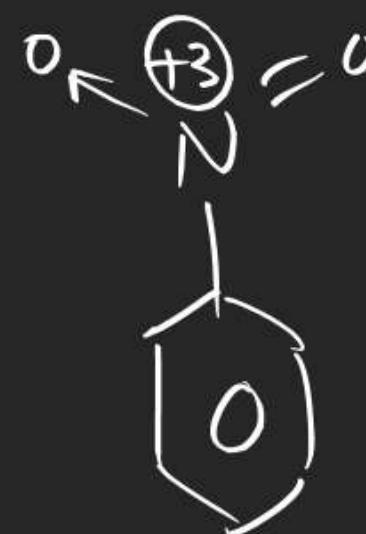
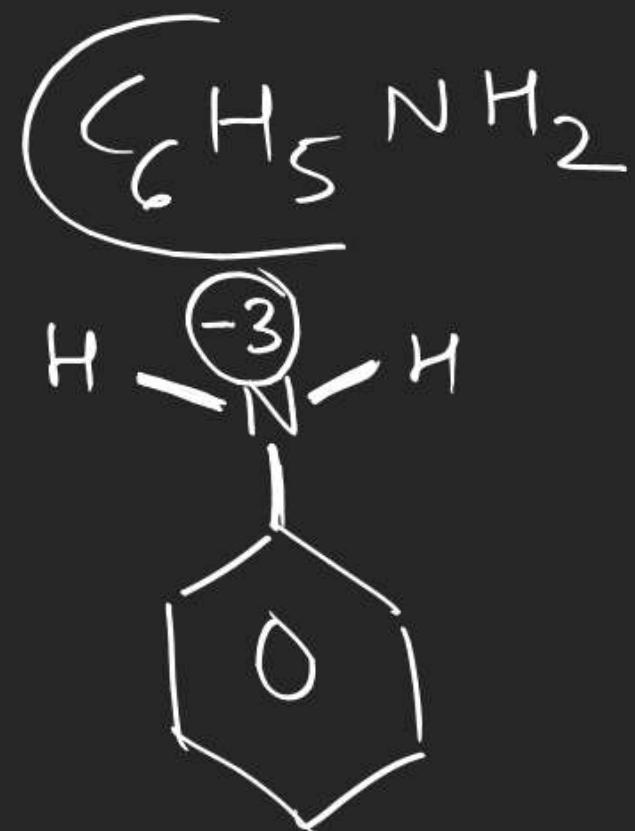


O-I 1-10
 S-I 1-3
 Redox



$$2 + 2x - 14 = 0$$

$$x = 6$$



$$x = 2.5^-$$

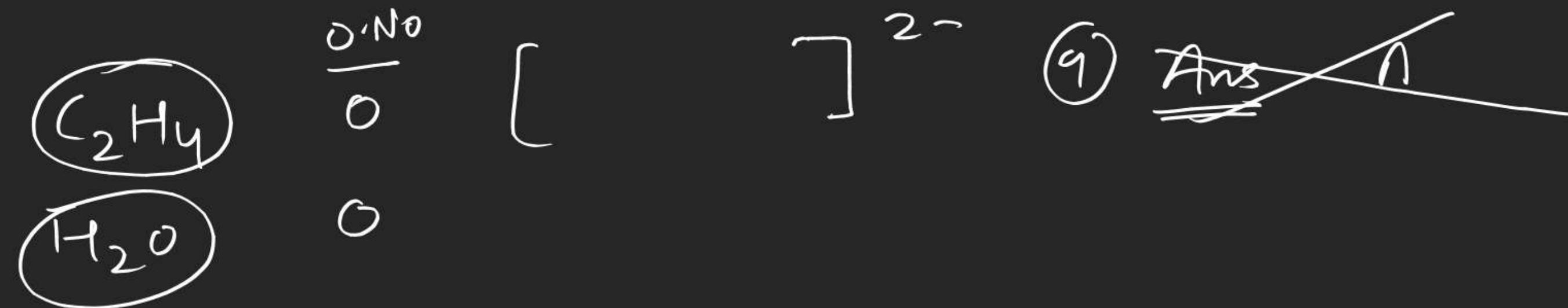


$$+2 + 2x - 6 = 0$$

$$x = 2$$

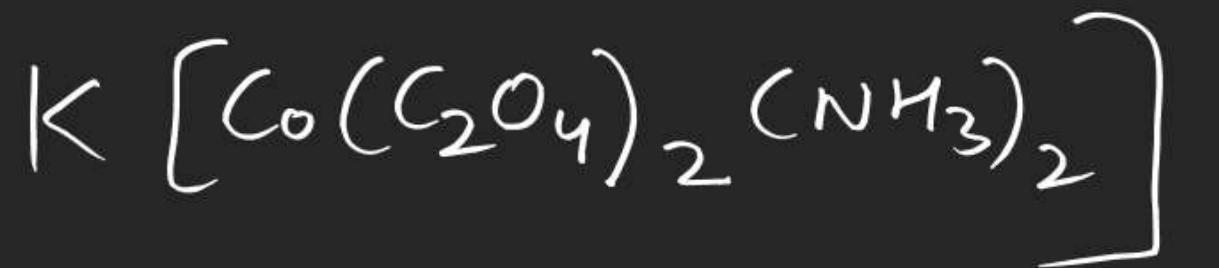
0

SO ²⁺	4
SO ₄ ²⁻	6



⑩ $2x - 8 + 0 + 0 = -2$

$x = 3$

S-1

①

$$+1 + x - 4 + 0 = 0$$

$$\underline{x = 3}$$

② (8)



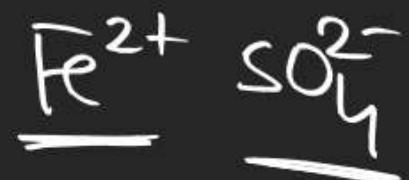
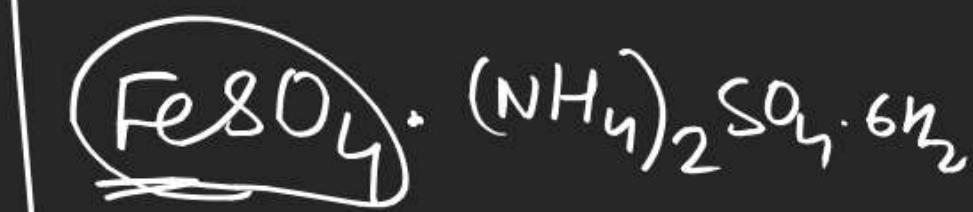
$$2x + 6 - 2 = 0$$

$$\underline{x = -2}$$



akk7007

$$\begin{array}{c} \textcircled{12} \\ \text{CH}_3^+ \quad \text{HSO}_3^- \\ \hline 1+x-6 = -1 \\ \textcircled{x = 4} \end{array}$$



$$\textcircled{13} \quad \underline{\underline{S-C=S}}$$

3



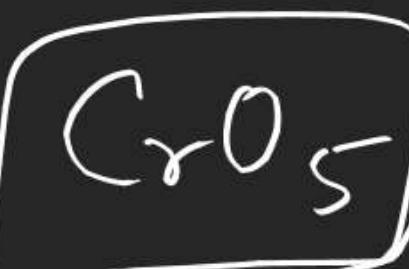
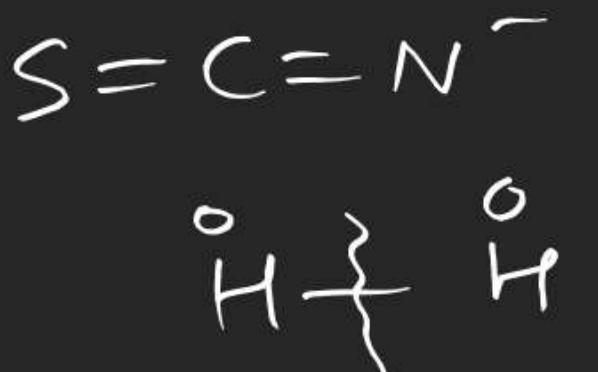
$$2 + 2x - 6 = 0$$

$$n = 2$$



$$\frac{6-2}{2} = \boxed{2}$$

④ Ba (SCN)₂



FIND OXIDATION NUMBER OF UNDERLINED ELEMENTS



$$3 + x - 6 = 0$$



$$1 + x - 4 + 0 = 0$$



$$x + 0 = 0$$



$$x + 0 = 0$$

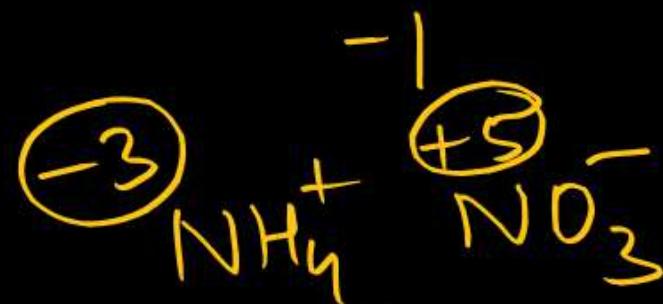


$$12x + 22 - 22 = 0$$



+2

FIND OXIDATION NUMBER OF UNDERLINED ELEMENTS



Avg +1



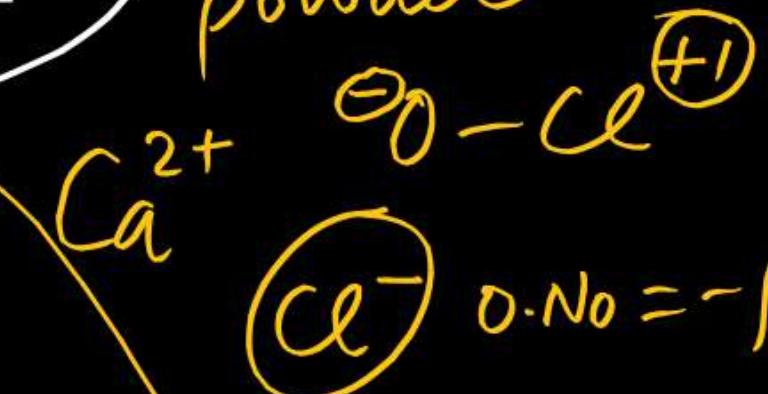
+6



+6



bleaching powder



+3 -2



+5 -2



+1 +2 -3



Q.S

$$+2 -2 + 2x = 0$$

$$\underline{x = 0}$$

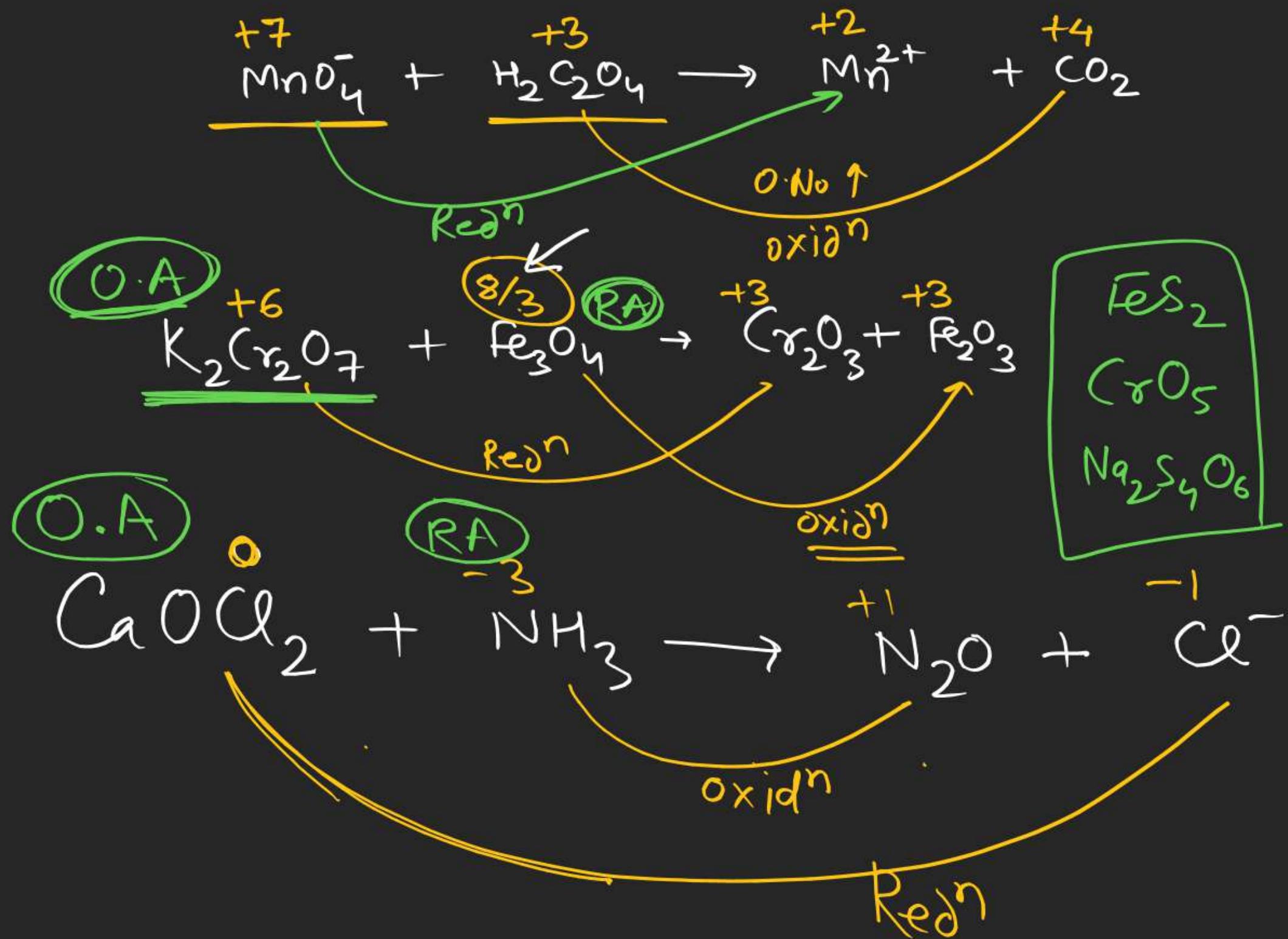
Application of Oxidation number

Oxidation \rightarrow O. No. \uparrow es (rise on atom \uparrow es), loss of e^-

Reduction \rightarrow O. No. \downarrow es (rise on atom \downarrow es) gain of e^-

Redox Rxn \Rightarrow Reduction + oxidation





Balancing of Redox Rxns: →

① Oxidation number

+2



-1



Oxid

0

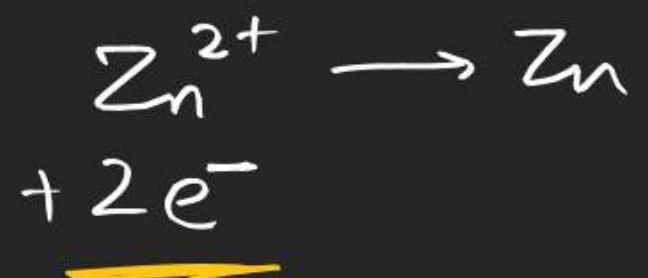


0

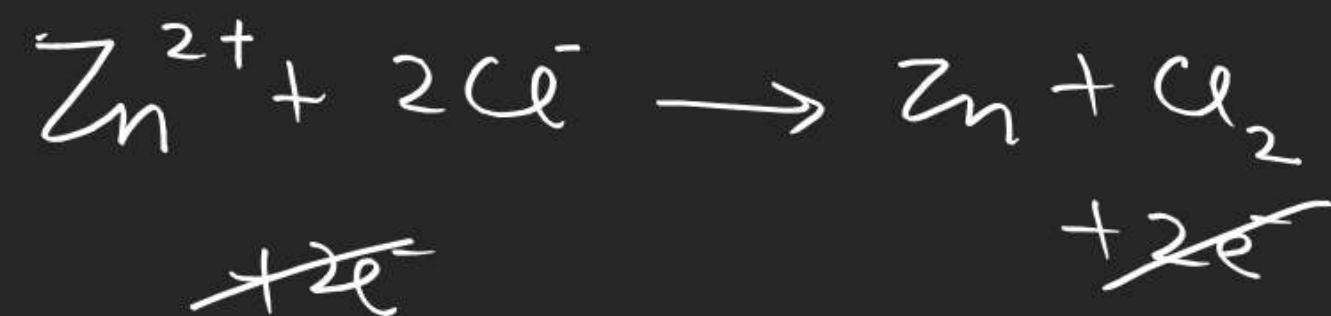


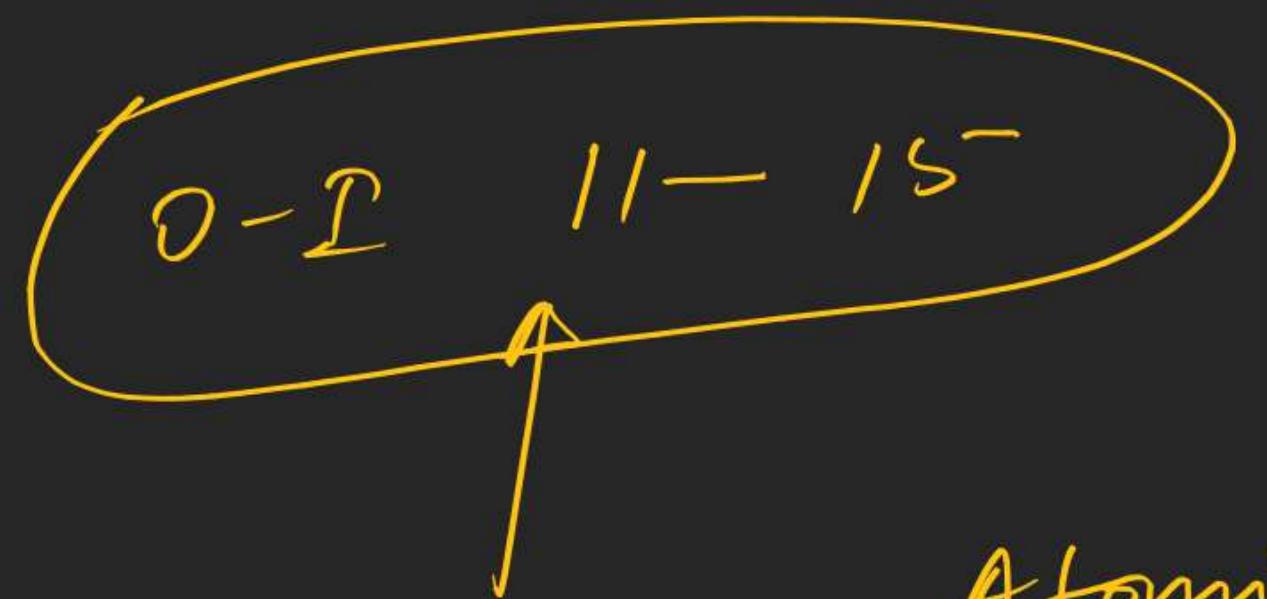
② Ion electron method

O. No method



Ion e⁻ method





Atomic

JEE - Adv

Remaining
JEE - M