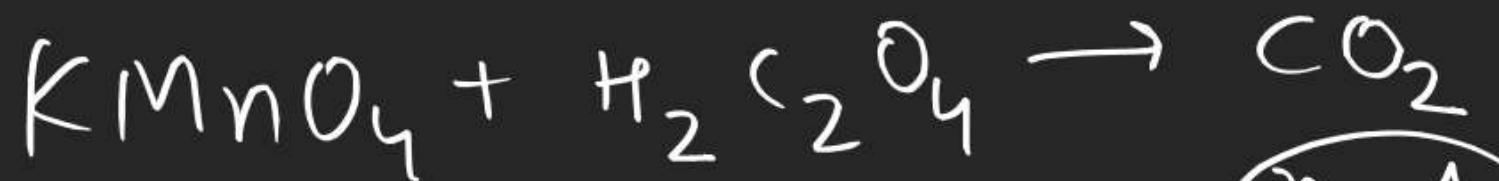
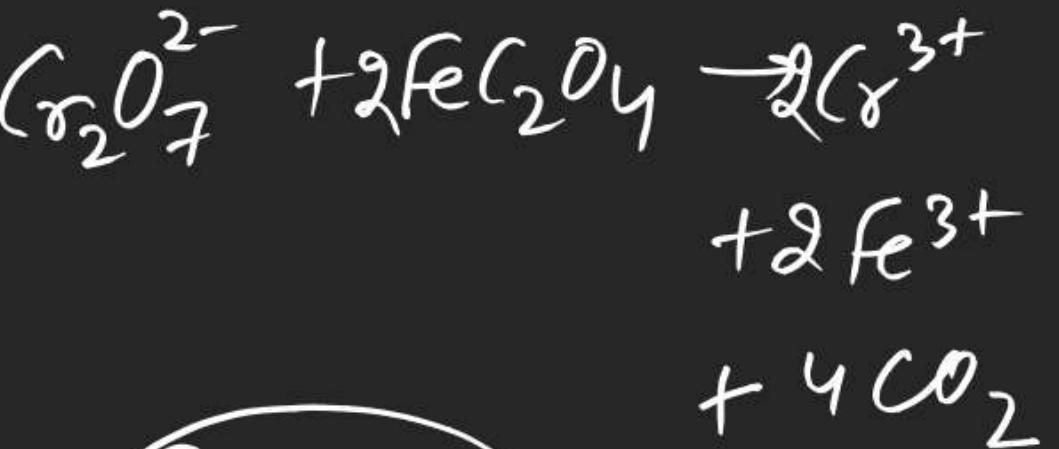


(10)
A, D

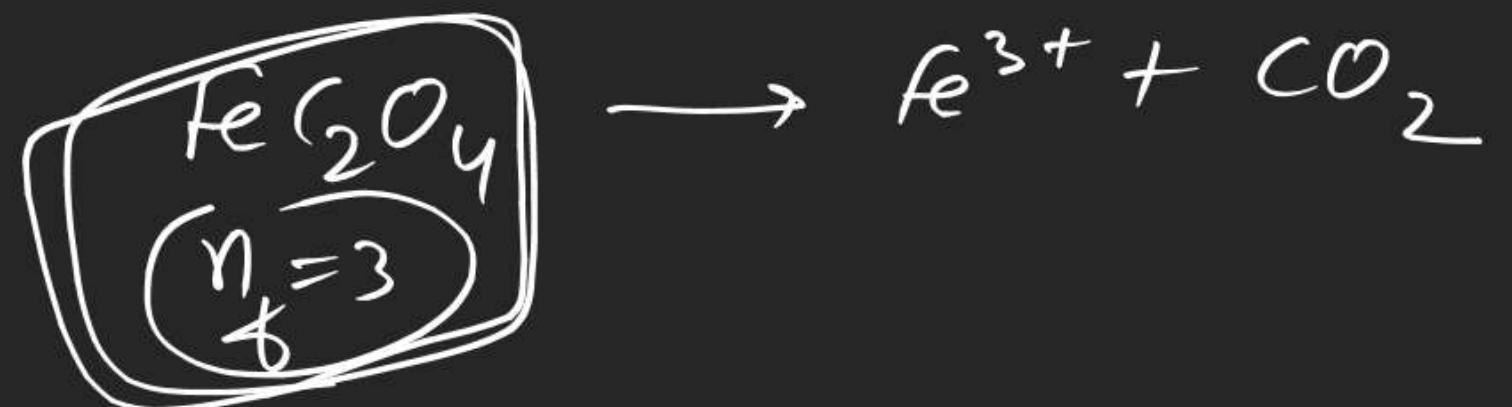


$n_f = 1$



$n_e = 6$

(11) & (15) Bonus



$$1 \times 3 + 1 \times 6 = n \times 5$$

$$\frac{9}{5} = n$$





$$n \times 2 = 60 \times 0.1 \times 1$$

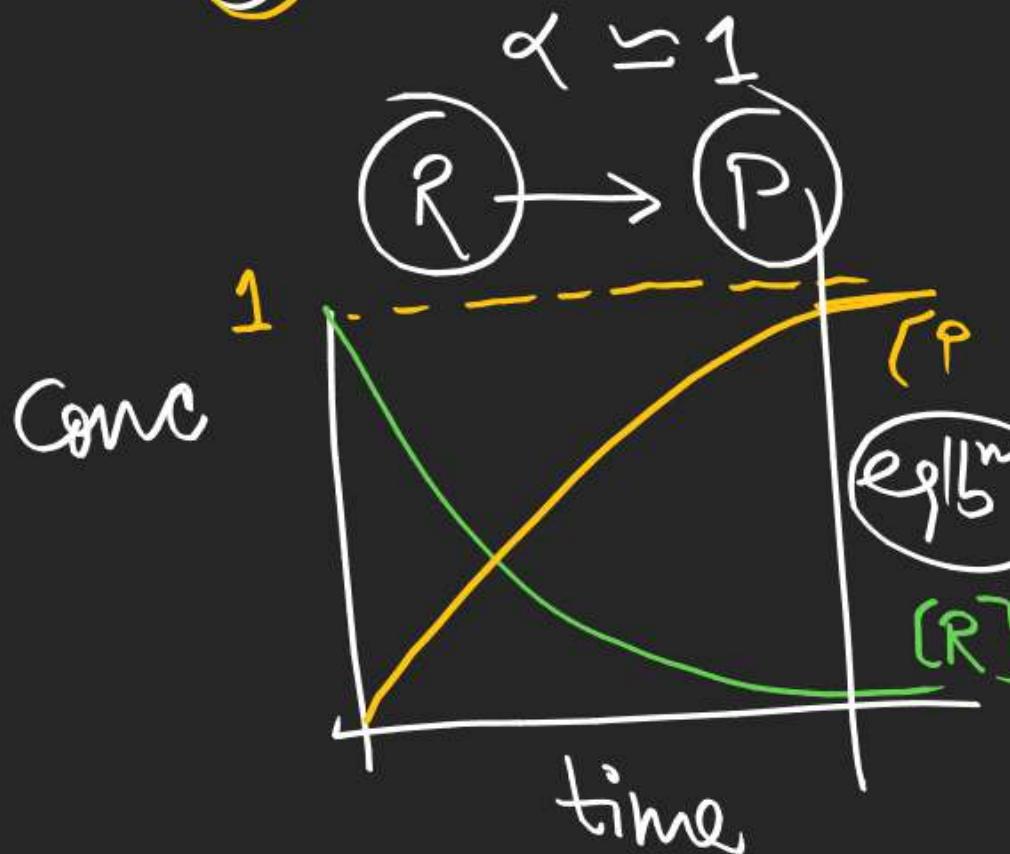


Chemical equilibrium

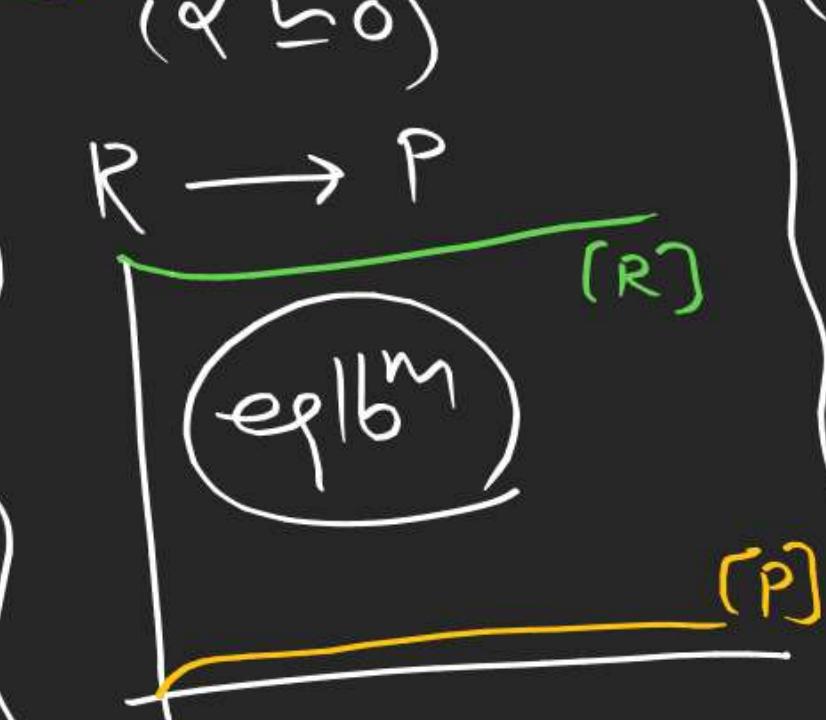
⇒ This chapter deals with extent of reaction and factors affecting extent of reaction.

⇒ Depending upon the extent of reaction, chemical reaction can be categorised into

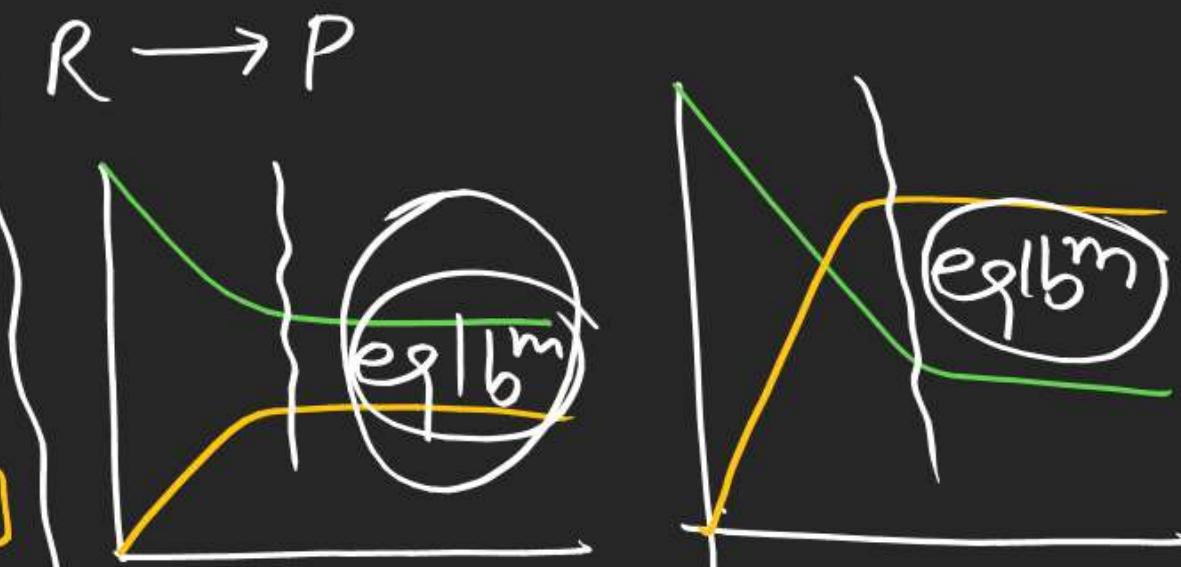
i) extent $\approx 100\%$.



ii) extent ≈ 0
($\alpha \approx 0$)



iii) $0 < \text{extent} < 100\%$.
($0 < \alpha < 1$)



a

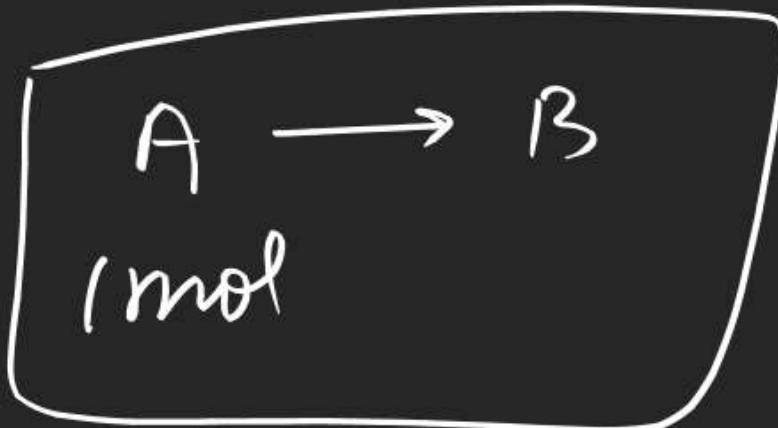
Characteristics of equilibrium state

- 1.) At equilibrium conc of reactants and products become constant.
 - 2) It is a dynamic eq^{b/m}
 - 3) forward rate of Rxn = backward Rate of rxn static
 $(\lambda_f = \lambda_b)$ dynamic



Law of chemical eqblm (or law of mass action)

4) Eq^lb^m can be achieved from any direction



5) Eq^lb^m can be achieved in closed container only



S-II

I-S'

J-Adv

Last 8 questions