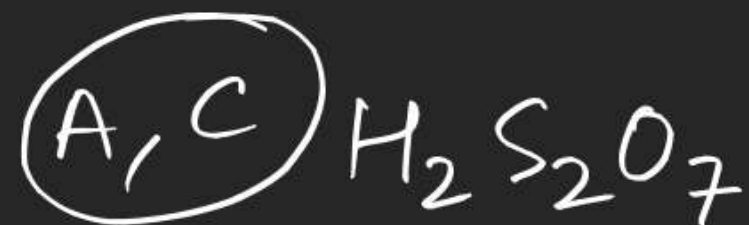
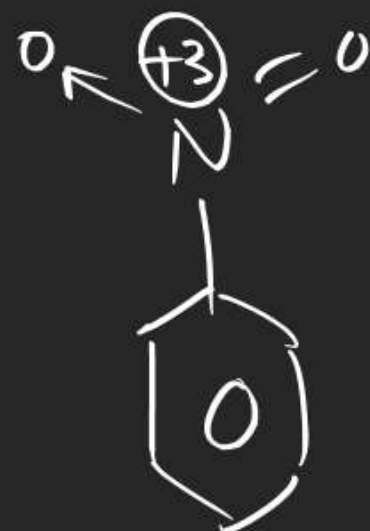
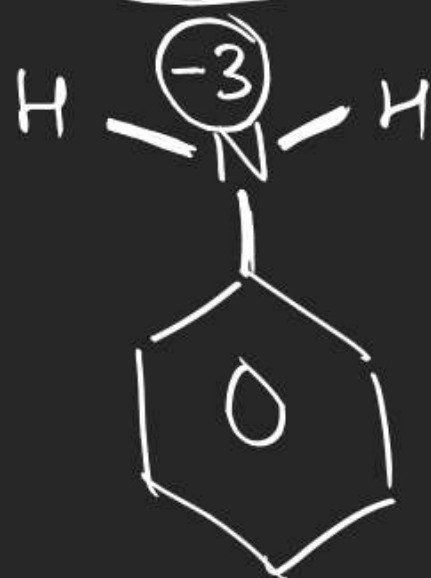


O-I 1-10

S-I 1-3

Redox



$$2 + 2x - 14 = 0$$

$$x = 6$$

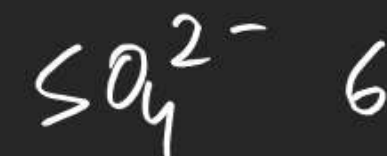
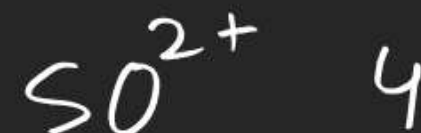


$$x = 2.5$$

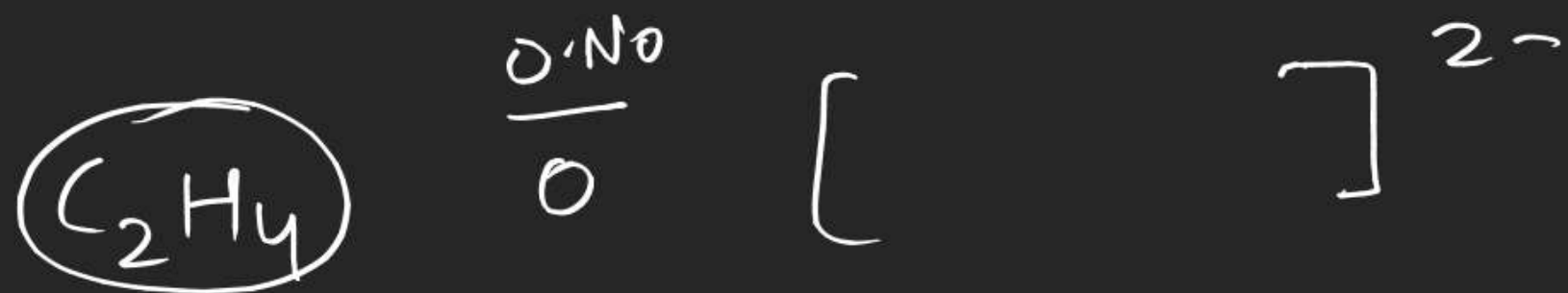


$$+2 + 2x - 6 = 0$$

$$x = 2$$



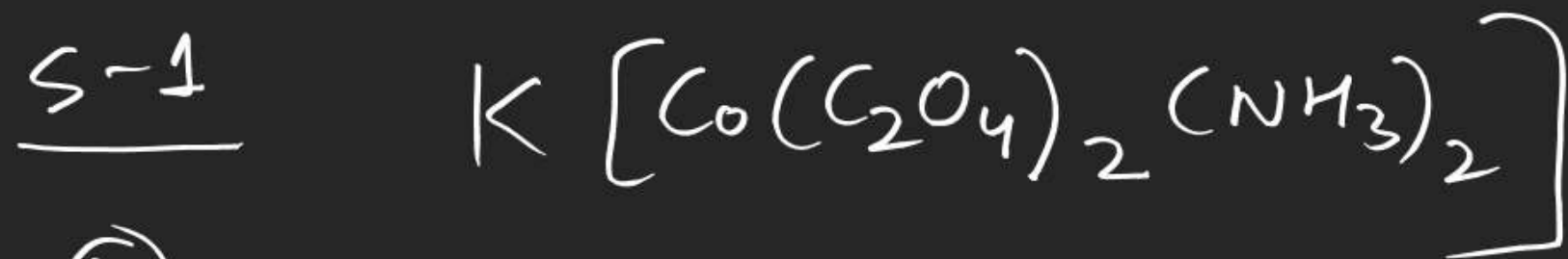
0



⑨ ~~Ans~~ 1

⑩ $2x - 8 + 0 + 0 = -2$

$x = 3$



①

$$+1 + x - 4 + 0 = 0$$

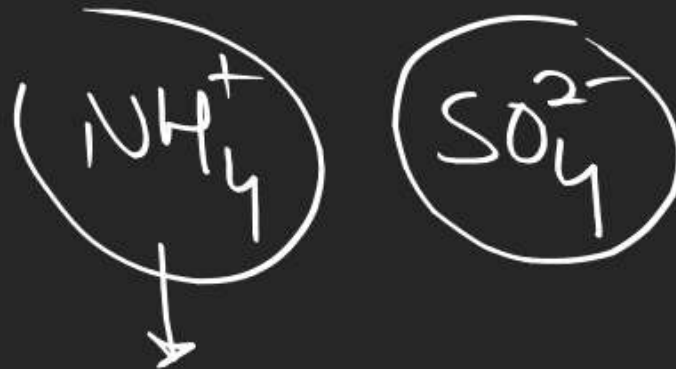
$$\underline{x = 3}$$

② (8)



$$2x + 6 - 2 = 0$$

$$\underline{x = -2}$$



akk 7007

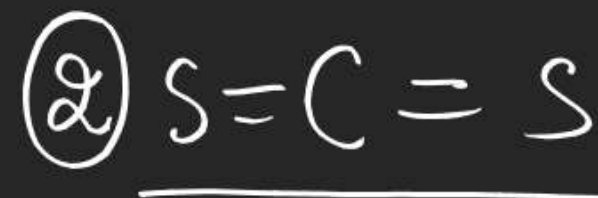
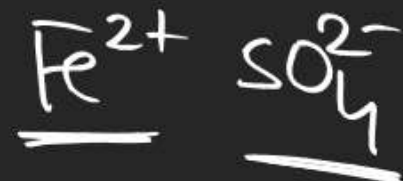
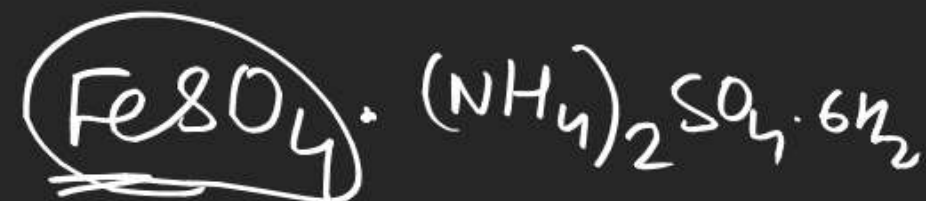
⑫



$$1 + x - 6 = -1$$

$$\textcircled{x = 4}$$

⑬



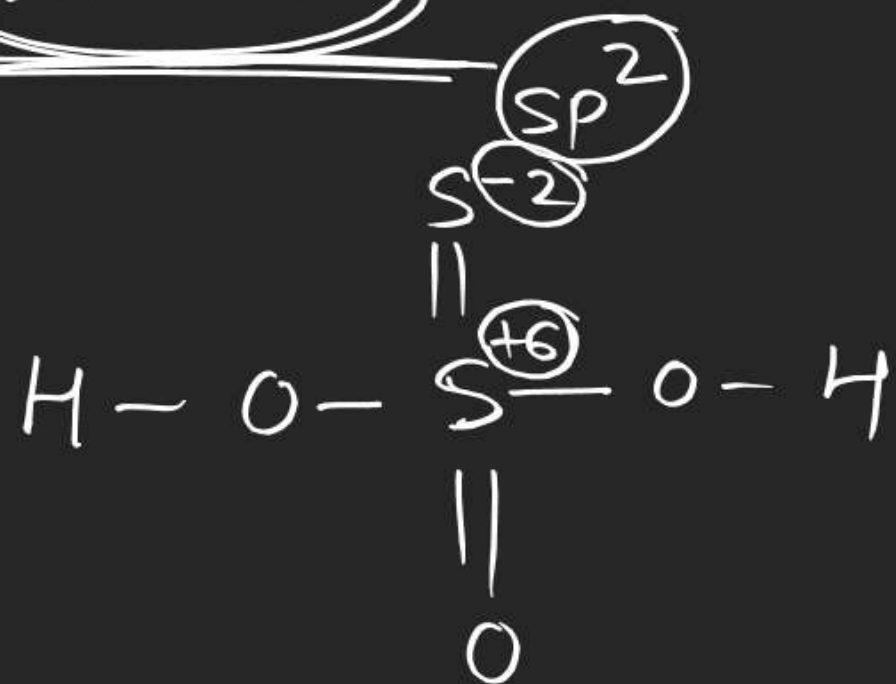
③

①



$$2 + 2x - 6 = 0$$

$x = 2$



$$\frac{6-2}{2} = 2$$

④



FIND OXIDATION NUMBER OF UNDERLINED ELEMENTS



$$3 + x - 6 = 0$$



$$1 + x - 4 + 0 = 0$$



$$x + 0 = 0$$



$$x + 0 = 0$$



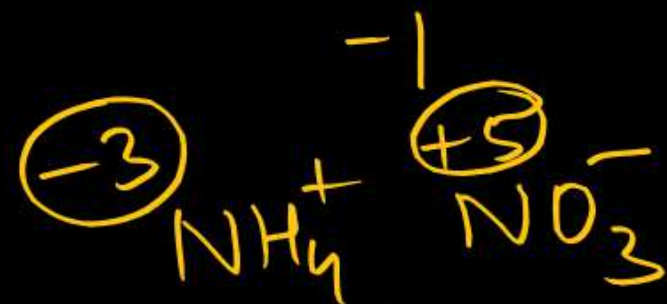
$$12x + 22 - 22 = 0$$



FIND OXIDATION NUMBER OF UNDERLINED ELEMENTS



bleaching powder



Avg +1

+2

+1 -3

+3 -2



+5 -2



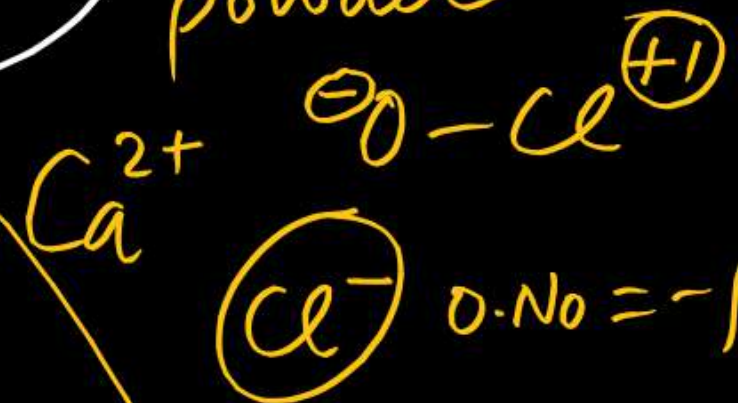
+1 +2 -3



2.5

$$+2 -2 + 2x = 0$$

$$\underline{x = 0}$$

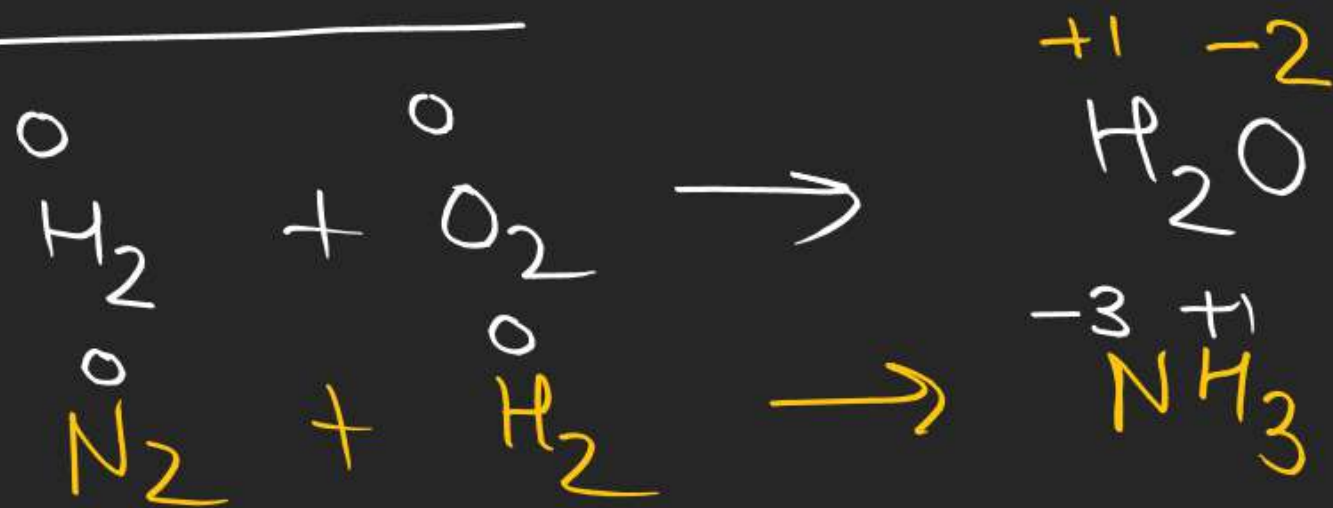


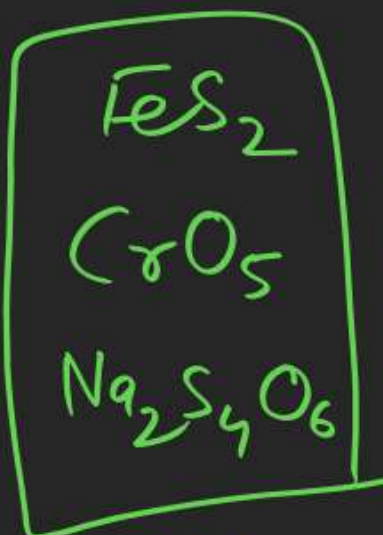
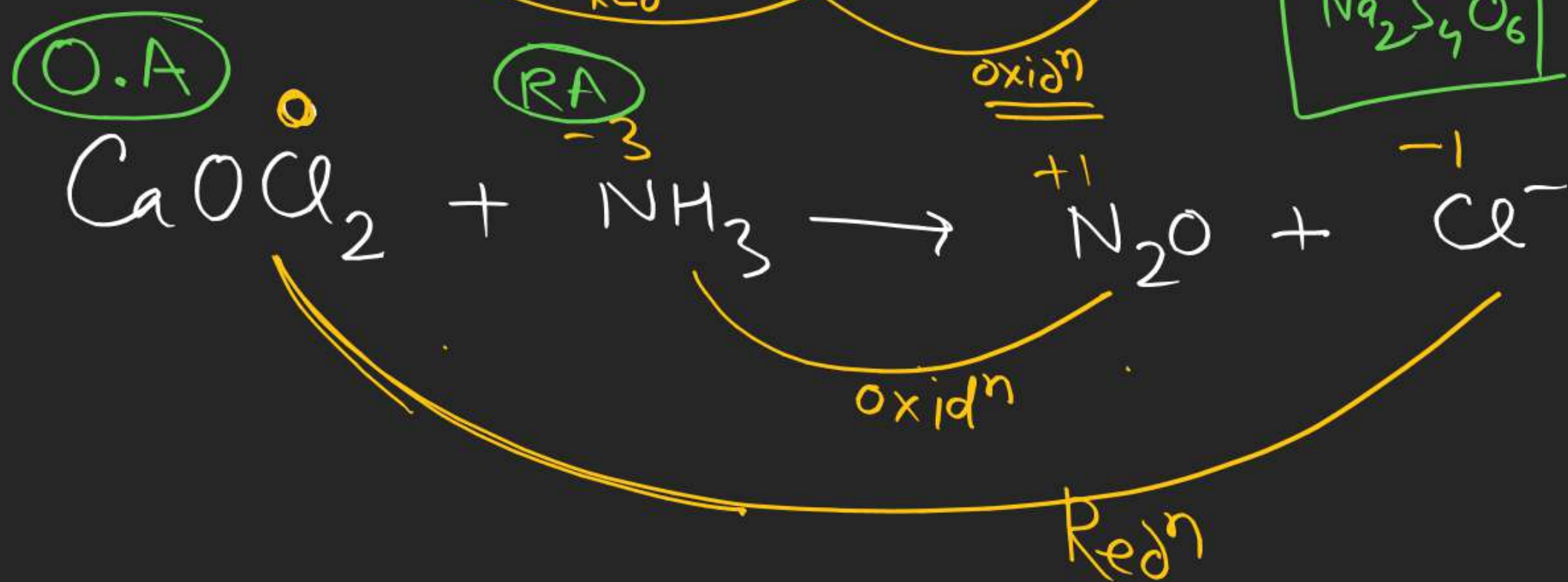
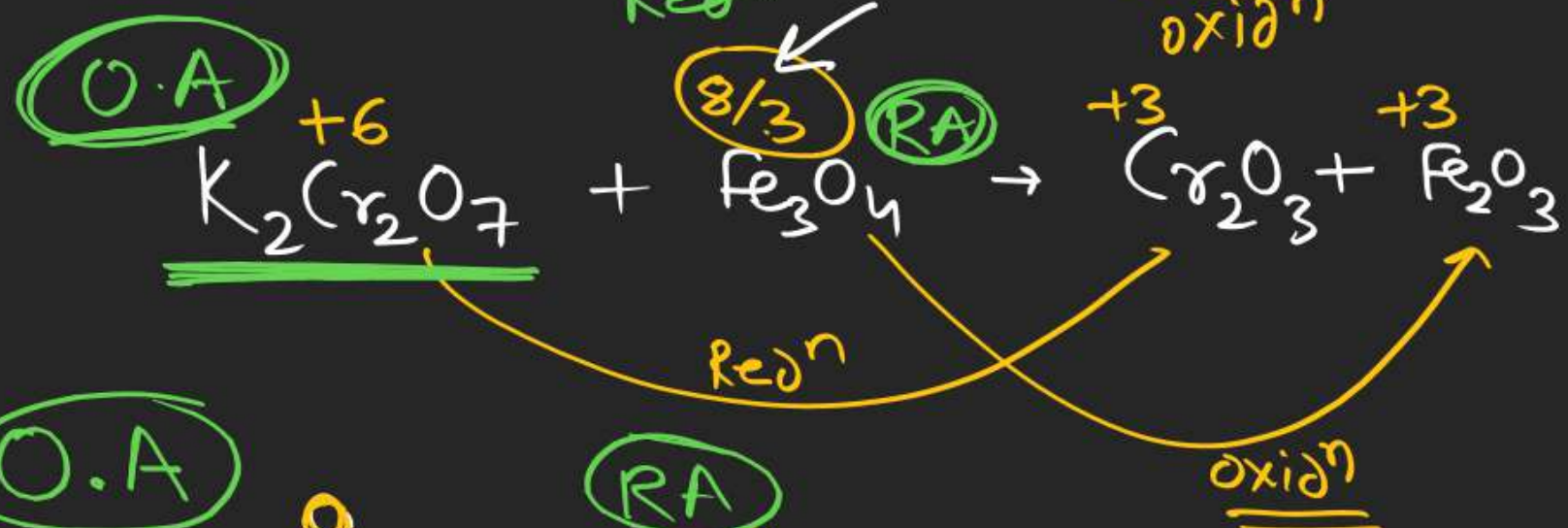
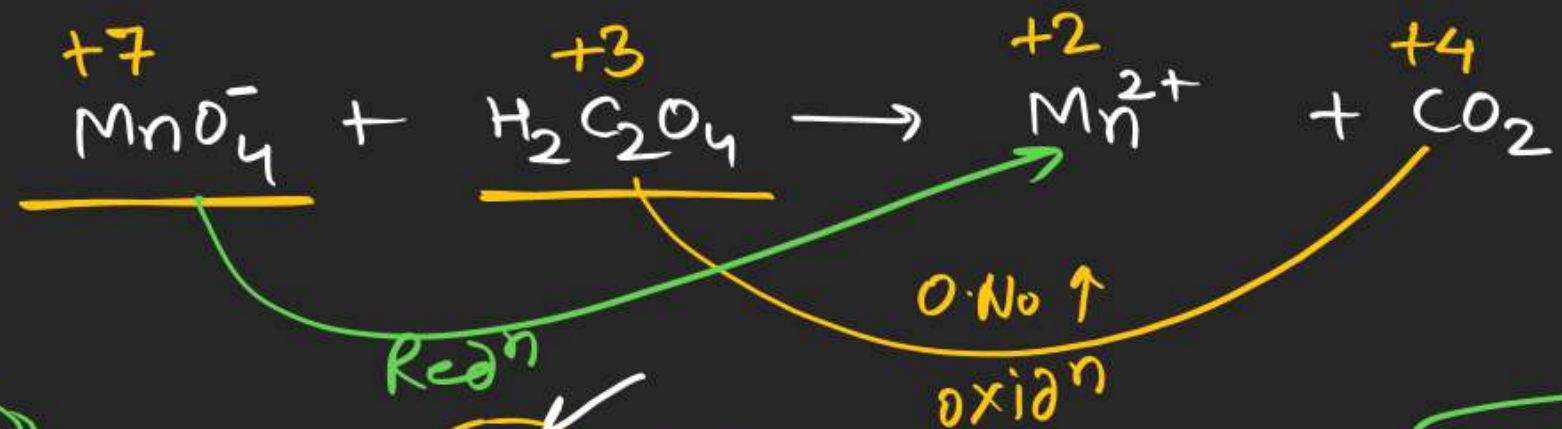
Application of Oxidation number

Oxidation \rightarrow O. No. \uparrow es (tive on atom \uparrow es); loss of e^-

Reduction \rightarrow O. No. \downarrow es (tive on atom \downarrow es) gain of e^-

Redox Rxn \Rightarrow Reduction + oxidⁿ





Balancing of Redox Rxns: →

① Oxidation number

+2

-1

oxid

0

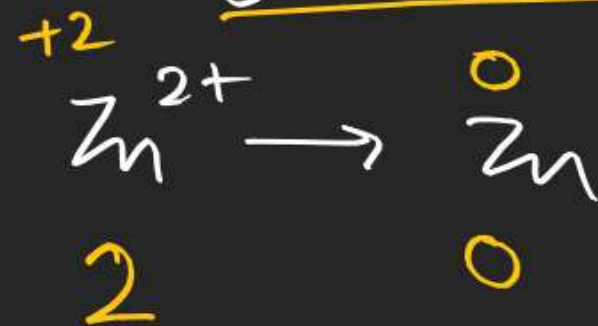
0



② Ion electron method

Ion e method

O. No method



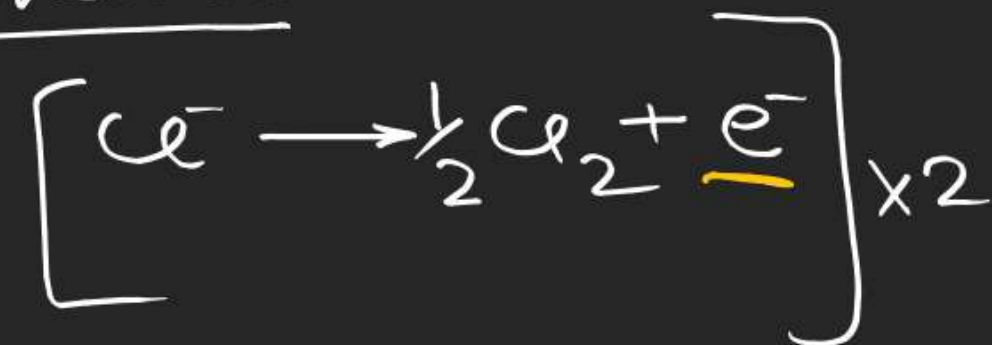
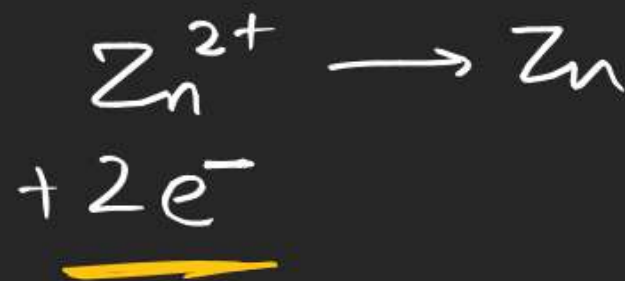
change

②



change

①



~~+2e^-~~

~~+2e^-~~

0-I 11-15

Atomic
JEE-Adv

Remaining
JEE-M