DPP-03

1. The double bond equivalent of C_4H_6 is:

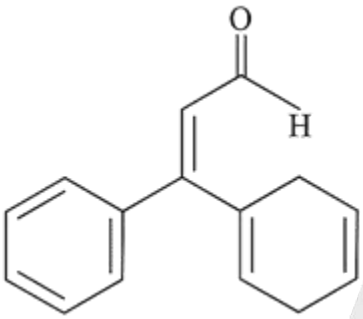
(A) 0

(B) 1

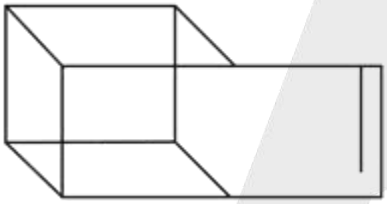
(C) 2

(D) 3

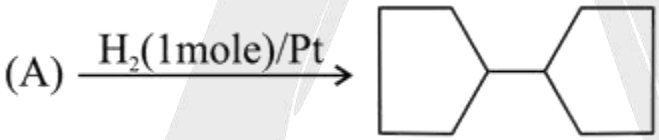
2. Find the double bond equivalent of the given compound.



3. Double bond equivalent of given compound is:



4. Double bond equivalent (degree of Unsaturation) of (A) is:



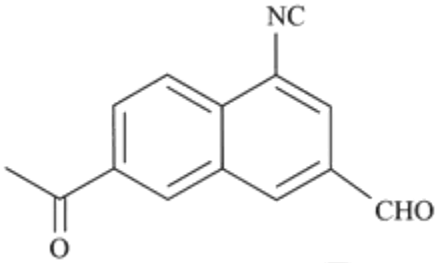
(A) 1

(B) 2

(C) 3

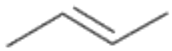
(D) 4

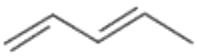
5. Double bond equivalent of the following is:

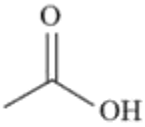


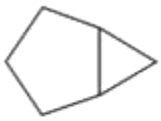
(A) 7 (B) 11 (C) 6 (D) None of these

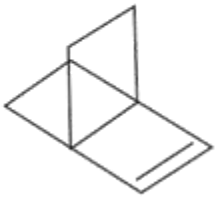
6. Look at the chemical structure below and calculate the DBE.

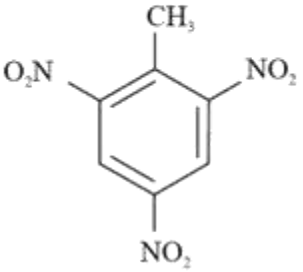
(1) 

(2) 

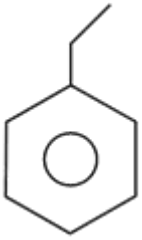
(3) 

(4) 

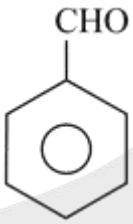
(5) 

(5) 

7. Which of following compound. has D.B.E is 5 :

(A) 

(B) 

(C) 

(D) Both (B) & (C)

8. Which of the following statements applies to $C_{10}H_{14}O_2$ compound?

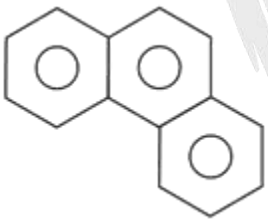
(A) It may have 2 double bonds and 2 rings.

(B) It may have 3 double bond and Oxygen ring.

(C) It may have 1 triple bond and 2 rings.

(D) It may have zero double bond and 3 rings

9. The difference in Double Bond Equivalent (DBE) value between

 and 

(A) 0 (B) 1 (C) 2 (D) 3

ANSWER KEY

1. (C) 2. (9) 3. (6) 4. (C) 5. (B) 7. (D)

8. (A,B, C) 9. (A)