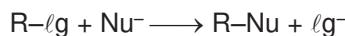


## Nucleophilic substitution reaction ( $S_N$ ) :

Replacement (displacement) of an atom or group by an other atom or group in a molecule is known as substitution reaction. If substitution reaction is brought about by a nucleophile then it is known as nucleophilic substitution reaction. Generally substitution takes place at  $sp^3$  carbon.



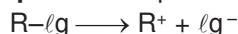
### Types of nucleophilic substitution reaction :

(I)  $S_N1$       (II)  $S_N2$       (III)  $S_Ni$

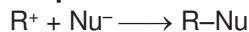
## Section (A) : Unimolecular nucleophilic substitution reaction ( $S_N1$ )

Nucleophilic substitution which involves two step process

(a) **First step** : Slow step which involves ionisation (to form carbocation)

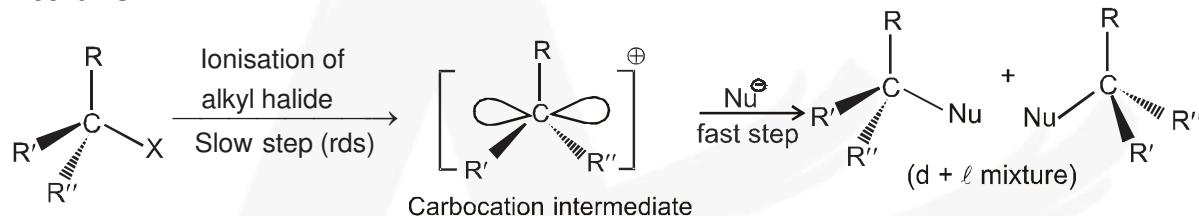


(b) **Second step** : Fast attack of nucleophile on carbocation results into product.



### (1) $S_N1$ Reaction of Alkyl halide

#### Mechanism :



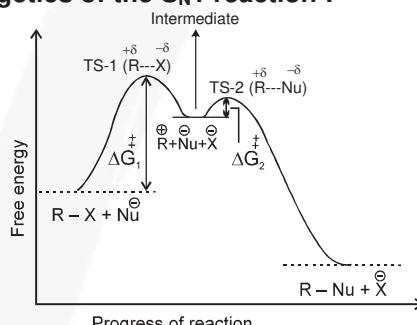
#### Characteristics of $S_N1$ reactions :

1. It is unimolecular, two step process.
2. Carbocation intermediate is formed, so rearrangement is possible in  $S_N1$  reaction.
3. It is first order reaction.

#### 4. Kinetics of the reaction      Rate $\propto$ [Alkyl halide]

Rate of  $S_N1$  reaction is independent of concentration and reactivity of nucleophile.

#### 5. Energetics of the $S_N1$ reaction :



**Figure :** Free energy diagram for the  $S_N1$  reaction.

#### 6. Factors affecting the rate of $S_N1$ reaction :

(i) **The structure of the substrate :** The rds of the  $S_N1$  reaction is ionization step, a carbocation is formed in this step. This ionisation is strongly endothermic process, rate of  $S_N1$  reaction depends strongly on carbocation stability because carbocation is the intermediate of  $S_N1$  reaction which determines the energy of activation of the reaction.

Reactivity of  $S_N1 \propto$  stability of carbocation.

**$S_N1$  reactivity :**  $3^\circ > 2^\circ > 1^\circ > \text{CH}_3-X$

**(ii) Concentration and reactivity of the nucleophile :** The rate of S<sub>N</sub>1 reaction is unaffected by the concentration and nature of the nucleophile.

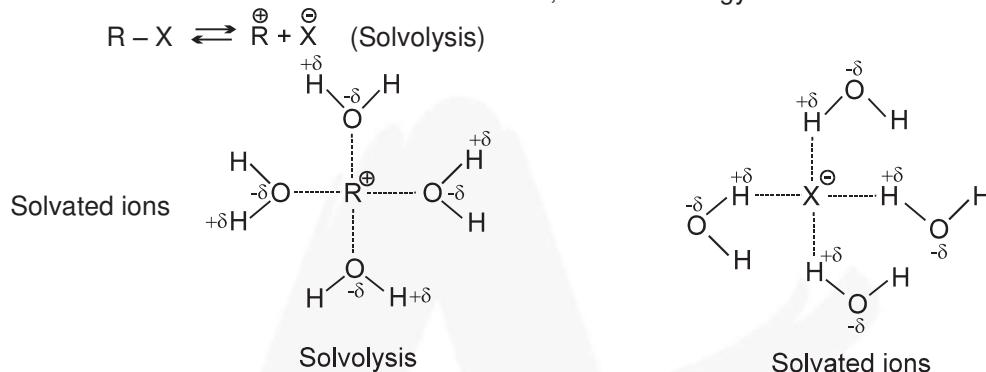
Weak and neutral nucleophile favours S<sub>N</sub>1 reaction.

Mostly solvents (protic) itself functions as nucleophiles in S<sub>N</sub>1 reaction, so S<sub>N</sub>1 reaction is termed as solvolysis reaction.



**(iii) Effect of the solvent :** (Ionising ability of the solvent)

The use of a polar protic solvent will greatly increase the rate of ionisation of an alkyl halide in any S<sub>N</sub>1 reaction because it solvate cations and anions so effectively and stabilises the transition state leading to the intermediate carbocation and halide ion, thus the energy of activation is lower.



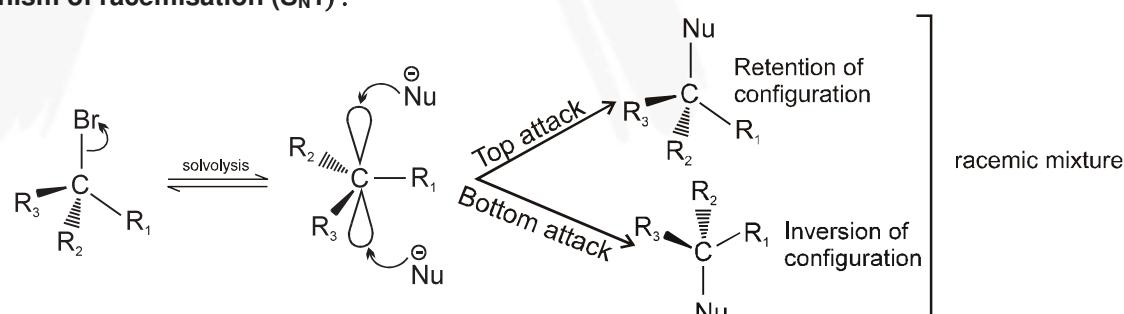
**Table :** Dielectric constants ( $\epsilon$ ) and ionisation rate of t-Butylchloride in few common solvents

Solvent	$\epsilon$	Relative rate
H <sub>2</sub> O	80	8000
CH <sub>3</sub> OH	33	1000
C <sub>2</sub> H <sub>5</sub> OH	24	200
(CH <sub>3</sub> ) <sub>2</sub> CO	21	1
CH <sub>3</sub> CO <sub>2</sub> H	6	—

**(iv) The nature of the leaving group :** In the S<sub>N</sub>1 reaction the leaving group begins to acquire a negative charge as the transition state is reached` stabilisation of this developing negative charge at the leaving group stabilises the transition state and this lowers the free energy of activation and there by increases the rate of reaction. Leaving ability of halogen is F<sup>-</sup> < Cl<sup>-</sup> < Br<sup>-</sup> < I<sup>-</sup>

**7. Stereochemistry of S<sub>N</sub>1 reactions :** In the S<sub>N</sub>1 mechanism, the carbocation intermediate is sp<sup>2</sup> hybridized and planar, A nucleophile can attack on the carbocation from either face, if reactant is chiral then attack of nucleophile from both faces gives enantiomers as the product, which is called racemisation.

**Mechanism of racemisation (S<sub>N</sub>1) :**



Reagents for alkyl halide are : H<sub>2</sub>O , RCOOH , ROH & RSH

## (2) S<sub>N</sub>1 Reaction of Alcohols

### (i) Reaction with hydrogen halides

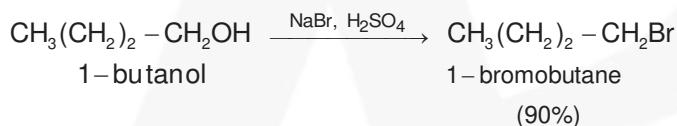
A common method is to treat the alcohol with a hydrohalic acid, usually HI or HBr. These acids are used to convert alcohols into the corresponding alkyl halides.

(i) In acidic solution, an alcohol is in equilibrium with its protonated form. Protonation converts the hydroxy group from a poor leaving group ( $\text{OH}^\ominus$ ) into a good leaving group ( $\text{H}_2\text{O}$ ). If the alcohol is protonated all the usual substitution and elimination reactions are feasible, depending on the structure ( $1^\circ$ ,  $2^\circ$ ,  $3^\circ$ ) of the alcohol.

(ii) Halides are anions of strong acids, so they are weak bases. Solutions of HBr and HI contain nucleophilic  $\text{Br}^-$  and  $\text{I}^-$  ions.

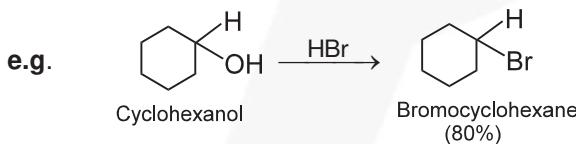
(iii) Concentrated hydrobromic acid rapidly converts t-Butyl alcohol to t-Butyl bromide. The strong acid protonates the hydroxyl group, converting it into a good leaving group. The hindered tertiary carbon atom cannot undergo  $S_N2$  displacement, but it can ionise to a tertiary carbocation. Attack by bromide ion gives the alkyl bromide. The mechanism is similar to  $S_N1$  mechanism.

(iv) 1-Butanol reacts with sodium bromide in concentrated sulfuric acid to give 1-Bromobutane by an  $S_N2$  displacement.



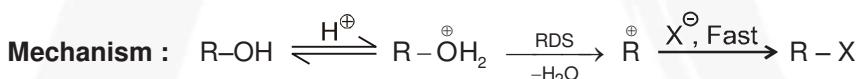
Protonation converts the hydroxy group to a good leaving group, but ionization to a primary carbocation is unfavourable. The protonated unbranched primary alcohol is well suited for the  $S_N2$  displacement.

(v) Secondary alcohols also react with HBr to form alkyl bromides usually by the S<sub>N</sub>1 mechanism.



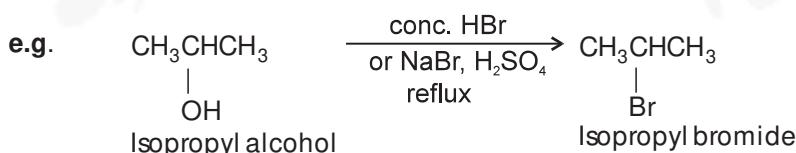
(vi) HCl (Hydrochloric acid) reacts with alcohols in much the same way that as the hydrobromic acid.

(vii) Chloride ion is a weaker nucleophile than bromide ion because it is smaller and less polarizable. Lewis acid, such as  $ZnCl_2$ , is sometimes necessary to promote the reaction of HCl with primary and secondary alcohols.



**Reactivity of HX :** HI > HBr > HCl

**Reactivity of ROH :**       $3^\circ > 2^\circ > 1^\circ$



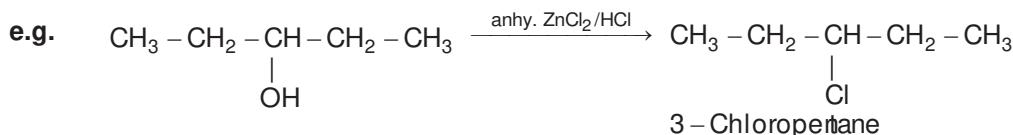
## **Lucas Reagent**

- (i) A mixture of concentrated hydrochloric acid and anhydrous zinc chloride is called the Lucas reagent.
  - (ii) Whether an alcohol is primary, secondary or tertiary is identified by the Lucas test, which is based upon the difference in reactivity of the three classes of alcohol towards hydrogen halides.
  - (iii) Alcohol (of not more than six carbons in their molecule) are soluble in the Lucas reagent. The corresponding alkyl chlorides are insoluble.
  - (iv) Formation of a chloride from an alcohol is indicated by the cloudiness that appears when the

# (Organic Chemistry) SUBSTITUTION ELIMINATION

chloride separates from the solution hence, the time required for cloudiness to appear is a measure of the reactivity of the alcohol.

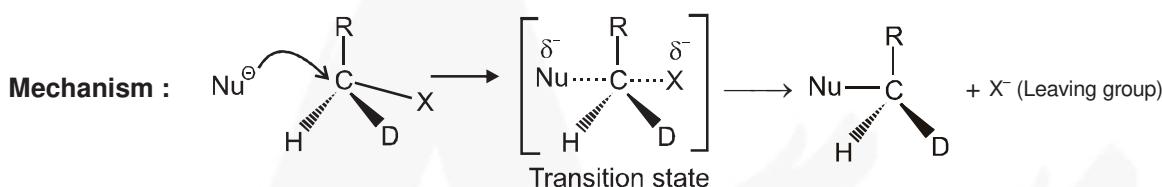
(v) A tertiary alcohol reacts immediately with the Lucas reagent, a secondary alcohol reacts within five minutes and a primary alcohol does not react appreciably at room temperature.



## Section (B) : Bimolecular nucleophilic substitution reaction ( $S_N2$ )

Nucleophilic substitution in which incoming group replaces leaving group in one step only.

### (1) $S_N2$ Reaction of Alkyl halide :



#### Characteristic of $S_N2$

1. It is bimolecular, one step concerted process
2. It is second order reaction because in the rds both species are involved
3. Kinetics of the reaction : rate  $\propto$  [alkyl halide] [nucleophile]  
rate =  $k[\text{alkyl halide}][\text{nucleophile}]$

If the concentration of alkyl halide in the reaction mixture is doubled, the rate of the nucleophilic substitution reaction is double. If the concentration of nucleophile is doubled the rate of reaction is also double. If the concentration of both are doubled then the rate of the reaction quadruples.

#### 4. Energetics of the reaction :

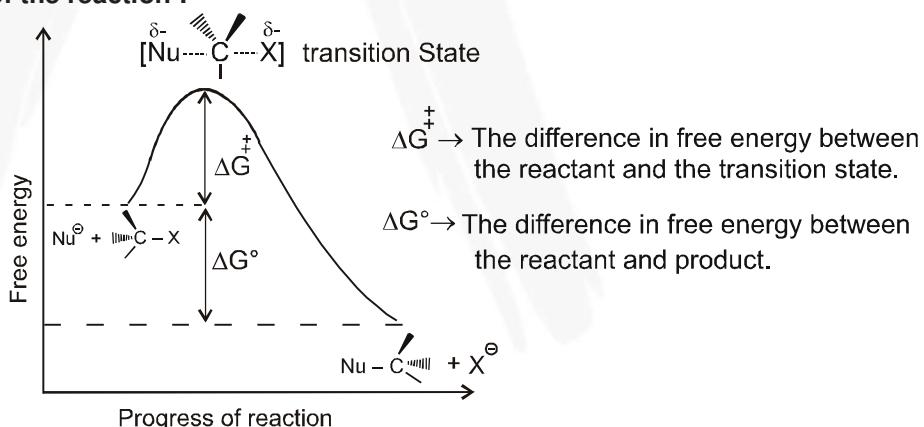
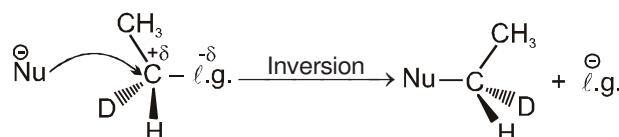


Figure : A free energy diagrams for  $S_N2$  reaction

5. No intermediates are formed in the  $S_N2$  reaction, the reaction proceed through the formation of an unstable arrangement of atoms or groups called transition state.
6. The stereochemistry of  $S_N2$  reaction : As we seen earlier, in an  $S_N2$  mechanism the nucleophile attacks from the back side, that is from the side directly opposite to the leaving group. This mode of attack causes an inversion of configuration at the carbon atom that is the target of nucleophilic attack. This inversion is also known as **Walden inversion**.





**7. Factor's affecting the rate of S<sub>N</sub>2 reaction :** Number of factors affect the relative rate of S<sub>N</sub>2 reaction, the most important factors are

**(i) Effect of the structure of the substrate :**



The important reason behind this order of reactivity is a steric effect. Very large and bulky groups can often hinder the formation of the required transition state and crowding raises the energy of the transition state and slow down the rate of reaction.

**Table :** Relative rate of reaction of alkyl halides by S<sub>N</sub>2 mechanism.

Substituent	Compound	Relative rate
Methyl	CH <sub>3</sub> X	30
1°	CH <sub>3</sub> CH <sub>2</sub> X	1
2°	(CH <sub>3</sub> ) <sub>2</sub> CHX	0.02
Neopentyl	(CH <sub>3</sub> ) <sub>3</sub> CCH <sub>2</sub> X	0.00001
3°	(CH <sub>3</sub> ) <sub>3</sub> CX	~ 0

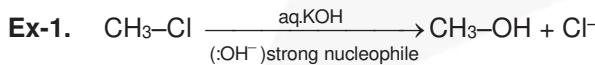
**(ii) Concentration and reactivity of the nucleophile :**

- As nucleophilicity of nucleophile increases rate of S<sub>N</sub>2 increases.
- Anionic nucleophiles mostly give S<sub>N</sub>2 reaction
- A stronger nucleophile attacks upon α-carbon with faster rate than the rate of departing of leaving group.

**(iii) The effect of the solvent :** Polar aprotic solvent have crowded positive centre, so they do not solvate the anion appreciably therefore the rate of S<sub>N</sub>2 reactions increased when they are carried out in polar aprotic solvent.

**(iv) The nature of the leaving group :** Weaker bases are good leaving groups. A good leaving group always stabilise the transition state and lowers its free energy of activation and there by increases the rate of the reaction. Order of leaving ability of halide ion F<sup>-</sup> < Cl<sup>-</sup> < Br<sup>-</sup> < I<sup>-</sup>

**Reagents for alkyl halide are :** OH<sup>-</sup>, SH<sup>-</sup>, I<sup>-</sup>, CN<sup>-</sup>, NH<sub>3</sub> ..... (strong anionic nucleophiles)



**(2) S<sub>N</sub>2 Reaction of Alcohol :**

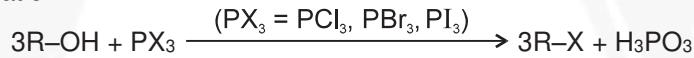
**(i) Reaction with HX :** The protonated β unbranched primary alcohol is well suited for the S<sub>N</sub>2 reaction.



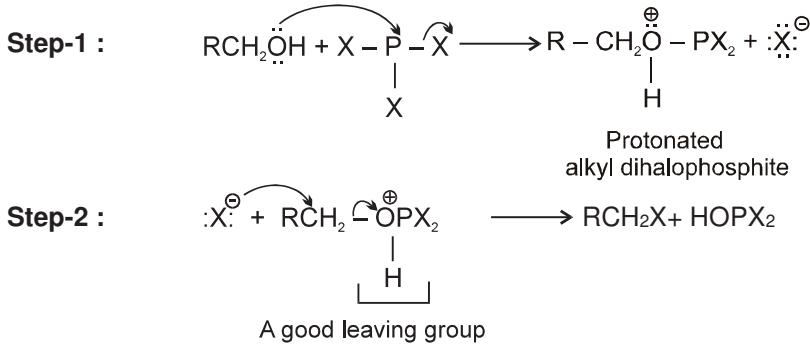
**(ii) Reaction with phosphorus trihalides**

Several phosphorus halides are useful for converting alcohols to alkyl halides. PBr<sub>3</sub>, PCl<sub>3</sub>, & PCl<sub>5</sub> work well and are commercially available.

Phosphorus halides produce good yields of most primary and secondary alkyl halides, but none works well with tertiary alcohols. The two phosphorus halides used most often are PBr<sub>3</sub> and the P<sub>4</sub>/I<sub>2</sub> combination.



**Mechanism :**



## Remarks

The mechanism for the reaction involves attack of the alcohol group on the phosphorus atom, displacing a halide ion and forming a protonated alkyl dihalophosphite

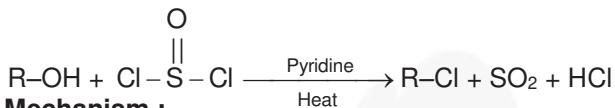
In second step a halide ion acts as nucleophile to displace  $\text{HOPX}_2$ , a good leaving group due to the electronegative atoms bonded to the phosphorus.

### (iii) Reaction with $\text{PCl}_5$

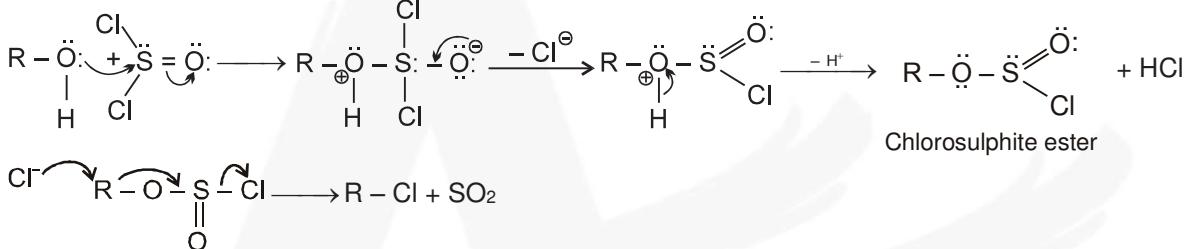


**(iv) Reaction with thionyl chloride in presence of pyridine**

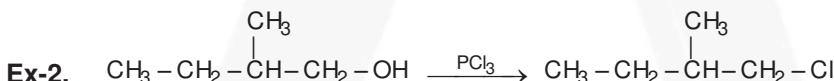
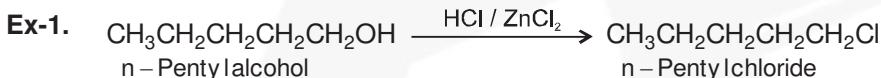
Thionyl chloride ( $\text{SOCl}_2$ ) is often the best reagent for converting an alcohol to an alkyl chloride. The by products (gaseous  $\text{SO}_2$  and  $\text{HCl}$ ) leave the reaction mixture and ensure that there can be no reverse reaction.



## Mechanism :



In the first step, the nonbonding electrons of the hydroxy oxygen atom attack the electrophilic sulphur atom of thionyl chloride. A chloride ion is expelled a proton and gives test of chloro sulphite ester. Second step is an  $S_N2$  mechanism.

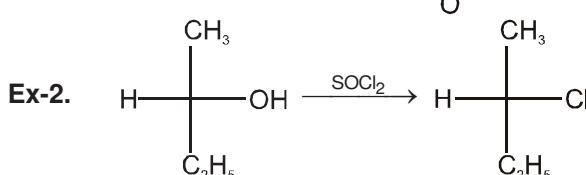
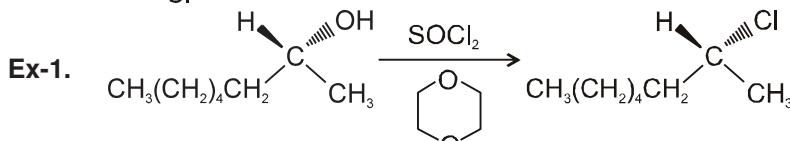
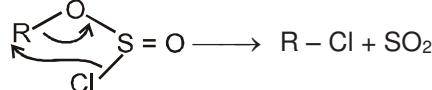
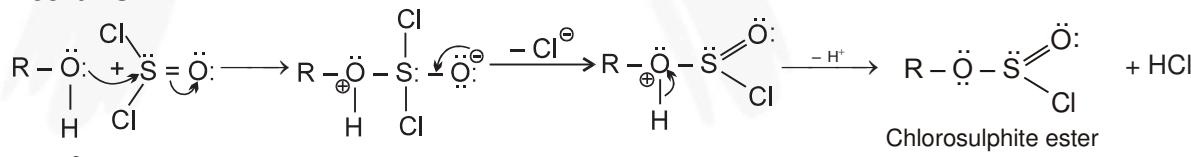


### (3) S<sub>N</sub>i Reaction :

In S<sub>i</sub>Ni mechanism an internal nucleophile attacks from the same side of leaving group, means retention of configuration. It is an S<sub>i</sub>Ni mechanism, where i means internal

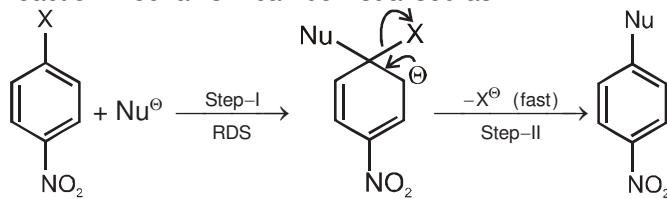


### Mechanism :

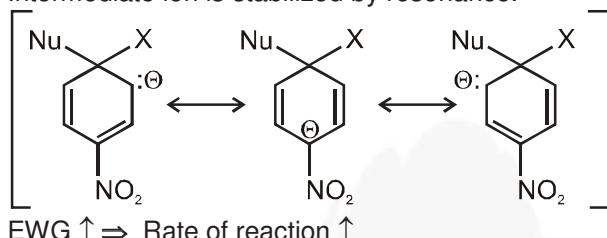


**Section (C) : Bimolecular aromatic nucleophilic substitution reaction ( $S_N2 Ar$ )**

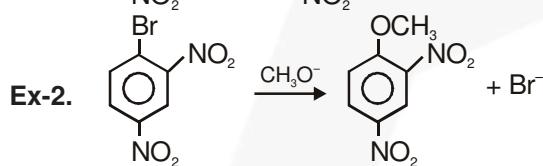
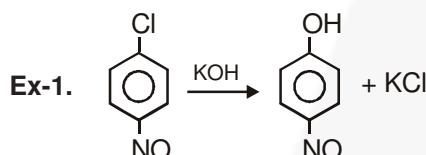
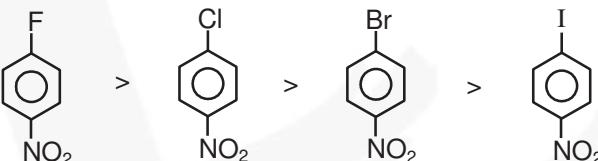
This is the characteristic reaction of arylhalides with ortho or para electron withdrawing substituent. The reaction mechanism can be visualised as :



Intermediate ion is stabilized by resonance.

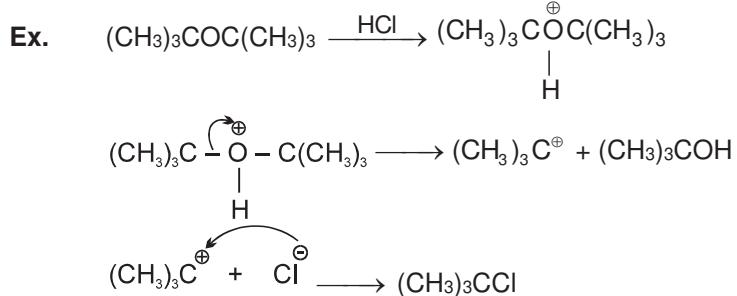
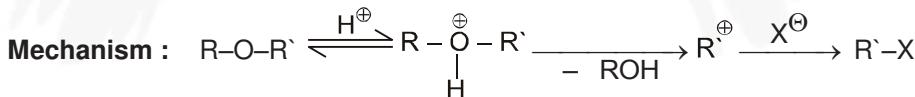


**Reactivity order towards  $S_N2Ar$  :**

**Section (D) : Nucleophilic substitution reaction of Ethers & Epoxides****(1)  $S_N1$  Reaction of Ethers****(i) Reaction with HX :**

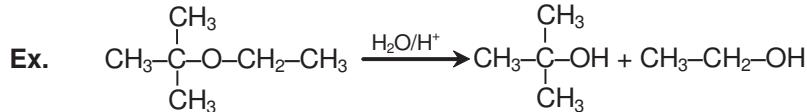
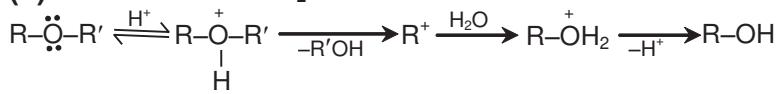
Ethers are unreactive towards most bases, but they can react under acidic conditions. A protonated ether can undergo substitution or elimination with the expulsion of an alcohol. Ethers react with conc. HBr and HI because these reagents are sufficiently acidic to protonate the ether, while bromide iodide are good nucleophiles for the substitution.

If R or R' is  $3^\circ$  then mechanism will be  $S_N1$  otherwise  $S_N2$ .



# (Organic Chemistry) SUBSTITUTION ELIMINATION

### (ii) Reaction with $\text{H}_2\text{O}/\text{H}^+$ :

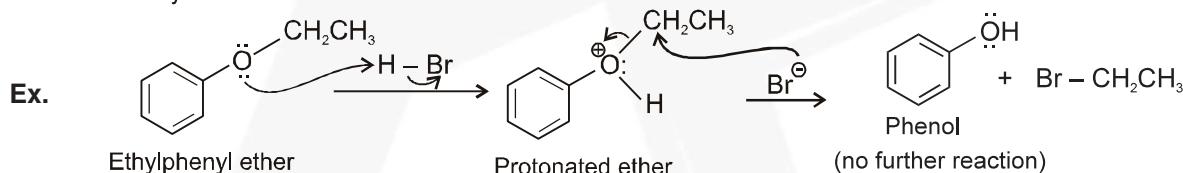
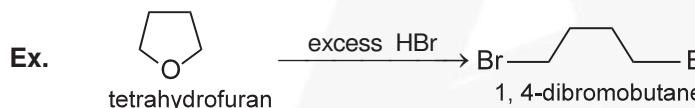
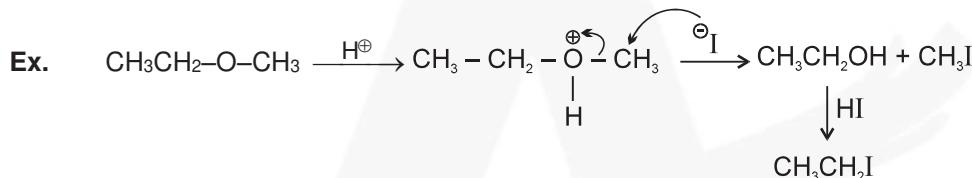


## (2) S<sub>N</sub>2 Reaction of Ethers

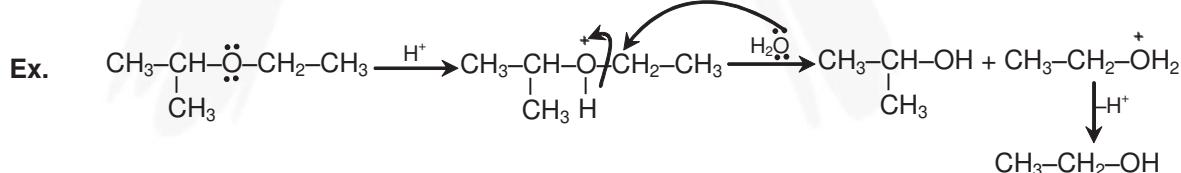
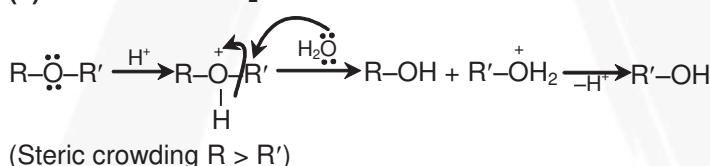
### (i) Reaction with HX :

A protonated ether can undergo substitution reaction. Ether react with conc. HBr and HI because these reagents are sufficiently acidic to protonate the ether. If R or R' is 3° then mechanism will be S<sub>N</sub>1 otherwise S<sub>N</sub>2.

### **Mechanism :**



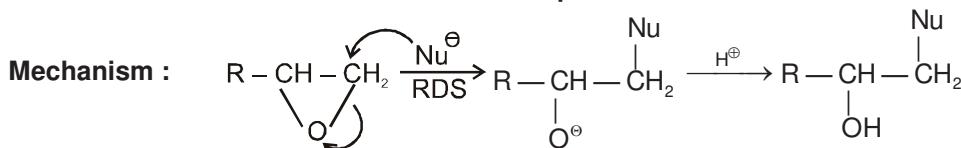
### (ii) Reaction with $\text{H}_2\text{O}/\text{H}^+$ :

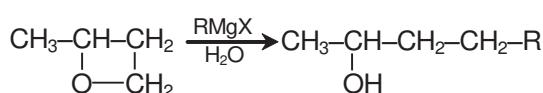
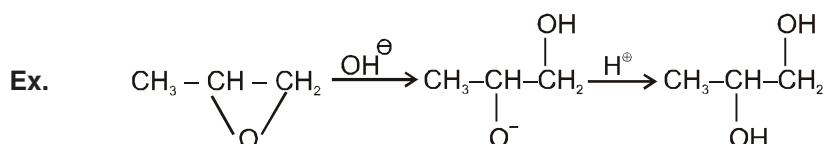


### (3) Nucleophilic substitution reaction of Epoxide :

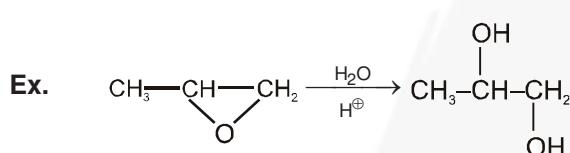
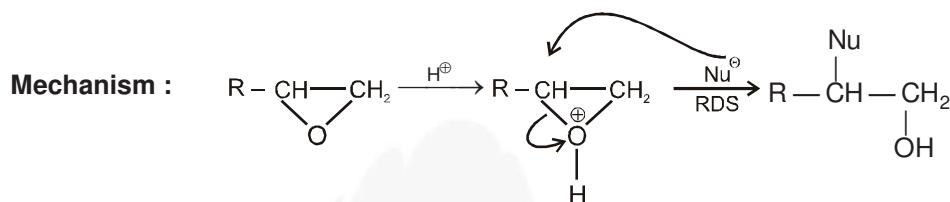
Epoxides are much more reactive than ether because of angle strain in three membered ring therefore epoxide readily undergo nucleophilic substitution reaction.

In basic medium mechanism is S<sub>N</sub>2. Nucleophile attacks on less hindered carbon.



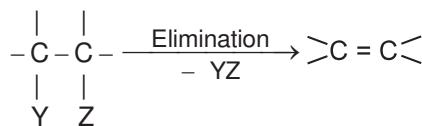


In acidic medium mechanism is S<sub>N</sub>1 type. Nucleophile attacks on more substituted carbon.

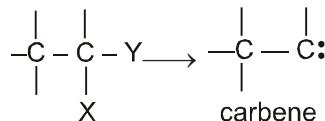


### Elimination reactions:

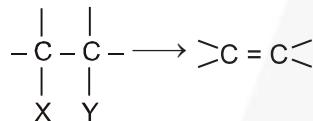
In an elimination reaction two atoms or groups (YZ) are removed from the substrate and generally resulting into formation of  $\pi$  bond.



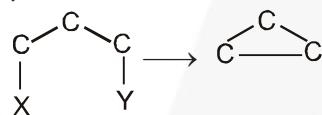
**$\alpha$ -elimination :** When two groups are lost from the same carbon atom to give a carbene (or nitrene). This is also called 1,1-elimination.



**$\beta$ -elimination :** When two groups are lost from adjacent atoms so that a new  $\pi$  bond is formed. This is also called 1,2-elimination.



**$\gamma$ -elimination :** It is also called 1,3-elimination, In this a three membered ring is formed.



**Type of  $\beta$ -elimination reactions :** There are three types of elimination reactions

- (I) E1              (II) E2              (III) E1cB

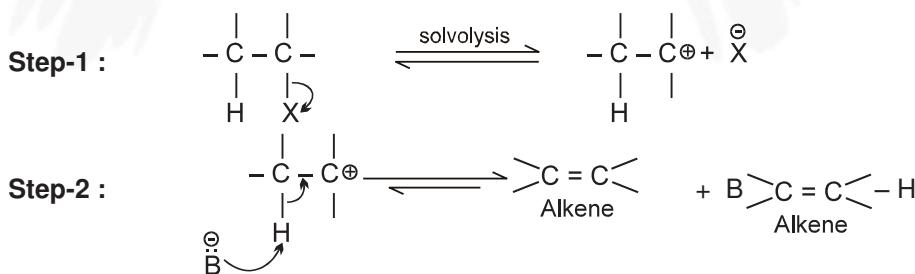
### Section (A) : Unimolecular elimination reaction (E1)

Proton and leaving group depart in two different step.

- (a) **First step :** Slow step involves ionisation to form carbocation  
 (b) **Second step :** Abstraction of proton

#### (1) E1 Reaction of Alkyl halides :

**Mechanism :**



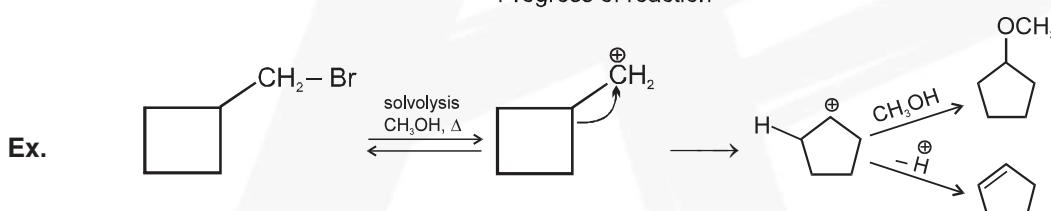
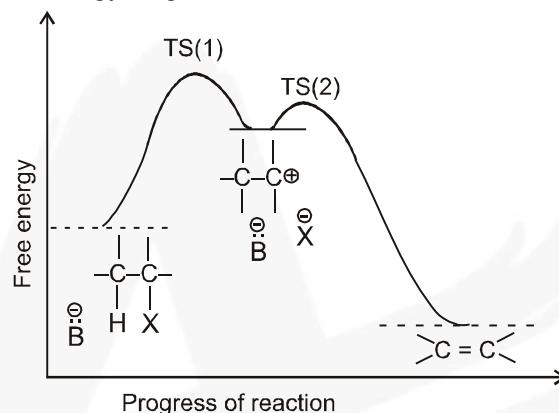
**Characteristics of E1 reaction :**

- It is unimolecular, two step process.
- It is a first order reaction.
- Reaction intermediate is carbocation, so rearrangement is possible
- In the second step, a base abstracts a proton from the carbon atom adjacent to the carbocation, and forms alkene.

**Rate of E1 = 3° Alcohol > 2° Alcohol > 1° Alcohol.****Rate of reaction  $\propto$  stability of carbocation****(v) Kinetics :**

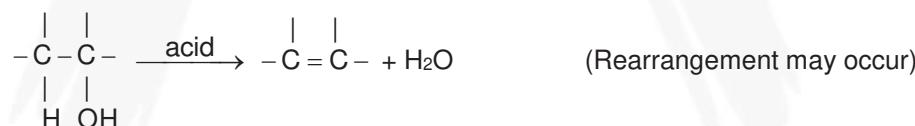
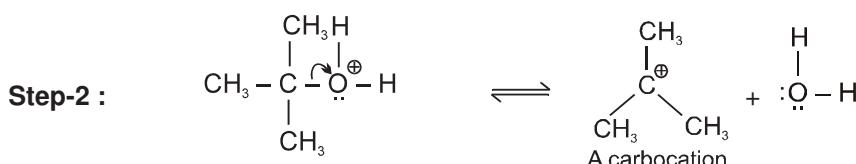
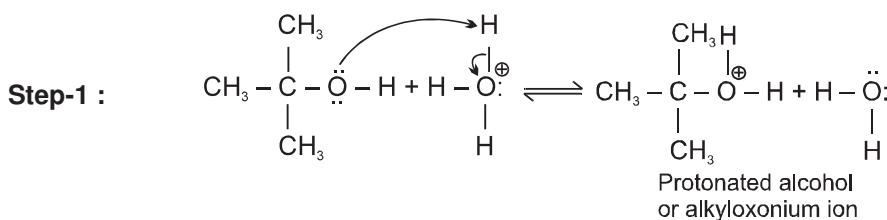
Rate  $\propto$  [Alkyl halide]

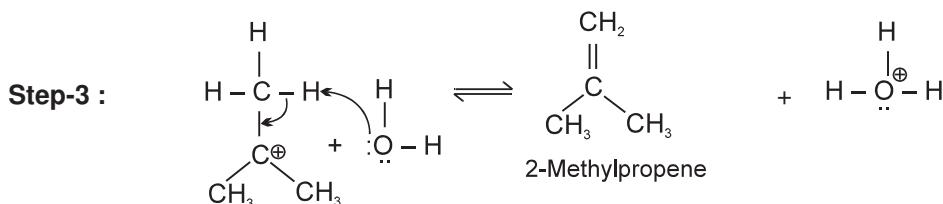
Rate =  $k$  [Alkyl halide]

**(vi) Energetics :** The free energy diagram for the E1 reaction is similar to that for the S<sub>N</sub>1 reaction.**(2) E1 Reaction of Alcohols :**

Dehydration requires an acidic catalyst to protonate the hydroxyl group of the alcohol and convert it to a good leaving group. Loss of water, followed by loss of proton, gives the alkene. An equilibrium is established between reactants and products. For E1 mechanism reagents are :

- H<sub>3</sub>PO<sub>4</sub>/Δ
- H<sub>2</sub>SO<sub>4</sub> / 160°

**Mechanism :**

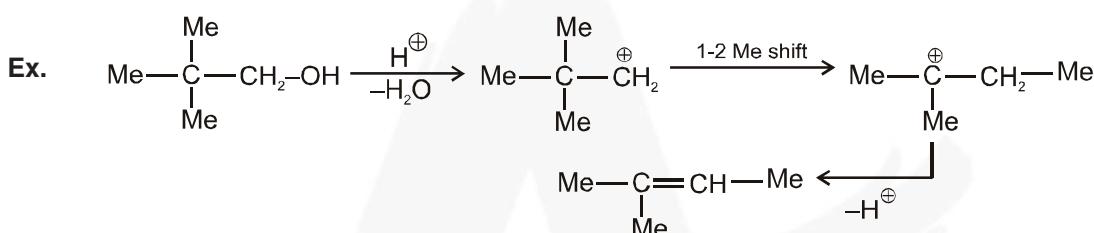
**Remarks :**

In first step, an acid-base reaction a proton is rapidly transferred from the acid to one of the unshared electron pairs of the alcohol.

In second step the carbon oxygen bond breaks. The leaving group is water molecule.

Finally, in third step the carbocation transfers a proton to a molecule of water. The result is the formation of a hydronium ion and an alkene.

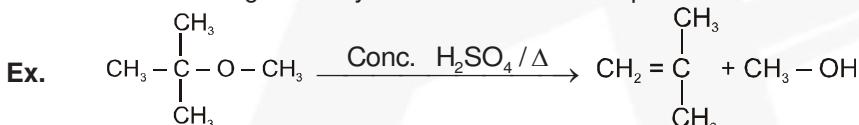
**Reactivity of ROH :  $3^\circ > 2^\circ > 1^\circ$**

**(3) E1 Reaction of Ether :**

Elimination is not a favourable reaction for ether, but however few reactions have been observed.

E1 Elimination takes place via formation of stable carbocation.

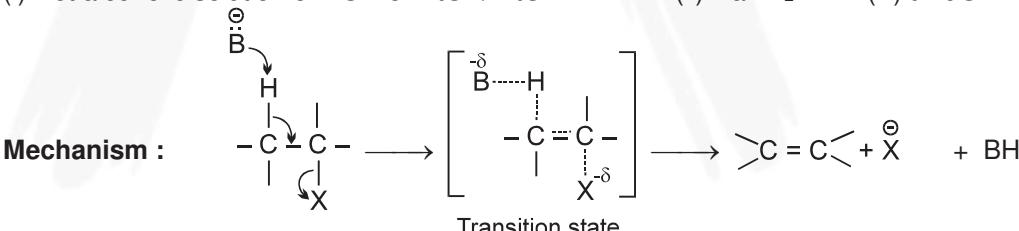
Ether undergoes dehydration reaction in the presence of conc.  $\text{H}_2\text{SO}_4 / \Delta$ .

**Section (B) : Bimolecular elimination reaction (E2)****(1) E2 Reaction of Alkyl halide :**

Dehydrohalogenation is the elimination of a hydrogen and a halogen from an alkyl halide to form an alkene. Dehydrohalogenation can take place by E1 and E2 mechanism.

**Reagent**

(i) Hot alcoholic solution of KOH or  $\text{EtO}^- / \text{EtOH}$       (ii)  $\text{NaNH}_2$       (iii)  $\text{t-BuO}^- \text{K}^+$  in  $\text{t-BuOH}$

**Characteristics of E2 reaction :**

1. This is a single step, bimolecular reaction
2. It is a second order reaction
3. Kinetics  $\rightarrow$  Rate  $\propto [R-X] [\text{Base}]$
4. Rearrangement is not possible
5. For the lower energy of activation, transition state must be stable
6. E2 follows a concerted mechanism
7. The orientation of proton & leaving group should be antiperiplanar.
8. Here  $\beta$ -H is eliminated by base hence called  **$\beta$  elimination**

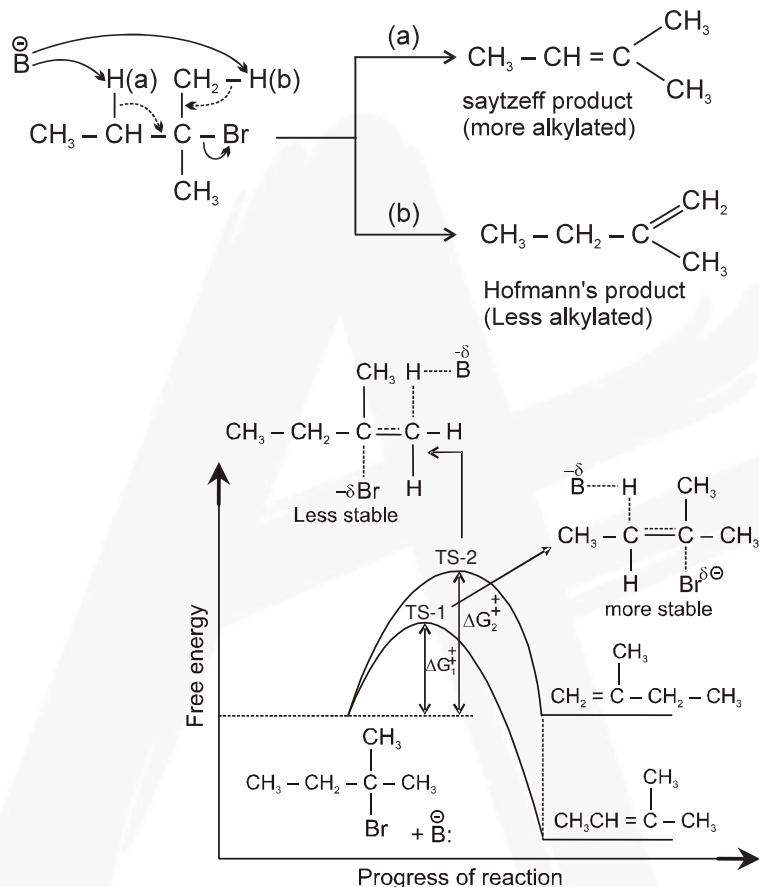
$$\text{Rate} = k [R-X] [\text{B}^-]$$

9. Positional orientation of elimination → In most E1 and E2 eliminations where there are two or more possible elimination products, the product with the most highly substituted double bond will predominate. This rule is called the saytzeff or zaitsev rule.

Reactivity towards E2 → R-I > R-Br > R-Cl > R-F

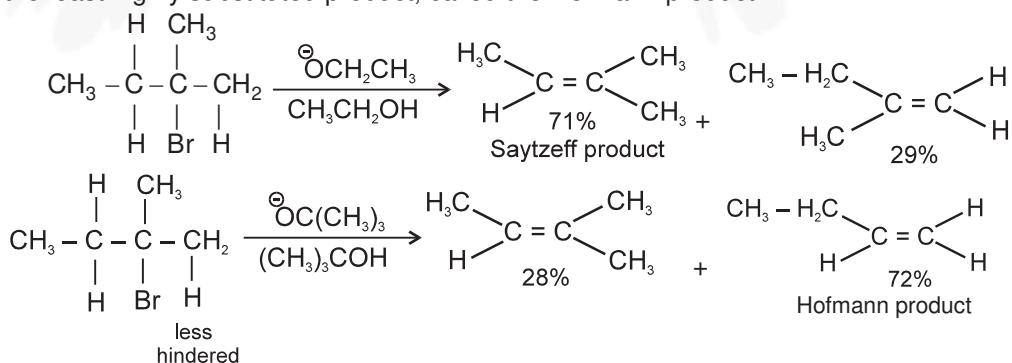
**Rate of E2 reaction = 3° Alkylhalide > 2° Alkylhalide > 1° Alkylhalide**

**Ex.** Dehydrohalogenation of 2-bromo-2-methylbutane can yield two products.

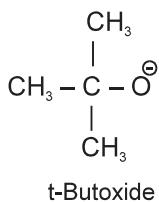


#### Formation of the Hoffmann product

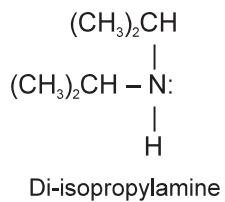
Bulky bases can also accomplish dehydrohalogenations that do not follow the Saytzeff rule. Due to steric hindrance, a bulky base abstracts a less hindered proton, often the one that leads to formation of the least highly substituted product, called the Hoffmann product.



**Bulky base :**



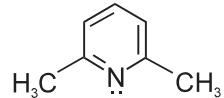
### t-Butoxide



### Di-isopropylamine



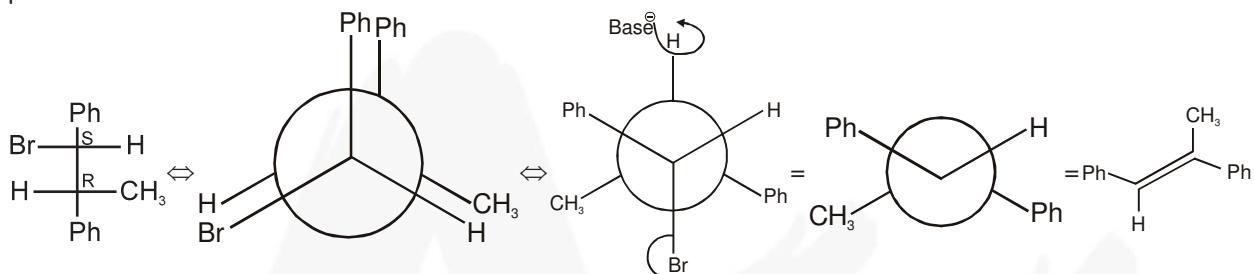
### Triethylamine



### 2, 6-dimethyl pyridine

## Stereospecific E2 reactions

The E2 is stereospecific because it normally goes through an anti periplanar transition state. The products are alkenes.



**Ex.**  $\begin{array}{c} \beta \\ | \\ \text{CH}_3 - \text{CH} - \text{CH} - \text{CH}_3 \\ | \\ \text{Br} \end{array} \xrightarrow{\text{alc. KOH}} \begin{array}{c} \text{CH}_3 - \text{CH} = \text{C} - \text{CH}_3 \\ (\text{major}) \\ (\text{more stable alkene}) \end{array} + \begin{array}{c} \text{CH}_2 = \text{CH} - \text{CH} - \text{CH}_3 \\ (\text{minor}) \\ (\text{less stable alkene}) \end{array}$

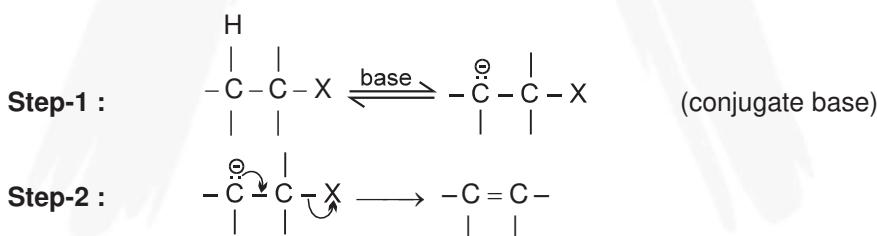
**Ex.**

$\text{Cl} \begin{array}{c} | \\ \diagdown \end{array} \text{CH}_3$   $\xrightarrow{\text{alc. KOH}}$ 
 $\begin{array}{c} \text{CH}_2 \\ \diagup \\ \text{C} \end{array}$  + 
  $\begin{array}{c} \text{CH}_3 \\ \diagup \\ \text{C} \end{array}$ 
  
 (minor) (Major)

## Section (C) : Unimolecular elimination by conjugate base (E1cB)

In the E1 cB, H leaves first and then the X. This is a two step process, the intermediate is a carbanion.

### **Mechanism :**



### **Remarks :**

First step consists of the removal of a proton ( $\text{H}^+$ ) by a base generating a carbanion.

In second step carbanion loses a leaving group to form alkene.

**Condition:** For the E1cB, (i) substrate must be containing acidic hydrogens and (ii) poor leaving groups.

**Ex.** 