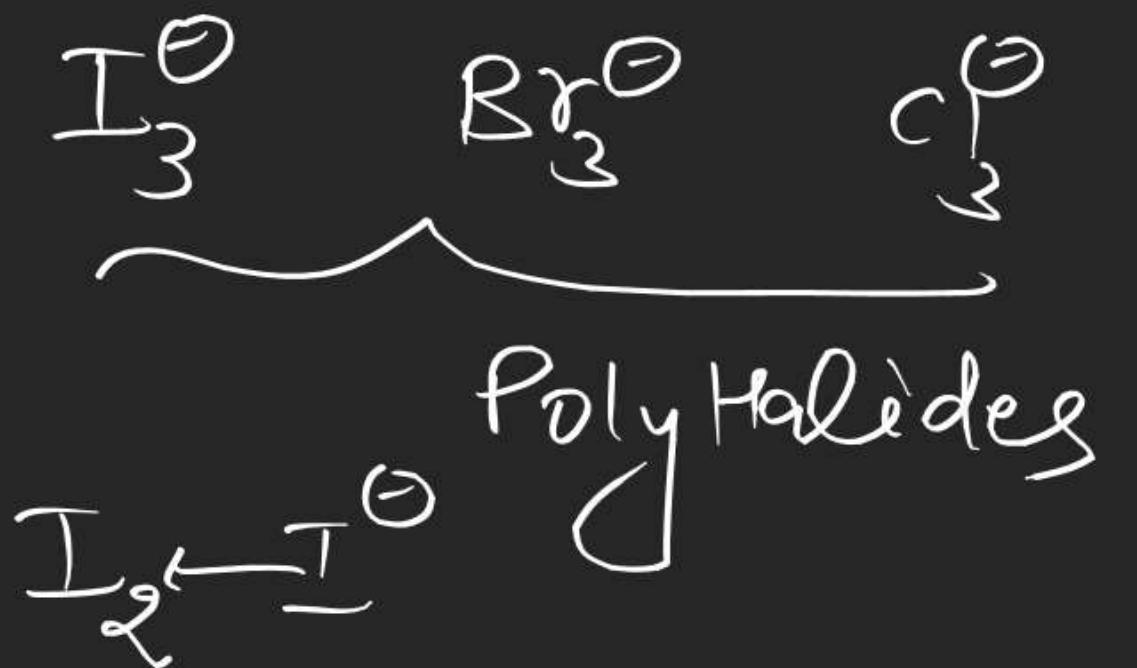


CHEMICAL BONDING

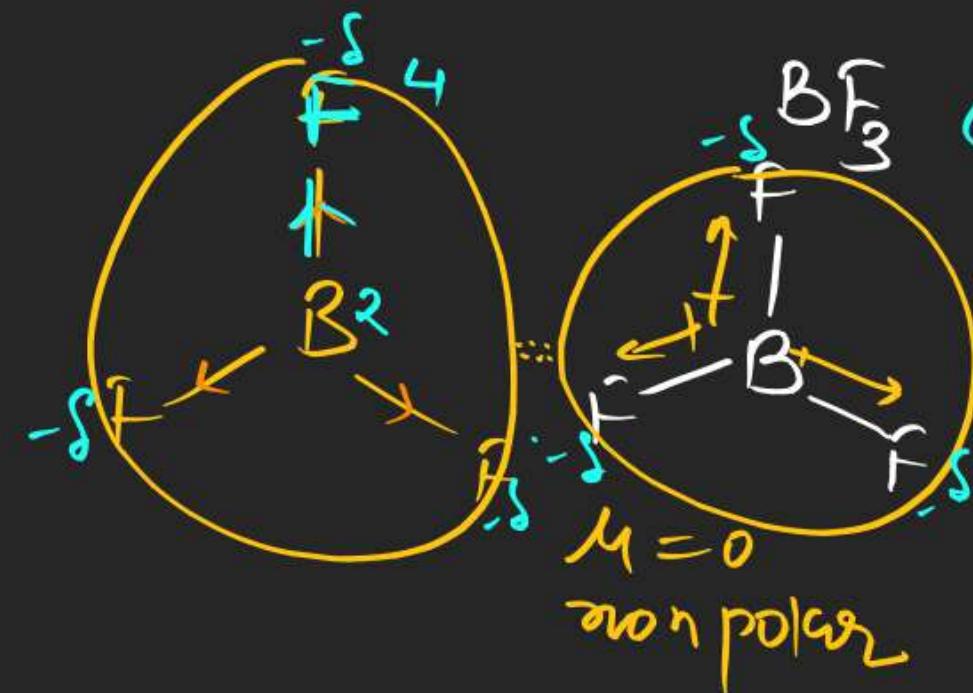


Ion - Induced dipole

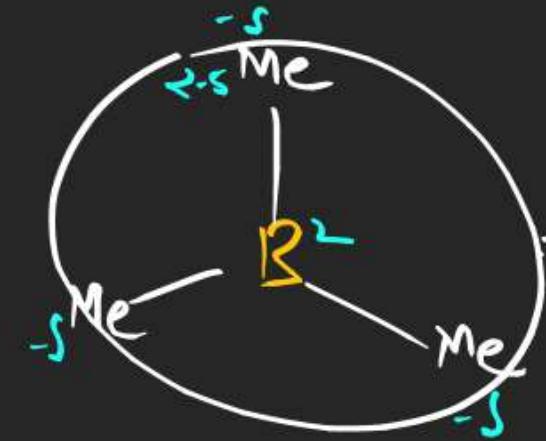
B.P \propto mol. wt.

B.P \propto att. b/w molecule

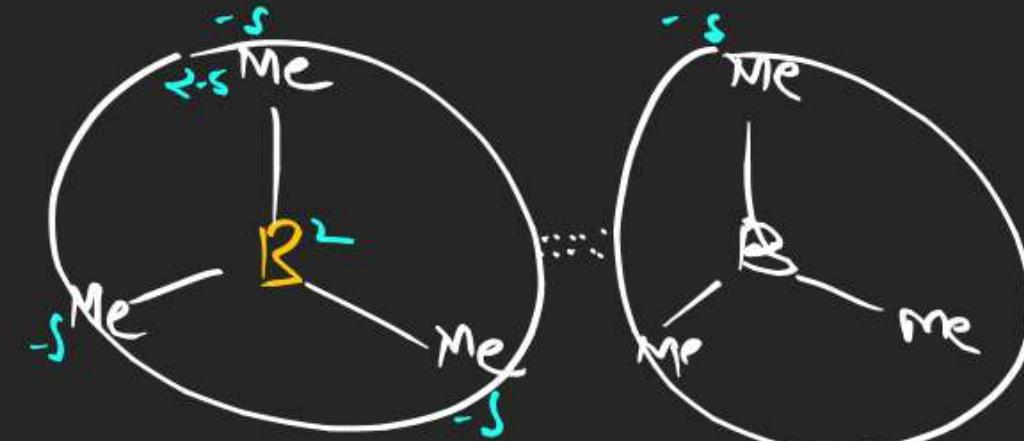
Order of B.P



L · D · F



$\mu = 0$
non polar
L · D · F

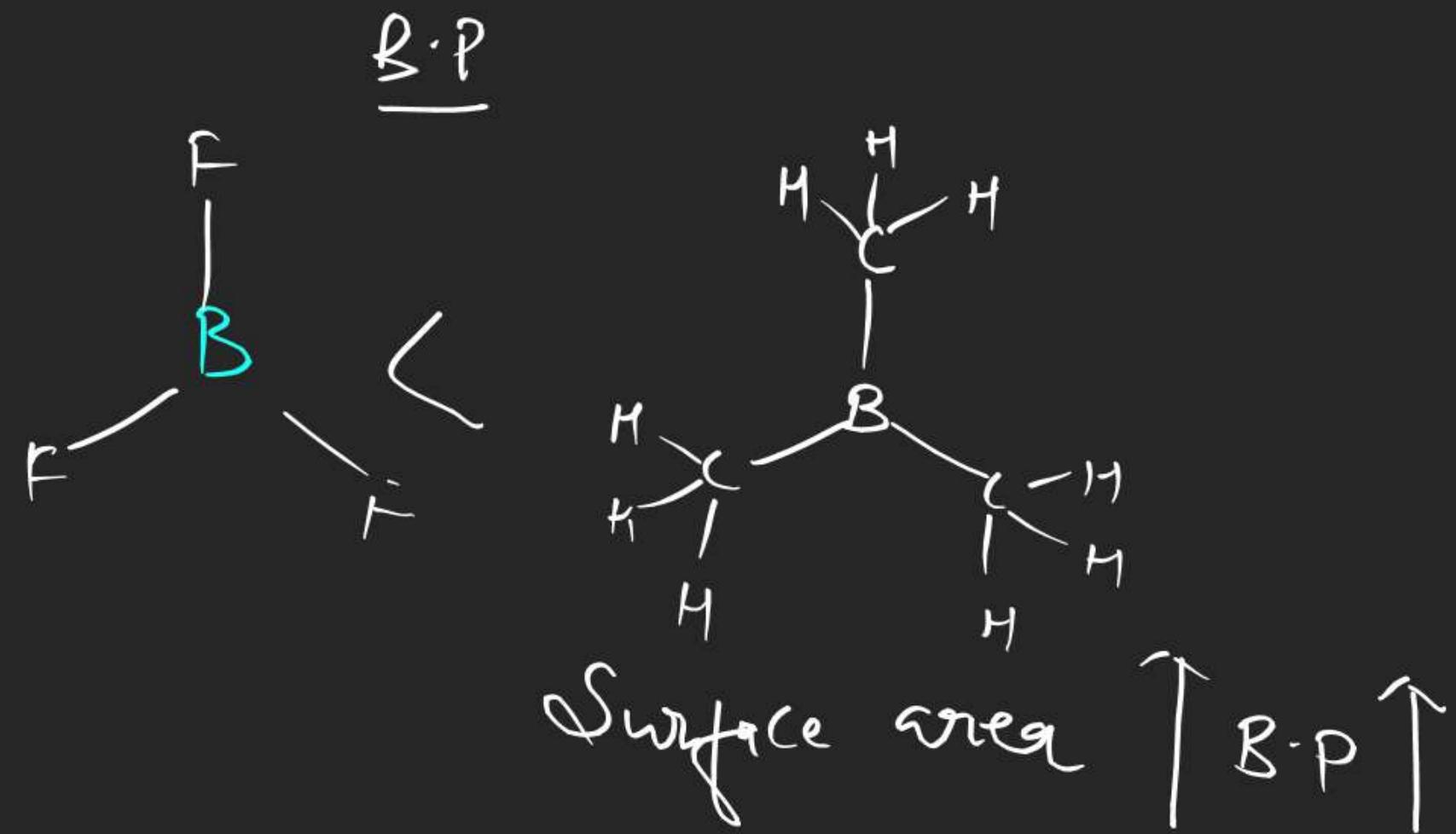


$\mu = 0$
non polar

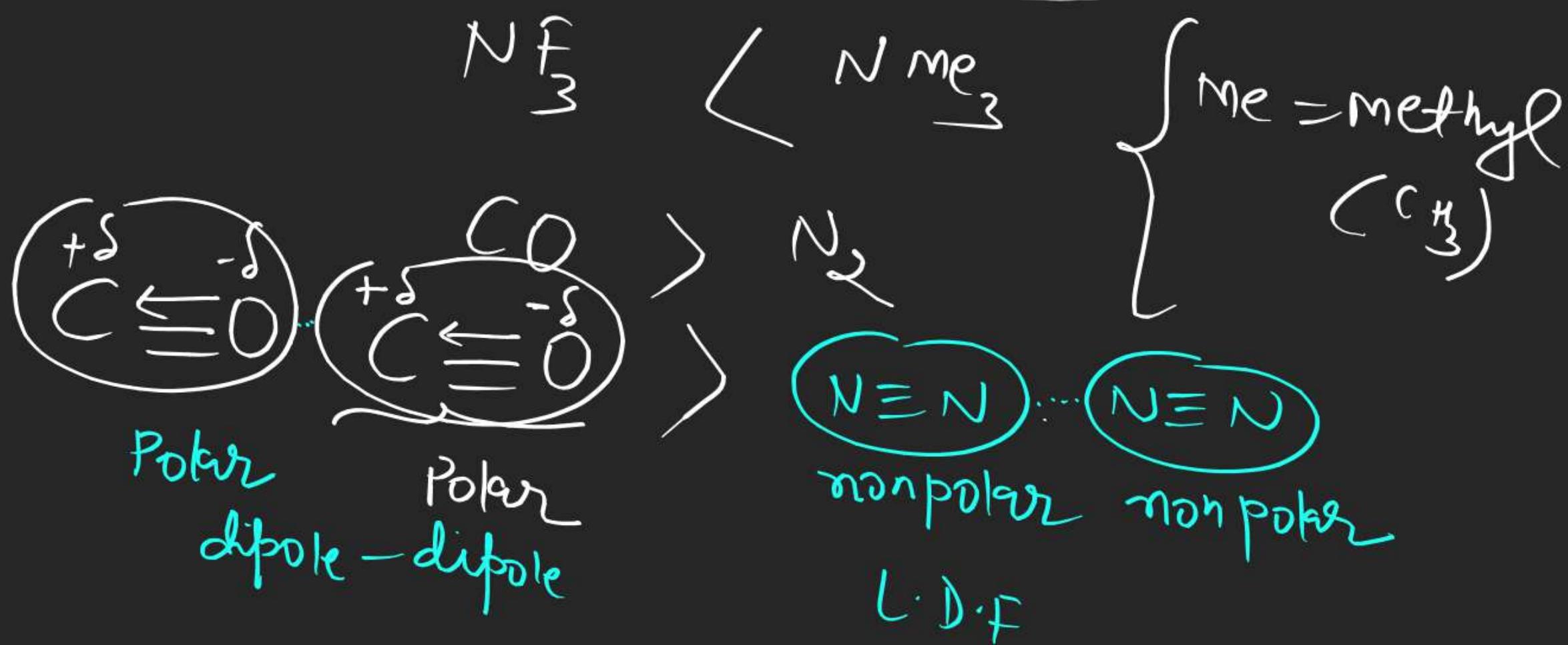
$\epsilon \cdot N$ values

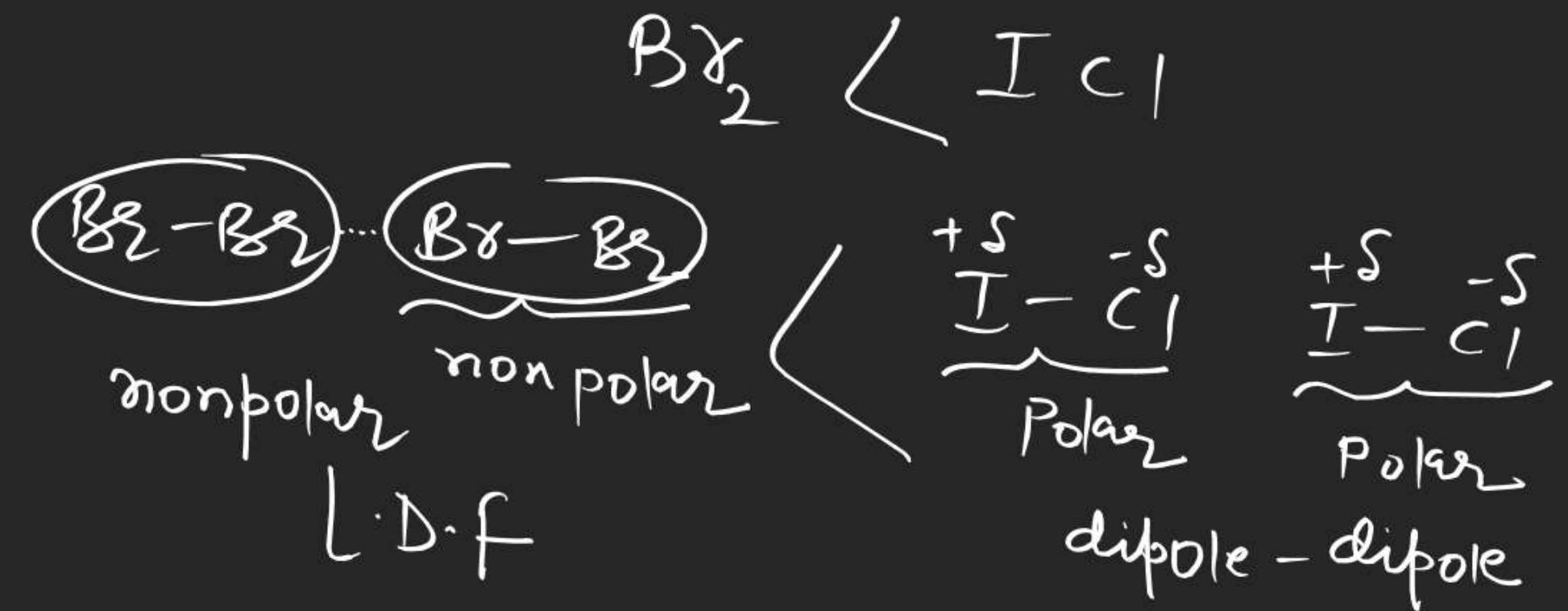
2.1 H	Li	Be	B	C	N	O	F
	1.5	1.5	2	2.5	3	3.5	4

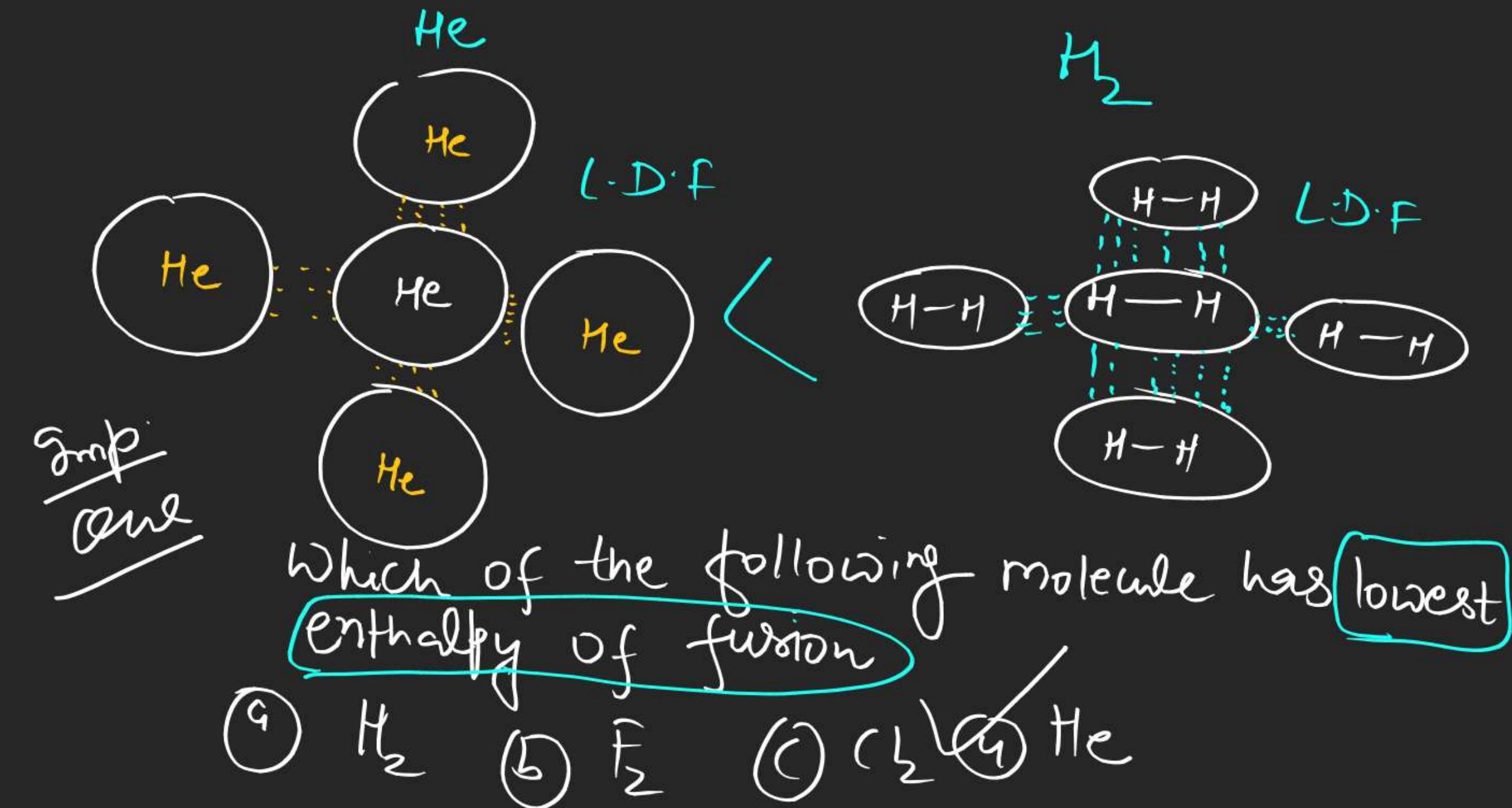
C

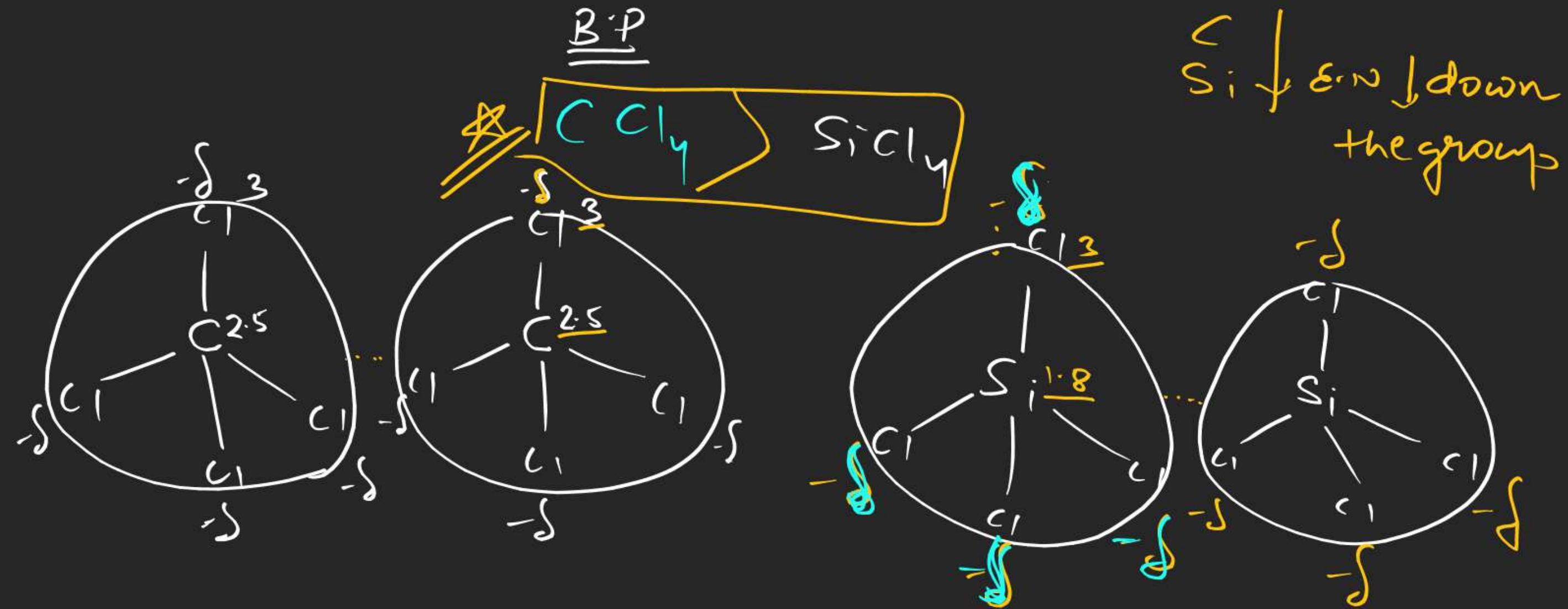


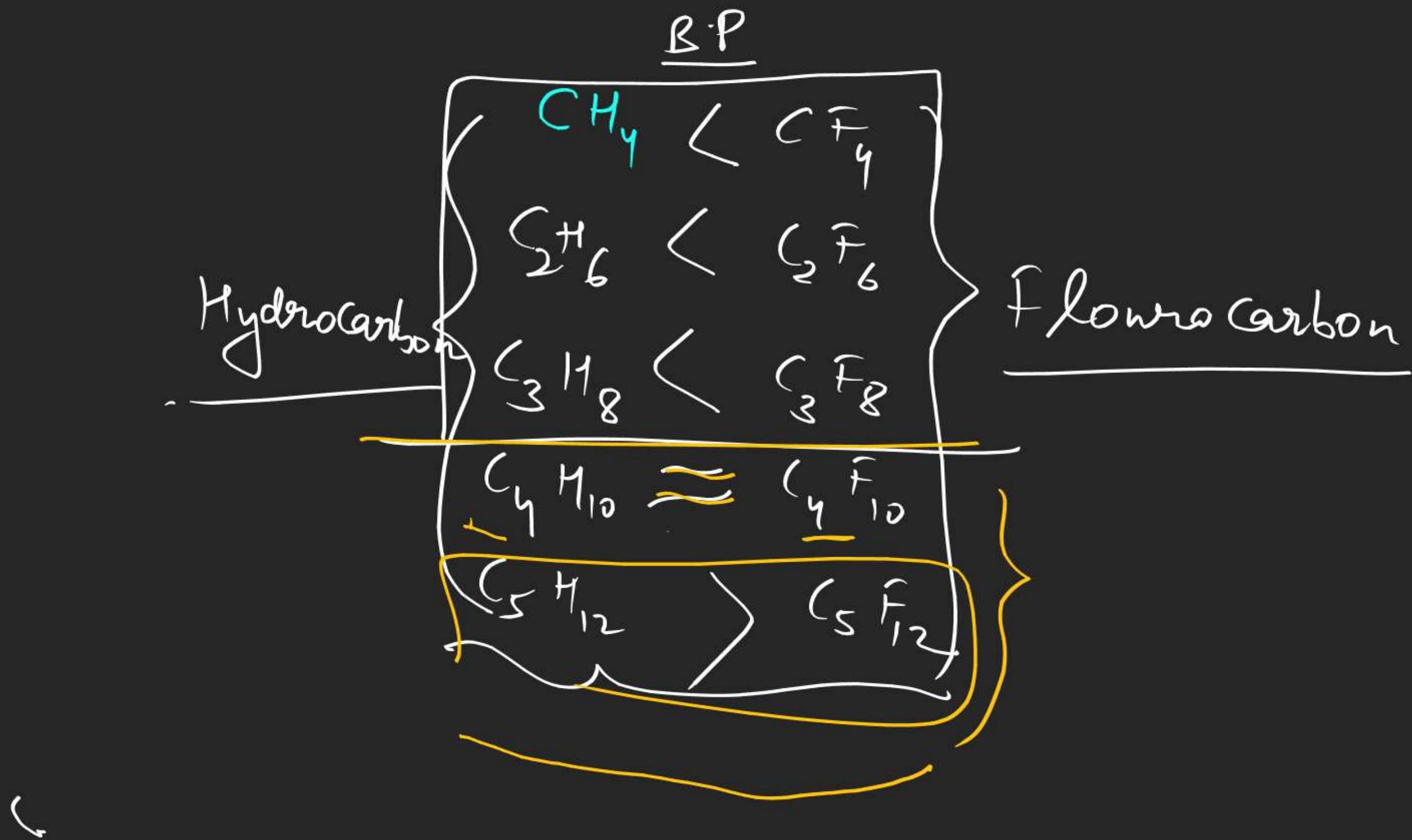
order of B.P



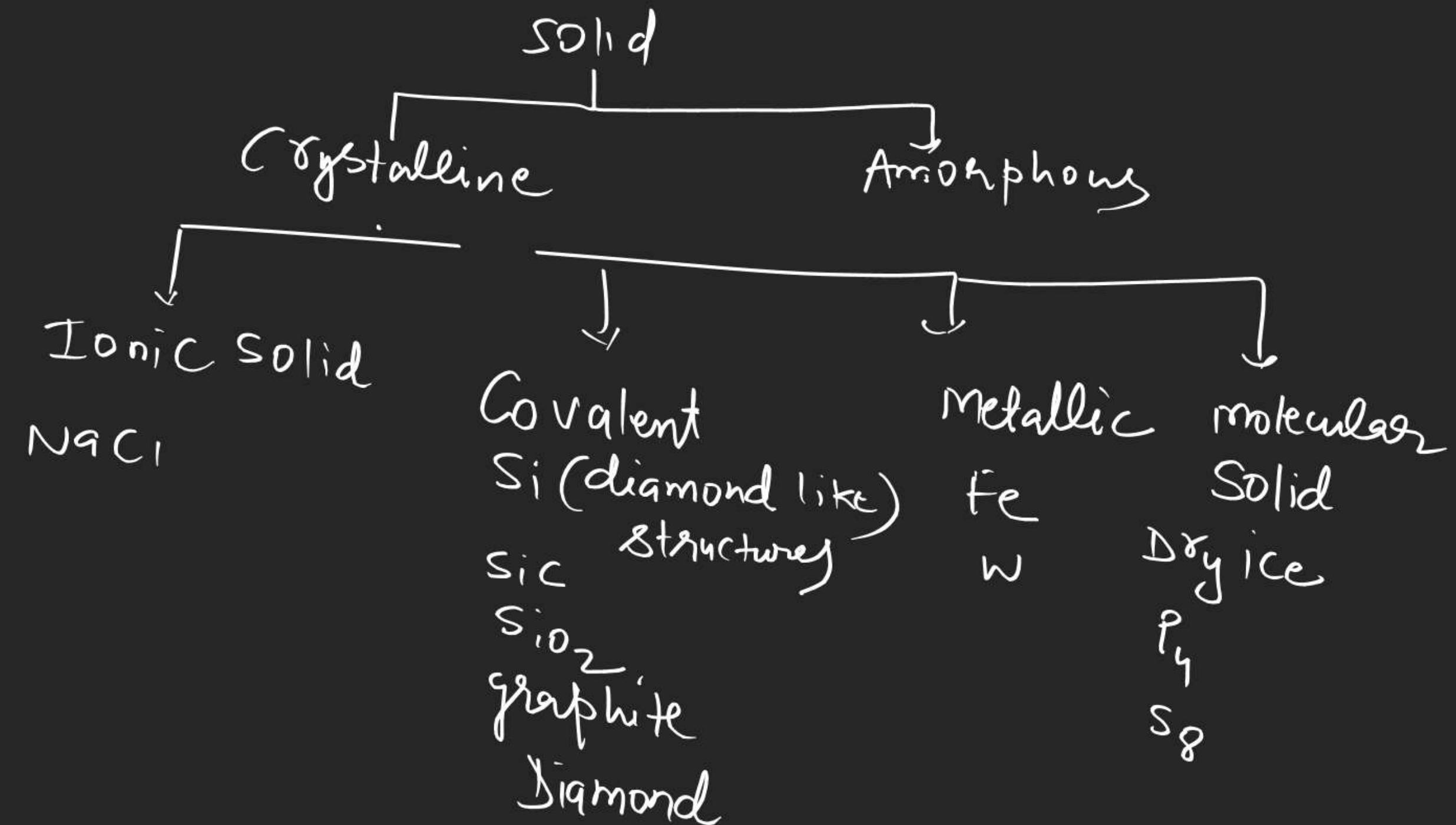


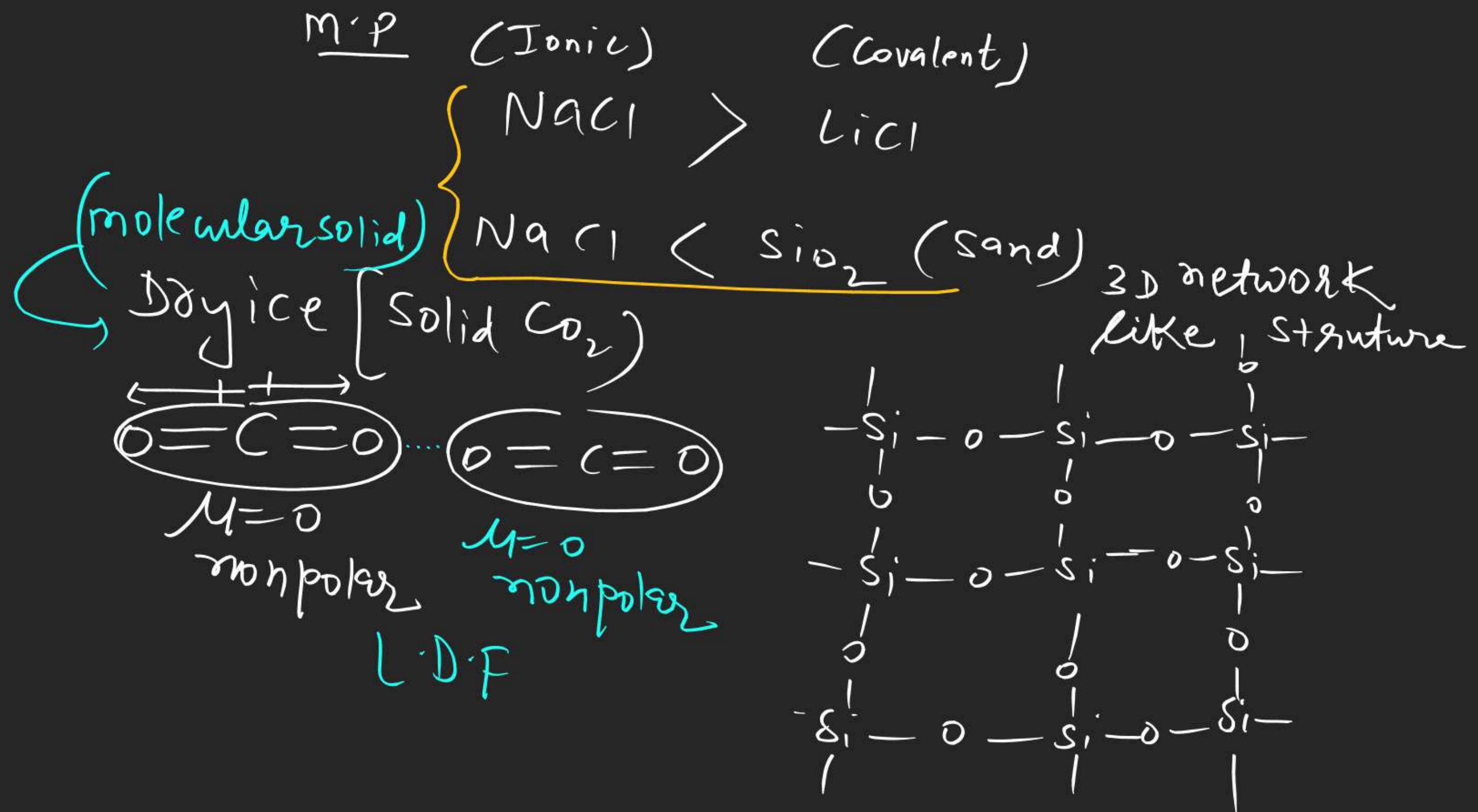






Note → fluorocarbon very often show
lower or equal B.P than
Corresponding hydrocarbon having
much lower molecular weight

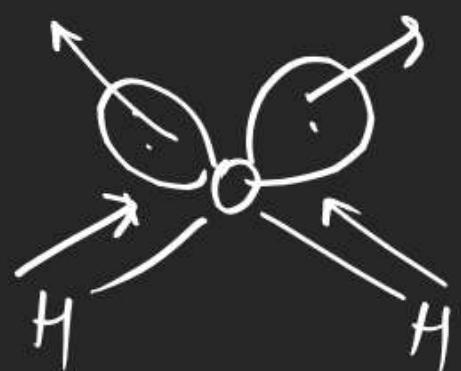






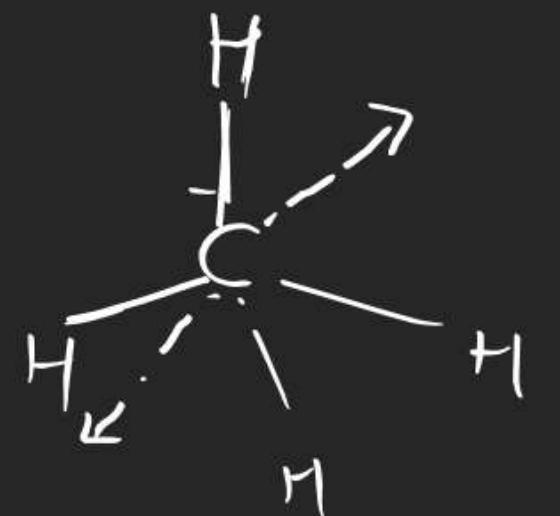
Dry ice < LiCl
and order of M.P.
 $\text{NaCl} \parallel \text{SiO}_2$ Dry ice LiCl
 $\text{SiO}_2 > \text{NaCl} > \text{LiCl} > \text{dry ice}$

$$\frac{\text{one}}{\text{B.P}} \frac{H_2 < D_2 < T_2}{D_2^o > H_2^o}$$

B.P

$$\mu \neq 0$$

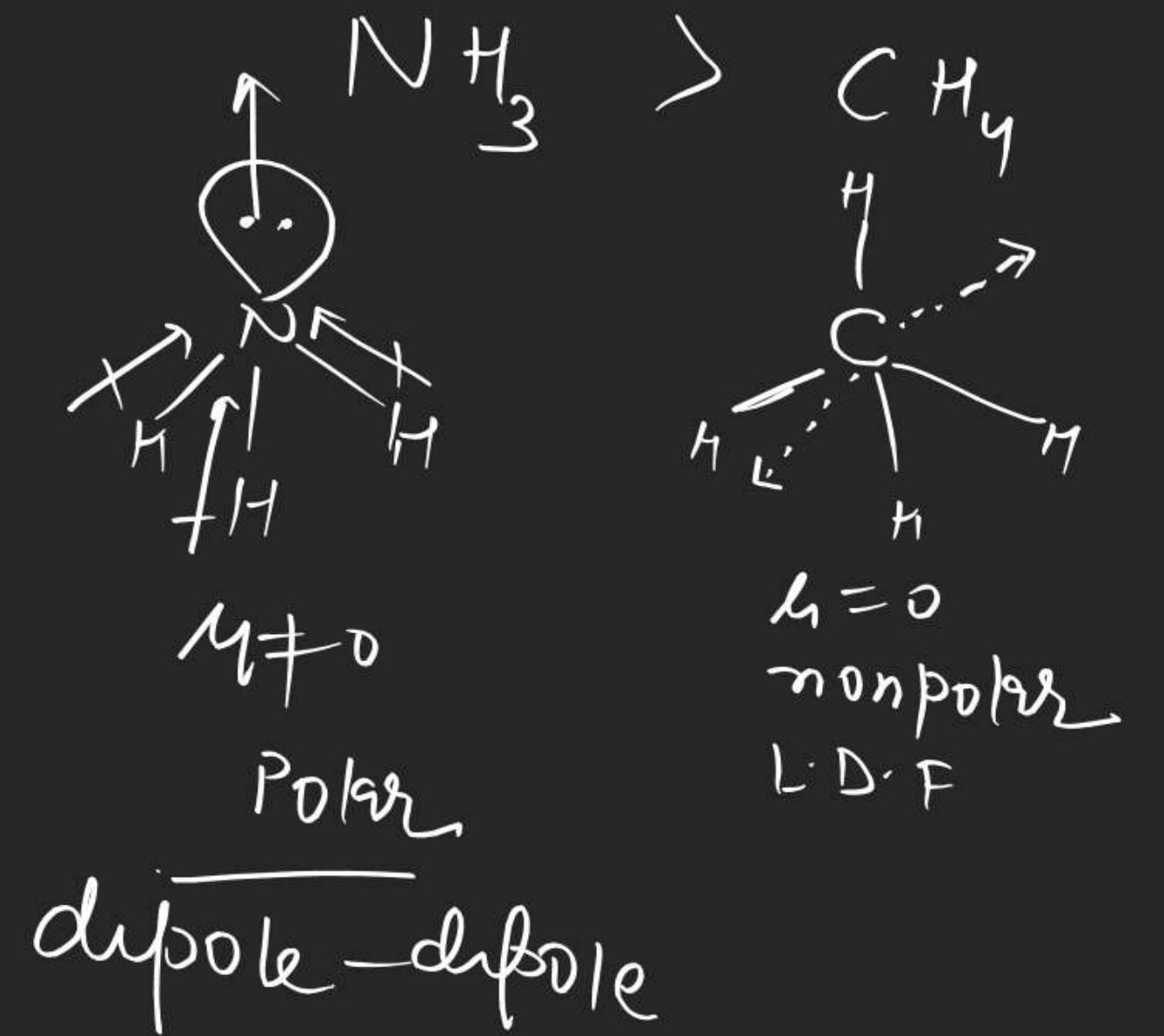
Polar
dipole-dipole



non polar

L.D.F

$$\underline{\mu = 0}$$



~~and~~ L.D.F depends upon He

- ① Molecular mass Ne
- ② Polarisable e- Ar
- (3) Size of molecule Kr
- ~~④~~ all of these - S, F, + S Xe



B · P

$$\frac{X_e}{L \cdot D \cdot F} > \frac{CH_y}{L \cdot D \cdot F}$$

JEE Mains PYQS

Ques Which of the following
relation is correct regarding
this

$$\text{interaction energy} \propto \frac{1}{r^3}$$

- (1) Ion-dipole (2) Ion-induced dipole
 (3) hydrogen bonding (4) none

Note \rightarrow Hydrogen bonding is special case of
dipole-dipole interaction.

one which of the following interaction is responsible for
 $Rn(PhOH)_2$ $Kr(PhOH)_2$

- Ⓐ dipole - dipole ✓ dipole - induced dipole
Ⓑ Ion-dipole Ⓑ nope



