Let's learn how to predict the sign of ΔS_{rxn} or ΔS_{o}

 $20_{9}(g) \longrightarrow 30_{2}(g)$ $\Delta n_{9}=3-2=+1$ increase in gas moler. $\Delta S^{\circ}=+\nu e$

 $N_2(g) + O_2(g) \longrightarrow 2NO(g) \Delta g = 2-2 = 0$ $\Delta S^{\circ} \approx 0$

 $2(06) + 026) \longrightarrow 2(026) \Delta n_3 = -1$ $\Delta S^\circ = -ve$

 $C_3H_8(g) + 5O_2(g) \longrightarrow 3(O_2(g) + 4H_2O(g))$ $\Lambda S^2 = -vr$ $2^{m}L_{ow}$: $\Delta S_{univ} \ge 0$ $\Delta S_{univ} = \Delta S_{mn} + \Delta S_{surr}$ $\Delta S_{univ} = \Delta S_{univ} + \Delta S_{surr}$ $\Delta S_{univ} = \Delta S_{univ} + \Delta S_{univ}$

How do we cale ΔS and ΔS_{sur} ? $S^{\circ}(P) - \Sigma S^{\circ}(R)$ Sourr = $\frac{1}{2}$ sur if we're Θ if we're Θ high?... a little bit of Qa little bit of Q makes a makes a large much smaller increase (Showhite in lib.)

$$\Delta S_{swr} = -\frac{q_{sys}}{T}$$

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N2(9) + 3H2(9)
$$\longrightarrow$$
 2NH3(9)
already saw: $\Delta S^{\circ} = -198.5 \text{ J/mol.K}$
however, $\Delta H^{\circ} = -92.6 \text{ KJ/mol}$
exothermic sm.
 $\Delta S_{univ} = \Delta S + \Delta S_{surr} = \Delta S - \Delta H$

criterion for whether a ron Defin: G = H-TS Gibb's Free Energy $\Delta G = \Delta H - T\Delta S$ (Tobes

if $\Delta G = 0$: @ com

△G>O: rxn is impossible (fund)

