot 16} = 8,82\%$$

$\Rightarrow M_R =$

### Câu 5:

- Hợp chất khí có dạng  $R_2O_5$
- $\Rightarrow R$  thuộc nhóm VA

- Hợp chất khí với  $H_2$  là:  $RH_3$

$$M_R = \frac{3}{M_R + 3} \cdot 100\% = 8,82\%$$

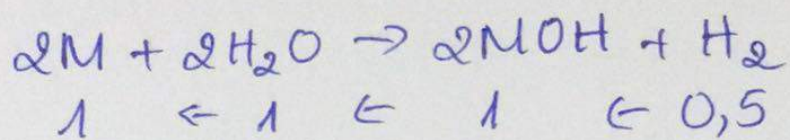
$\Rightarrow M_R = 31$

$\Rightarrow R$  là photpho



Câu 6: Gọi M là hỗn hợp 2 kim loại

$$n_{H_2} = \frac{11,2}{22,4} = 0,5 \text{ mol}$$



$$M_M = \frac{29,4}{1} = 29,4$$

$$M_1 < \bar{M} < M_2$$

$$\Rightarrow \begin{cases} M_1 = 23 (\text{Na}) \\ M_2 = 39 (\text{K}) \end{cases}$$