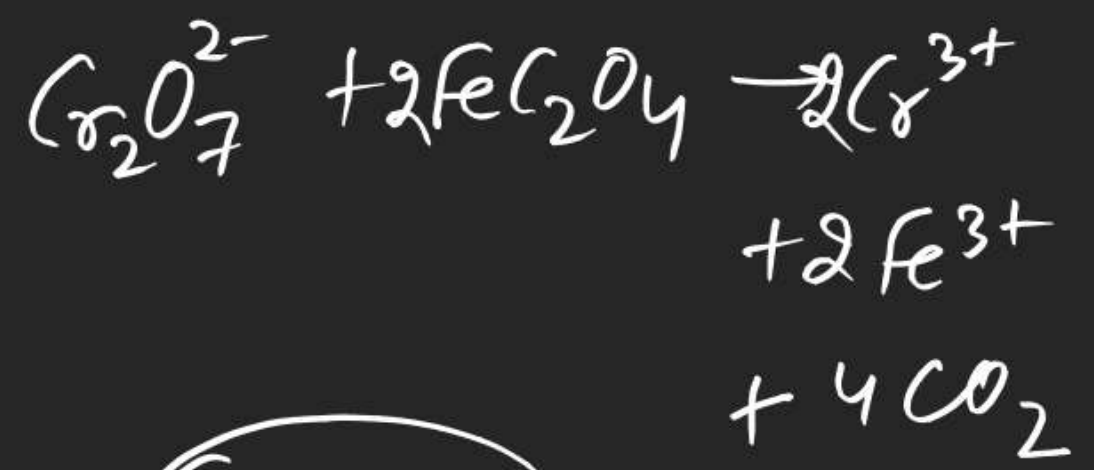
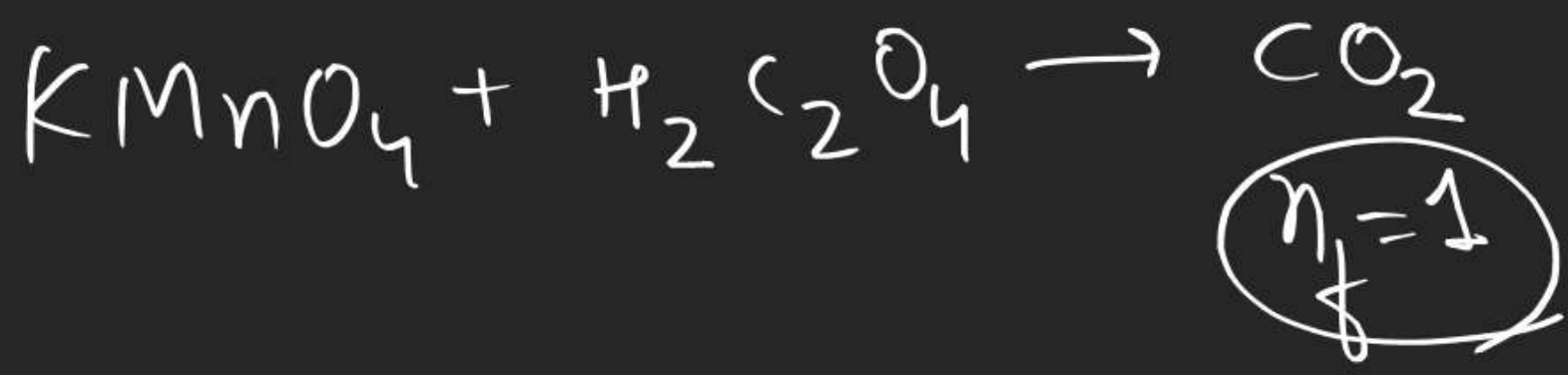
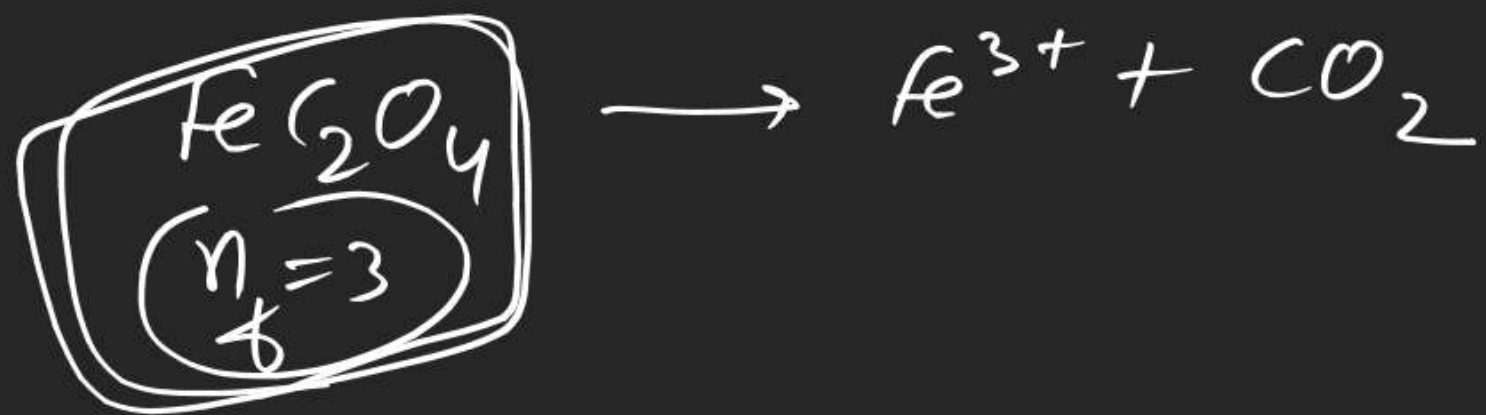


10  
A, D



$n_e = 6$

~~11 & 15~~ Bonus



$$1 \times 3 + 1 \times 6 = n \times 5$$

$$9/5 = n$$



$$\textcircled{25} \quad \text{eq of MnO}_2 = \text{eq of hypo}$$



$$n \times 2 = 60 \times 0.1 \times 1$$

$$\textcircled{26} \quad \text{eq of MnSO}_4 = \text{eq of hypo}$$

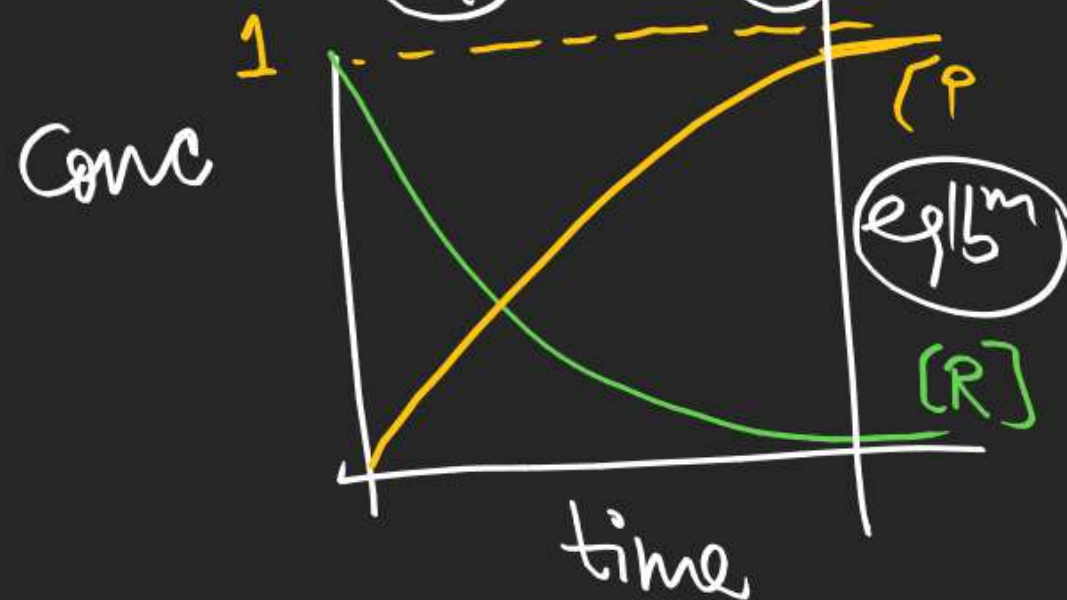
# Chemical equilibrium

⇒ This chapter deals with extent of reaction and factors affecting extent of reaction.

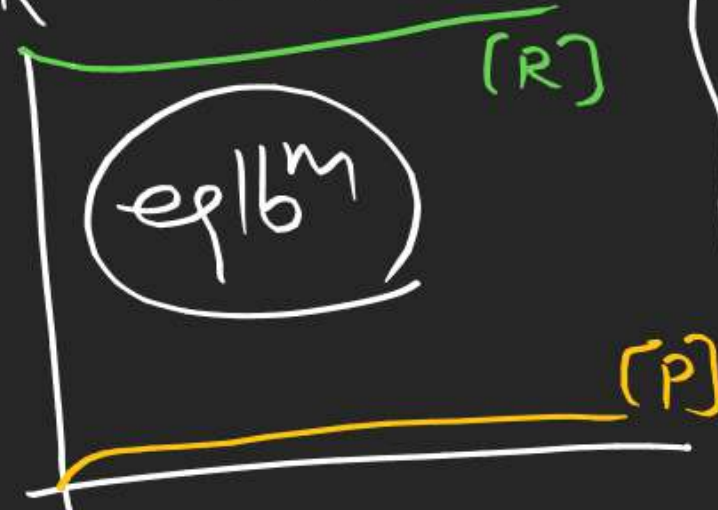
⇒ Depending upon the extent of reaction, chemical reaction can be categorised into

(i) extent = 100%.

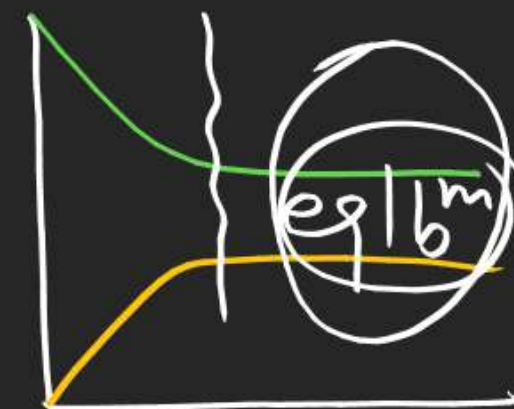
$$\alpha = 1$$



(ii) extent = 0  
( $\alpha = 0$ )



(iii)  $0 < \text{extent} < 100\%$   
( $0 < \alpha < 1$ )



## Characteristics of $eg1b^m$ state

- 1.) At equilibrium conc of reactants and products become  
constant.

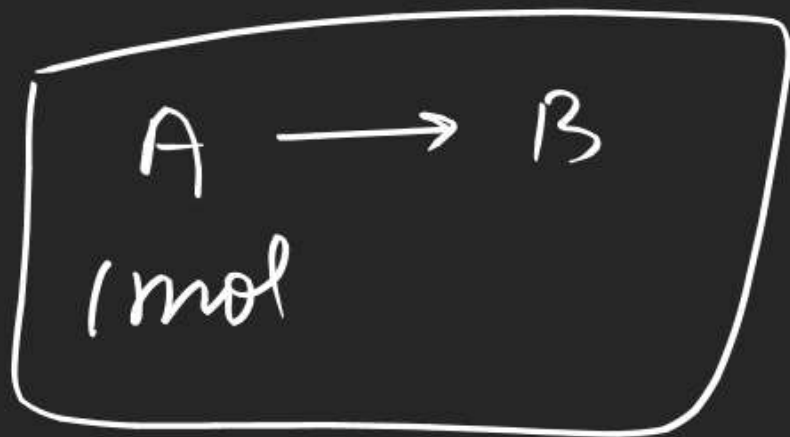
② Conclusion It is a dynamic eqn

(3) forward rate of  $R \times n$  = backward rate of  $r \times n$       static  
dynamic



# Law of chemical eqbm (or Law of mass action)

④  $E_{q/b}^m$  can be achieved from any direction



⑤  $E_{q/b}^m$  can be achieved in closed container only



S-II

1-5

J-Adv

last 8 questions