# Microscopic World I – Relative Atomic Mass

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| How to calculate the relative atomic mass. |
| Calculate the Relative Atomic Mass of W:     |  | | --- | |  | |  | |  | |  | |  | |
| Given that Element Y has 3 isotopes and the relative atomic mass of Y is 61.8. The relative abundance of 60Y is 30%.  Calculate the relative abundance of 62Y and 64Y   |  | | --- | |  | |  | |  | |  | |  | |  | |  | |  | |

# Microscopic World I – Ion

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| **The Force between the Electron and the Proton is Electrostatic Force**  Where it can be expressed as  When the distance between electron and the proton (r) increase, the electrostatic force decrease.   * When the number of occupied electron shell increase:   + The Attraction between the proton and the outermost electron is smaller     - It is easier to **loss one more** electron     - It is more difficult to **gain one more** electron * When the number of occupied electron shell decrease:   + The attraction between the proton and the outermost electron is greater     - It is easier to **gain one more** electron     - It is more difficult to **loss one more** electron |
| **How to predict an atom is reactive or not?**  GCSE CHEMISTRY - What is a Sodium Ion? - How do you Draw a Sodium Ion?-  What is the Electronic Structure of a Sodium Ion? - GCSE SCIENCE.   * Octet rule:   + There is a strong tendency of all atoms to attain the stable electronic arrangement of a noble gas, where there is 8 electrons in the outermost electron shell |
| **How to be stable – I**  By Lossing one or more electron, the ion with positive charge is formed = Cation  By Gaining one or more electron, the ion with negative charge is formed = Anion |