

An assessment of structural enthalpy and crystallisation pathways of $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$ bulk metallic glass and amorphous films

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ABSTRACT

The structural nature and thermal stability of amorphous alloys is highly dependent on the method by which they are produced, i.e. their relaxation rate upon cooling. Both bulk samples and metallic glass films of $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$ were produced by copper mold casting and direct current (DC) magnetron sputtering onto aluminium substrates, respectively. Comparisons between structural enthalpy, crystallisation pathways, relaxation and crystallisation kinetics of the bulk samples and films were examined by elevated temperature XRD and DSC. Compared with equivalent experiments on the bulk alloy, results for the thin films show distinct differences in structural enthalpy and deviations from the expected crystalline phase evolution, displaying minor peak shifts, failure of some phases to evolve, and variations in the evolution rates.

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1 INTRODUCTION

The structural nature and thermal stability of amorphous alloys is highly dependent on the method by which they are produced, i.e. their relaxation rate upon cooling. Both bulk samples and metallic glass films of $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$ were produced by copper mold casting and direct current (DC) magnetron sputtering onto aluminium substrates, respectively. Comparisons between structural enthalpy, crystallisation pathways, relaxation and crystallisation kinetics of the bulk samples and films were examined by elevated temperature XRD and DSC. Compared with equivalent experiments on the bulk alloy, results for the thin films show distinct differences in structural enthalpy and deviations from the expected crystalline phase evolution, displaying minor peak shifts, failure of some phases to evolve, and variations in the evolution rates.

2 METHOD

2.1 Master alloy

The master alloy of $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$ was produced using high-purity elements of Mg (99.85 wt%), Zn (99.995 wt%), and Ca (99.8 wt%). The alloy was prepared with an in-house induction melting furnace within boron nitride coated graphite crucibles, purged with Ar (99.997 vol.% purity) five times, and protected with a circulating Ar atmosphere. Alloy homogeneity was ensured by heating and cooling through a cycle of 700°C, 385°C, 650°C, 385°C, 650°C to a casting temperature of 500 °C and 450°C for injection and gravity casting respectively. Bulk amorphous $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$ rods of 2.5mm diameter and plates of thickness of XX μm were produced by copper mold injection casting. The 25.4mm diameter targets were prepared from a cylindrical copper mold gravity castings sectioned to thicknesses of 3.25mm. All samples and targets were stored under Ar when not being examined or used.

measure
thickness

2.2 DC magnetron sputtering

Films were produced from an in-house DC magnetron sputtering facility with Ar working gas (99.997 vol.% purity). The power was 15W, typical voltage of 290 – 350V, nominal chamber pressure of 1 bar, substrate temperature of 25°C, and Ar flow of 3.01 SCCM. Films were deposited directly onto to Al DSC lid substrates. Depositions were for a period of 35 minutes. Deposition rate was estimated at 1.2nm/s.

2.3 Stylus profiler analysis

Nominal film thickness was measured by a stylus profiler (Dektak 2A, Bruker, Germany). A glass slide was placed under the substrates within the sputtering chamber, allowing the substrates to act as a mask. Profile measurements were taken by measuring the height difference between the bare glass and the film coated glass. This film thickness was used to estimate the sputter deposition rate.

2.4 EDS analysis

Alloy composition and homogeneity were confirmed by SEM-EDS (S3400, Hitachi, Japan). Hyper-maps were collected with an accelerating voltage of 15 – 20 kV, and a probe current of 50 μ A. (Conditions; counts were 5000 kps or better, dead time was less than 20 %, and working distance was 10 mm).

2.5 DSC characterization

Isochronic DSC (204 F1 Phoenix, Netzsch, Selb, Germany) was carried out in Al crucibles under a protective Ar atmosphere (99.997 vol.% purity). Scans were performed at heating rates (β s) of 5 to 100 K/min.

Isothermal relaxation DSC was performed by heating samples at 20 K/min to the desired annealing temperature, holding for desired time, and Ar quenching to room temperature.

For annealed XRD the samples were heat treated in the DSC by heating to the desired temperature at 20 K/min followed by Ar quenching to room temperature.

2.6 XRD characterization

Annealing XRD (Empyrean, PANalytical, Cu K_{α} X-ray source, $\lambda = 1.541 \text{ \AA}$) was performed at room temperature. (Generator Voltage 45, Tube Current 40, Scan Step Size 0.0262606, Time per Step 397.29).

Dynamic XRD (D8, Bruker, Cu K_{α} X-ray source, $\lambda = 1.541 \text{ \AA}$) was performed by raising temperature at a rate of 20 K/min and performing scans *in situ*. The first scan was performed at 35°C, then 75°C, after which temperature was raised in 5 K increments until reaching the peak temperature at 185°C. The 2θ scans from 31 – 60° were completed within 1092 sec.

(18min, 12sec) to minimise the effects of recrystallisation during the experiment. (Generator Voltage 45, Tube Current 100, Scan Step Size 0.02, Time per Step 134.4).

3 RESULTS

3.1 Alloy composition

From the 35 minute depositions a nominal film thickness of $2.5\mu m$ was obtained, giving a deposition rate of approximately $1.2nm/s$. The temperature within the chamber was found to rise $3 - 4^{\circ}C$, significantly less than the expected $20K$ suggested by similar setups [1].

EDS analysis shows good agreement in the nominal composition for both the bulk and film $Mg_{65}Zn_{30}Ca_5$, see Table 1.

EDS Analysis	Bulk (at%)	Film (at%)
Mg	64.85 ± 3.18	62.92 ± 3.24
Zn	29.55 ± 0.82	31.17 ± 0.95
Ca	5.60 ± 0.17	5.91 ± 0.19

Table 1: EDS composition of bulk and film $Mg_{65}Zn_{30}Ca_5$ in atomic weight percent.

3.2 DSC

3.2.1 Isochronic DSC

Isochronic DSC was performed on the bulk and film $Mg_{65}Zn_{30}Ca_5$ to examine the thermal properties. The bulk alloy was relaxed at $120^{\circ}C$ for 10 minutes before DSC measurements to ensure the T_g was clearly visible. The film was not relaxed as unlike the bulk the loss in free volume from relaxation would be significant and make differences between the samples much more difficult to observe [source needed???].

The bulk $Mg_{65}Zn_{30}Ca_5$ was examined at heating rates (β s) of 5, 10, 15, 20, 30, 40, 60, 80, and $100 K/min$ to observe changes in the T_g and the T_x s with β . As expected greater β resulted in greater signal strength, exothermic peaks shifting to higher start temperatures, and an increase in thermal lag resulting in later exothermic finish temperatures and curve convolution. With this convolution the T_g and T_{x1} remained clearly visible for all β s, but $T_{x2,4,5}$ were only visible at low β s, and T_{x3} was not clear at any β , see Figure 1.

The film was examined at β s of 15, 20, 30, 40, 60, 80, and 100 K/min. The lower β s of 5 and 10 K/min were not utilised owing to the lower film signal compared to the bulk. The reduced signal was likely from the low mass of the film, about $\frac{1}{10}$ that of the bulk. The film showed the expected variable relationships with increasing β as observed in the bulk. The signal intensity increased at a compatible rate to bulk up until β s of 80 and 100 K/min. These final two β s showed great increases in the signal intensity. The exothermic peaks all convoluted together making many of the thermodynamic events difficult to observe. It also appeared that all exothermic events shifted to lower temperatures as compared to the bulk. The T_g and T_{x1} s were less defined than for the bulk, but could still be identified for all β s. For all β s the T_{x2-5} onsets could not be easily identified, see Figure 2. fig:DSC_vHeatingRate_Film

3.2.2 Fragility

Using the isochronic DSC data the fragility (m) of the $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$ system could be established for both the bulk and film. Numerical solutions were used to fit the DSC variant of the Vogel–Fulcher–Tammann (VFT) relationship for β [2]. Busch1998

$$\beta^{-1} = \tau_0 e^{\left(\frac{D^* T_0}{T_g - T_0}\right)} \quad (1) \quad \text{equ:VFT}$$

Where τ_0 is a pre-exponential factor, D^* is the liquid fragility parameter, and T_0 is the VFT temperature where the barrier to flow becomes infinite.

The m could then be calculated from Equation 2 [3, 4]. edding2002, Wei2014

$$D^* = 590 / (m - 16) \quad (2) \quad \text{equ:dSta}$$

Using these two equations for the bulk it was found $\beta^{-1} = 1.338E - 16 e^{5274\left(\frac{1}{T-T_0}\right)}$ with an Adj. $R^2 = 0.972$. This gave a $D^* = 20.4$, and a $m = 44.9$. The film was fitted to $\beta^{-1} = 5.921E - 11 e^{2766\left(\frac{1}{T-T_0}\right)}$ with a lower confidence of Adj. $R^2 = 0.861$, likely owing to the reduced number of data points. This gave a $D^* = 10.0$, and $m = 75.0$, see Figure 3. fig:Fragility_BulkFilm_mValue

3.3 DSC deconvolution

3.3.1 Onset determination

Numerical solutions were used to deconvolute the isochronic DSC data so the various T_x onsets could be accurately determined. This numerical fitting utilised a summation of skewed Gaussian curves to fit a target curve corresponding to the original data; as is a common method [5–9]. This fitting summation takes the form of Equation 3.

$$f(x) = \sum_{n=i}^n h_i e^{-\left(\frac{(x - T_i)^2}{(2c_j)^2}\right)} \quad (3)$$

Where h is the enthalpy peak intensity, T is the temperature at the enthalpy peak centre, and c is the Gaussian RMS width.

The final converged solutions of this fitting for both the bulk and film are shown in Figures 4 and 5 respectively. These results are tabulated in Table 2 for the bulk and Table 3 for the film. Note T_g and T_{x1} are obtained from the original raw data, not the deconvolution.

It is worth noting the deconvolution fitted 5 crystallisation events for the bulk $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$, but only 3 events for the film. This occurred because the bulk had 5 well defined events, whereas the film was largely convoluted together. Thus unique solutions could not be obtained by the lesser T_{x2} and T_{x4} events of the film.

Heating Rate β K/min	T_g	T_{x1}	T_{x2}	T_{x3}	T_{x4}	T_{x5}
100	136.1	164.8	193.4	201.8	240.2	262.4
80	132.0	160.0	194.4	201.9	238.2	260.3
60	129.6	157.7	190.0	197.8	232.9	259.0
40	126.6	155.2	189.0	200.0	226.4	254.7
30	126.2	151.5	187.0	198.4	221.0	251.1
20	125.1	149.8	188.4	197.0	216.0	246.8
15	123.8	148.3	186.2	195.6	212.2	243.9
10	123.5	144.5	183.4	192.9	207.4	239.8
5	120.5	141.1	179.7	187.5	199.8	232.7

Table 2: Bulk $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$ alloy onset temperatures for the various DSC heating rates (β s). All temperatures are in $^{\circ}\text{C}$.

Heating Rate β K/min	T_g	T_{x1}	T_{x2}	T_{x3}	T_{x4}	T_{x5}
100	108.5	128.6		177.3		240.3
80	106.0	121.2		165.6		238.8
60	107.3	134.0		176.1		237.8
40	100.2	119.8		170.7		234.2
30	95.3	110.4		169.5		232.5
20	95.5	115.2		170.5		229.4
15	92.5	113.5		168.8		224.0

Table 3: Film $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$ alloy onset temperatures for the various DSC heating rates (β s). All temperatures are in $^{\circ}\text{C}$.

3.3.2 Reaction enthalpy

The deconvolution fits were integrated to find the area under each curves. This information provides the specific enthalpy (h) of the crystallisation of each phase. This is presented in Tables 4 and 5 for the bulk and film respectively. Figure 8 shows the T_x onsets and specific enthalpy (h) for both the bulk and film plotted together.

Heating Rate β K/min	$h_{T_{x1}}$ J/g	$h_{T_{x2}}$ J/g	$h_{T_{x3}}$ J/g	$h_{T_{x4}}$ J/g	$h_{T_{x5}}$ J/g
100	59.59	6.97	49.16	22.84	46.08
80	42.61	6.08	32.33	18.27	31.25
60	30.02	4.05	25.41	16.76	19.81
40	16.93	4.36	12.44	11.13	11.68
30	12.03	3.68	9.32	9.18	9.02
20	7.18	2.21	4.99	5.67	5.78
15	5.48	2.01	3.65	4.69	4.43
10	3.45	1.43	2.28	3.14	2.92
5	1.65	0.69	1.09	1.47	1.42

Table 4: Bulk $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$ alloy specific enthalpy (h) of crystallisation formation for T_{x1-5} for the various DSC heating rates (β s); h is in J/g.

3.3.3 Relaxation enthalpy

For next paper on relaxation / rejuvenation.

Heating Rate β K/min	$h_{T_{x1}}$ J/g	$h_{T_{x2}}$ J/g	$h_{T_{x3}}$ J/g	$h_{T_{x4}}$ J/g	$h_{T_{x5}}$ J/g
100	48.24		49.85		43.38
80	43.27		53.56		36.18
60	15.5		8.78		22.4
40	16.22		9.13		16.27
30	13.72		7.16		11.81
20	6.16		2.3		7.45
15	6.99		3.66		6.57

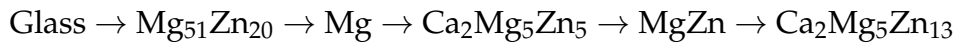
Table 5: Film $Mg_{65}Zn_{30}Ca_5$ alloy specific enthalpy (h) of crystallisation formation for T_{x1-5} for the various DSC heating rates (β s); h is in J/g .

3.4 XRD

3.4.1 Annealing XRD

The crystallisation events observed and deconvoluted from the DSC were further examined with annealing XRD. Both the bulk and films were heat treated by annealing to 120, 140, 145, 170, 200, 250, 290, and 320°C before XRD. This allowed for the location of the T_g and T_x s to be confirmed as well as for the crystallisation phases to be identify.

From these experiments 5 previously observed crystallisation phases of the $MgZnCa$ system [10–14] were characterised in the bulk and film $Mg_{65}Zn_{30}Ca_5$. This allowed the crystallisation process of $Mg_{65}Zn_{30}Ca_5$ from the fully amorphous glass to fully crystalline metal to be identified as;



The annealing XRD results are shown in Figures 9 and 10 for the bulk and film respectively. In these figures each phase is identified with a tracer at the temperature it was most strongly observed. Note the Al substrates peaks have been faceted in Figure 10 as to not dwarf the other peaks.

In Figures 9 and 10 it can be observed that $Mg_{51}Zn_{20}$ comes out at lower temperature in the film compared to the bulk, while the other phases nucleate and grow at similar rates for both the bulk and film. The temperature each phase is first and last observed are tabulated in Table 6.

3.4.2 Dynamic XRD

The annealing XRD was useful for identifying the crystal phases present, but not many difference between the evolution rates of the bulk or films could be observed. Thus samples

Phase	Bulk		Film	
	First Temp	Last Temp	First Temp	Last Temp
Glass	35	200	35	200
Mg ₅₁ Zn ₂₀ [10, 13]	170	200	140	200
Mg	170	320	140	320
Ca ₂ Mg ₅ Zn ₅ [10, 14]	200	250	200	250
MgZn [13]	250	250	250	250
Ca ₂ Mg ₅ Zn ₁₃ [10–12]	290	320	290	320

Table 6: Temperatures at which each crystallisation phase is first observed and last observed in the annealing XRD for both the bulk and film. All temperatures are in °C.

were subjected to dynamic XRD over their most active range of 31 – 60° to observe changes *in-situ*. This allowed the crystallisation to be actively observed over the range of 35 – 185 °C, showing how phases evolved over time as the temperature was raised in 5°C increments.

4 DISCUSSION

The use of a 60K DSC heating rate compared to the more commonly used 20K rate [sources] shifts peaks for the bulk Mg₆₅Zn₃₀Ca₅ alloy about 8 - 15 degrees higher. This higher heating rates were used because crystallisation events for the films were difficult to differentiation at the lower heating rate. Films show little shift to high temperature peaks with increases heating rates, but large shifts with relaxation. Bulk show the opposite behaviour, larger peaks shifts with higher heating rates and little shift with relaxation.

5 CONCLUSIONS

6 ACKNOWLEDGEMENTS

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7 REFERENCES

- [1] J. Q. Wang, N. Chen, P. Liu, Z. Wang, D. V. Louzguine-Luzgin, M. W. Chen, and J. H. Perepezko. The ultrastable kinetic behavior of an Au-based nanoglass. *Acta Materialia*,

79(0):30–36, 2014.

- usch1998 [2] R. Busch, W. Liu, and W. L. Johnson. Thermodynamics and kinetics of the Mg₆₅Cu₂₅Y₁₀ bulk metallic glass forming liquid. *Journal of Applied Physics*, 83(8):4134–4141, 1998.
- gell2002 [3] C. A. Angell and S. Borick. Specific heats C_p , C_v , C_{conf} and energy landscapes of glassforming liquids. *Journal of Non-Crystalline Solids*, 307 - 310:393–406, 2002.
- Wei2014 [4] Shuai Wei, Zach Evenson, Isabella Gallino, and Ralf Busch. The impact of fragility on the calorimetric glass transition in bulk metallic glasses. *Intermetallics*, 55:138–144, 2014.
- hour2010 [5] Samir K. Ashour and Mahmood A. Abdel-hameed. Approximate skew normal distribution. *Journal of Advanced Research*, 1(4):341–350, 2010.
- moto2007 [6] Yuzo Yamamoto, Yuichi Inoue, Teruaki Onai, Chikashi Doshu, Hiroshi Takahashi, and Hiroki Uehara. Deconvolution analyses of differential scanning calorimetry profiles of β -crystallized polypropylenes with synchronized x-ray measurements. *Macromolecules*, 40(8):2745–2750, 2007.
- pink2008 [7] Charles H. Spink. *Differential Scanning Calorimetry*, volume Volume 84, pages 115–141. Academic Press, 2008.
- pink2015 [8] Charles H. Spink. The deconvolution of differential scanning calorimetry unfolding transitions. *Methods*, 76:78–86, 2015.
- ffer2005 [9] Balázs Schäffer, Béla Schäffer, and D. Lőrinczy. Decomposition of dsc curves of dairy products with gaussian functions. *Journal of Thermal Analysis and Calorimetry*, 82(2):531–535, 2005.
- hang2013 [10] Y. N. Zhang, G. J. Rocher, B. Briccoli, D. Kevorkov, X. B. Liu, Z. Altounian, and M. Medraj. Crystallization characteristics of the Mg-rich metallic glasses in the Ca-Mg-Zn system. *Journal of Alloys and Compounds*, 552:88–97, 2013.
- hang2012 [11] Yi-Nan Zhang, Dmytro Kevorkov, Xue Dong Liu, Florent Bridier, Patrice Chartrand, and Mamoun Medraj. Homogeneity range and crystal structure of the Ca₂Mg₅Zn₁₃ compound. *Journal of Alloys and Compounds*, 523:75–82, 2012.
- hang2011 [12] Yi-Nan Zhang, Dmytro Kevorkov, Florent Bridier, and Mamoun Medraj. Experimental study of the Ca-Mg-Zn system using diffusion couples and key alloys. *Science and Technology of Advanced Materials*, 12(2):025003, 2011.
- Khan1989 [13] Y. Khan. Dynamic temperature crystallization behaviour of amorphous and liquid Mg₇₀Zn₃₀ alloy. *Journal of Materials Science*, 24(3):963–973, 1989.

- Cao2016 [14] Jake D. Cao, Thomas Weber, Robin Schaublin, and Jörg F. Löffler. Equilibrium ternary intermetallic phase in the Mg-Zn-Ca system. *Journal of Materials Research*, 31(14):2147–2155, 2016.



Figure 1: Bulk Mg₆₅Zn₃₀Ca₅ relaxed at 120 °C for 10 minutes and heated at various heating rates (β s) from 5 to 100K/min. The insert stacks the differential scanning calorimetry (DSC) curves and labels the T_g and T_x s of the $\beta = 100K/min$ sample.



Figure 2: Unrelaxed film Mg₆₅Zn₃₀Ca₅ heated at various heating rates (β s) from 15 to 100K/min. The insert stacks the DSC curves and labels the T_g and T_x s of the $\beta = 100K/min$ sample.

./images/Bulk_Film_Fragility.png

Figure 3: Fitted fragility (m) for the $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$ system obtained from the T_g of DSC at various heating rates (β s) for the bulk and film.

m_mValue



Figure 5: DSC deconvolution for the film at various heating rates (β s). From left to right, top to bottom, $\beta = 15, 20, 30, 40, 60, 80, 100 \text{ K/min}$.

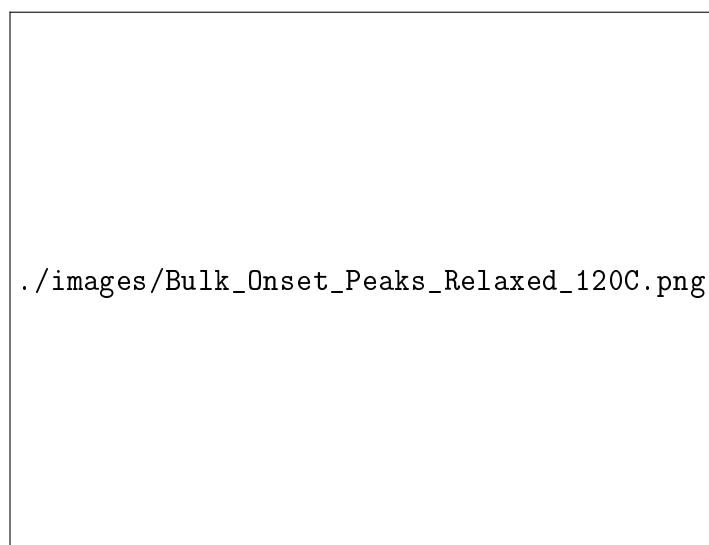


Figure 6: The T_g s and T_x s plotted at each DSC heating rate (β) for the bulk $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$.

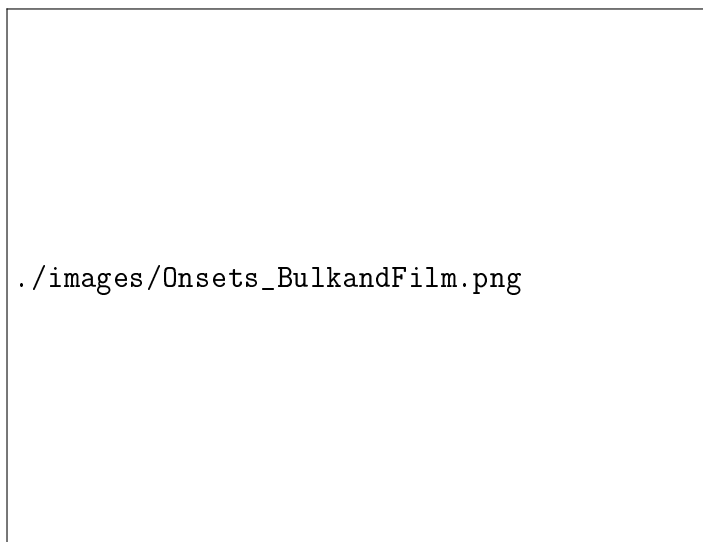


Figure 7: The T_g s and T_x s plotted at each DSC heating rate (β) for both the bulk and film $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$. Bulk is shown in black, and film in red.

BulkFilm

./images/Decon_Onsets_BR.png

./images/Decon_peak_area_BR.png

Figure 8: DSC T_x onset temperatures and specific enthalpy (h) of crystallisation formation for the bulk and film at each heating rate (β). Bulk is shown in black, and film in red.

SC_Decon

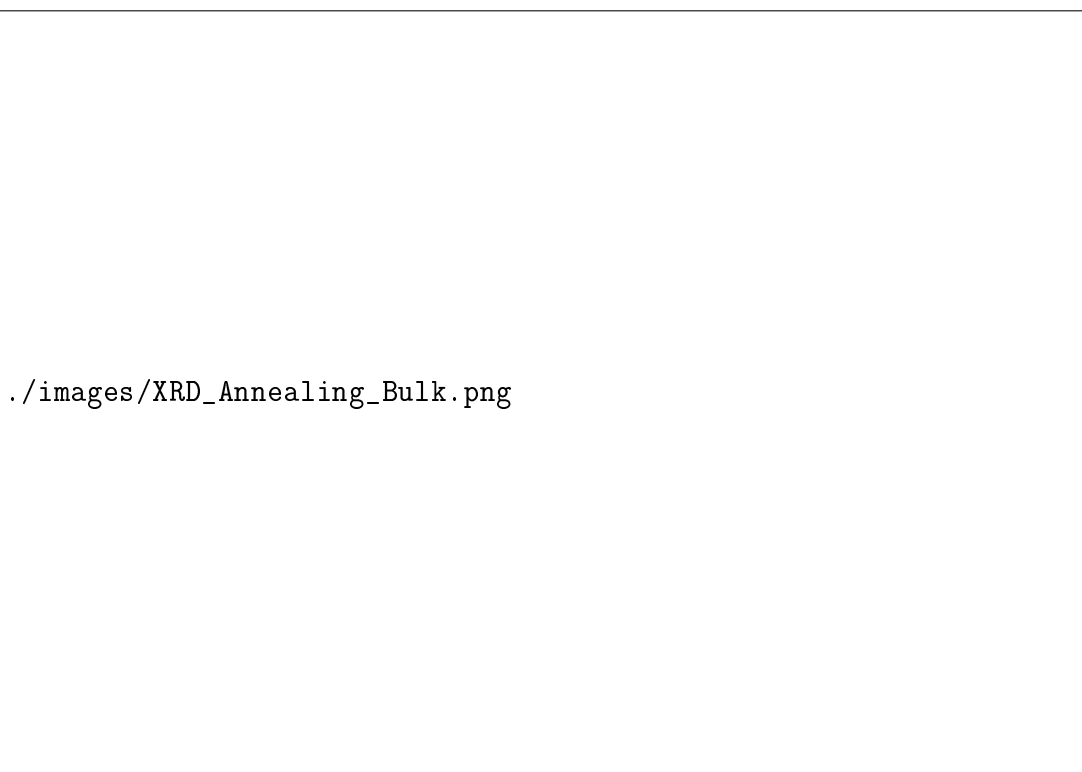


Figure 9: XRD pattern for bulk $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$ heated treated to several temperatures for crystallisation peak identified from DSC.

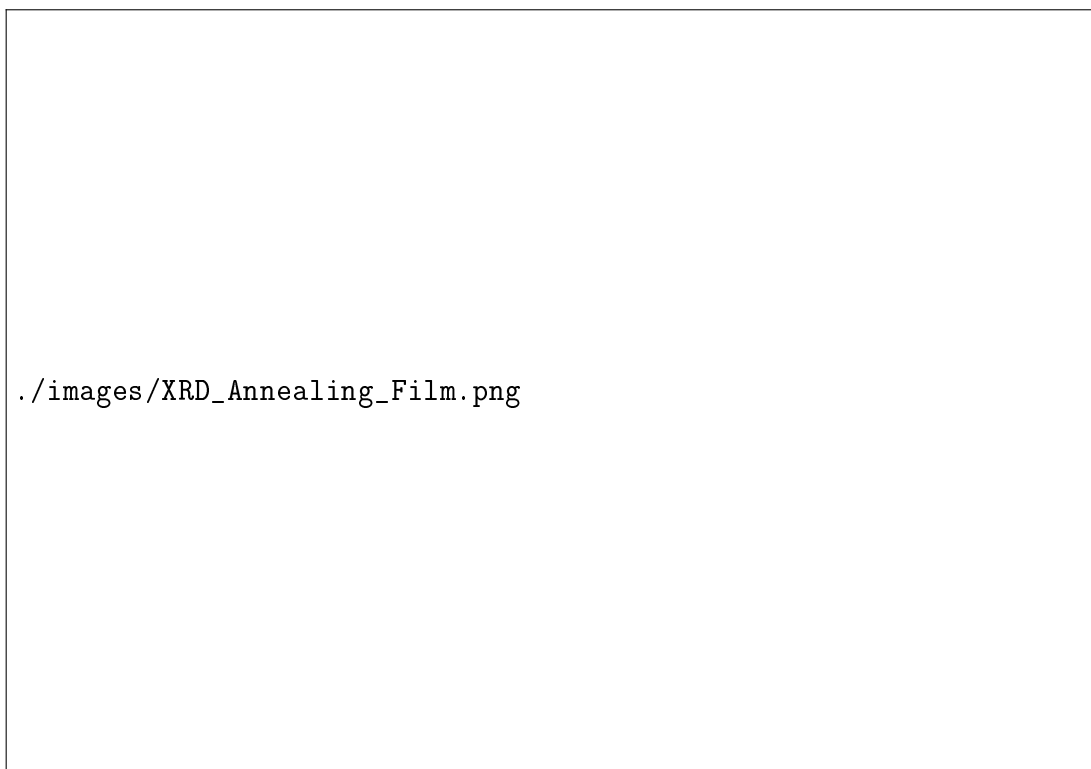


Figure 10: XRD pattern for film $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$ heated treated to several temperatures for crystallisation peak identified from DSC. Note the Al substrate peaks have been faceted as to not dwarf all other peaks.

./images/XRD_Dynamic_Bulk.png

(a)

./images/Bulk_Heated_XRD_Waterfall3D_Smooth2.png

(b)

Figure 11: (a) Stacked XRD patterns from the incremental dynamic *in-situ* heating of bulk $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$. Note the peak around 51° is attributed to the AlN heating element. (b) The same XRD patterns as (a) presented in a cascading layout.

amic_FullStack_Film

./images/XRD_Dynamic_Film.png

(a)

amic_WaterFall_Film

./images/TF_Facet_HeatXRD_Waterfall3D_Smooth.png

(b)

mic_Film

Figure 12: (a) Stacked XRD patterns from the incremental dynamic *in-situ* heating of film $\text{Mg}_{65}\text{Zn}_{30}\text{Ca}_5$. Note the peak around 51° is attributed to the AlN heating element, and the Al substrate peaks have been faceted as to not dwarf all other peaks. (b) The same XRD patterns as (a) presented in a cascading layout.