

CHEMISTRY

CLASS – X

CHAPTER – 2

Acid, Base & Salt

Activity 1- Acids

Take some acid solution in a test tube. Add a few drops of litmus solution. It turns blue. Add a few drops of sodium hydroxide solution. The colour changes to red. This shows that acid reacts with base to form salt and water.

Activity 2- Bases

Take some base solution in a test tube. Add a few drops of litmus solution. It turns red. Add a few drops of hydrochloric acid solution. The colour changes to blue. This shows that base reacts with acid to form salt and water.

Activity 3- Reaction of Zinc with Hydrochloric Acid

Take some zinc metal in a test tube. Add some hydrochloric acid solution. Bubbles of gas are evolved. The gas burns with a pop sound. This shows that zinc reacts with hydrochloric acid to form zinc chloride and hydrogen gas.



Activity 4- Properties of Acids

Take some acid solution in a test tube. Add a few drops of litmus solution. It turns blue.

- Acid solution turns blue litmus red.
- Acid reacts with metal to form salt and hydrogen gas.
- Acid reacts with metal oxide to form salt and water.
- Acid reacts with metal hydroxide to form salt and water.
- Acid reacts with carbonate to form salt, water and carbon dioxide gas.

izu 5- {*lkj dsikp xqk dksfy/l*

*m*lkj & {*lkj dsxqk fuufyf/kr g%*

- (i) *buck Lokn dMek glrk g*
- (ii) ; *s l k t sfpdus glrsg rFlk Ropk dks glfu igpkrsg*
- (iii) ; *syky fyVel dksulyk dj nrsg*
- (iv) ; *sgYnh dsjæ dksHyk yky dj nrsg*
- (v) ; *svEyk dsl kfk fØ; k djdsyo. k rFlk ty cukrsg*

izu 6- {*lkj dsmi; lx erlo*

*m*lkj & {*lkj dsmi; lx fuufyf/kr g%*

- (i) *bl dk mi; lx l k t cukuseafd; k t lrk g*
- (ii) *blga {lkj; cVj; laeiz Ør fd; k t lrk g*
- (iii) *buck mi; lx iVky fjQlbfuæ vlf dkxt m/lx eaz Ør glrk g*
- (iv) *dBkj ty dksenqcukuseabl dk mi; lx fd; k t lrk g*

izu 7- *vEY rFlk {lkj eavaj Li "V dj*

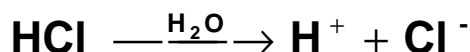
*m*lkj & *vEY rFlk {lkj eafuufyf/kr varj g%*

| <i>vEY</i> | <i>{lkj</i> |
|---|--|
| (i) <i>bl dk Lokn /KVk glrk g</i> | (i) <i>bl dk Lokn dMek glrk g</i> |
| (ii) ; <i>g uhysfyVel i= dksyky dj nrk g</i> | (ii) ; <i>g yky fyVel i= dksulyk dj nrk g</i> |
| (iii) ; <i>g ty eafosy glkj glbM u vk u (H⁺) nrk g</i> | (iii) ; <i>g ty eafosy glkj glbM lbM vk u (OH⁻) nrk g</i> |
| (iv) ; <i>g {lkj d dksmkl hu dj nrk g</i> | (iv) ; <i>g vEY dksmkl hu dj nrk g</i> |
| (v) <i>bl dkpH eku 7 l sde glrk g</i> | (v) <i>bl dkpH eku 7 l svfæcl glrk g</i> |

izu 8- *vljgsul }lkj nhxbZvEY rFlk HLe dhifjHk nk n*

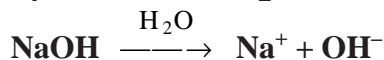
*m*lkj & *vEY* & *vEY* og *inkfZgSt ksty ea?lydj glbM u vk u (H⁺) inku djrk g*

t s s HCl, H₂SO₄, HNO₃, CH₃COOH *br k nA*



HLe & *HLe* og *inkfZgSt ksty ea?lydj glbM lbM (OH⁻) vk u nrk g*

t s s NaOH, KOH, NH₄OH, Ca(OH)₂ *br k nA*



izu 9- vk uhdj. k ds vlekj ij vEyladk foHkt u fdu oxkææfd; k t krk gS¹ o. kã dja

mũkj & vk uhdj. k ds vlekj ij vEyladk foHkt u nksoxkææfd; k t krk gS²

- (i) **izy vEy (Strong Acid)** – t k vEy ty ea?lydj yxHx i vLZ% vk fur ghdj gbmMt u vk u(H^+) izku dgrsg t s& gbmMdykjd vEy(HCl), ulbfVd vEy(HNO_3), l YÝ; fjd vEy(H_2SO_4) bR, kñA
- (ii) **nqZ vEy (Weak Acid)** – osvEy t k ty ea?lydj fl QZvL'kd : i l s vk fur gkrs gS ml snqZ vEy dgrsg t s& dkchud vEy(H_2CO_3), , l hfVd vEy(CH_3COOH), nqZ vEy gS chjd vEy(H_3BO_3) Hh, d nqZ vEy gS ft l dk mi; kx , Jh fVd ds: i ea gkrg gS

izu 10- foy; u eamifLFkr vEy dh ek=k ds vuq kj vEyladk foHkt u fdu oxkææfd; k t krk gS¹ o. kã dja

mũkj & fo; yu eamifLFkr vEy dh ek=k ds vuq kj vEyladk nksoxkææfd; k t krk gS²

- (i) **l k vEy (Concentrated Acid)** – t c foy; u eavEy dh vfkd ek=k mifLFkr jgrh gS rks ml s l k vEy dgrsg l k vEy eat y dh ek=k de jgrh gS
- (ii) **ruqvEy (Dilute Acid)** – t c foy; u eavEy dh ek=k de jgrh gS rks ml s ruqvEy dgrsg ruqvEy eat y dh ek=k vfkd jgrh gS

izu 11- vEy rFlk {kij ds vjgfu; l fl) kR dh l hekvladk mYysk dja mũkj & vjgfu; l fl) kR ds nk'k fuEufyf[kr gS²

- (i) bl fl) kR ds vuq kj vEy H ; Ør ; kxd gS vkS {kij OH ; Ør ; kxd gS yfdu dN, l s {kij gft ueOH ugh jgrk bu vEylarFlk {kij dh Q kj; k bl fl) kR ds vlekj ij ugh dh t k l drhA
- (ii) æo vekfu; ke NH_4NO_3 dk vlpj. kvEylt gkrg gS bl dh Q kj; k bl fl) kR ds vlekj ij ugh dh t k l drhA
- (iii) tyh foy; u eaHCl rks vEy ekuk t krk gS fda rqs l vofLFk ea; k vl foyk d t s& ea hu/ea vEy ugh ekuk t krhA

izu 12- vk uhdj. $k ds vlekj$ ij $HLeakdsfdrushkx$; $kg\%$ o. kz dja
 $mukj \& vk$ uhdj. $k ds vlekj$ ij $HLeakdsnksoxkx$; $kg\%$

- (i) **izy HLe (Strong base)**- $osHLe$ tkt ylt foy; u eai vk fur
 $gkdj$ $dkQh$ $ek=k$ $eaglbM$ lbM vk u (OH^-) inku $djrs g$ $ml s$
 $izy H'e$; $k izy$ $\{kij$ $dgrsg$ tS $NaOH$, KOH $izy H'e$ g
(ii) **ngZ HLe (Weak base)** & $osHLe$ tkt ylt foy; u eaf QZ $vakr\%$
 vk fur $gkdj$ de $ek=k$ $eaglbM$ lbM (OH^-) inku $djrs g$ $ml s$
 $ngZ HLe$; $k ngZ$ $\{kij$ $dgyksg$ tS $\&veku$; e $glbM$ lbM
 (NH_4OH) , $dS'k$ e $glbM$ lbM $Ca(OH)_2$ |

izu 13- **l pd (Indicator)** fdl $sdgrsg$; $sfdrusizlkj$ $dsgksg$ $ifjH'$ kr dja
 $mukj \& l$ pd , s $inkfZ$ $gksg$ t ks $viusj$ $ifjorZ$ ds $\{kij$ $inkfZ$ ds $vEyl$
; k $\{kij$; k mkl lu $gkudh$ l puk $nsrg$
 rhu l $lekl$ l pd $fyVel$ $i=$ $feFkby$ vkj rFk $Qhuk$ $QFkyu$ g
 l pd dk $vEyl$ $\{kij$ l pd Hh $dgrsg$
; $snksizlkj$ $dsgksg\%$

- (i) **ikNfrd l pd** & **ikNfrd l pd** ds $varxZ$ $fyVel$ $i=$ rFk $gYnh$ $vkrs$
 g $pqlhj$ yky $xkHh$ i UkA
(ii) **layf'kr l pd** & bl ds $varxZ$ $feFkby$ vkj rFk $fQuk$ $QFkyu$ g

izu 14- **xg fufeZ l pd** vki dS s cuk xs |

$mukj$ $\& gYnh$ ds NkV $\& NkV$ VqM $dkst$ y ds l kFk xje dj Nku $ysrg$ $gYnh$
 l $siHr$ foy ; u dk nk $vyx \& vyx$ ij $[kufy$; $keay$ $ysrg$, d ij $[kuyh$
 eal kcq dk foy ; u $Myrsg$ ij $[kuyh$ ds foy ; u dk i lyk j $ky \& Hjis$
 j $acny$ t $krkgs$ ft l l sl kcq ds $\{kij$ $gkudh$ $igplu$ $gkrg$ g $nWjs$
 ij $[kuyheaf$ j dk $Myrsg$ fl j dk ij $[kuyheaf$ $[k$ foy ; u ds j $eadk$ bz
 $ifjorZ$ ugl $djrka$ $vr\%$ fl j dk $vEyl$ g

izu 15- $vEyl$ dh $'kDr$ ds $ckjs$ $eavki$ D ; k t $kursg$ fd $u \& fd$ u $rjhd$ l s
 $vEyl$ dh $'kDr$ dh $ryuk$ dh t k l drh gs |

$mukj$ $\& vEyl$ dst yht foy ; u $eavEyl$ $\{kij$ iHr $glbM$ u vk $uladh$ $ek=k$ l s
 ml dh $vEyl$ $'kDr$ dk $fuekl$ k $gkrg$ g

निम्न तरीकों से अम्लों की तुलना की जा सकती है—

- (i) **$vEyl$ ds kr ds $vlekj$** & tS $\&dkZud$; kxd & $ilkl$ l $siHr$ $vEyl$
& l kb Vel $vEyl$ $vkut$ Syd $vEylA$

- (ii) *vl. od l'apuk ds vleljij ij vEj & glbM vEj & HCl, HBr, HI*
 (iii) *izy vEj ds vleljij ij & vEj t ylr foy; u esiwlz%vk fur glrk gSft l ds vk uhclj. k dk vak yxHx 100% glrk gA*

izu 16- yo. k fdl sdgrsgA ; sfdrusizlj dsgrsgA ifjHf'kr djA mUlj & yo. kos; l'xd gSft udkfuekzkfdl hvEj dkfdl h{kljd dsvfHfO; k dsQyLo: i glrk gS, oaf l eavEj v. kqds; l'xd eami fLFkr glbM u i jek kqfdl h ekkq}kjk foLFHf'ir glrk gA

mnl hulclj. k vfhfO; k eadkZvEj fdl h{kljd dsl kfk vfhfO; k dj yo. k, oat y cukrk gA

tS s&l kM, e glbM kM, oaglbMdykjd vEj dsmnl hulclj. k vfhfO; k eal kUlj. k ued yo. k, oat y cukrk gA



yo. k ds izlj fuEufyf'kr gA&%

- (i) *l'ekl' yo. k (Normal Salt) & og yo. k ft l l svk ulNfr H i jek kq; k glbM d l y l eg ughajgrk gS ml sl'ekl' yo. k dgrsgA ; g vEj , oafHle dsiwlzmnl hulclj. k dsQyLo: i curk gA tS NaCl, HCl, NaNO₃, Na₂SO₄ br' knA*

- (ii) *vEylr yo. k (Acidic Salt) & fdl h vEj ds v. k eami fLFkr foLFHf'ir u ; l' glbM u i jek kqdkkkq}kjk vakr%foLFHf'ir dju dsQyLo: i cus yo. k dks vEylr yo. k dgrsgA*

vFlok

os yo. k t k fdl h Hle }kjk fdl h vEj ds viwlzmnl hulclj. k dsQyLo: i curs gA ml svEylr yo. k dgrsgA bl eafLFHf'ir u ; l' glbM u glrk gA tS NaHSO₄, KHSO₄, br' kn A

- (iii) *Hlelr yo. k (Basic Salt) & osHle ft u ds v. k ea, d l svfekl OH l eg glrk gA vEylr }kjk vl'xd : i l smnl hu glclj Hf'led yo. k cukrk gA tS s&Pb(OH)NO₃. bl eafLFHf'ir u ; l' glbM kM ewd glrk gA*

izu 17- pH ds vleljij ij yo. k adk oxhclj. k djA

mUlj & vEj vl' Hle dhiNfr ; k pH vleljir yo. k ds foy; u rlu rjg ds glrk gA&%

- (i) *mnl hu yo. k foy; u (Natural Salt Solution) & izy vEj , oaizy Hle ds yo. k foy; u mnl hu glrk gA buclpH eku 7 glrk gA ; s*