

# Doodle Notes

## Endo/Exothermic Reactions

By Jady Thone

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### Resource Explanation

A brief description of the resource and teacher preparation instructions, including suggestions for classroom use.

2

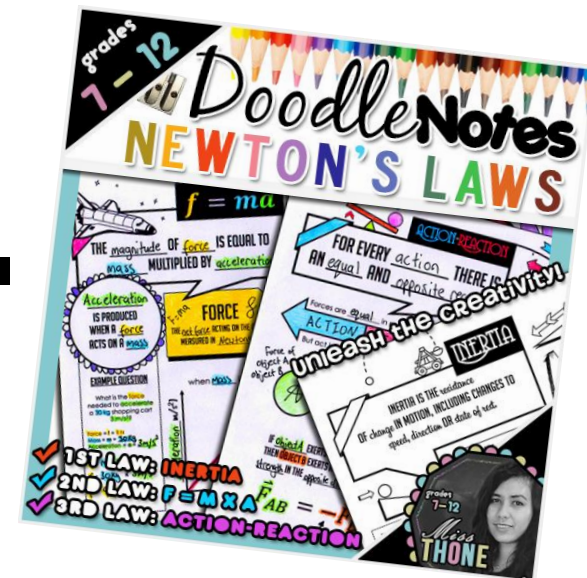
### Doodle Notes Graphic Organizer

3 levels of the graphic organizer, completed for differentiation (fully completed, fill-in-the-blanks and "empty") + extra note-taking pages!

3

### Teacher "Answer Key"

Stuck for design ideas? Show students the answer key first (only quickly), so they can gain an idea of what the finished product will look like.



**Thank you!**

I hope this resource is useful in your classroom. Please don't hesitate to email me with questions at [jadythone@outlook.com](mailto:jadythone@outlook.com)

# Doodle Notes

## Resource Explanation

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By Jadyn Thone

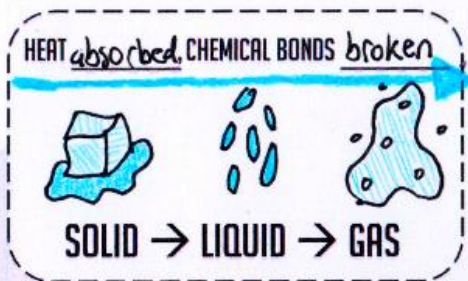
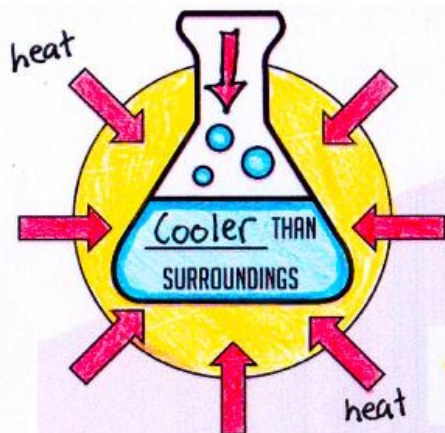
### Description

This is a Newton's First Law of Motion graphic organizer that allows students to compile their knowledge and understanding in a doodle notes style! Give your students a welcome change of pace by leading them through the ideas and concepts behind Newton's First Law of Motion (Inertia) and then gift them with time to doodle and make those connections! This graphic organizer would be great for an interactive notebook, too!

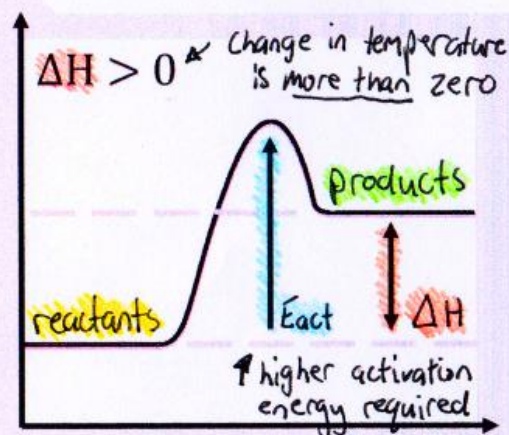
### Teacher Preparation Instructions

Simply print off as many copies as necessary, provide students with colored pencils, textas or crayons and watch them engage the creative side of their brain! This single-sided sheet doodle notes could be used as an introduction to Newton's First Law of Motion or as independent/group review of inertia/friction/gravity/Newton's Laws.





ENDOTHERMIC REACTIONS take in (OR absorb) ENERGY FROM THE SURROUNDINGS. THE REACTION IS cold.



IN AN endothermic reaction, THE ENERGY increases OVER TIME AS HEAT IS absorbed INTO THE SYSTEM. THEREFORE, THE OVERALL TEMPERATURE DECREASES HEAT ENERGY INCREASES

# ENDOTHERMIC

in, absorb

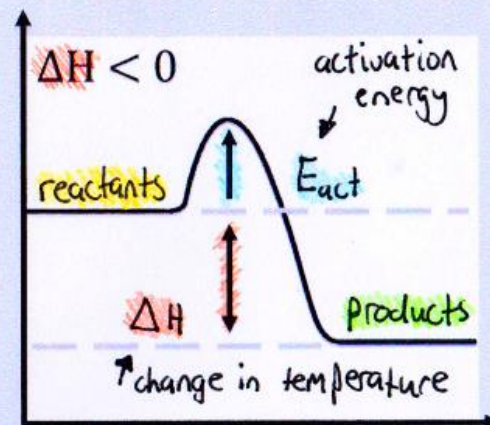
heat energy

out, release

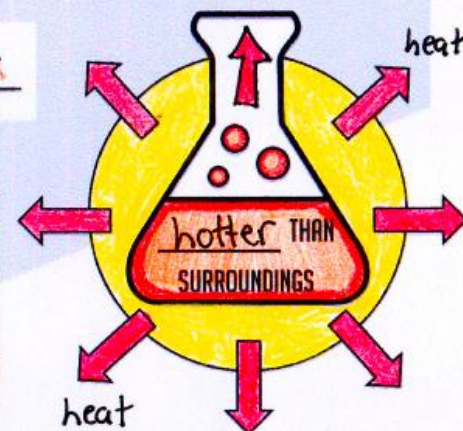
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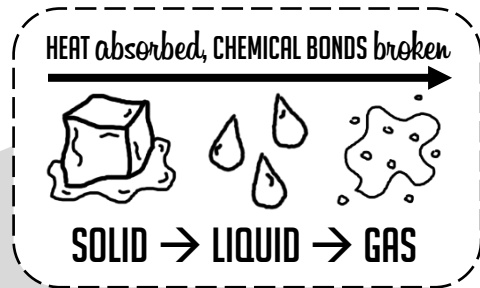
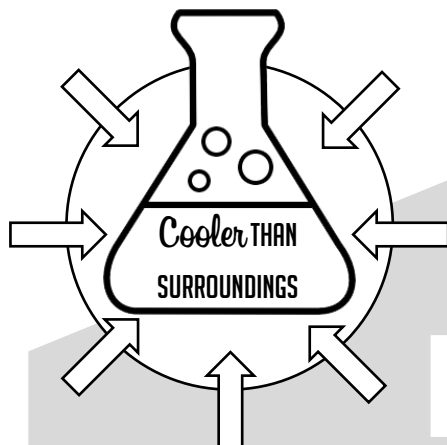
# EXOTHERMIC

IN AN exothermic reaction, THE ENERGY decreases OVER TIME AS HEAT IS released FROM THE SYSTEM. THEREFORE, THE OVERALL TEMPERATURE INCREASES HEAT ENERGY DECREASES



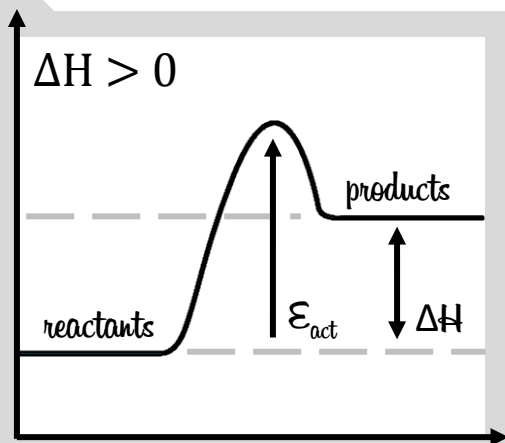
EXOTHERMIC REACTIONS give out (OR release) ENERGY TO THE SURROUNDINGS. THE REACTION IS hot.





REACTANTS + Heat  $\rightarrow$  PRODUCTS

ENDOTHERMIC REACTIONS take in (OR absorb) ENERGY FROM THE SURROUNDINGS. THE REACTION IS cold.



IN AN endothermic reaction, THE ENERGY increases OVER TIME AS HEAT IS absorbed INTO THE SYSTEM. THEREFORE, THE OVERALL heat energy increases.

# ENDOTHERMIC

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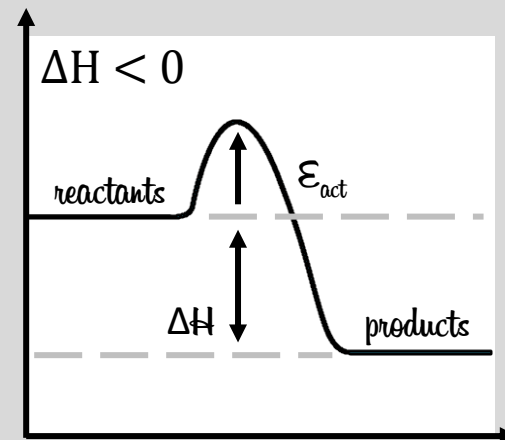
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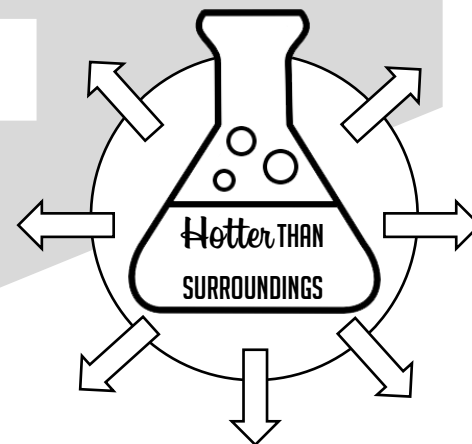
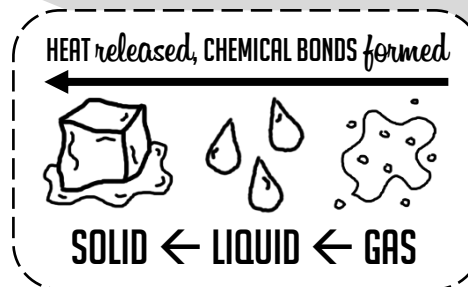
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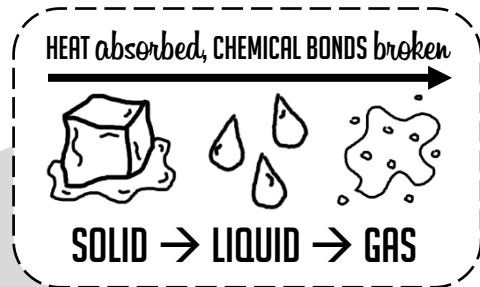
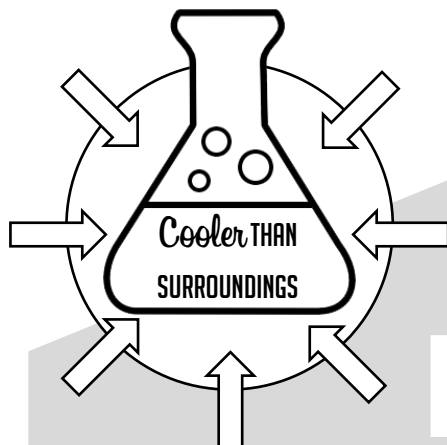


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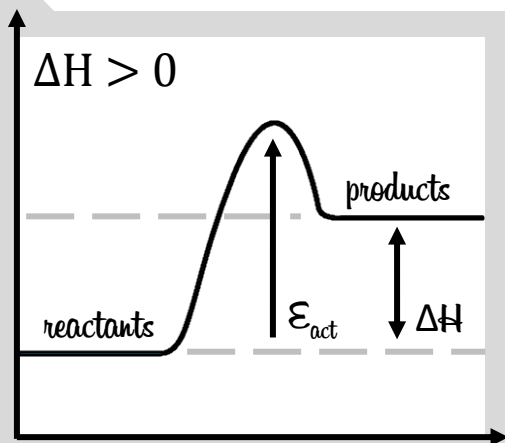






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IN AN endothermic reaction, THE ENERGY increases OVER TIME AS HEAT IS absorbed INTO THE SYSTEM. THEREFORE, THE OVERALL temperature decreases.

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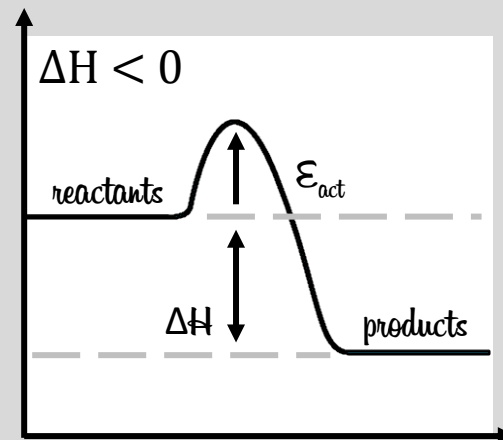
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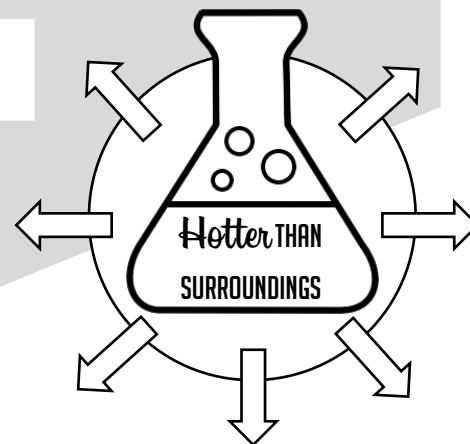
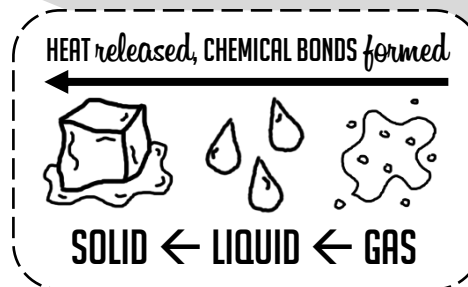
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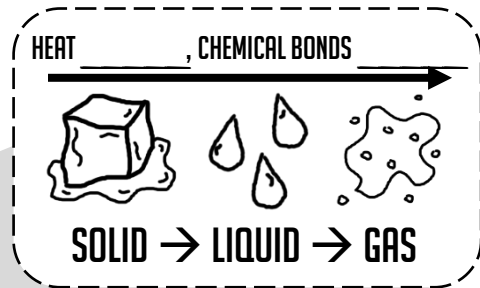
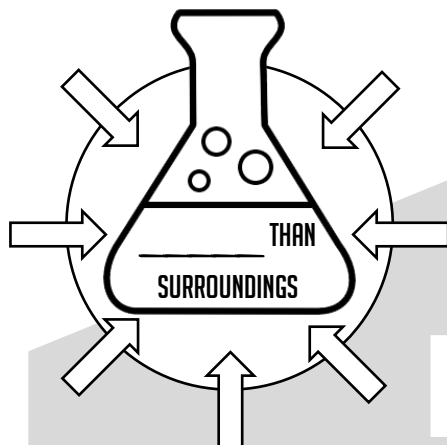
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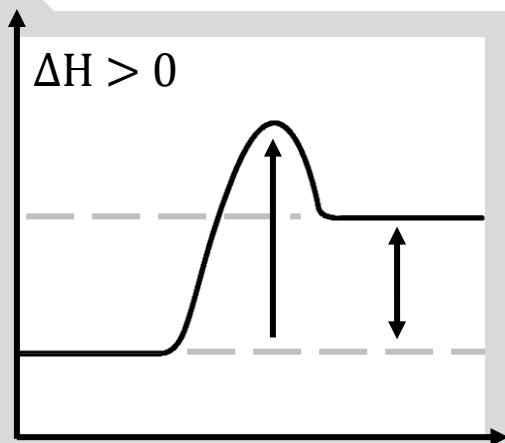
REACTANTS → PRODUCTS + Heat





REACTANTS +  $\xrightarrow{\hspace{1cm}}$  PRODUCTS

ENDOTHERMIC REACTIONS  $\xrightarrow{\hspace{1cm}}$  (OR  $\xrightarrow{\hspace{1cm}}$ )  
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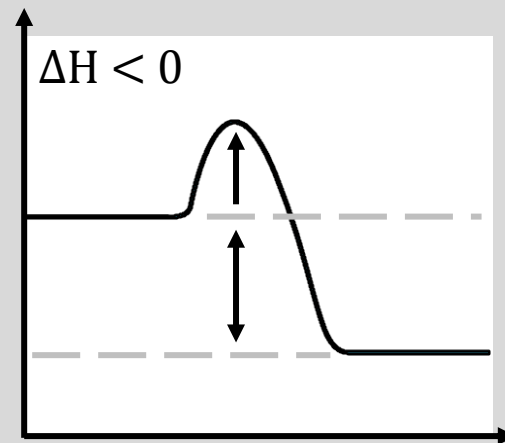


IN AN  $\xrightarrow{\hspace{1cm}}$ , THE ENERGY  
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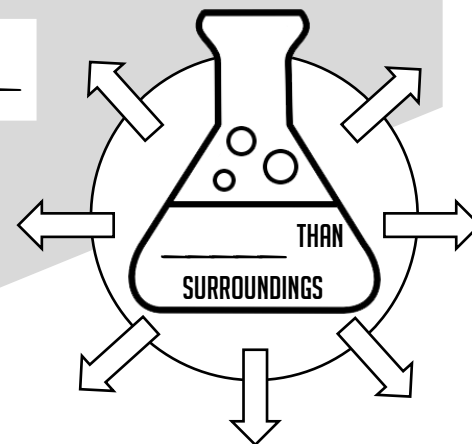
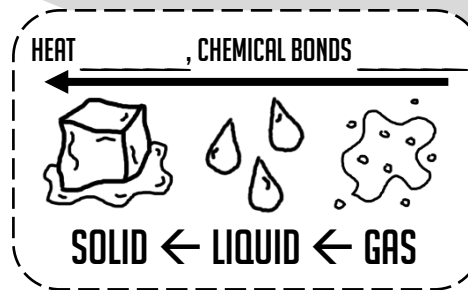
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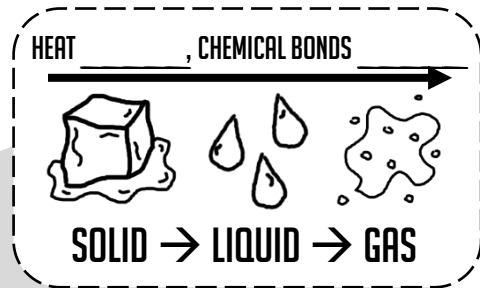
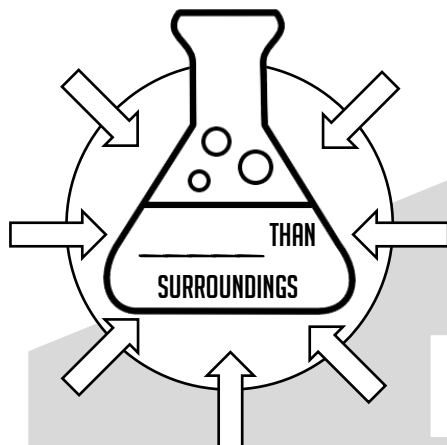
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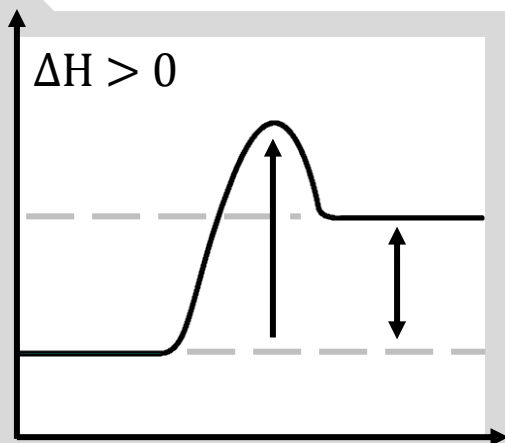
REACTANTS  $\rightarrow$  PRODUCTS +  $\xrightarrow{\hspace{1cm}}$





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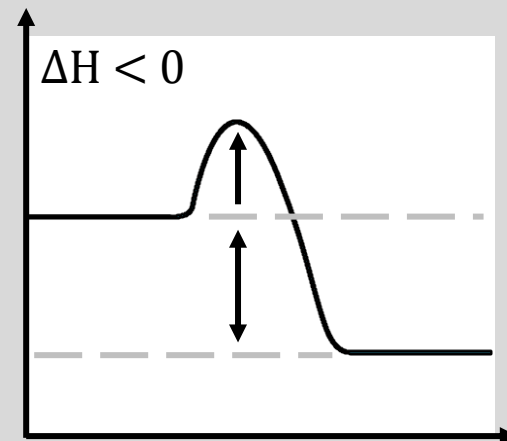


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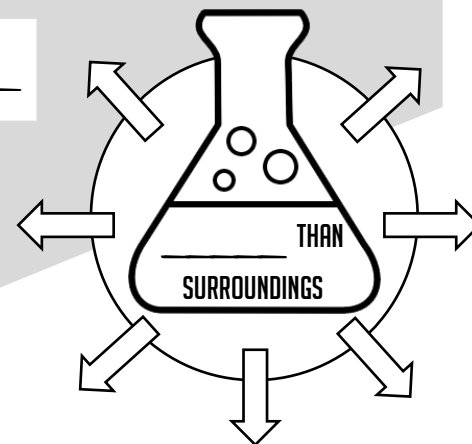
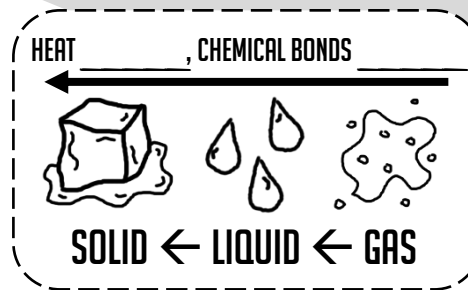
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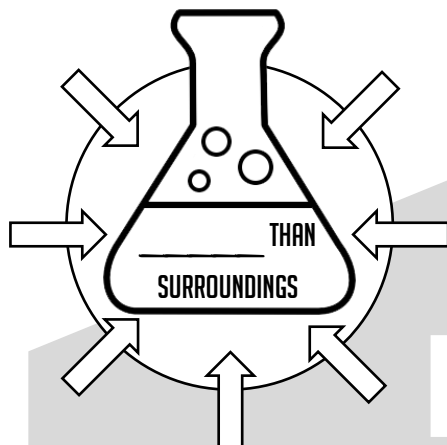
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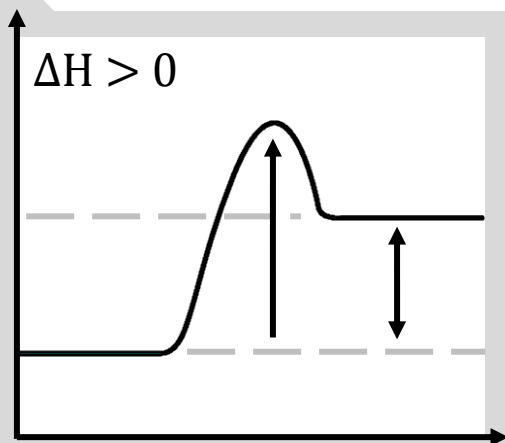


HEAT \_\_\_\_\_, CHEMICAL BONDS \_\_\_\_\_

SOLID → LIQUID → GAS

REACTANTS + \_\_\_\_\_ → PRODUCTS

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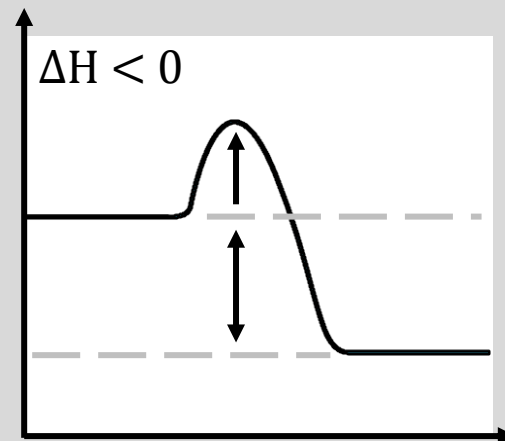


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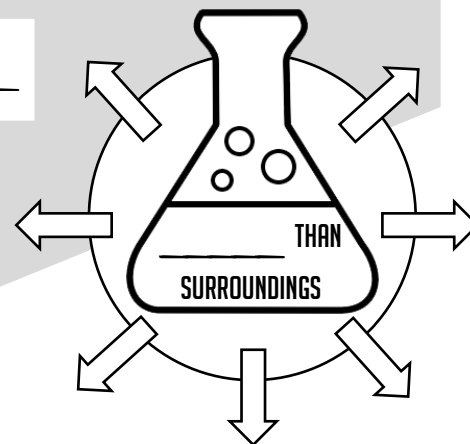


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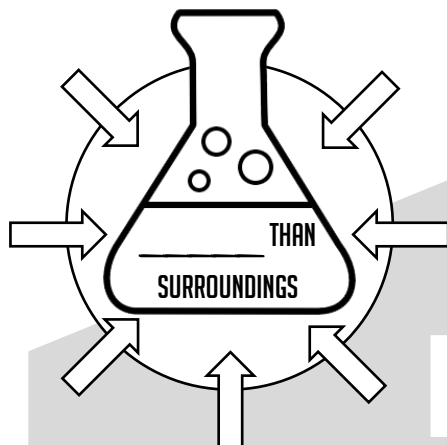
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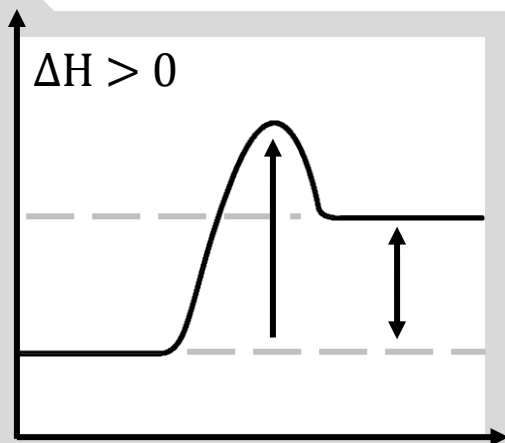


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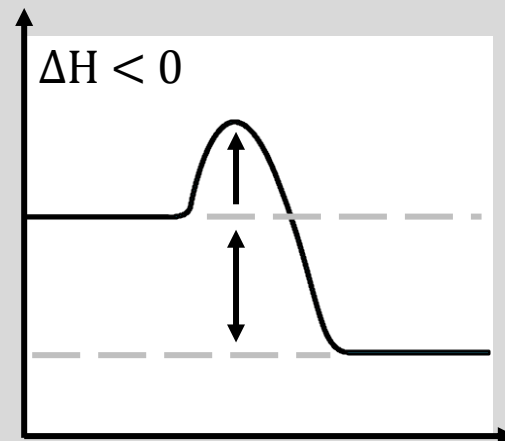


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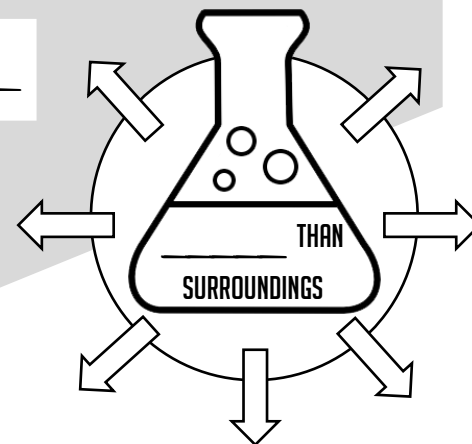


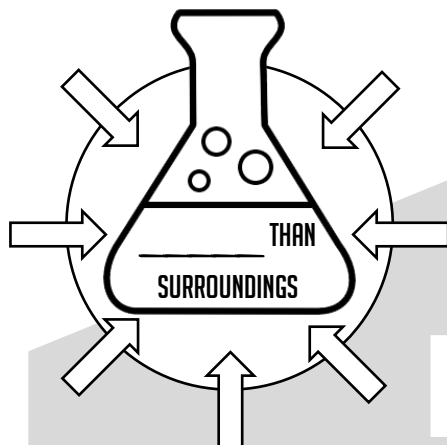
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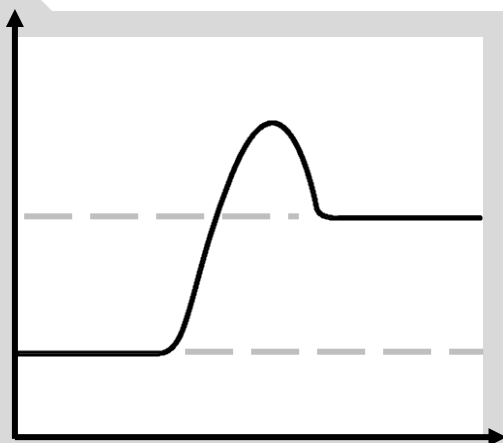
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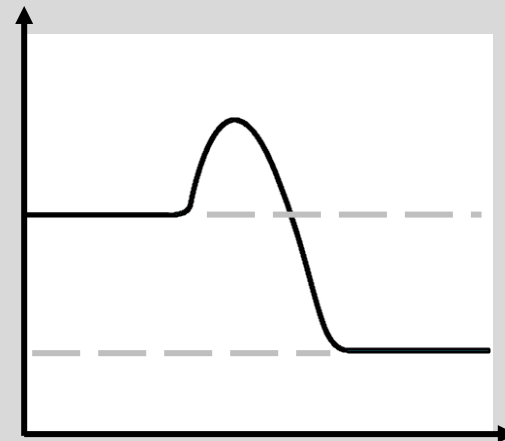


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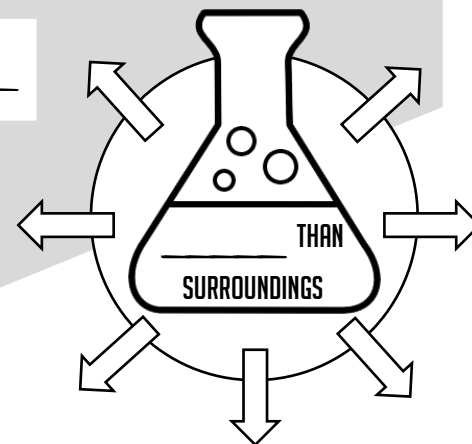
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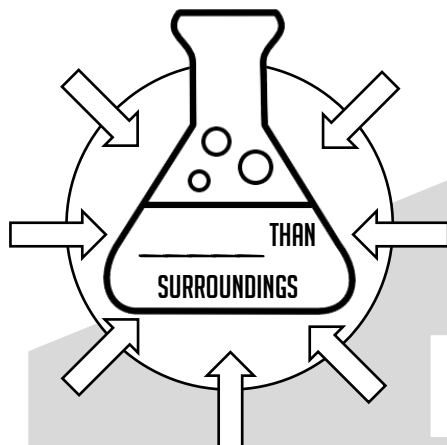
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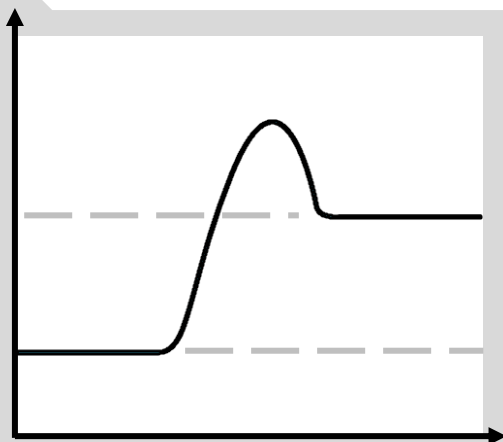


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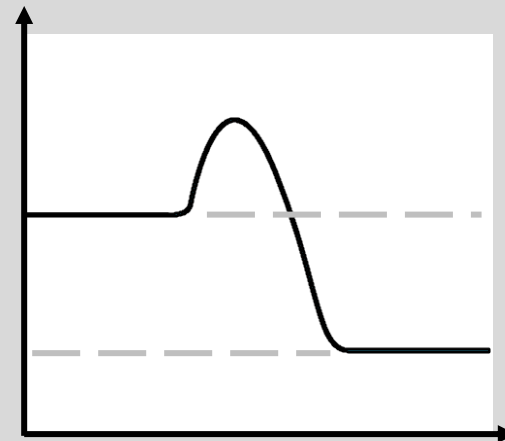


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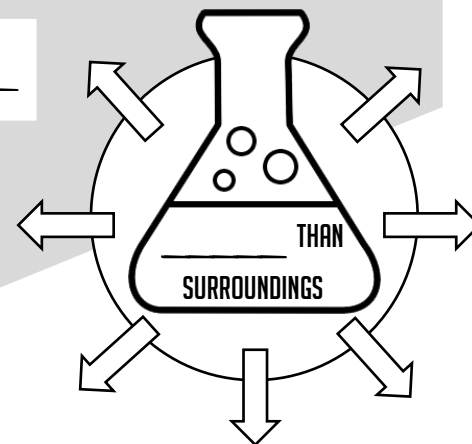


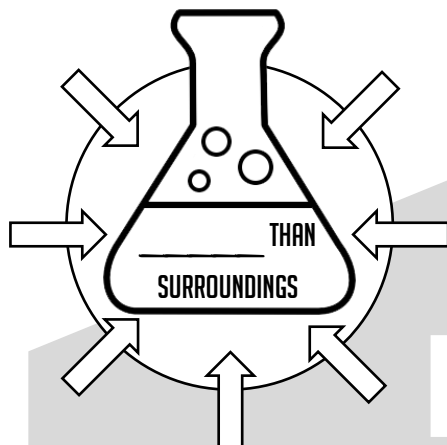
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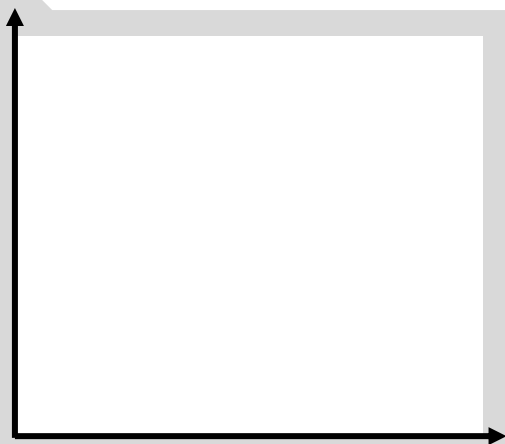
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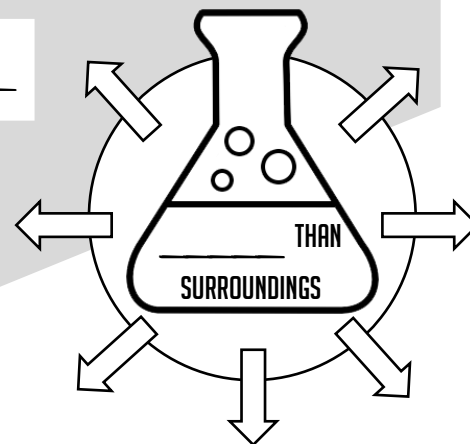
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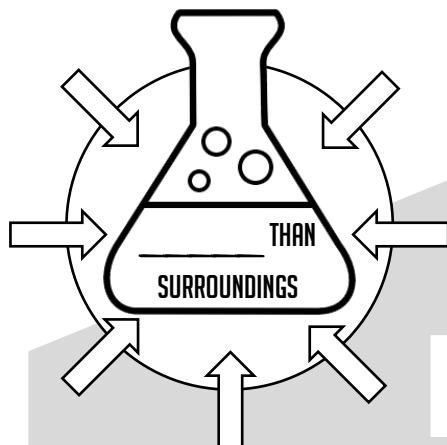
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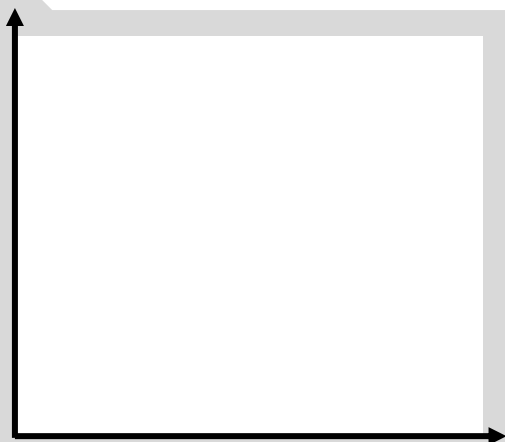


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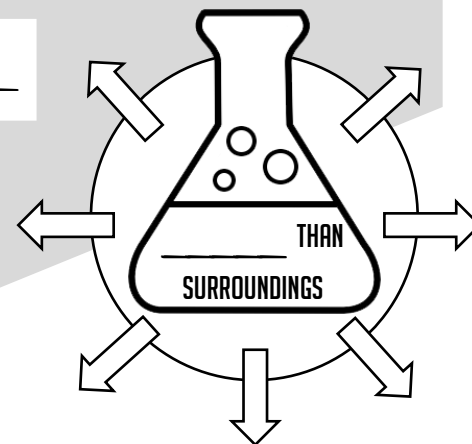


EXOTHERMIC REACTIONS \_\_\_\_\_ (OR \_\_\_\_\_)  
ENERGY TO THE SURROUNDINGS. THE REACTION IS \_\_\_\_\_.

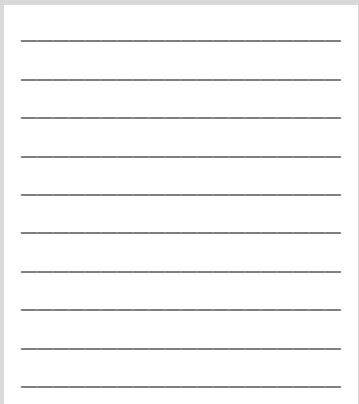
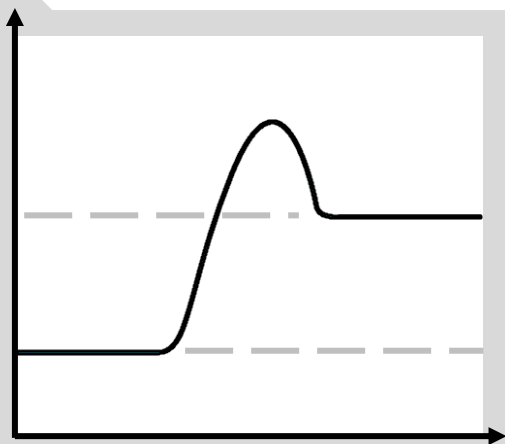
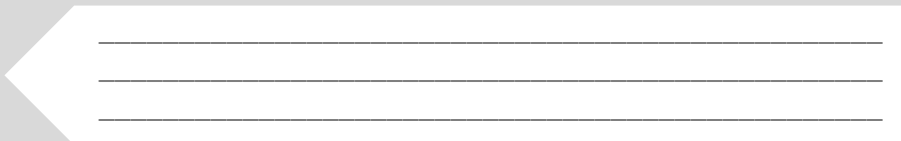
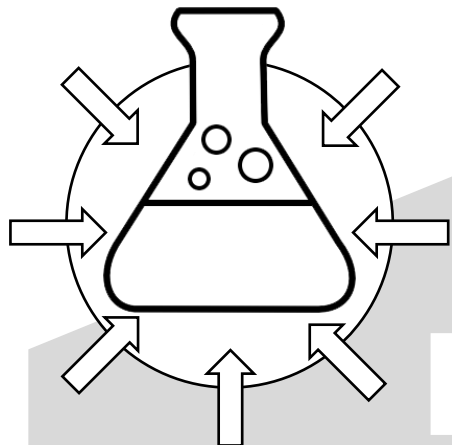
REACTANTS      PRODUCTS \_\_\_\_\_

HEAT \_\_\_\_\_, CHEMICAL BONDS \_\_\_\_\_

SOLID      LIQUID      GAS



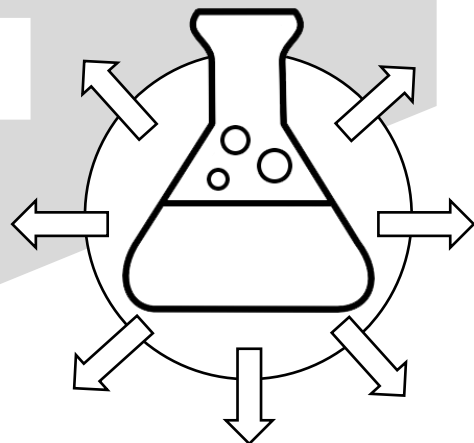
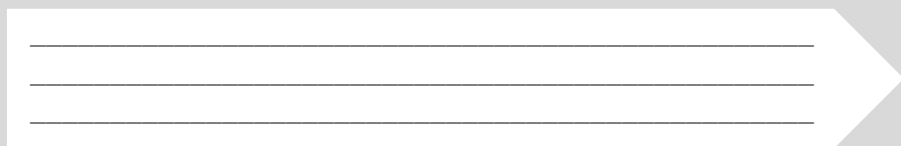
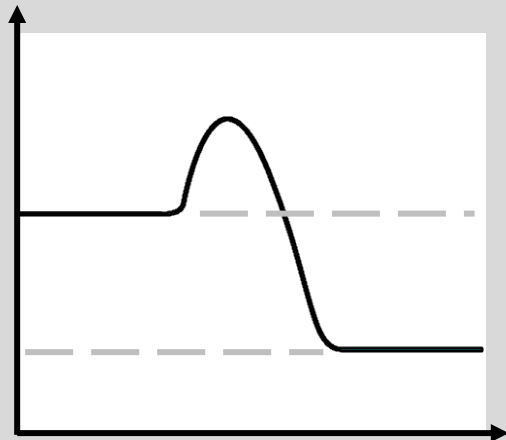
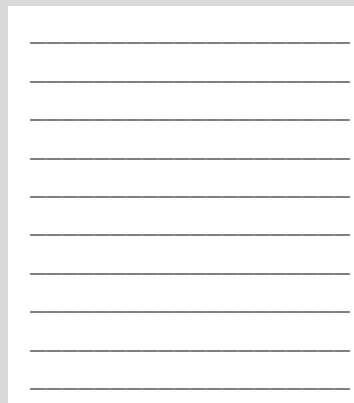


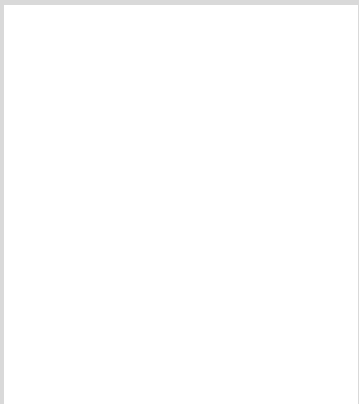
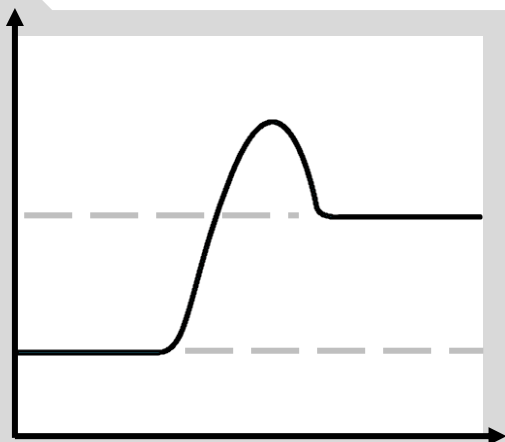
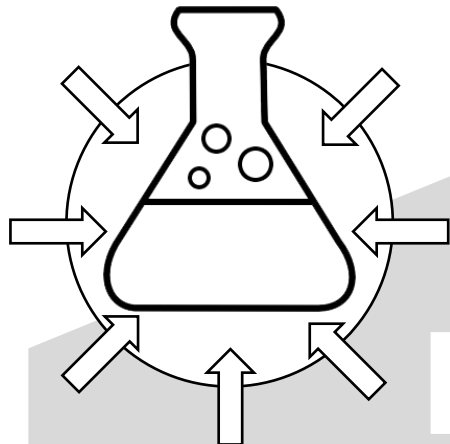


**ENDOTHERMIC**



**EXOTHERMIC**

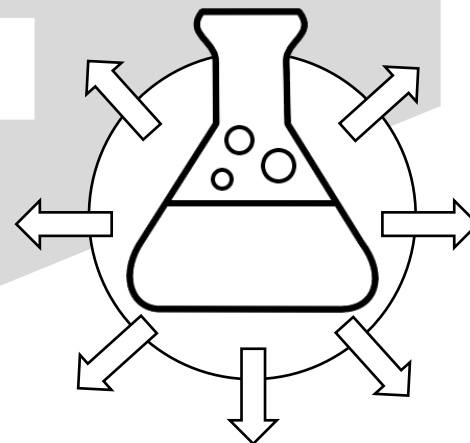
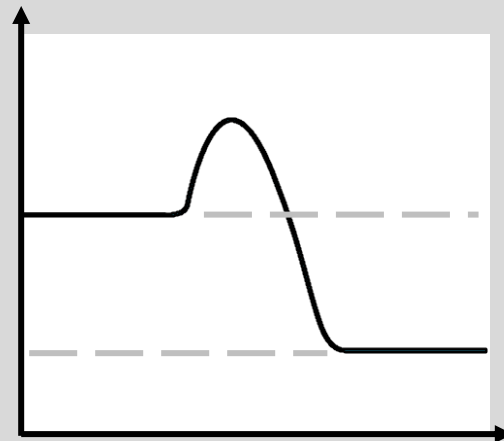
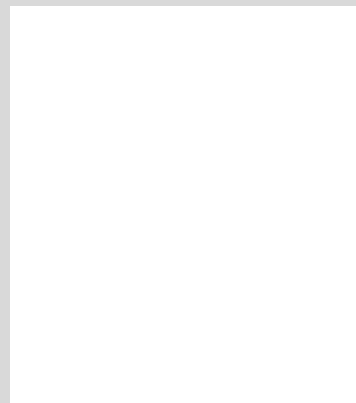


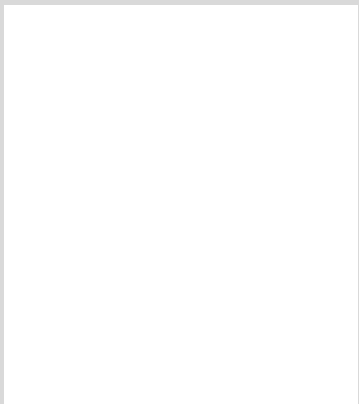
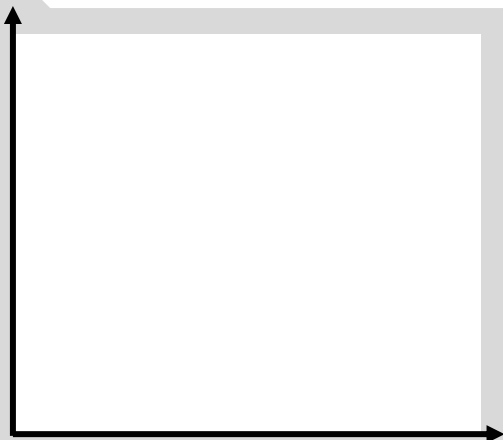
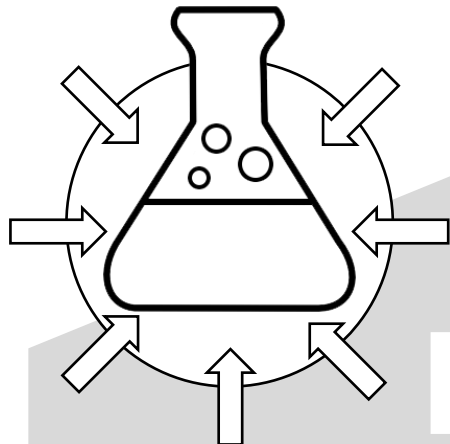


**ENDOTHERMIC**



**EXOTHERMIC**

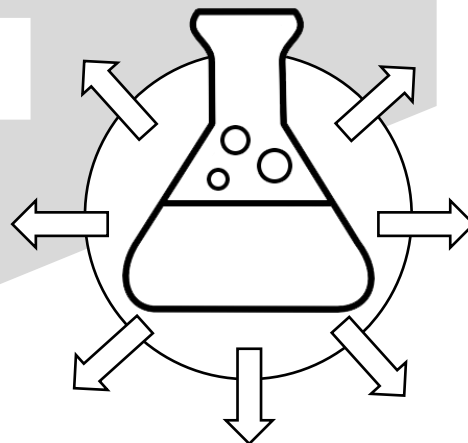
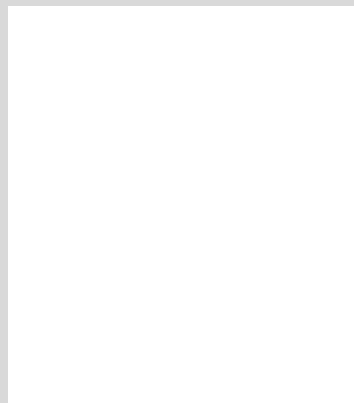


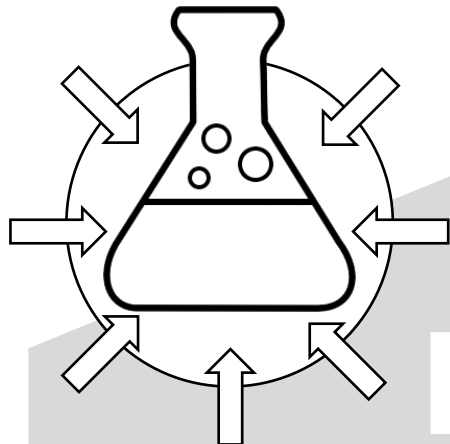


**ENDOTHERMIC**



**EXOTHERMIC**





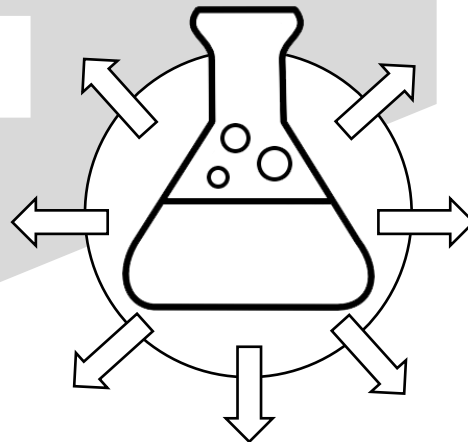
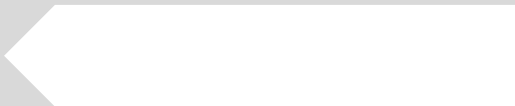
Lined area for writing notes about endothermic reactions.

**ENDOTHERMIC**



**EXOTHERMIC**

Lined area for writing notes about exothermic reactions.



# Jadyn Thone

## Who Am I?

My name is **Jadyn Thone** and I am an Australian Secondary teacher of Science, Technology, Math and English. During my time as a teacher I have taught many age groups of varying skill and ability. I found that I needed to create resources that suited my students. All my resources are tried and tested in authentic school settings. I upload new materials all the time, so don't forget to check back in! All the best!



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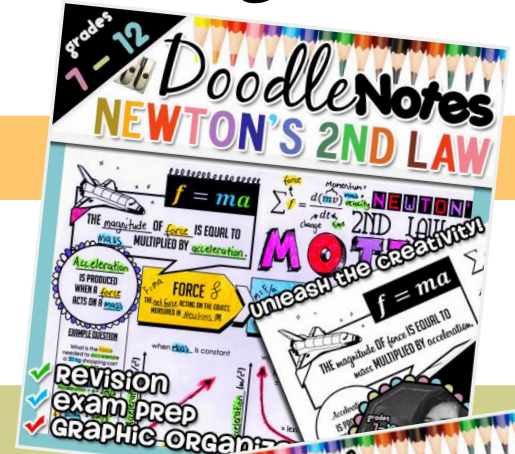
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