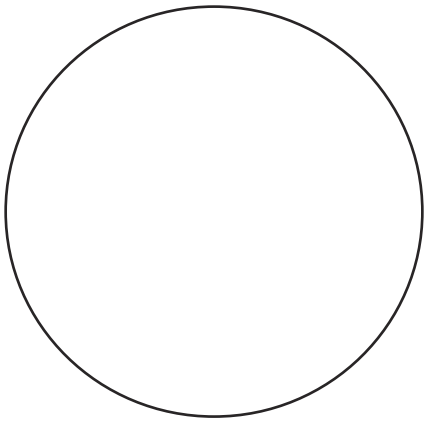
Ionic Compounds Observations

Observation 1

Place some salt on a slide and view it under a microscope at different magnifications.
Draw your observations below, adding any detail you think necessary.



Observation 2

Ensure your working area is clear. Tie long hair up and wear safety goggles. Using a Bunsen burner on a blue roaring flame, gently heat a small amount of salt in a test tube.

What happens?

Why do you think this is?

Challenge

One grain of salt contains, on average, over a trillion atoms of sodium and chlorine.

Thinking back to last lesson on ionic bonding, can you predict the arrangement of the sodium and chlorine ions? Use your observations from step 1 to think about the overall shape, then think how the ions would be arranged within that shape.

Draw your ideas here:

Observation 3

Place a small amount of salt in a beaker.
Set up a circuit with a cell, bulb and wires.
Incorporate the beaker with salt **into** the circuit.

Does the bulb light up?

Why do you think this is?

Observation 4

Now add water to the beaker of salt and repeat.

Does the bulb light up?

Why do you think this is?