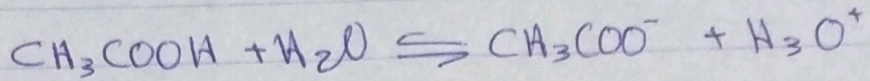


في قحيطه

والجبه كيمياء



هو = pH

بعد اضافة الماء  $[\text{CH}_3\text{COO}^-] = 1. \times 10^{-3}$  و  $[\text{H}_3\text{O}^+] = 1. \times 10^{-3}$

$$0.1 \times 1.1 \times 10^{-3} = \frac{1. \times 10^{-3} \times 1. \times 10^{-3}}{[\text{CH}_3\text{COOH}]} \rightleftharpoons \frac{[\text{CH}_3\text{COO}^-][\text{H}_3\text{O}^+]}{[\text{CH}_3\text{COOH}]} = K_a$$

$$1.1 \times 10^{-4} = \frac{1. \times 10^{-3}}{[\text{CH}_3\text{COOH}]} \rightleftharpoons [\text{CH}_3\text{COOH}] = 0.91 \times 10^{-3}$$

بعد اضافة الماء

$[\text{CH}_3\text{COO}^-]$  و اضافة الماء تؤدي الى زيادة

$[\text{H}_3\text{O}^+]$  تفر

pH تزداد

$$\frac{[\text{CH}_3\text{COOH}]}{[\text{CH}_3\text{COO}^-]} K_a = [\text{H}_3\text{O}^+]$$

$$\rightleftharpoons \text{هو} + \text{هو} = \text{pH} \rightleftharpoons$$

$$\frac{[\text{CH}_3\text{COO}^-][\text{H}_3\text{O}^+]}{[\text{CH}_3\text{COOH}]} = K_a \rightleftharpoons$$

$$1.1 \times 10^{-4} = \frac{1. \times 10^{-3} \times 1. \times 10^{-3}}{[\text{CH}_3\text{COOH}]} \rightleftharpoons [\text{CH}_3\text{COOH}] = 0.91 \times 10^{-3}$$

ن ل  $\text{HCOOH}$  = 1.1

1.1 x 10<sup>-3</sup> =

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الاجابة

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