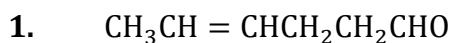


**DPP-1****SOLUTION**

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Compounds A, C and D have the same molecular formula as that of $\text{CH}_3 - \text{CO} - \text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$ with a formula of $\text{C}_6\text{H}_{12}\text{O}$. But compound B has $\text{C}_6\text{H}_{10}\text{O}$ and hence it is not an isomer of $\text{CH}_3 - \text{CO} - \text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$

3. Same molecular formula but different physical and/or chemical properties.

4. $\text{CH}_3\text{CH}(\text{Cl})\text{CH}_3$ and $\text{CH}_3\text{CH}_2\text{CH}_2\text{Cl}$

6. $(\text{CH}_3)_3\text{C} - (\text{C}_2\text{H}_6)_2\text{C} - \text{CH}_2 - \text{CH}_3$, or $\text{C}_{11}\text{H}_{24}$.

7. Bond line structure are those where C, H bond are not shown, lines representing C – C bond are drawn in a zig-zag fashion. terminal group are methyl group unless indicated otherwise by O, N, Cl etc. The correct answer is option (a).

8. The bond line formulae of the given compounds are:

