


## DPP-2

## SOLUTION

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1.  $\alpha$ -particle
2. For nitrogen: atomic number = 7 and mass number = 14  
 $\therefore$  Number of neutrons =  $14 - 7 = 7$   
For silicon: atomic number = 14 and mass number = 28  
 $\therefore$  number of neutrons =  $28 - 14 = 14$   
 $\therefore$  Ratio of number of neutrons in nitrogen and silicon =  $7:14 = 1:2$
3. Isoelectronic ions having same no. of electrons.
4. (C) Same no. of electrons
5. (A, B, C, D)
6. (B, C, D)
- Sol. Isotones have same number of neutrons.
7. (A, B, C)
8. 0
- Sol.  $A = 2Z$   $A - 2Z = 0$
9. 5
10. (1)