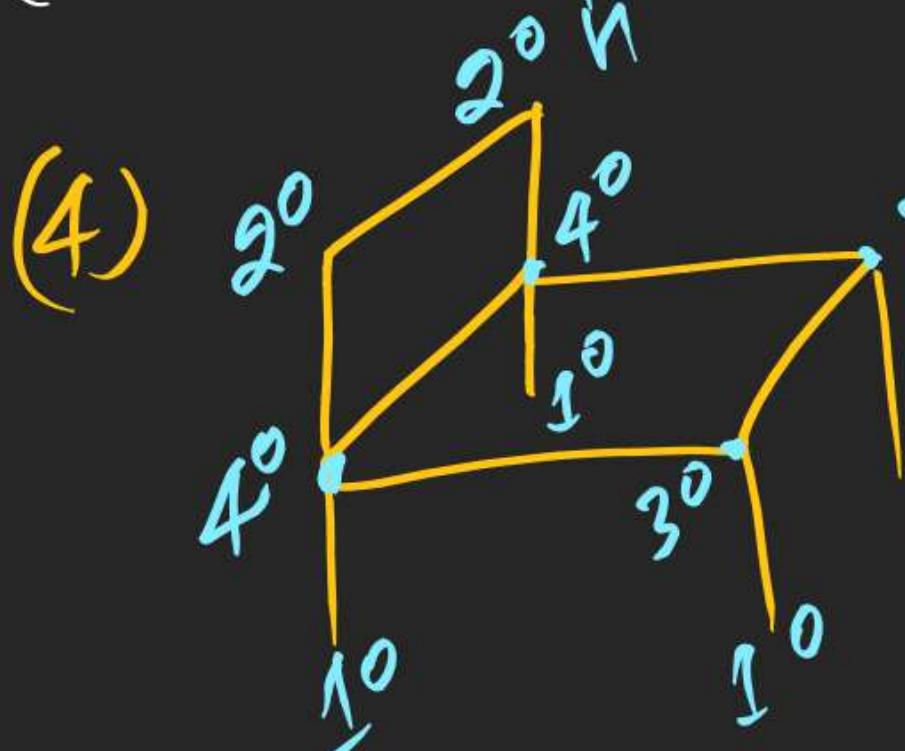
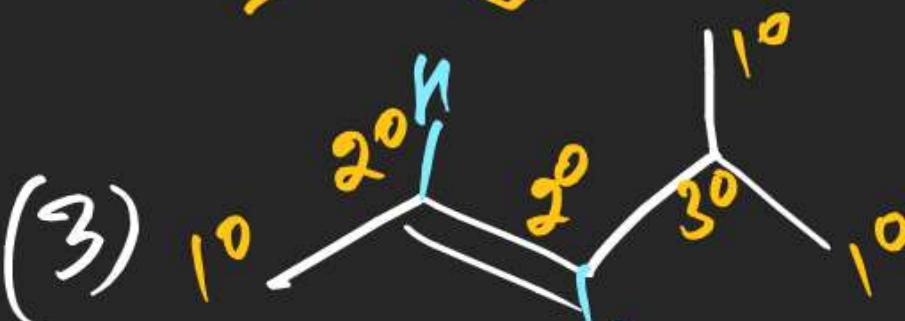
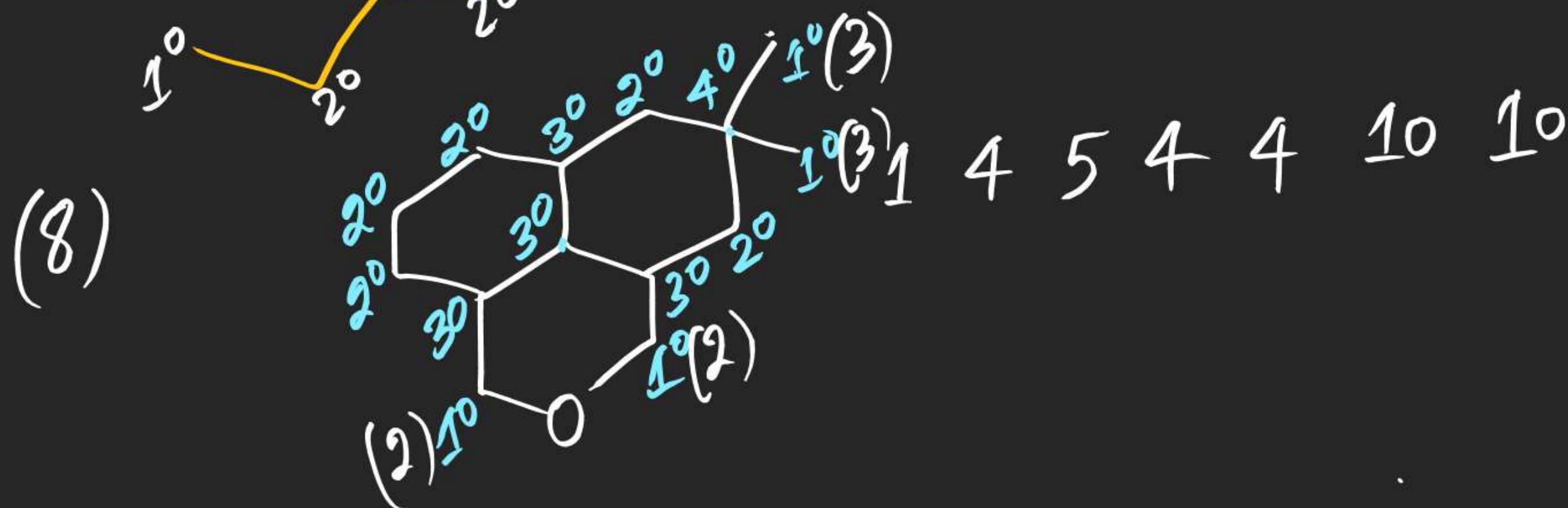


Basic Organic Chemistry

HW Discussion (Theory Copy)



	4°C	3°C	2°C	1°C	3°H	2°H	1°H
(1)	0	1	1	3	1	2	9
(3)	0	1	2	3	1	2	9
(4)	2	2	4	2	4	12	



shant Jindal

(q) 1 2 3 4 7 4 3 11 10

$$(10) \quad \begin{array}{c} 2^{\circ} \\ \swarrow \end{array} \quad \begin{array}{c} 1^{\circ} \\ \nearrow \end{array} \quad \begin{array}{c} 2^{\circ} \\ \text{---} \\ 2^{\circ} \end{array} \quad \begin{array}{c} 2^{\circ} \\ \text{---} \\ 2^{\circ} \end{array} \quad 0 \quad 0 \quad 8 \quad 2 \quad 0 \quad \boxed{13} \quad \underline{4}$$

$$(10) \quad 2^{\circ} \quad 1^{\circ} \quad 1^{\circ} \quad 1^{\circ} \quad 2^{\circ} \quad 2^{\circ} \quad 0 \quad 0 \quad 8 \quad 2 \quad 0 \quad (13) \quad 4$$

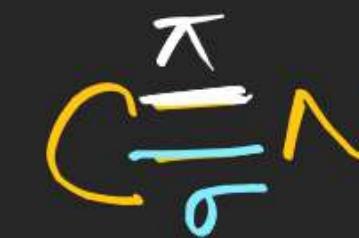
(11)

(#) Sigma & Pi Bond:

\Rightarrow All Single Bonds are Sigma Bonds (σ -Bond)



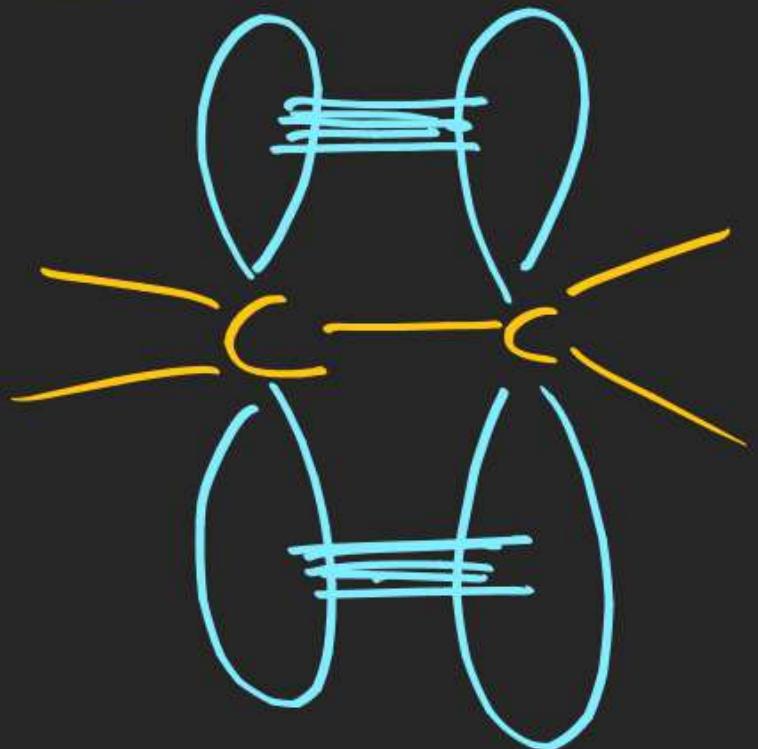
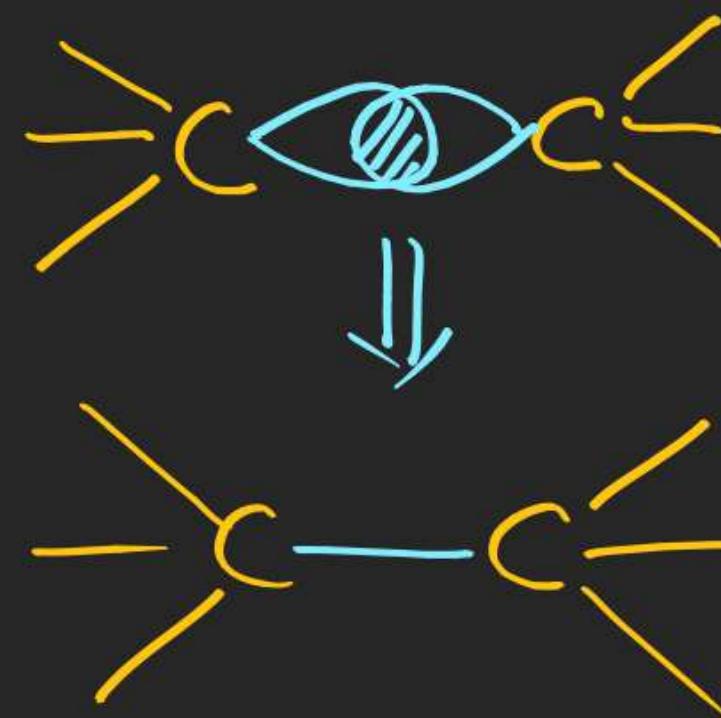
\Rightarrow Each Double Bond contains (1σ & 1π Bond)



\Rightarrow Each Triple Bond contains (1σ & 2π Bonds)



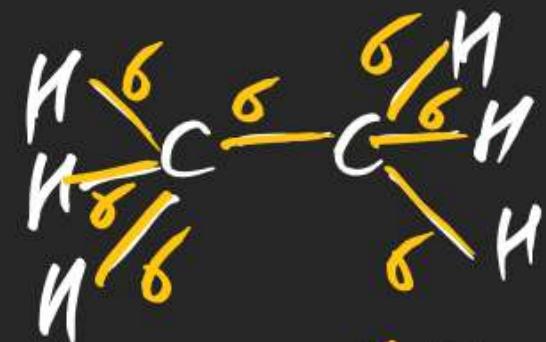
(*) Sigma Bonds are stronger Bonds than π Bonds



wd Study in
IOC in detail

Nishant Jindal
Calculate σ Bonds & π Bonds

(1)



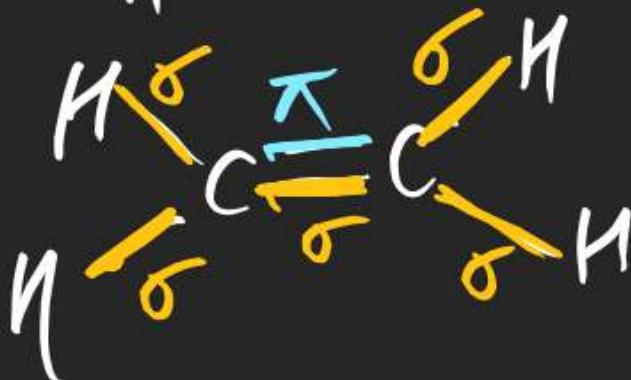
σ Bond

7

π Bond

0

(2)



5

1

(3)



3

2

(4)



(5)



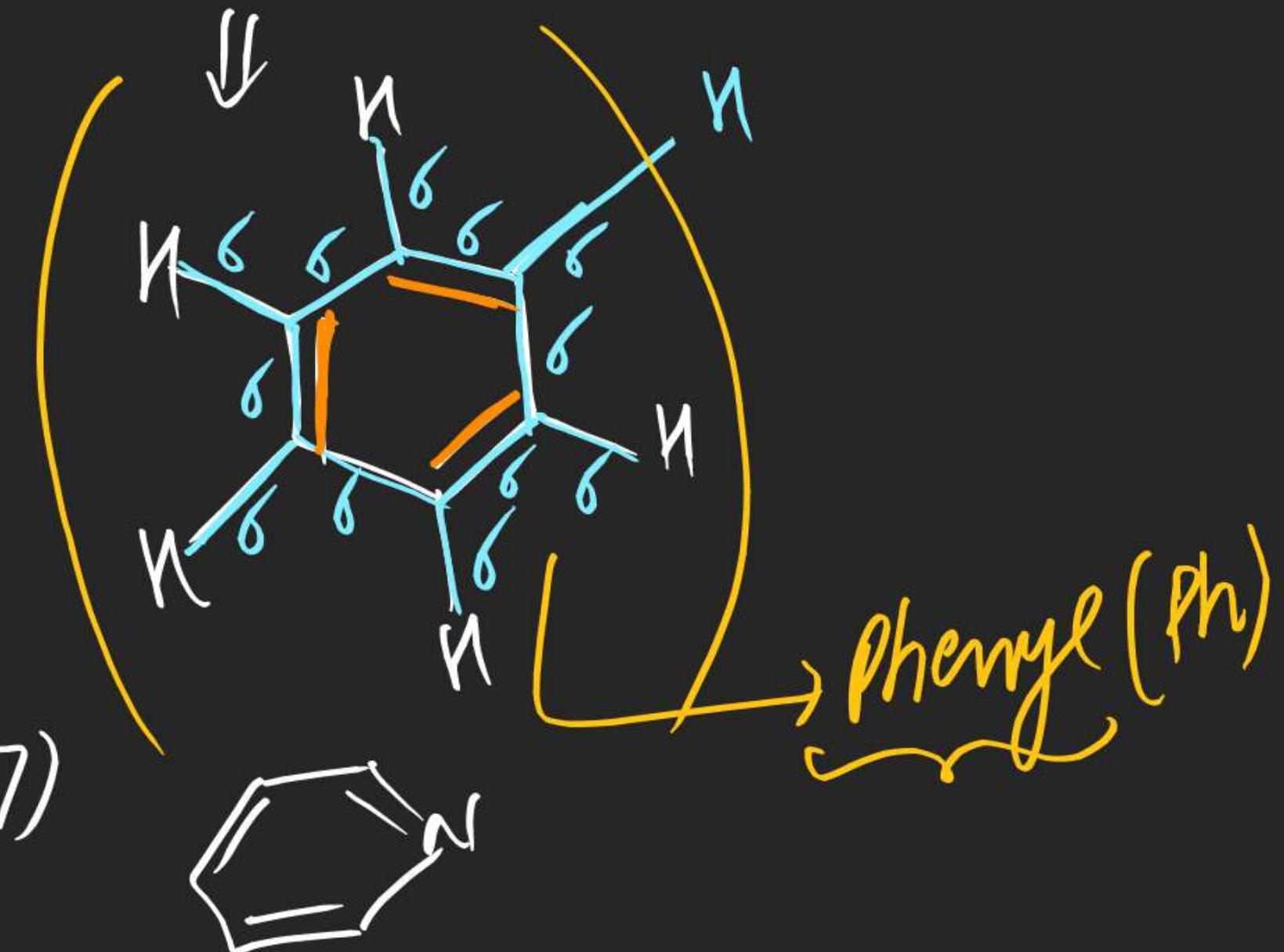
σ Bond π -Bond

(6)



12

3

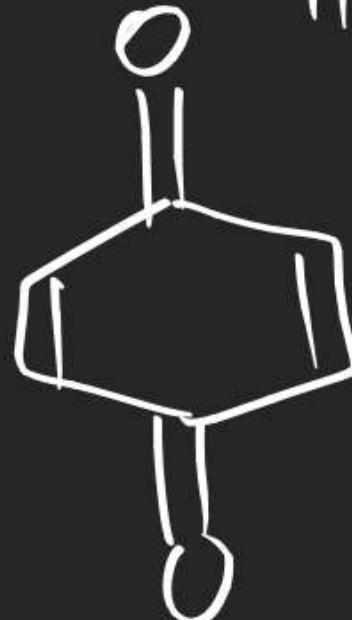


(7)

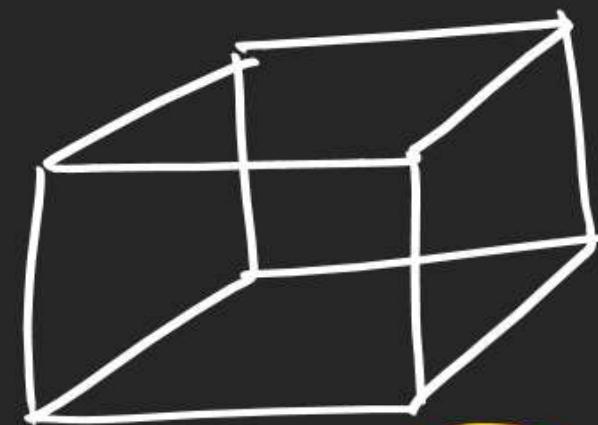
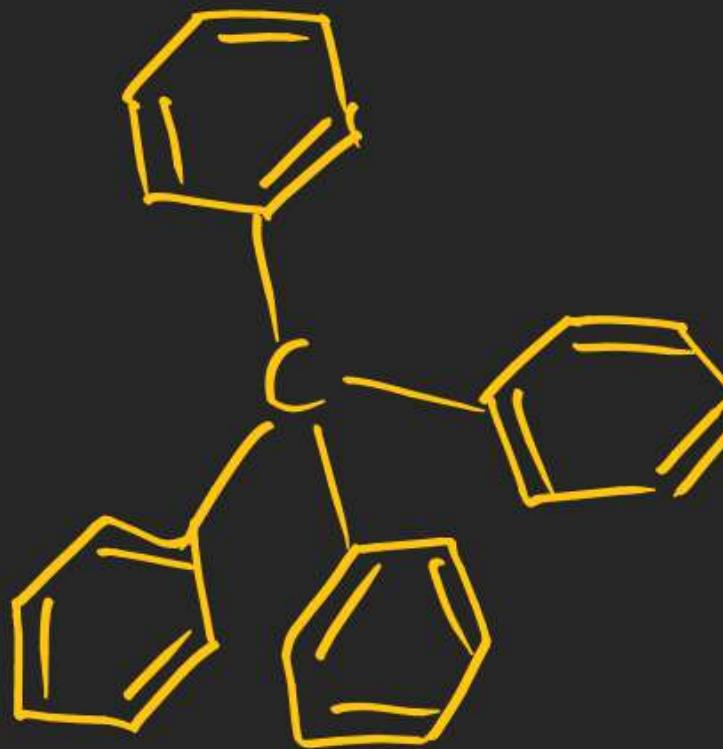
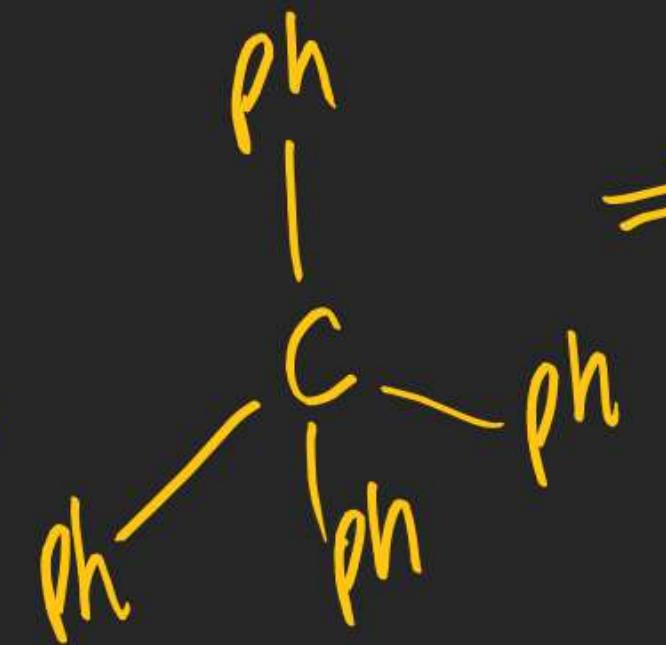
(8)

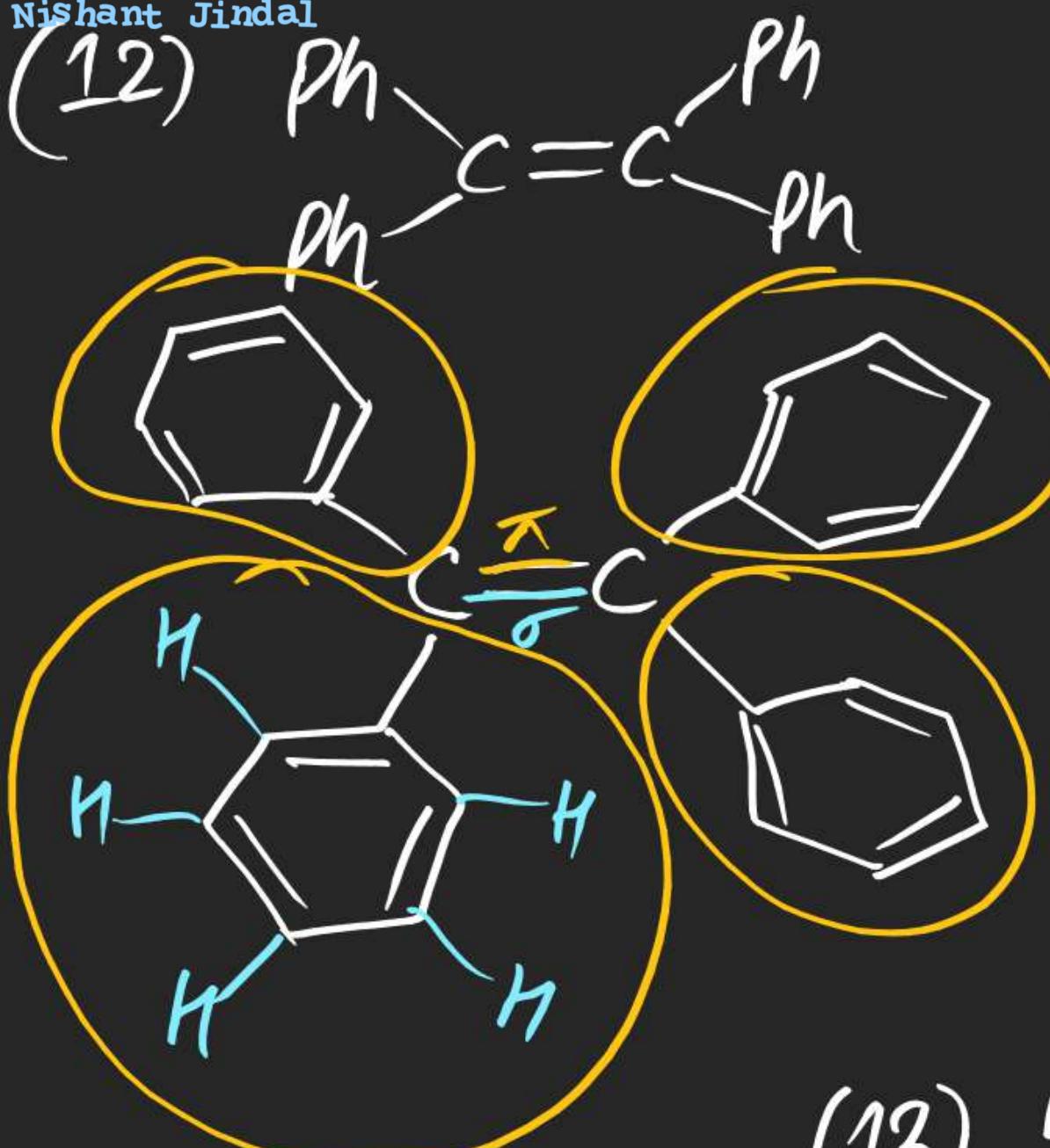


(9)

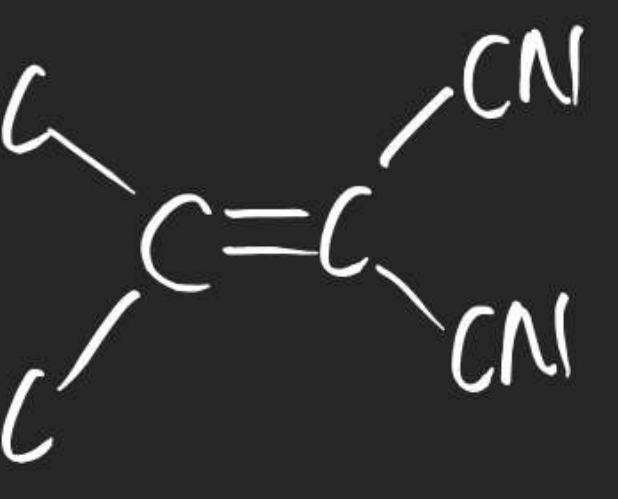


(10)

(11) $\text{Ph}_4\text{C} \Rightarrow$ 



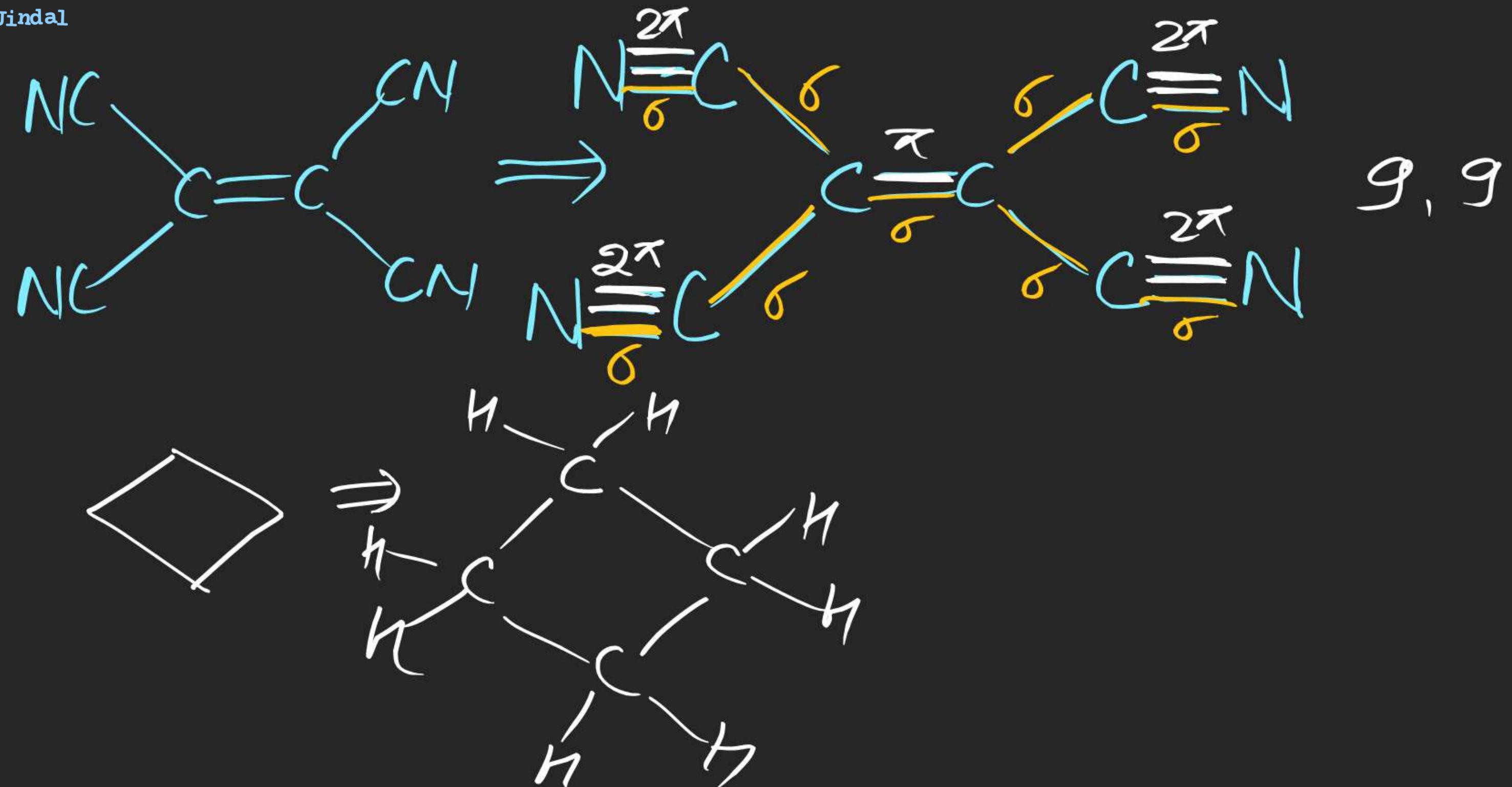
$12\sigma + 3\pi$

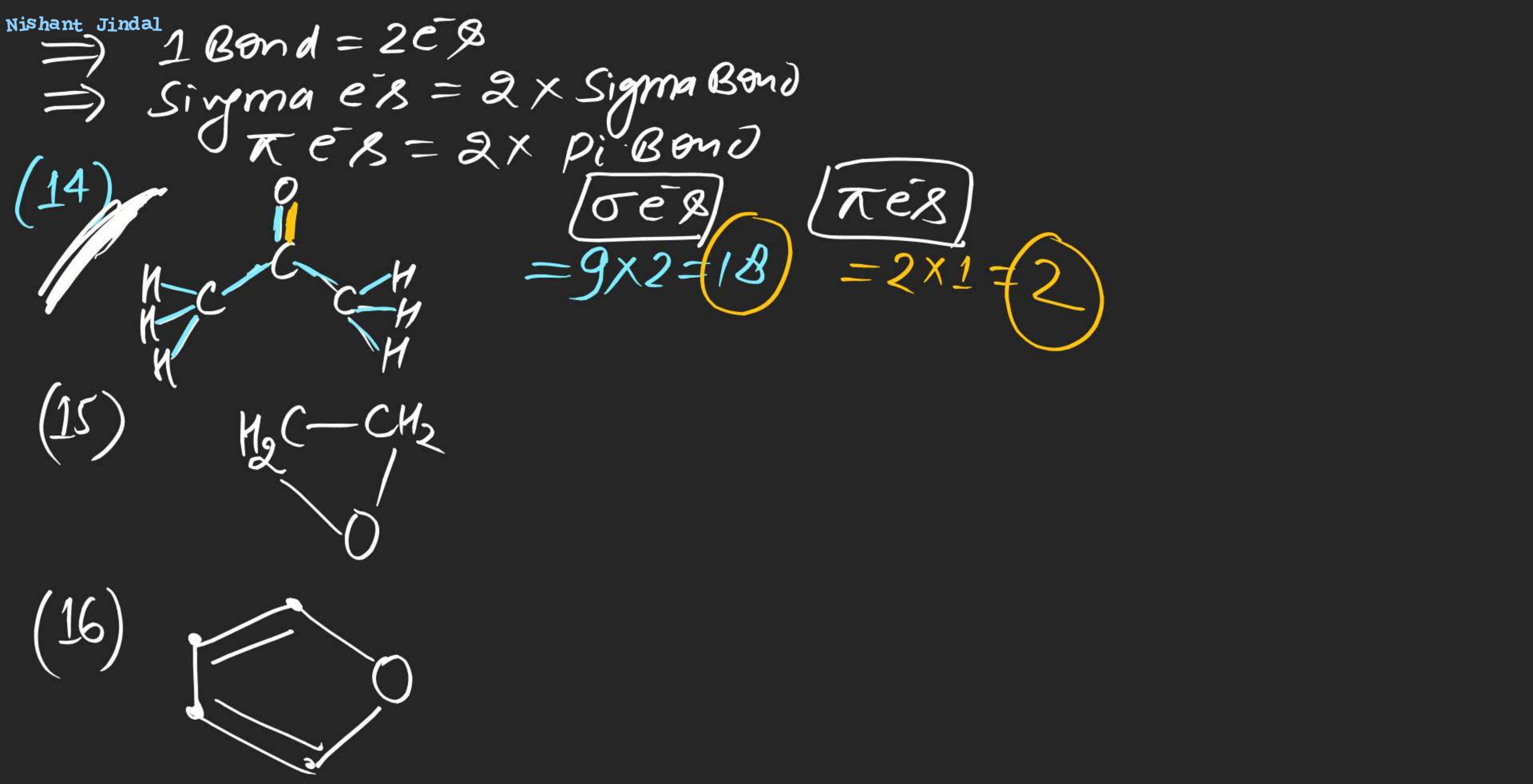


$$\begin{array}{l}
 49 \\
 = 12 \times 4 + 1 \\
 = 48 + 1 \\
 = 49
 \end{array}
 \left| \begin{array}{l}
 13 \\
 = 3 \times 4 + 1 \\
 = 12 + 1 \\
 = 13
 \end{array} \right.$$

σ
g
 π
g

C_6H_6 Benzene
 C_6H_5 Phenyl





~~Books~~
~~Nomenclature~~
~~GOC~~ Book for Organic chemistry

- * No Theory Book is Reqd.
- * Practice Book

~~Blue Book~~

~~BB~~

(i) Theory Copy
(ii) DPP
(iii) Sheet
(iv) Cengage Publication (3e)
(Problems & Solution) in
organic chemistry

(Suresh K. Mishra)

IIT Adravne

- AIR 1. →
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Basic Organic Chemistry

