

Teacher Information

Thank you for your download! I hope you find this resource useful in your classroom and for your students. I would love to hear your feedback. You can leave me a comment and rating on my Teachers Pay Teachers store.

<http://www.teacherspayteachers.com/Store/Elly-Thorsen>

© Copyright 2015 by Elly Thorsen. All rights reserved. This resource is copyrighted so please only use it for your classroom and do not use it for any commercial purposes. Permission is granted to copy pages specifically designed for student or teacher use by the original purchaser or licensee. The reproduction of any other part of this product is strictly prohibited. **Copying any part of this product and placing it on the Internet in any form (even a personal or classroom website) is strictly forbidden.** Doing so makes it possible for an Internet search to make the document available on the Internet, free of charge, and is a violation of the Digital Millennium Copyright Act (DMCA).

The digital papers in this resource are from CarrieStephensArt.

<https://www.teacherspayteachers.com/Store/Carriestephensart>

Take a look at the many other Physical Science resources in my store.

<https://www.teacherspayteachers.com/Store/Elly-Thorsen/Category/Physical-Science-187510>

You can find more teaching information and notices of sales on my blog and Facebook page.

<http://ellythorsenteaching.blogspot.kr/>

<https://www.facebook.com/EllyThorsenTPT>

Thanks again!

Elly Thorsen



Law of Conservation of Mass Exit Ticket

Name: _____

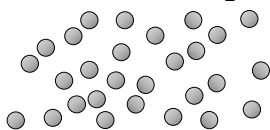
Date: _____

Hour: _____

1. What does the Law of Conservation of Mass state?
 - a. Mixtures combined create new matter.
 - b. Matter is not created or destroyed in any chemical or physical change.
 - c. Chemical and physical changes can create or destroy matter but not at the same time.
 - d. Matter can be created in chemical changes only.
2. What happened to the atoms in a piece of wood that was burned?
 - a. They are grouped differently, but all of the atoms are still there.
 - b. They are grouped differently, but some of the atoms were burned and destroyed in the fire.
 - c. All of the atoms were destroyed in the fire.
 - d. The atoms in the wood moved really fast and then stopped moving forever.
3. A student placed fruit in a tightly sealed bag and found the mass of the fruit every week. She found the mass did not change over time even though the fruit was going through a chemical change. Is this an example of the Law of Conservation of Mass? Why or why not? Answer in sentences.

4. Examine the pictures of particles below. Is this possible after a chemical change? Why or why not? Explain in sentences.

Before the Change



After the Change

Law of Conservation of Mass Exit Ticket

Name: _____

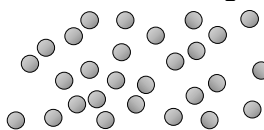
Date: _____

Hour: _____

1. What does the Law of Conservation of Mass state?
 - a. Mixtures combined create new matter.
 - b. Matter is not created or destroyed in any chemical or physical change.
 - c. Chemical and physical changes can create or destroy matter but not at the same time.
 - d. Matter can be created in chemical changes only.
2. What happened to the atoms in a piece of wood that was burned?
 - a. They are grouped differently, but all of the atoms are still there.
 - b. They are grouped differently, but some of the atoms were burned and destroyed in the fire.
 - c. All of the atoms were destroyed in the fire.
 - d. The atoms in the wood moved really fast and then stopped moving forever.
3. A student placed fruit in a tightly sealed bag and found the mass of the fruit every week. She found the mass did not change over time even though the fruit was going through a chemical change. Is this an example of the Law of Conservation of Mass? Why or why not? Answer in sentences.

4. Examine the pictures of particles below. Is this possible after a chemical change? Why or why not? Explain in sentences.

Before the Change



After the Change

Law of Conservation of Mass Exit Ticket

Name: _____

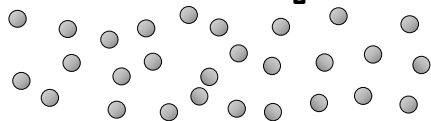
Date: _____

Hour: _____

1. What does the Law of Conservation of Mass state?
 - a. Mixtures combined create new matter.
 - b. Matter is not created or destroyed in any chemical or physical change.
 - c. Chemical and physical changes can create or destroy matter but not at the same time.
 - d. Matter can be created in chemical changes only.
2. What happened to the atoms in a piece of wood that was burned?
 - a. They are grouped differently, but all of the atoms are still there.
 - b. They are grouped differently, but some of the atoms were burned and destroyed in the fire.
 - c. All of the atoms were destroyed in the fire.
 - d. The atoms in the wood moved really fast and then stopped moving forever.
3. A student placed fruit in a tightly sealed bag and found the mass of the fruit every week. She found the mass did not change over time even though the fruit was going through a chemical change. Is this an example of the Law of Conservation of Mass? Why or why not? Answer in sentences.

4. Examine the pictures of particles below. Is this possible after a chemical change? Why or why not? Explain in sentences.

Before the Change



After the Change

Law of Conservation of Mass Exit Ticket

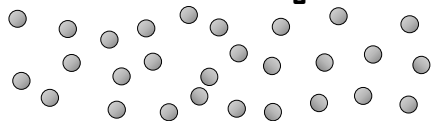
Answer Key

1. What does the Law of Conservation of Mass state?
 - a. Mixtures combined create new matter.
 - b. Matter is not created or destroyed in any chemical or physical change.**
 - c. Chemical and physical changes can create or destroy matter but not at the same time.
 - d. Matter can be created in chemical changes only.
2. What happened to the atoms in a piece of wood that was burned?
 - a. They are grouped differently, but all of the atoms are still there.**
 - b. They are grouped differently, but some of the atoms were burned and destroyed in the fire.
 - c. All of the atoms were destroyed in the fire.
 - d. The atoms in the wood moved really fast and then stopped moving forever.
3. A student placed fruit in a tightly sealed bag and found the mass of the fruit every week. She found the mass did not change over time even though the fruit was going through a chemical change. Is this an example of the Law of Conservation of Mass? Why or why not? Answer in sentences.

Example Answer: This is an example of the Law of Conservation of Mass because it shows no matter (or mass) is lost even during a chemical change. All of the matter is still in the bag, just in a different form.

4. Examine the pictures of particles below. Is this possible after a chemical change? Why or why not? Explain in sentences.

Before the Change



After the Change

Example Answer: This change is not possible because of the law of Conservation of Mass. The picture "After the Change" shows all of the atoms have disappeared. We know from the law that atoms are never destroyed or created in chemical changes; they are only rearranged.
