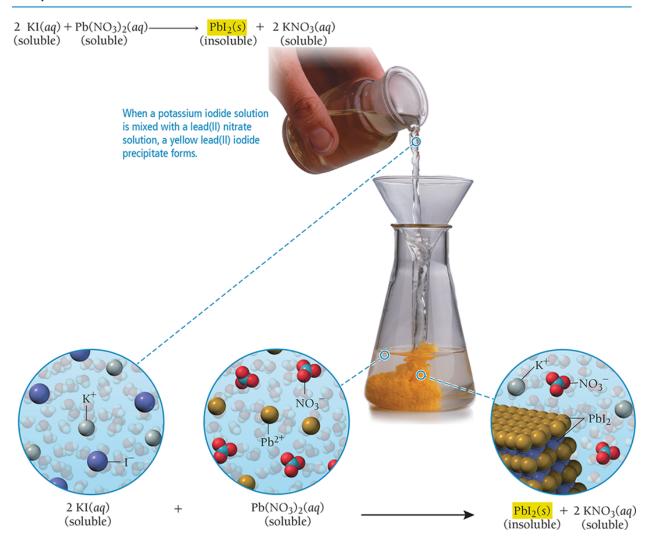
Molecular ea 10/4/2019 2 KI(ag) + Pb(NO3)2(ag) -> PbI2(s) + 2KNO3(ag). FULL IONIC Ca - show reparated ions for dissolved compounds (not the ppt) 2Ktaq) + 2I(aq) + Pbaq) + 2NO3(aq) -> PbIz(s) -1Ktaq)+)NQ(aq) NET IONIC EQ
- remove the spectator ins K+, NO3-Pb2+(ag) + 2 I (ag) -> Pb Iz(s) 1 for you: molernler ea: 3 LiOH(ag) + AL(NO3)3(ag) -> 3 LiNO3(ag) + AL(OH)3(s) Lit NO3 A134 OHT complete ionic: , specketor ion 3 Litar + 30 Hiar + Al 3tags + 3NO3 (ags -> 3 Litags + 3NO3 (ags) + A1(OH)2(5) L ret-ionic! A13+(aq) + 30+ (aq) -> Al(OH)3(5)

Precipitation of Lead (II) lodide

Precipitation Reaction





	Acid-Base rins
	Arrhenius definition: { acids form H+ ions in water bases "OH- "
	bases " OH " "
	when they need: (neutralization!) water, a molerule!
	doesn't break
	H+(ag) + OH(ag) -> H2O(e) into ion!
	(acia) (base) H+OH (non-electrolyte)
	HOH
	molecular ea
	ex: H2SO4(ag) + 2LiOH(ag) -> 2H2O(R) + Li2SO4(ag)
diprohic acid	H+ SO2- Lit OH- H+OH- Lit So2-
2H+/molecul	e! H+
	НОН
	complete ionic:
	2H+(ag) + Sou (ag) + 2Lit(ag) + 2OH (ag) -> 2H, U(e) + 2Lit(ag) + Sout(ag)
	net ionic
	2Htag) + 20Htag) -> 2H20(8).

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