



DPP-2

SOLUTION

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1. α -particle
 2. For nitrogen: atomic number = 7 and mass number = 14
 \therefore Number of neutrons = $14 - 7 = 7$
For silicon: atomic number = 14 and mass number = 28
 \therefore number of neutrons = $28 - 14 = 14$
 \therefore Ratio of number of neutrons in nitrogen and silicon = $7:14 = 1:2$
 3. Isoelectronic ions having same no. of electrons.
 4. (C) Same no. of electrons
 5. (A, B, C, D)
 6. (B, C, D)
- Sol.** Isotones have same number of neutrons.
7. (A, B, C)
 8. 0
- Sol.** $A = 2Z$ $A - 2Z = 0$
9. 5
 10. (1)