DRUGS II

Drugs we learnt last week :

Opiates – Morphine, heroin, fentanyl

Hallucinogens -THC, LSD, mescaline, PCP, psilocybin, and MDMA (Ecstasy), ergot alkaloids, mushroom ID

Depressants - Alcohol (ethanol), barbiturates, tranquilizers, and various substances that can be sniffed

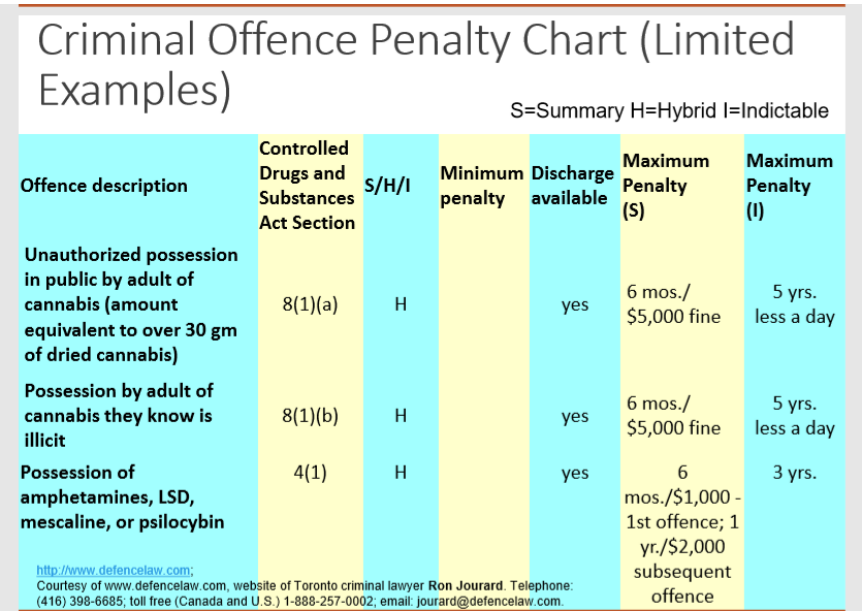
Stimulants - Amphetamines, cocaine

Club Drugs - MDMA, GHB, Rohypnol, ketamine, and methamphetamine

Anabolic Steroids

➔**Illicit Drug :** a controlled substance or precursor the import, export, production, sale or possession of which is prohibited or restricted pursuant to the "Controlled Drugs and Substance Act:

➔**Illicit drug use :** the importation, exportation, production, sale or possession of a controlled substance or precursor contrary to the Controlled Drugs and Substances Act or a regulation made under that Act



Criminal Offence Penalty Chart (Limited Examples)

S=Summary H=Hybrid I=Indictable

Offence description	Controlled Drugs and Substances Act Section	S/H/I	Minimum penalty	Discharge available	Maximum Penalty (S)	Maximum Penalty (I)
Unauthorized possession in public by adult of cannabis (amount equivalent to over 30 gm of dried cannabis)	8(1)(a)	H		yes	6 mos./ $5,000 fine	5 yrs. less a day
Possession by adult of cannabis they know is illicit	8(1)(b)	H		yes	6 mos./ $5,000 fine	5 yrs. less a day
Possession of amphetamines, LSD, mescaline, or psilocybin	4(1)	H		yes	6 mos./$1,000 - 1st offence; 1 yr./$2,000 subsequent offence	3 yrs.

http://www.defencelaw.com;
Courtesy of www.defencelaw.com, website of Toronto criminal lawyer **Ron Jourard**. Telephone: (416) 398-6685; toll free (Canada and U.S.) 1-888-257-0002; email: jourard@defencelaw.com.

Collection and preservation of drug evidence

➔The field investigator has the responsibility of ensuring that the evidence is properly packaged and labeled for the laboratory.

➔Generally common sense is the best guide, keeping in mind that the package must prevent the loss of the contents and/or cross-contamination. Often the original container in which the drug was seized will suffice.

➔make sure as the person collecting it, the packages must be marked properly with information that is sufficient to ensure identification by the officer in future and stablish chain of custody

➔white powder can be anything, therefore our notes should be properly done and provide any background information that may relate to a drug's identity.

Drug Identification :

When identifying the drug, the problem comes on what analytical procedure to use that will ensure a specific identification of a drug.

The plan is divided into 2 parts :

1) Screening tests that are non-specific and preliminary in nature, they reduce the possible drug IDs to a manageable number.
2) Confirmation tests that are single tests that specifically identify a substance.

Preliminary Analysis

We are faced with the prospect that the unknown substance may be one of thousand or more commonly encountered drugs, the analyst must employ screening tests to reduce these possibilities to a small manageable number.

Thee objective is often accomplished by subjecting the material to a series of color tests that will produce characteristic colors for more commonly encountered illicit drugs.

Microcrystalline test can also be used to identify specific drug substances by studying the size and shape of crystals formed when the drug is mixed with specific reagents.

Conformational Determination

After the preliminary analysis is done, a conformational determination is pursued.

Forensic chemists will employ aa specific test to identify a drug substance to the exclusion of all other known chemical substances.

Typically infrared spectrometry or gas chromatography -mass spectrometry is used to specifically identify a drug substance

Qualitative vs Quantitative

➔there are different types of analytical technique that has to be done with the drug : qualitative or quantitative

➔The former relates just to the identity of the material, whereas the latter requires the determination of the percent composition of the components of a mixture

Chromatography

➔separating and tentatively identifying the components of mixture

➔The theory of chromatography is based on the observation that chemical substances have a tendency to partially escape into the surrounding environment when dissolved in a liquid or when absorbed on a solid surface.

➔Those materials that have a preference for the moving phase will slowly pull ahead and separate from those substances that prefer to remain in the stationary phase.

➔depends on what the chemical's affinity is, with liquid or solid : chromatography. We have a solid phase and then a stationary phase.

Thin layer chromatography

➔solid stationary phase : coated onto a glass plate and a mobile liquid phase to separate the components of the mixture.

• The liquid will slowly rise up the plate by capillary action causing the sample to become distributed between the solid stationary phase and the moving liquid phase.

• Because most compounds are colorless, the materials must be visualized by placing the plates under ultraviolet light or spraying the plate with a chemical reagent.

• The distance a spot travels up a thin-layer plate can be assigned a numerical value known as the Rf (rate of flow) value.

Gas Chromatography

➔instead of liquid moving over solid, we have gas moving through a small tube, so the gas is movable state : carrier gas flowing through a column

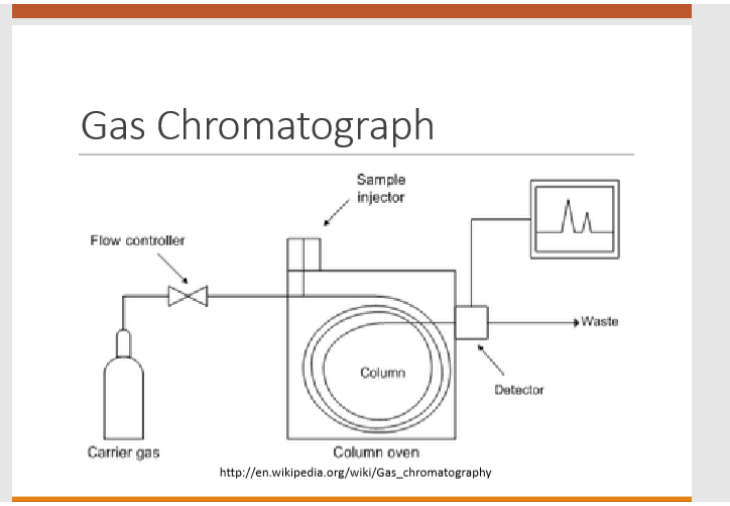
➔ thin film of liquid contained with the column is a stationary phase

➔After a mixture has traversed the length of the column, it will emerge separated into its components.

➔The written record of this separation is called a chromatogram.

➔The time required for a component to emerge from a GC column is known as retention time.

➔gas chromatograph can't be moved so we get the sample to it and not carry it to the crime scene(just for knowledge XD)



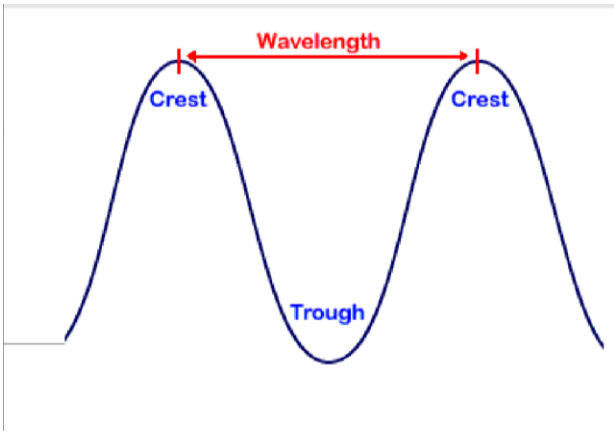
Theory of light

Light : continuous wave

White light disperses into a continuous spectrum of colors on passing through a prism

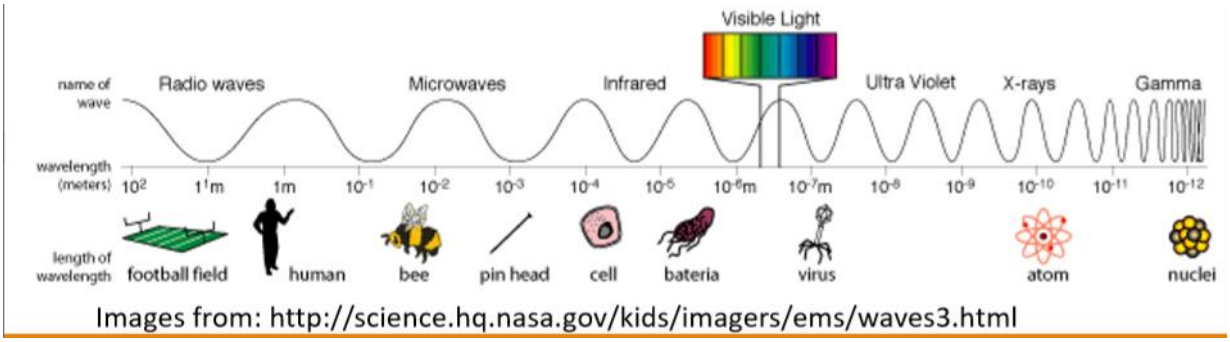
Waves are described in terms such as:

➔Wavelength, the distance between two successive crests (or one trough to the next trough).
➔Frequency, the number of crests (or troughs) passing any one given point per unit of time



➔The electromagnetic spectrum is the entire range of radiation energy from the most energetic cosmic rays to the least energetic radio waves.

➔Visible light is only a small part of the electromagnetic spectrum.



Spectrophotometry

➔Just as a substance can absorb visible light to produce color, many of the invisible radiations of the electromagnetic spectrum are likewise absorbed.

➔Spectrophotometry, an important analytical tool, measures the quantity of radiation that a particular material absorbs as a function of wavelength and frequency.

Spectrophotometer

The spectrophotometer is the instrument used to measure and record the absorption spectrum of a chemical substance.

The components of a spectrophotometer are:

➔ A radiation source

➔ A monochromator or frequency selector

➔ A sample holder ◦ A detector to convert electromagnetic radiation into an electrical signal

➔ A recorder to produce a record of the signal

Absorption spectra can be recorded for the visible, ultraviolet (UV) or infrared (IR) regions.

UV and IR Spectrophotometry

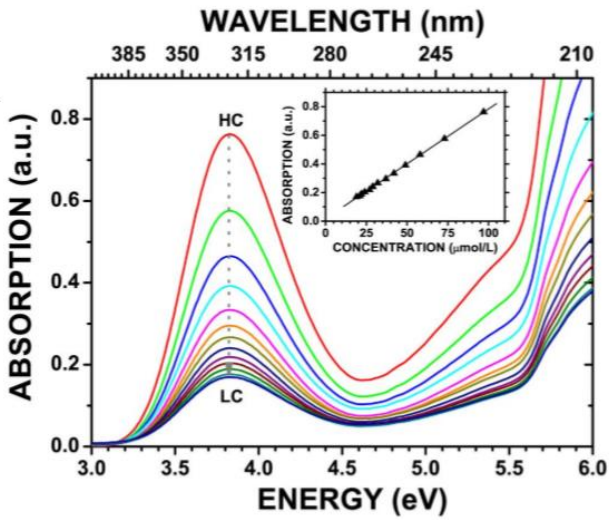
Currently, most forensic laboratories use UV and IR spectrophotometers to characterize chemical compounds.

The simplicity of the UV spectrum facilitates its use as a tool for determining a material's probable identity, although it may not provide a definitive result.

The IR spectrum provides a far more complex pattern.

Different materials always have distinctively different infrared spectra; each IR spectrum is therefore equivalent to an IR "fingerprint" of that substance.

UV Spectra



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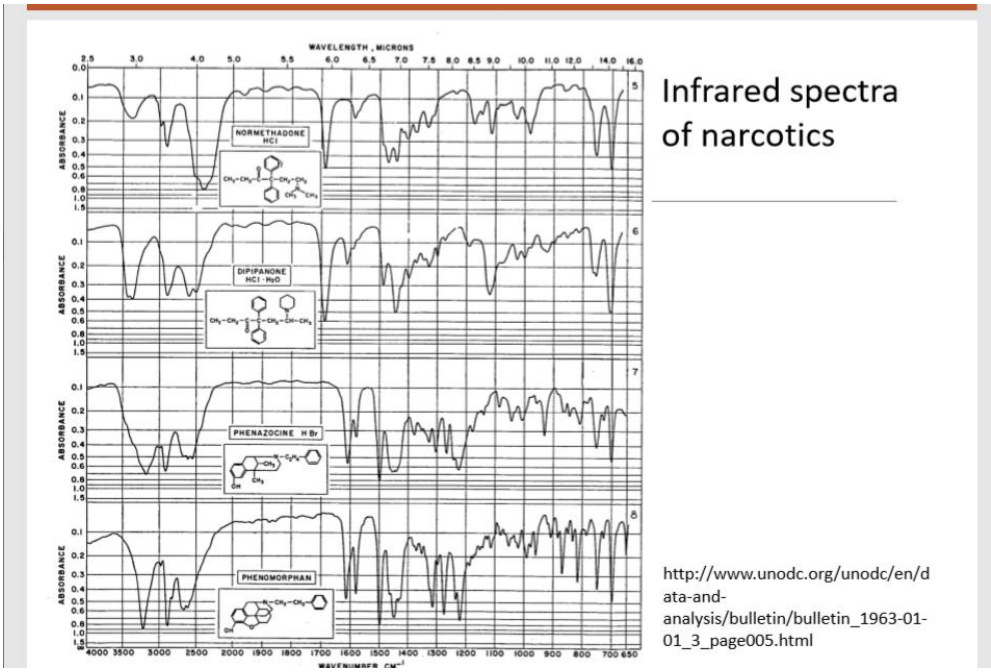
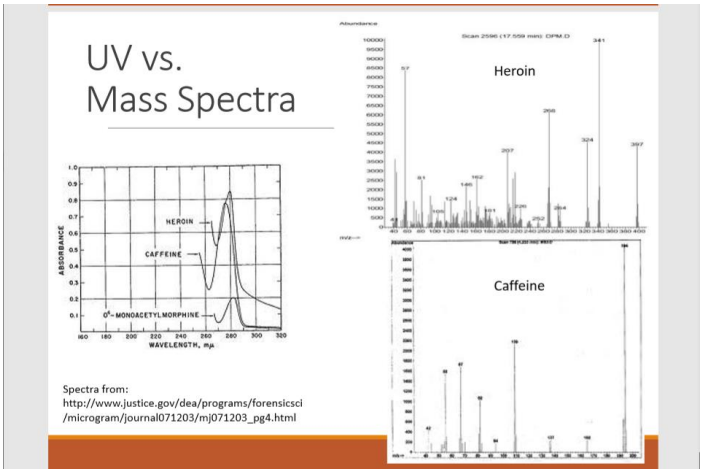


Mass Spectrometry

In the mass spectrometer, a beam of high-energy electrons collide with a material, producing positively charged ions.

These positive ions almost instantaneously decompose into numerous fragments, which are separated according to their masses.

The unique feature of mass spectrometry is that under carefully controlled conditions, no two substances produce the same fragmentation pattern.



GC and Mass Spectrometry

A direct connection between the GC column and the mass spectrometer allows each component to flow into the mass spectrometer as it emerges from the GC.

The separation of a mixture's components is first accomplished by the GC.

Then, fragmentation of each component by highenergy electrons in the mass spectrometer, will produce a distinct pattern, somewhat like a "fingerprint", of the substance being examined.