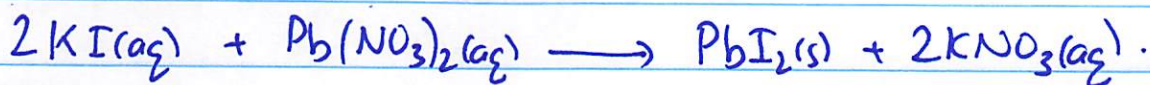


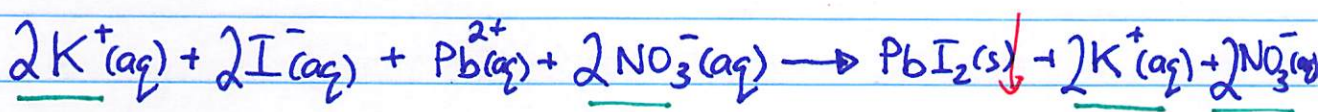
10/4/2019

Molecular eq



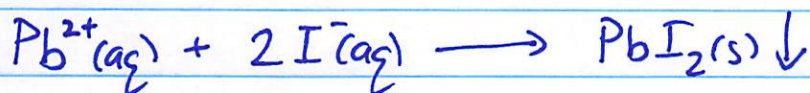
Full Ionic eq

- show separated ions for dissolved compounds (not the ppt)

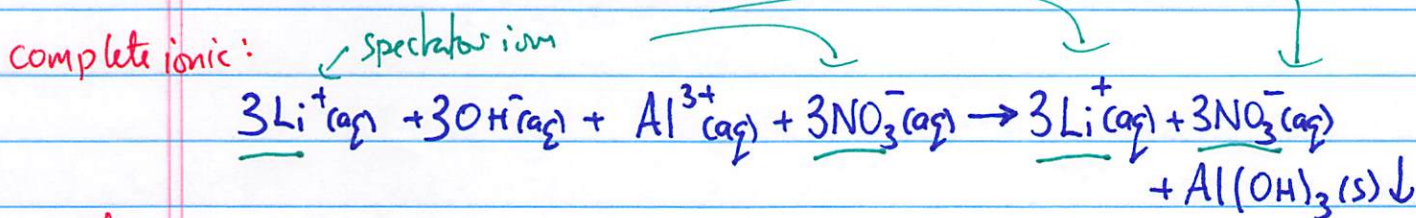
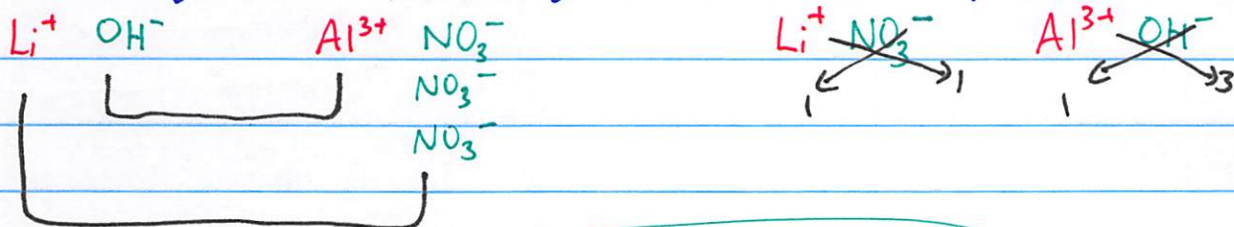
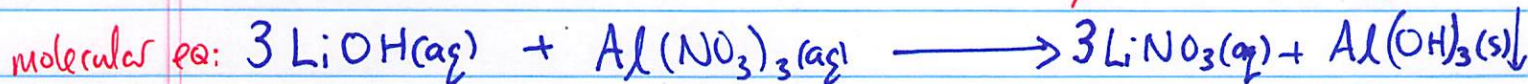


NET IONIC EQ

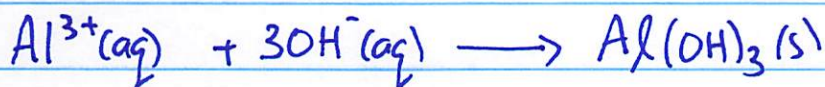
- remove the spectator ions $\text{K}^+, \text{NO}_3^-$



For you:



net-ionic:

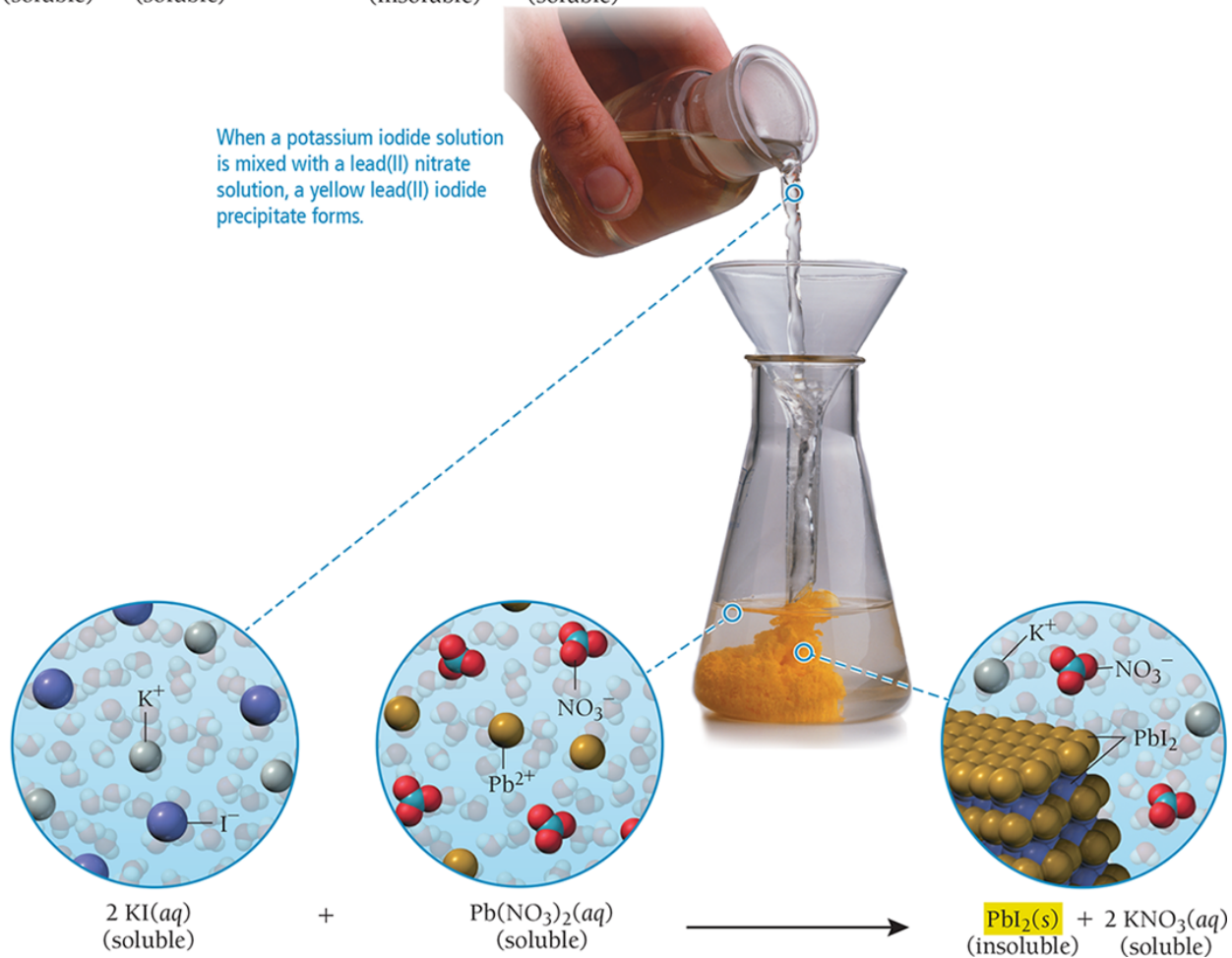


Precipitation of Lead (II) Iodide

Precipitation Reaction



When a potassium iodide solution is mixed with a lead(II) nitrate solution, a yellow lead(II) iodide precipitate forms.

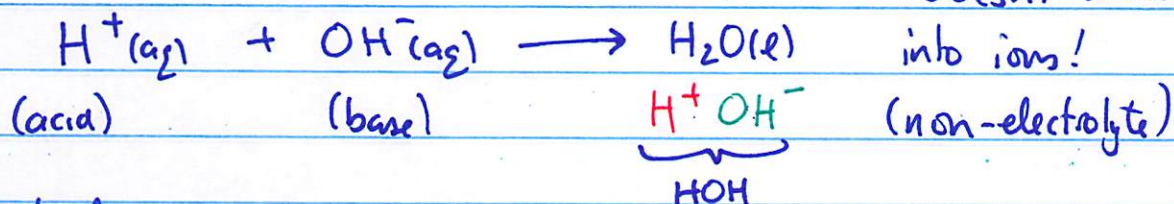


Acid-Base rxns

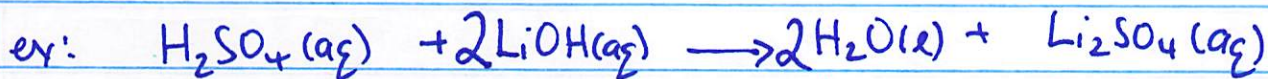
Arrhenius definition: $\begin{cases} \text{acids form } H^+ \text{ ions in water} \\ \text{bases " } OH^- \text{ " " " "} \end{cases}$

when they meet: (neutralization!)

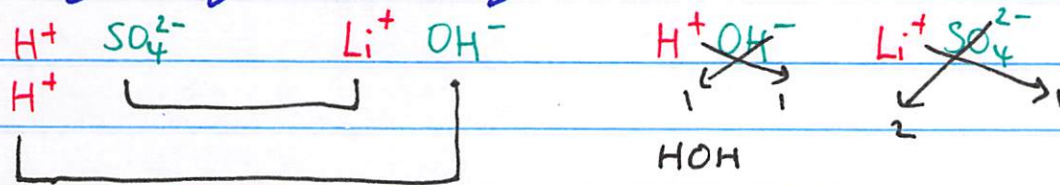
water, a molecule!
... doesn't break



molecular eq



diprotic acid
 $2H^+$ /molecule!



complete ionic:



net ionic

