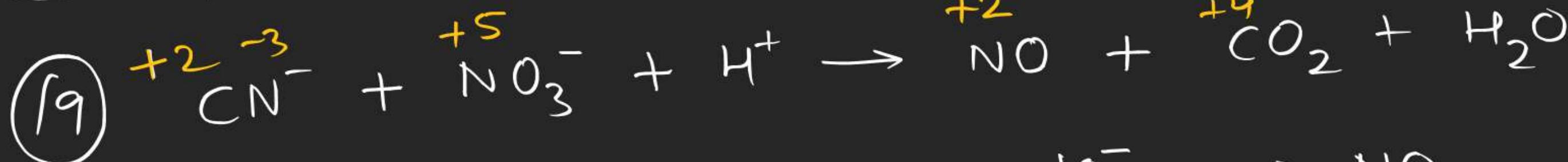
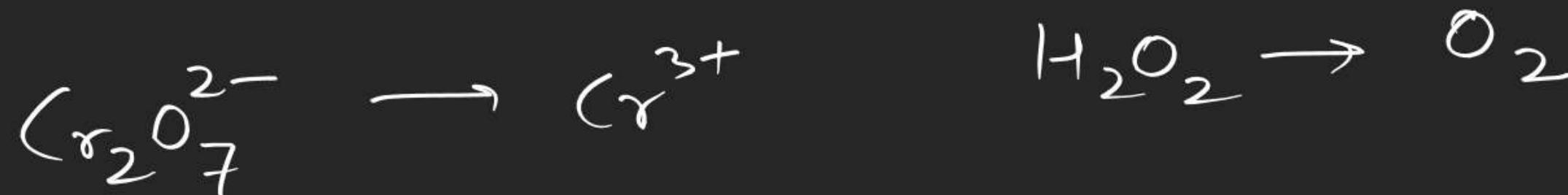
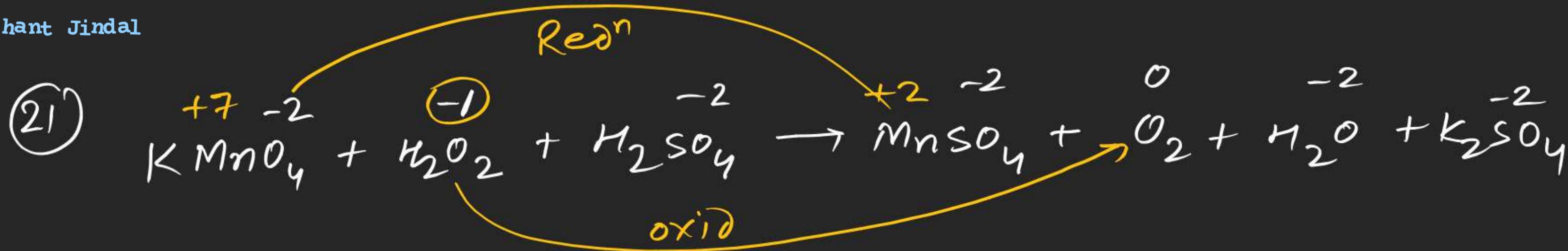


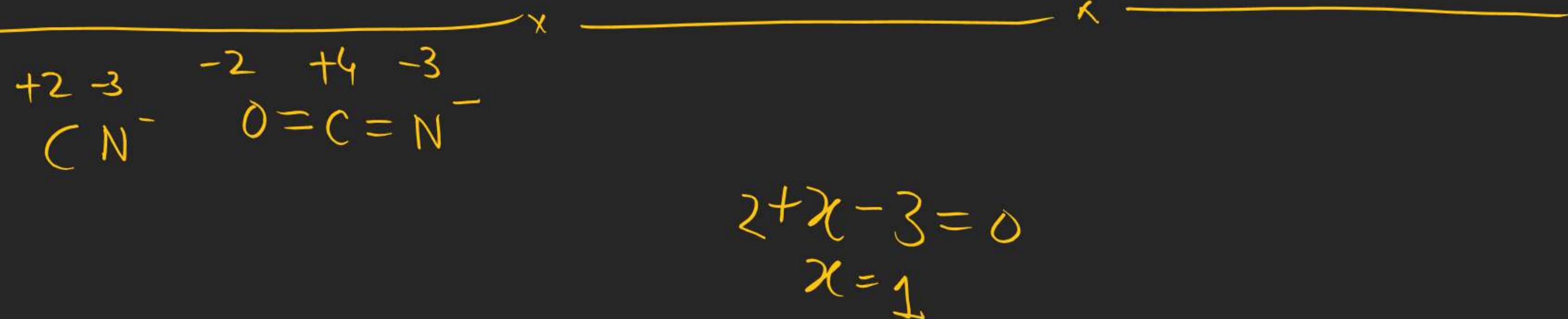
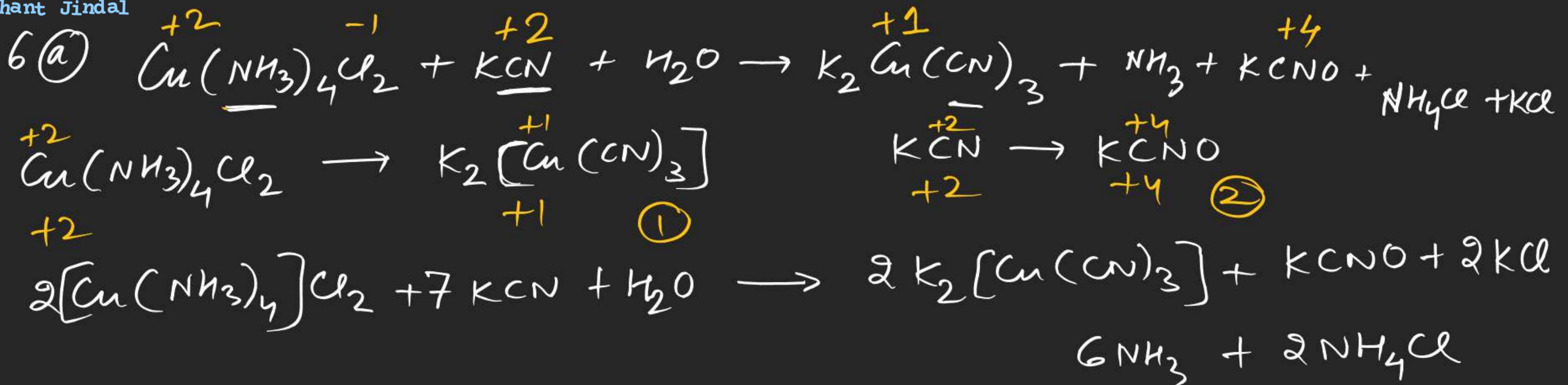
O-I Up to 22

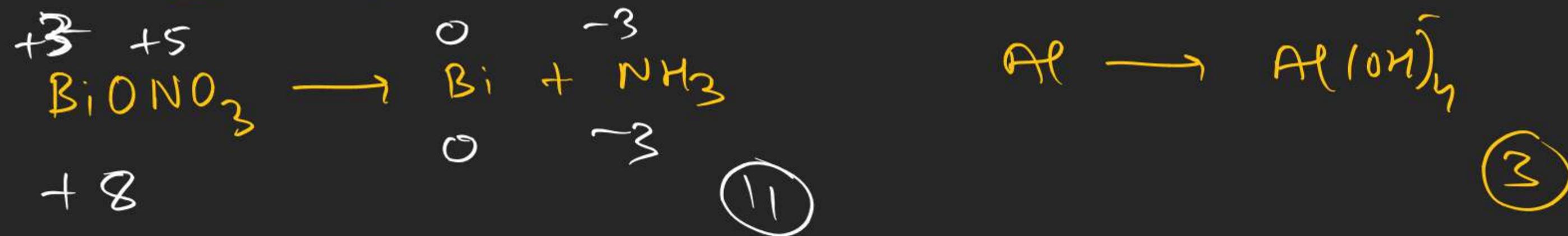
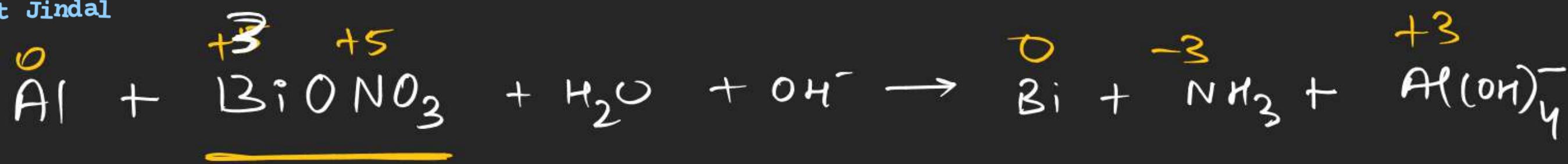
S-II 6, 7, 8

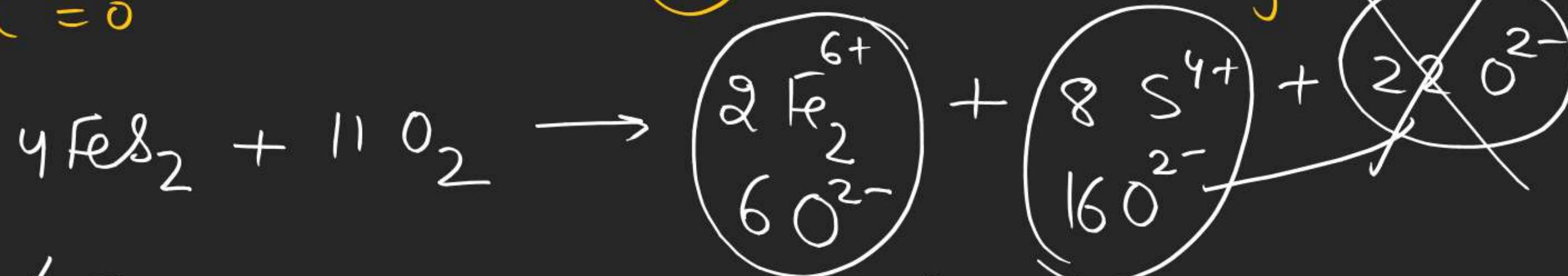
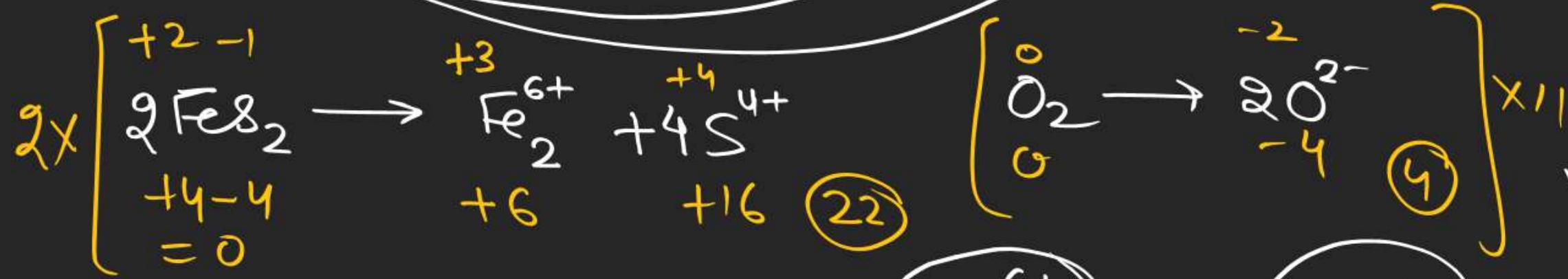
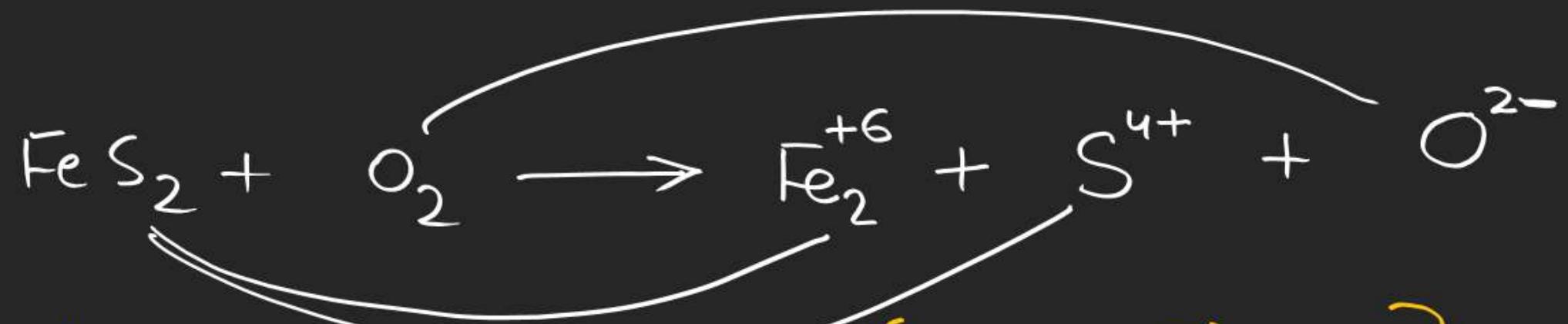
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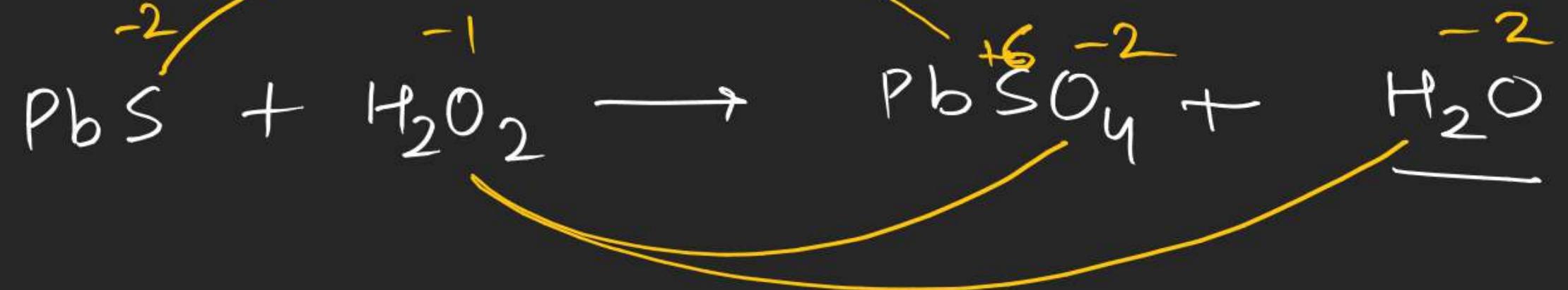












Nishant Jindal

Equivalent concept



$$\left[\begin{array}{l} \frac{\text{no. of equivalents}}{\text{of 'A' reacted}} = \frac{\text{no. of equivalents}}{\text{of 'B' reacted}} = \frac{\text{no. of equivalents}}{\text{of 'C' formed}} = \frac{\text{no. of equivalents}}{\text{of 'D' formed}} \\ \text{mass} \\ \text{molar mass} \end{array} \right]$$

$$\frac{\text{no. of moles}}{\text{mass}} = \frac{\text{mass}}{\text{molar mass}}$$
$$\text{no. of equivalents} = \frac{\text{mass}}{\text{Equivalent mass}} = \frac{\text{mass}}{\text{molar mass}} \times n\text{-factor}$$

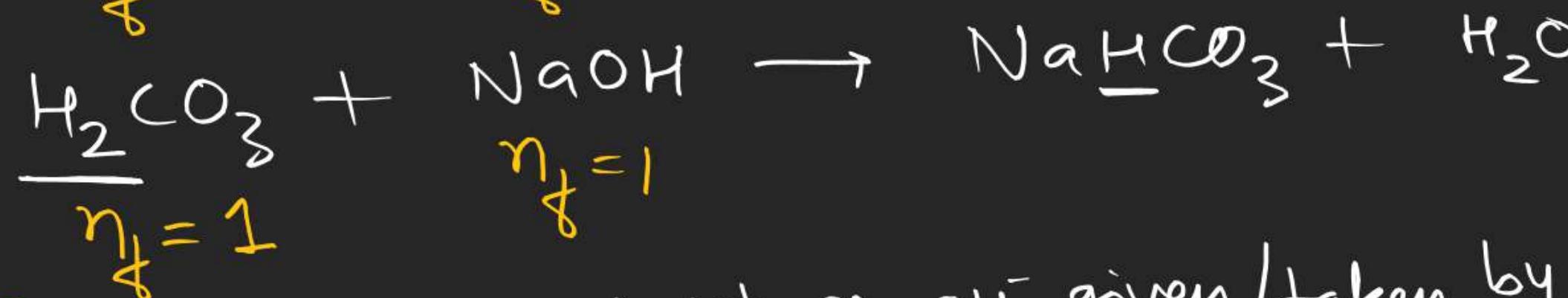
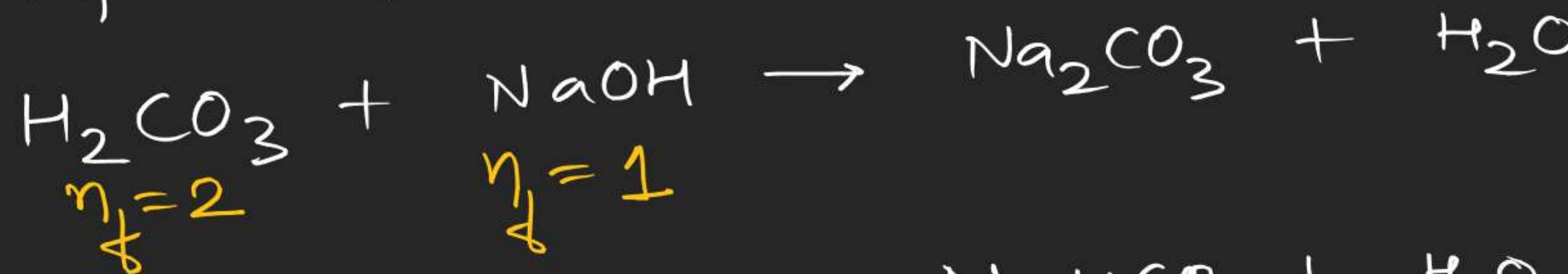
$$\text{Equivalent mass} = \frac{\text{molar mass}}{n\text{-factor}}$$

no. of equivalent = no. of moles \times n-factor

n-factor calculation :-

① For non-redox Rxn

(a) n-factor for acid and base



n-factor = no. of H^+ or OH^- given/taken by per mole
acid or base.

