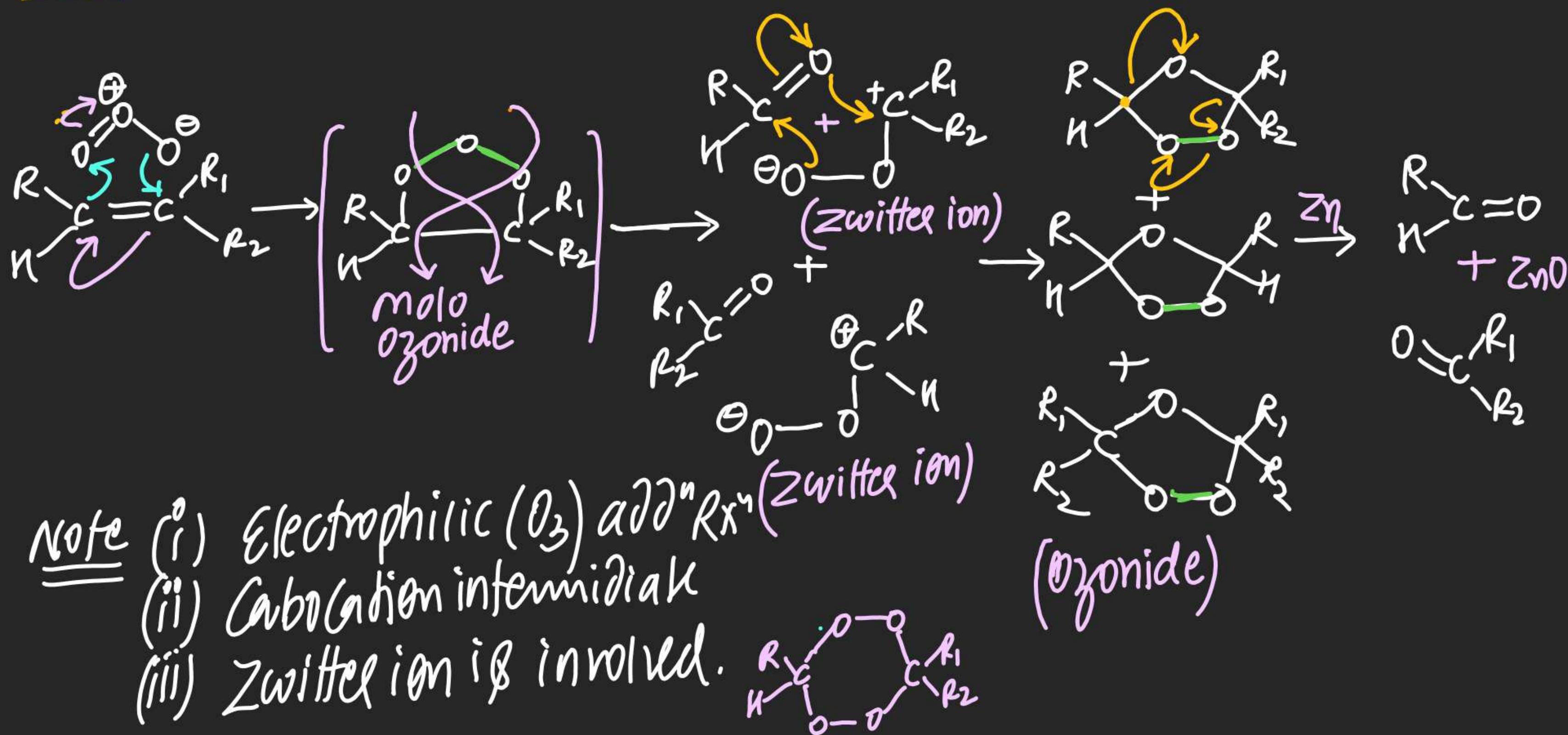


Mechanism

(6) $m_2 = m_2$

(7) 

(8)

(g) >=

(10) } =

A diagram consisting of two pairs of white lines forming brackets or parentheses, and a single green oval shape.



100504

(ii) KN802

(ijm-CPBA)

(ii) 130

m-CPB

23

03/2

03 | 420

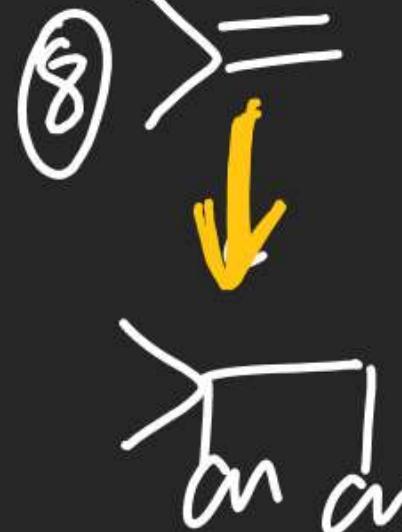
Not $\text{KMnO}_4 \rightarrow \text{K}_2\text{MnO}_4 + \text{MnO}_2$

$$2\text{Na}_2\text{O}_2 + 2\text{CO}_2 \rightarrow 2\text{Na}_2\text{CO}_3 + \text{O}_2$$

$M_2 = 0 + M_3 - M = 0$

$$\text{Diagram: } \text{A trapezoid with a diagonal line from top-left to bottom-right. A horizontal line segment connects the top vertex to the midpoint of the bottom side. The left side is labeled 'a' and the right side is labeled 'b'.}$$

Nishant Jindal



1



$$y_0 + a_2 =$$

$\omega_0 + \omega$

$$y = 0 + \omega_2$$

$$(9) \quad > \begin{matrix} \diagup \\ \diagdown \end{matrix} \downarrow \quad \begin{matrix} * \\ \text{on on} \end{matrix} \quad 2(\pm)$$

11

$\alpha(\pm)$

3

The diagram shows two chemical structures. The left structure is a branched chain with a terminal asterisk (*). The right structure is a more complex branched chain with multiple asterisks (*).

$$y = 0 + \text{no } u$$

$\geq 0 + 10$

(10) >=<
no → ← on
))

11

11

1

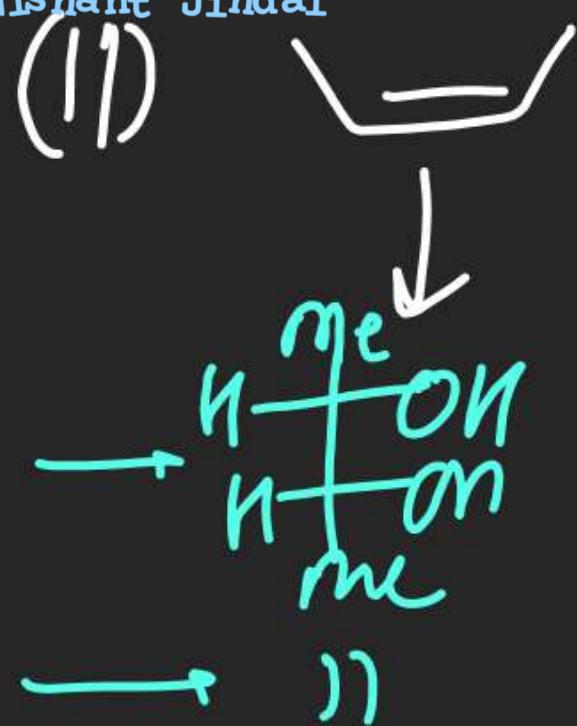
1

1



1

2



(12)



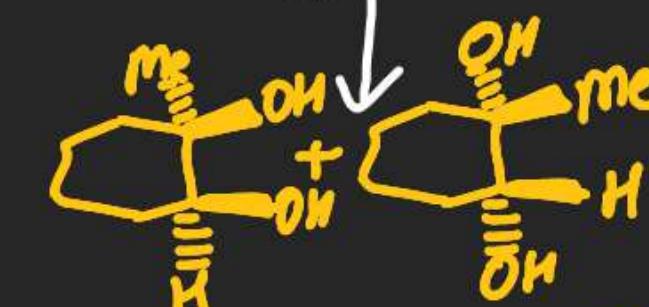
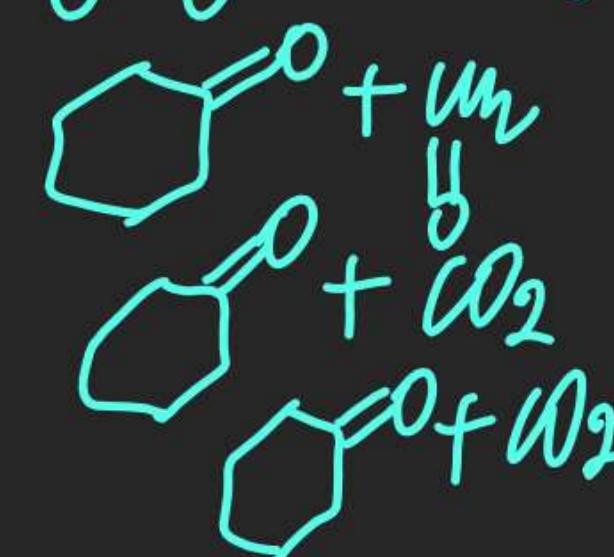
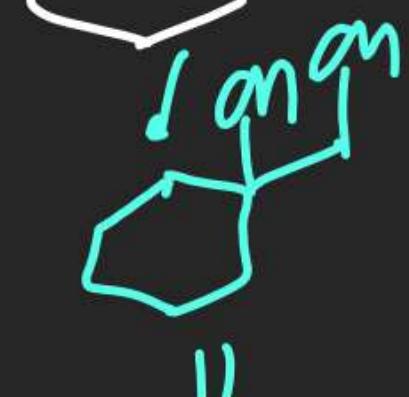
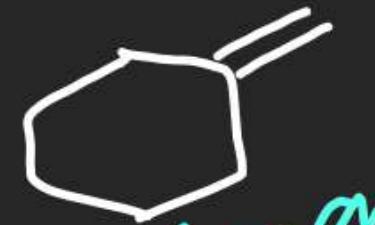
(13)



(14)



(15)



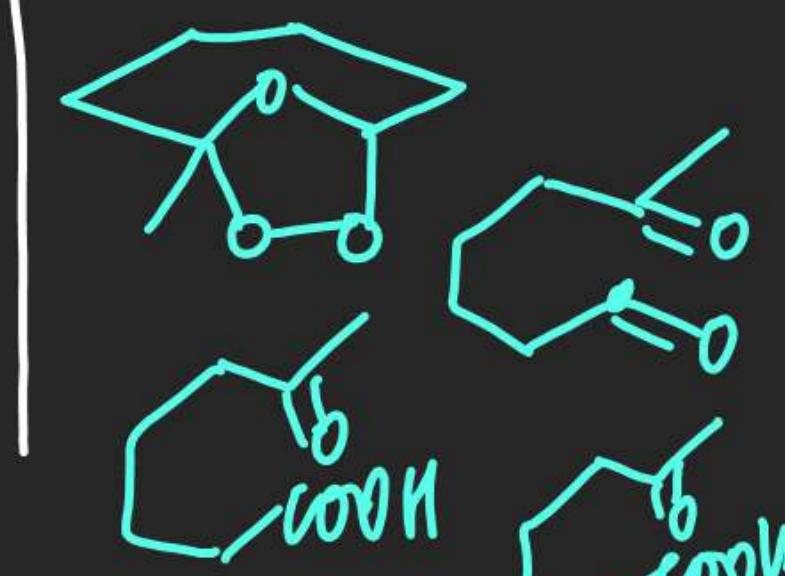
2(±)



2(±)

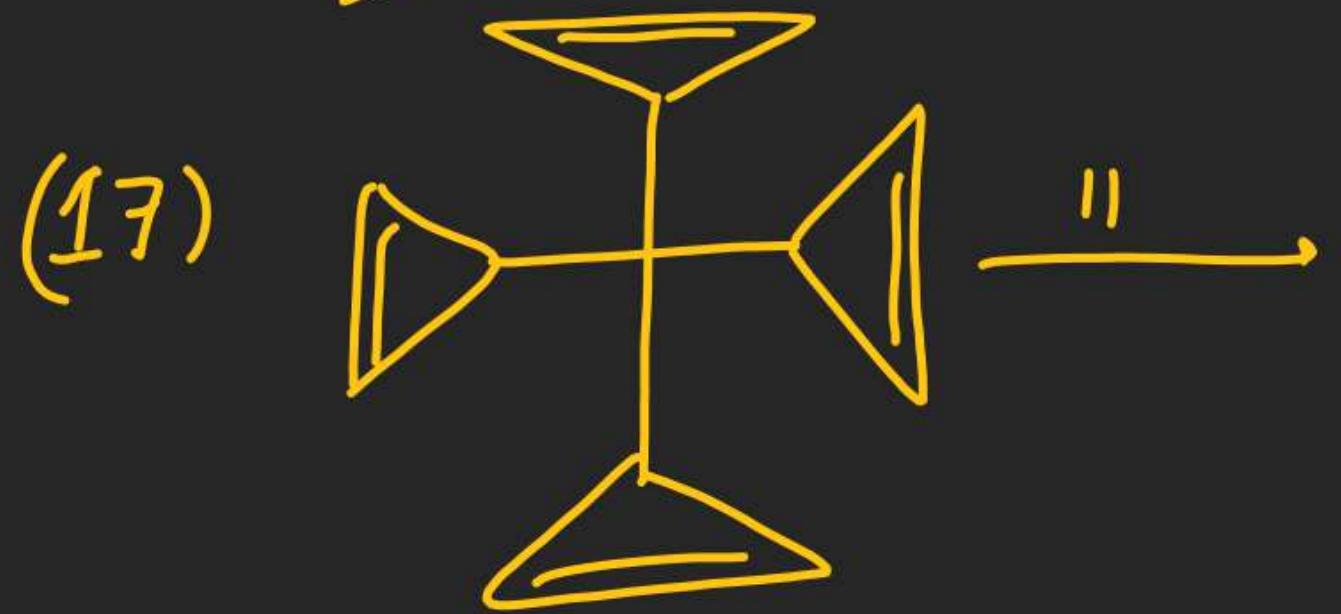


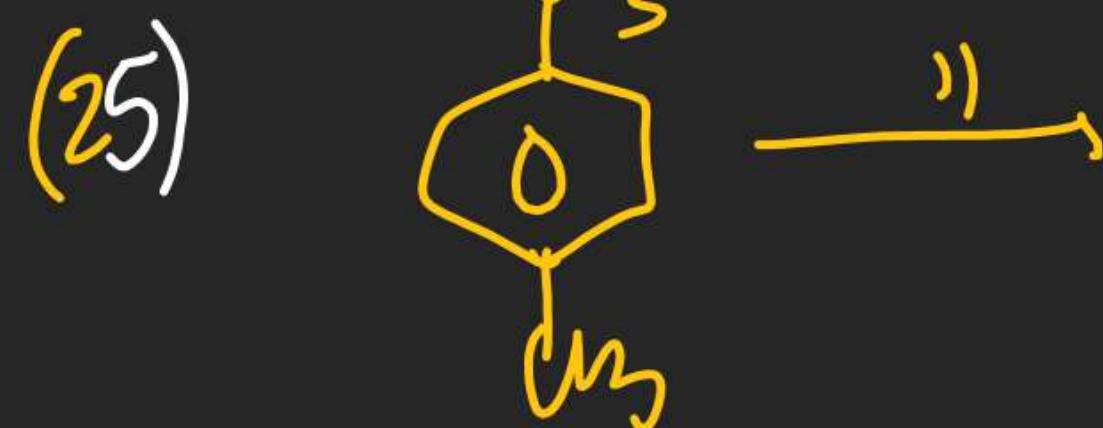
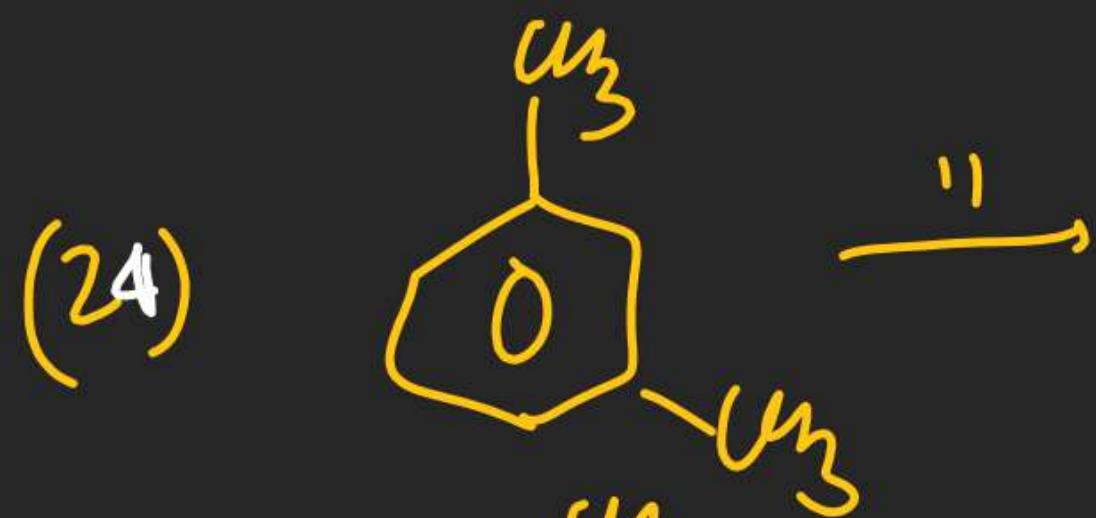
2(±)



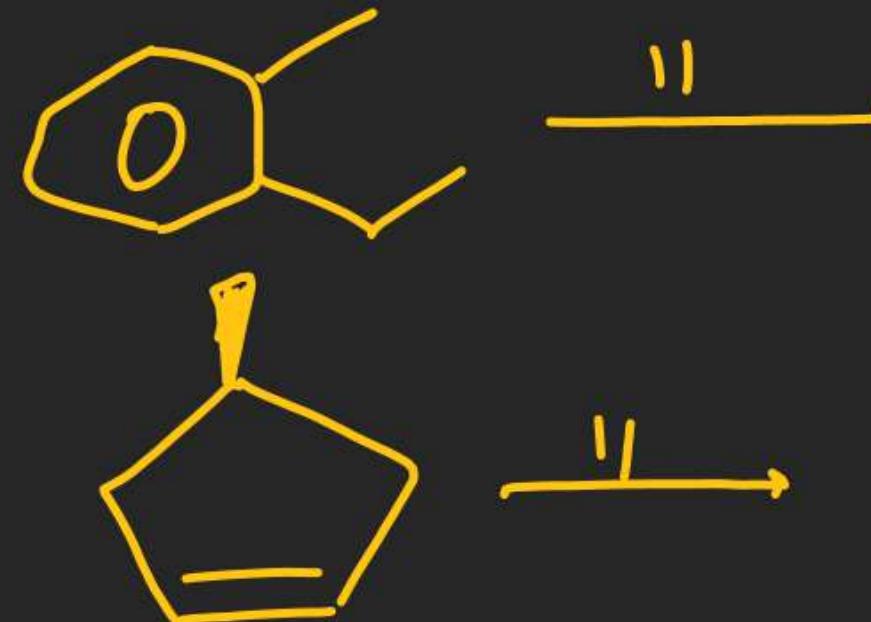
COOH

COOH

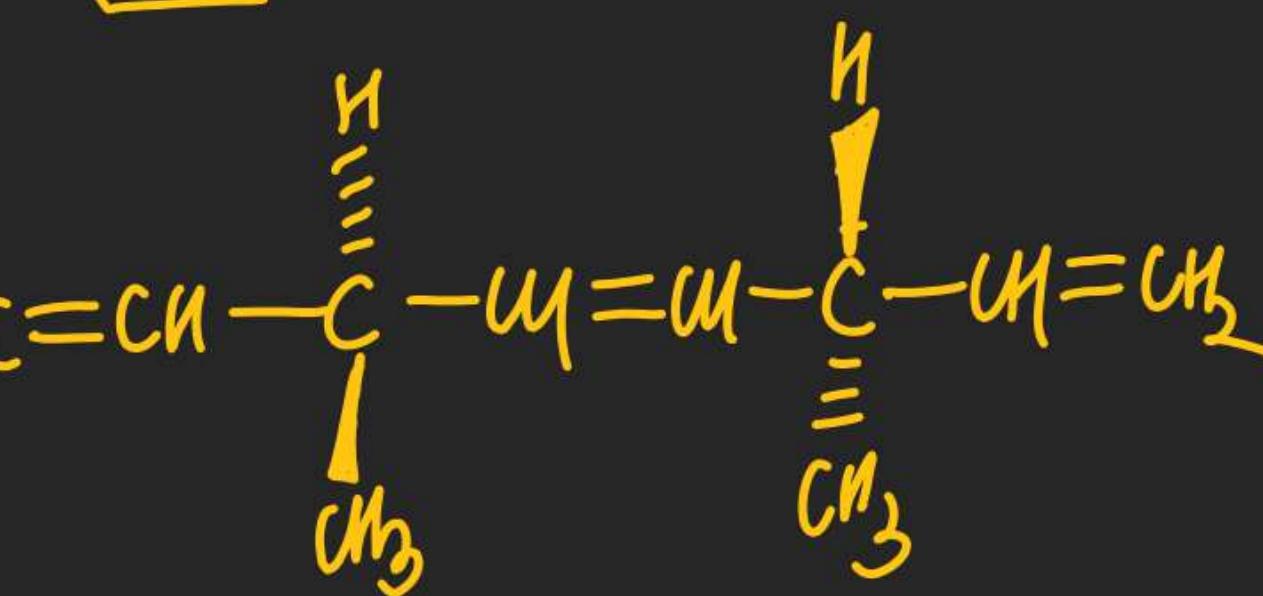




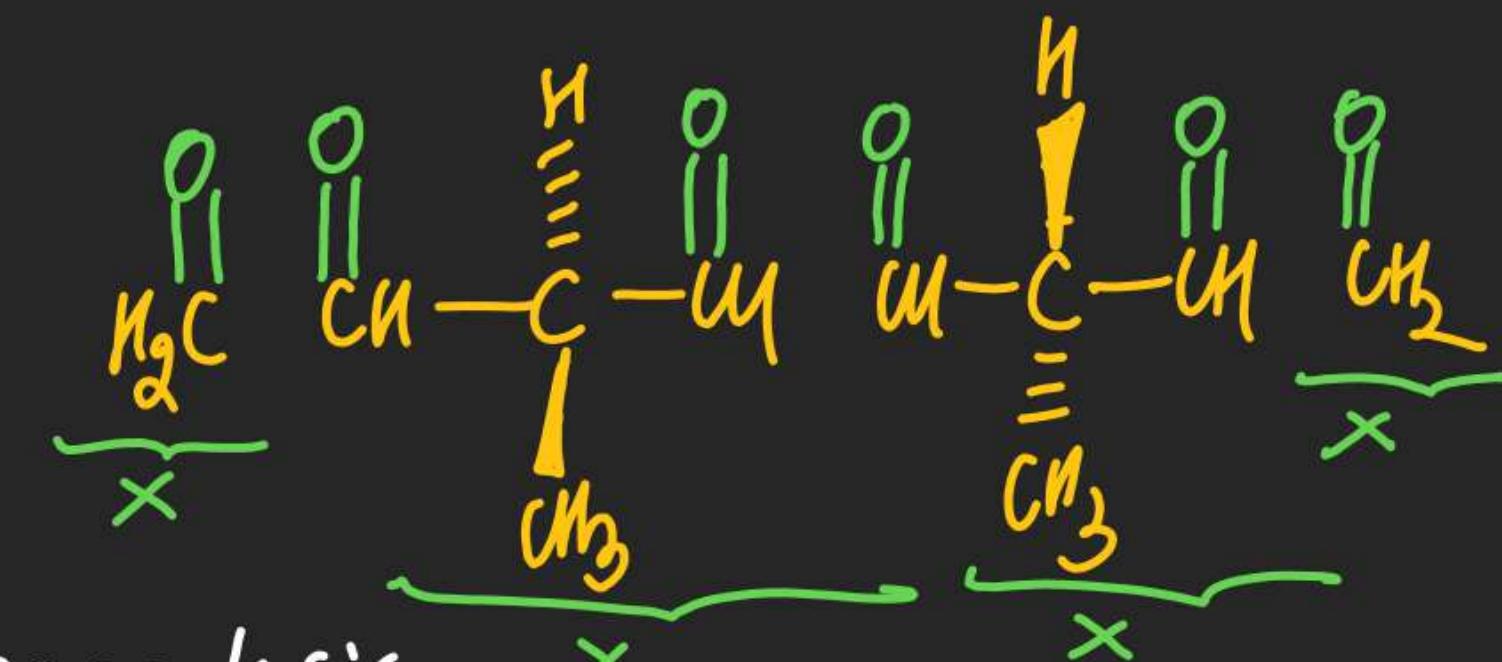
(26) 



(27)



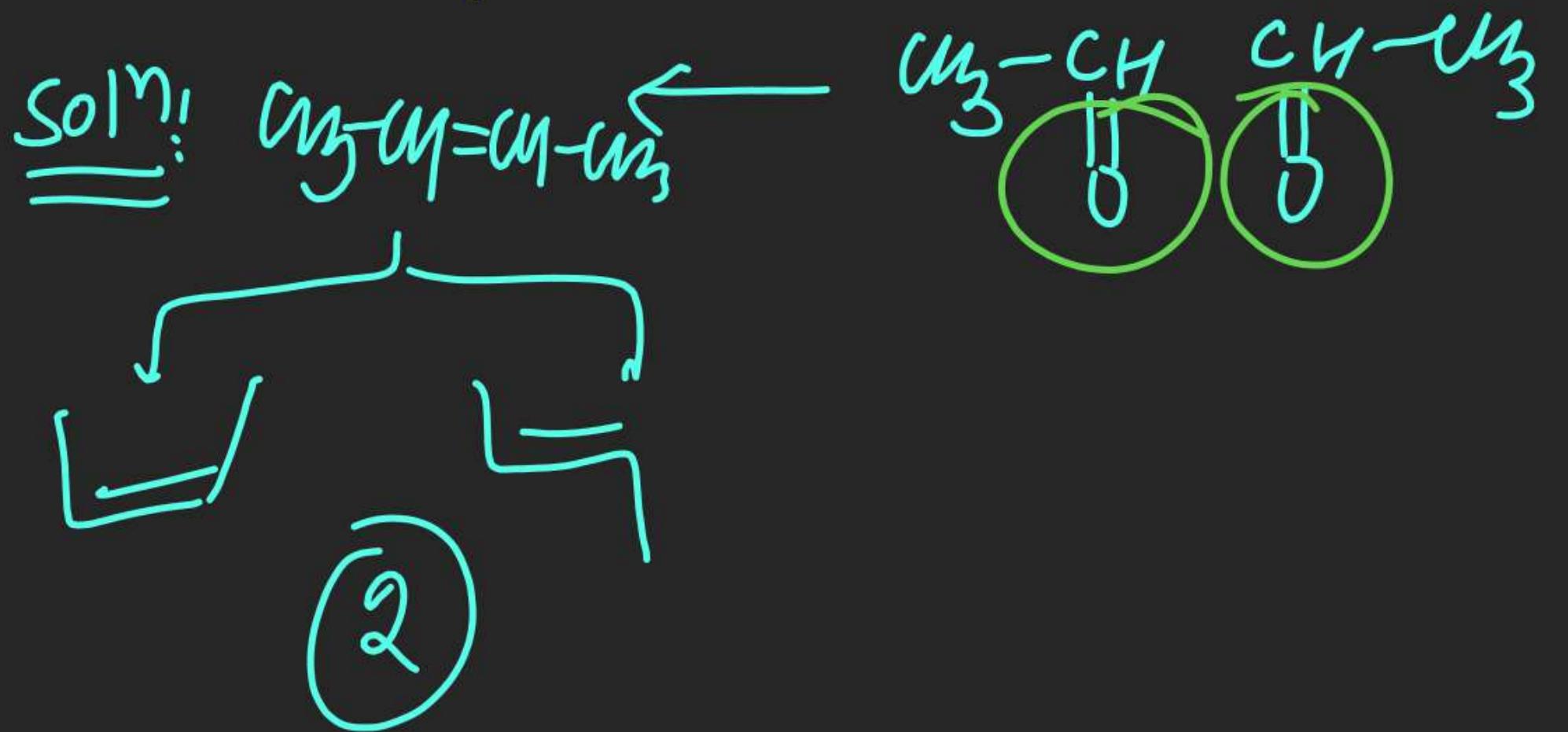
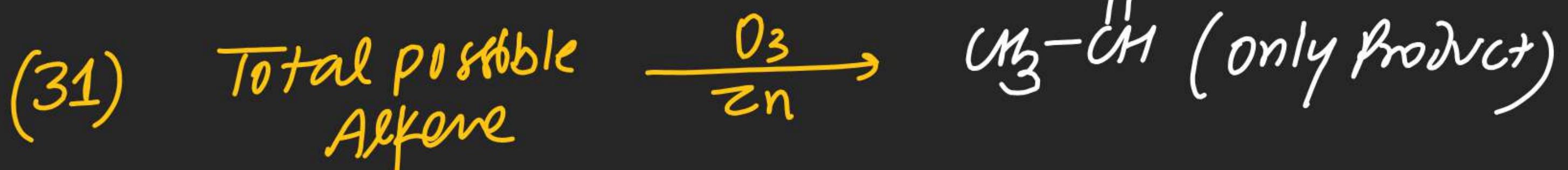
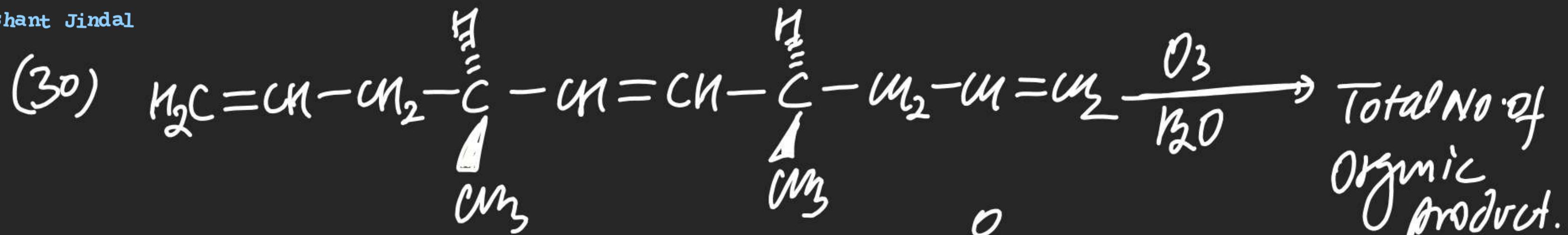
$$(29) \quad M_2 = M - \frac{M}{C} \Delta M$$

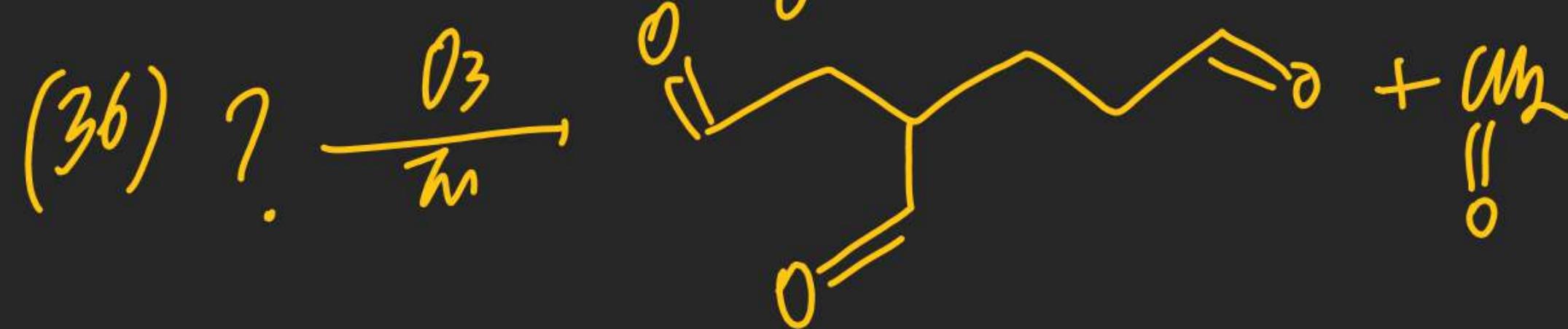
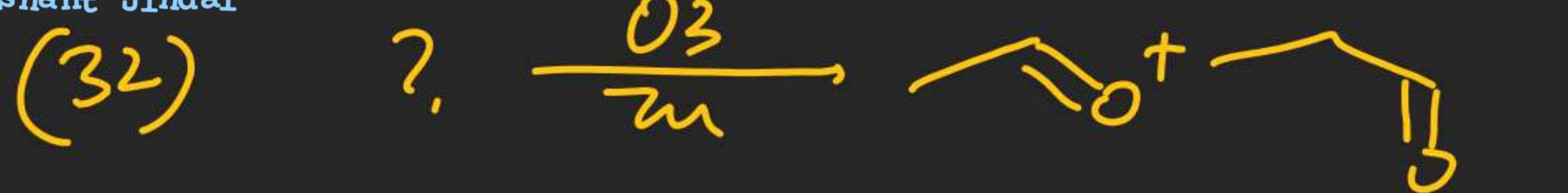


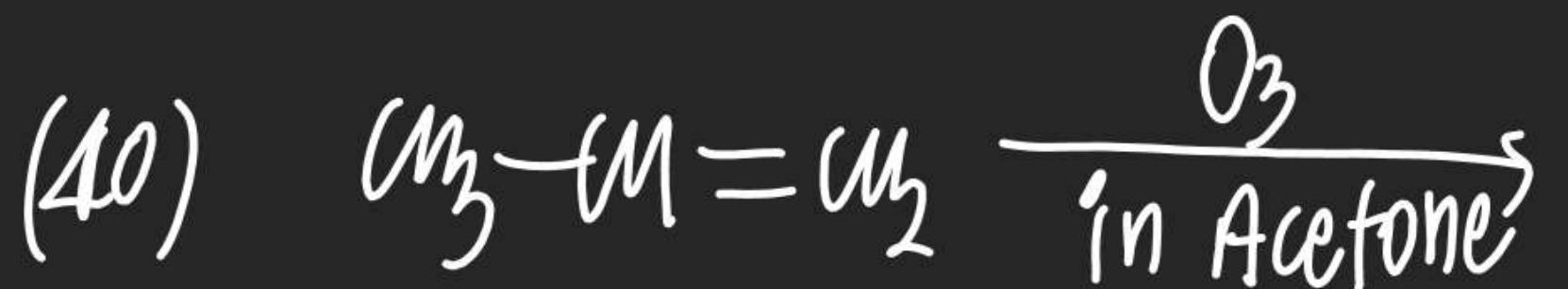
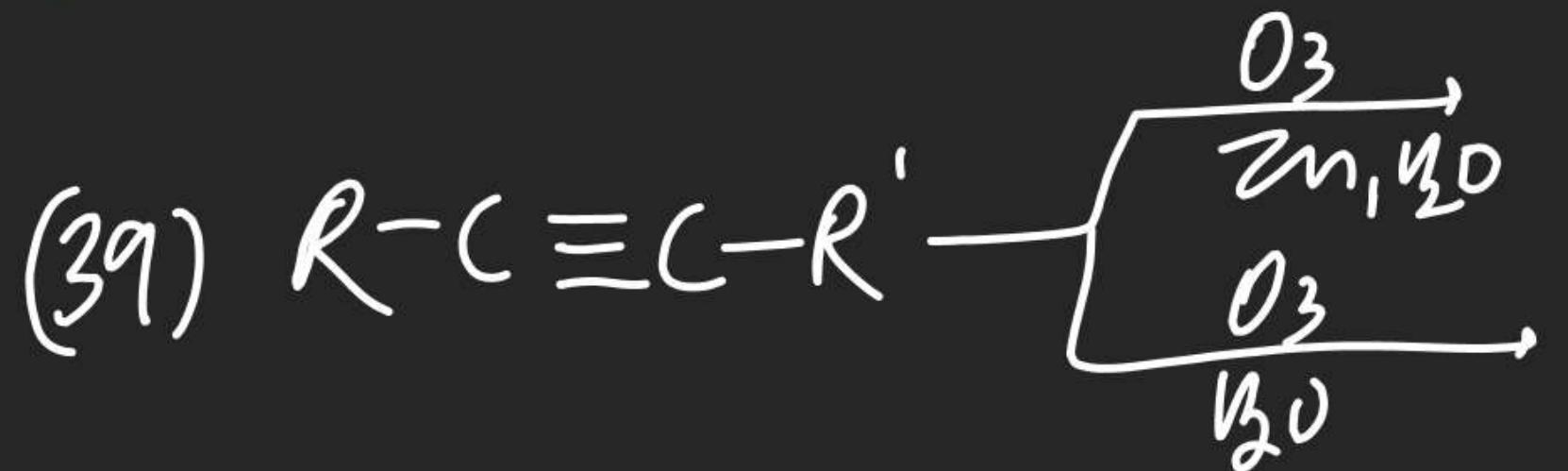
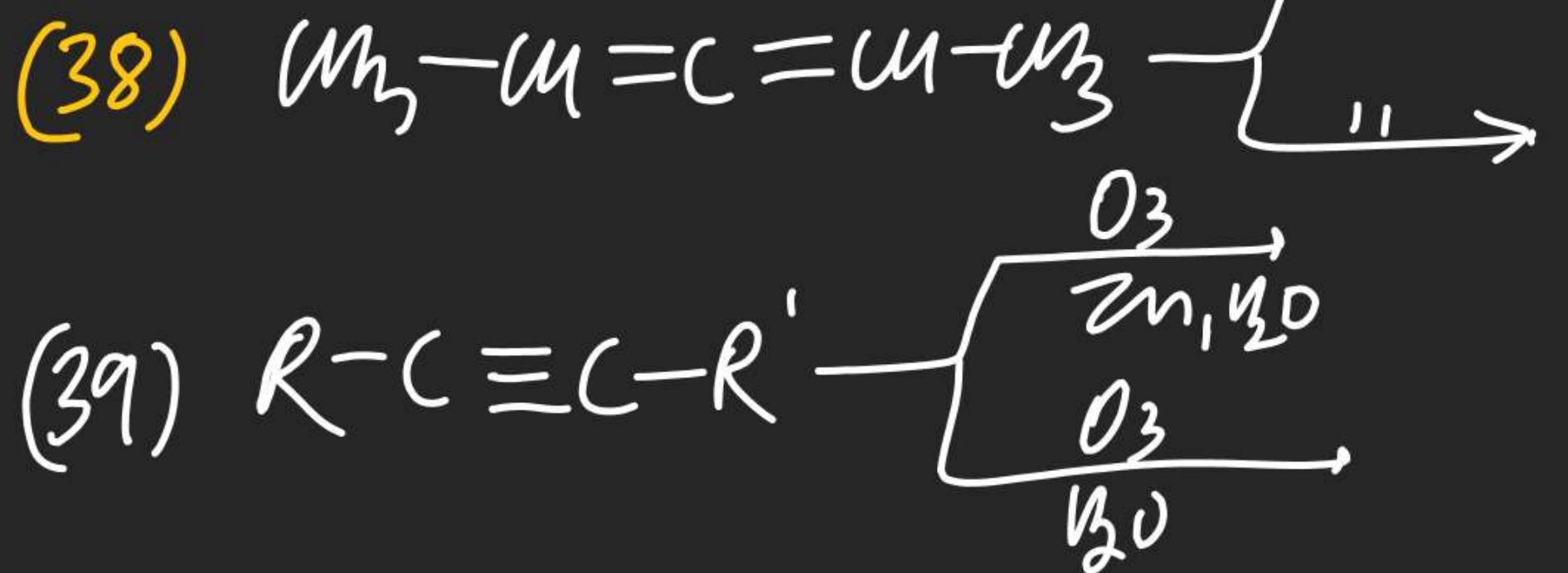
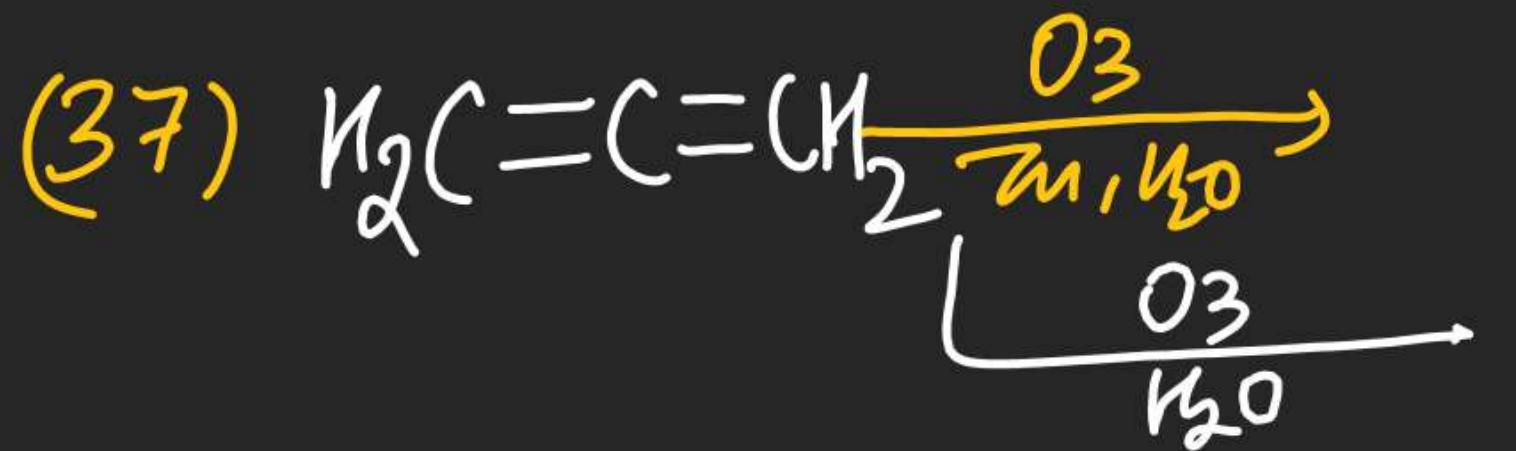
Ozonolysis \rightarrow Total no. of optically Active product.

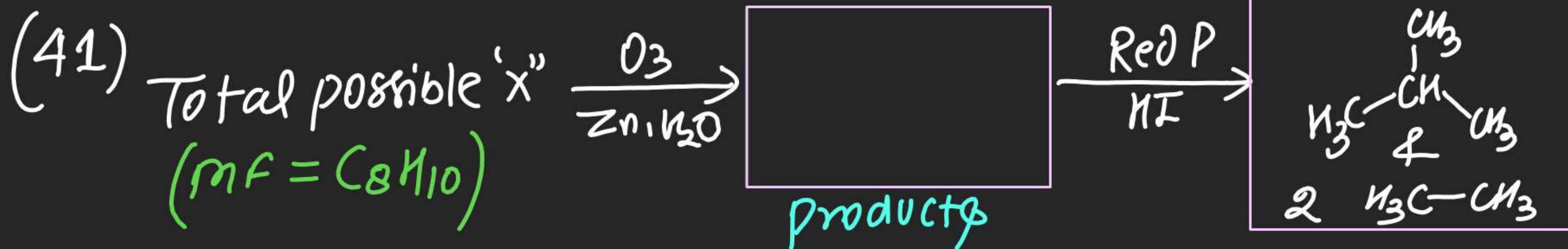
Ans "D":

$$(29) \quad M_2 = M - \frac{M}{C} - M_1 - m = M - \frac{M}{C} - M_1 - \underline{\underline{m}} = M_2 \xrightarrow{\text{Ozone - } 14\text{ days}}$$

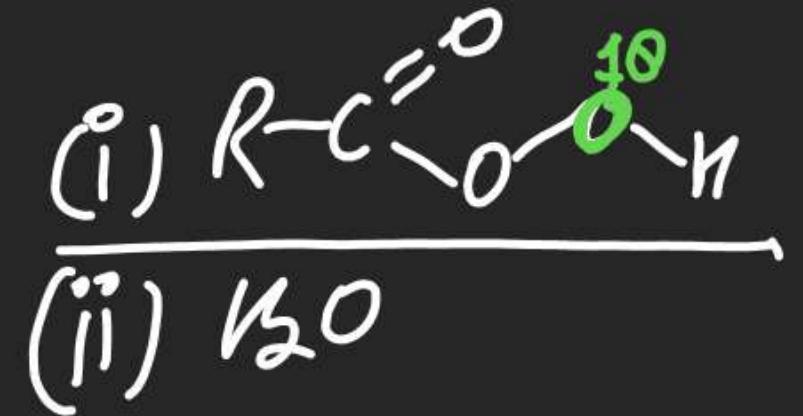




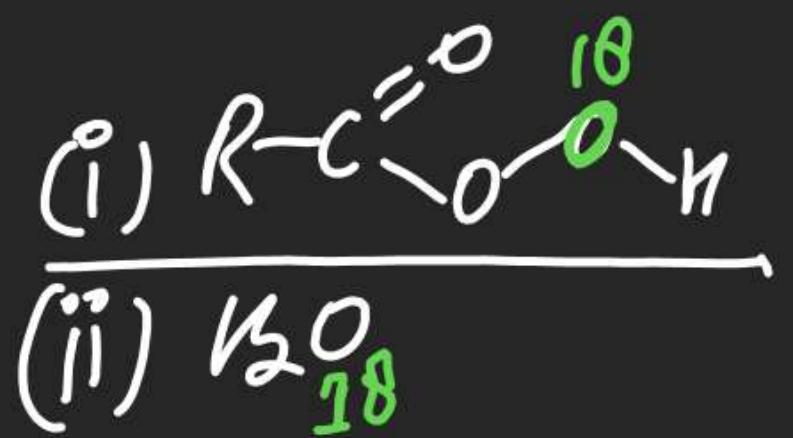




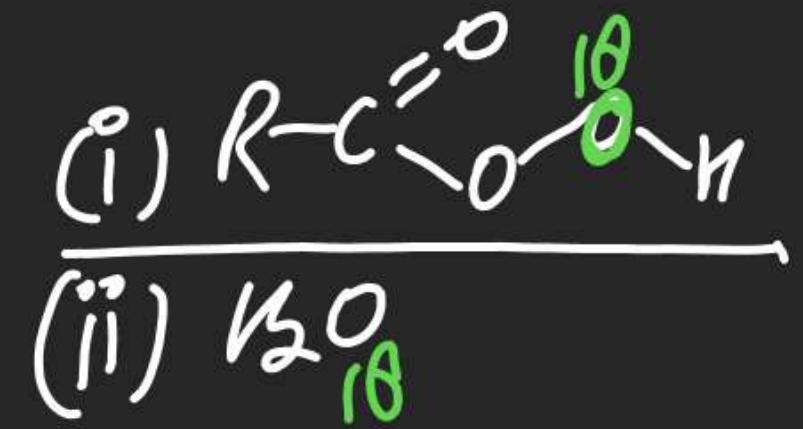
(42)



(43)

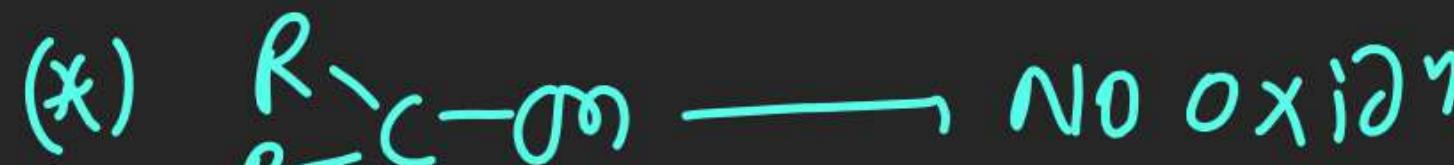
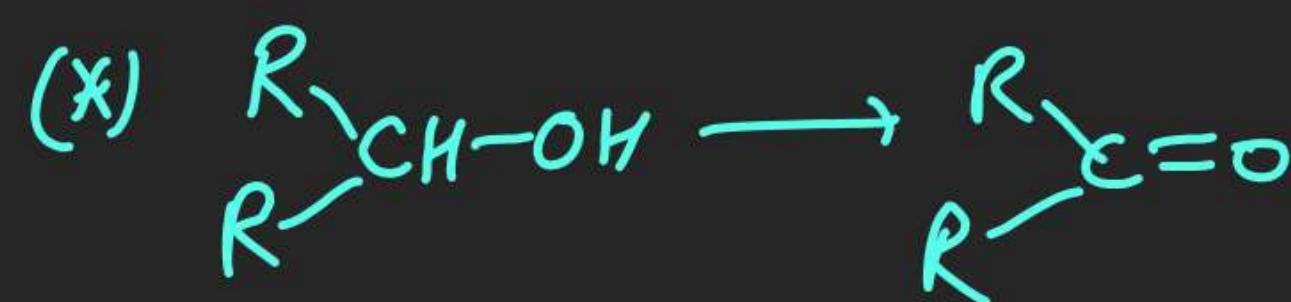
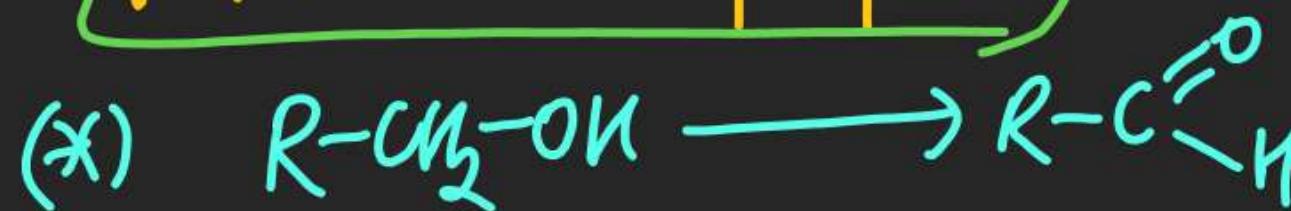


(44)



Oxidation of Alcohol:-

mild oxidising agent



PCC

PDC

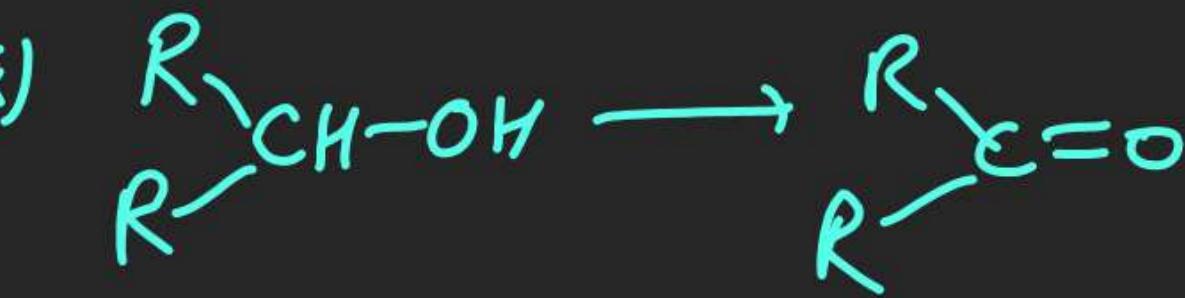
Collins

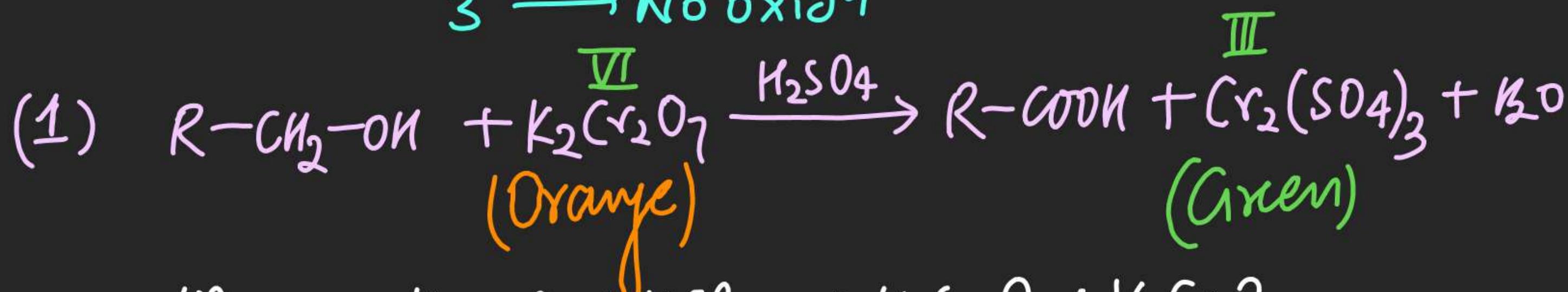
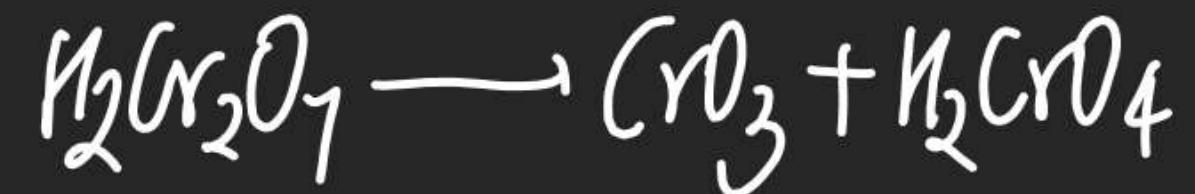
NBS

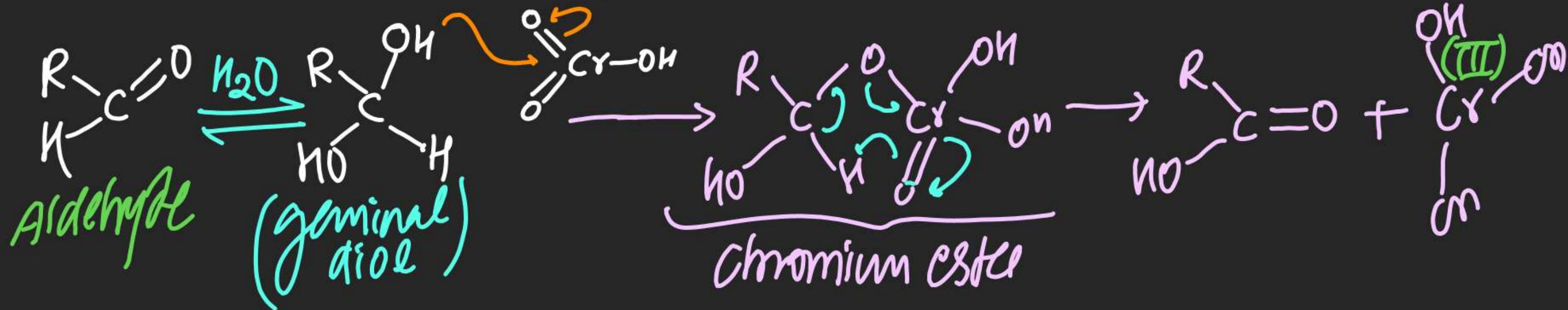
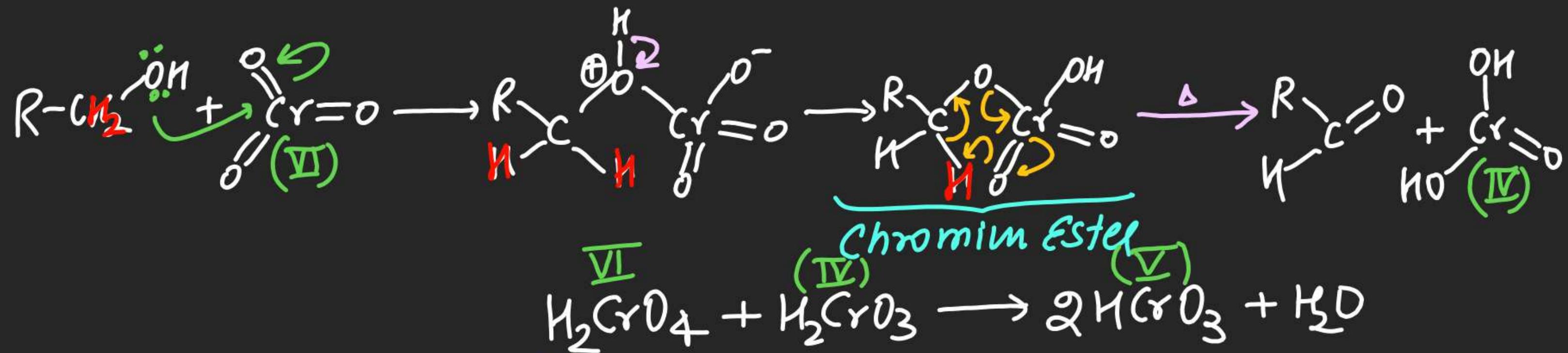
Cu, Δ

 MnO_2 TSL, NaHCO_3 , Δ

Strong oxidising agent

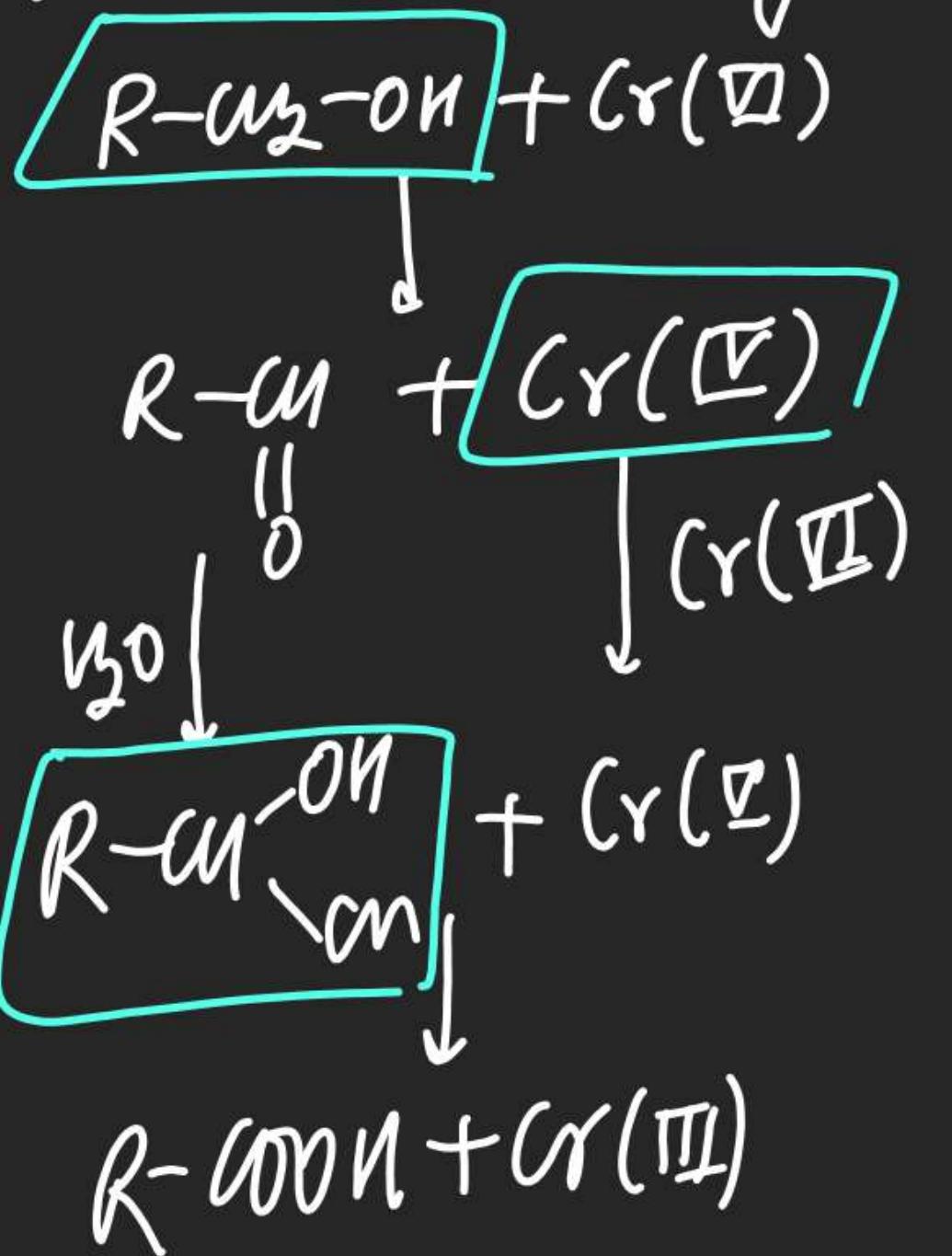
Acidic $\text{K}_2\text{Cr}_2\text{O}_7$ KMnO_4 $\text{CrO}_3/\text{H}_2\text{SO}_4$

(+) Oxidation By Acidic $K_2Cr_2O_7$: \Rightarrow Strong O⁻ Agent \Rightarrow Oxidises $1^\circ \rightarrow$ Acid $2^\circ \rightarrow$ Ketone $3^\circ \rightarrow$ No oxidnmech:



Note (i) Cr-Ester intermediate

(ii) Species oxidised during Rx^n



(iii) O.S change



(iv) Color change

~~Cr(IV)~~ Orange \rightarrow Green

(v) Rx^n is used in POC for
distinguishion of Alcohol.