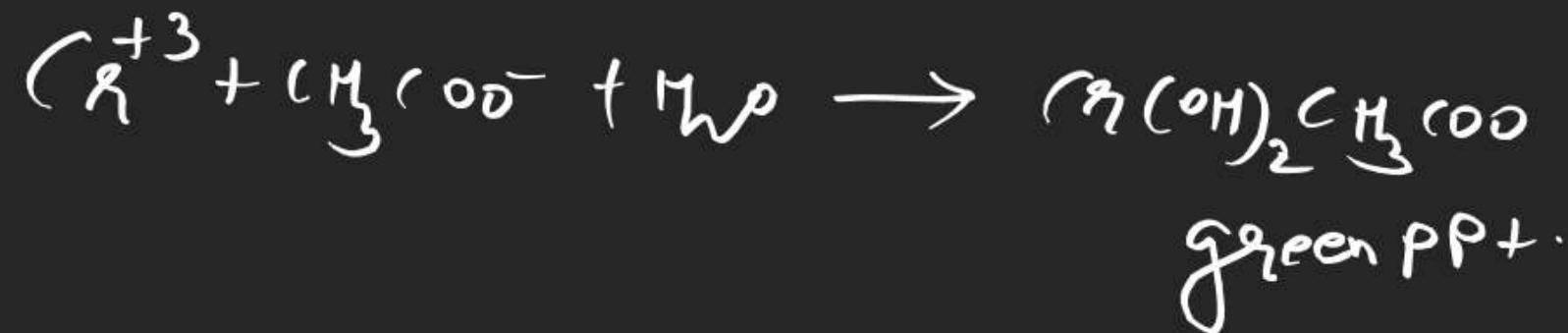
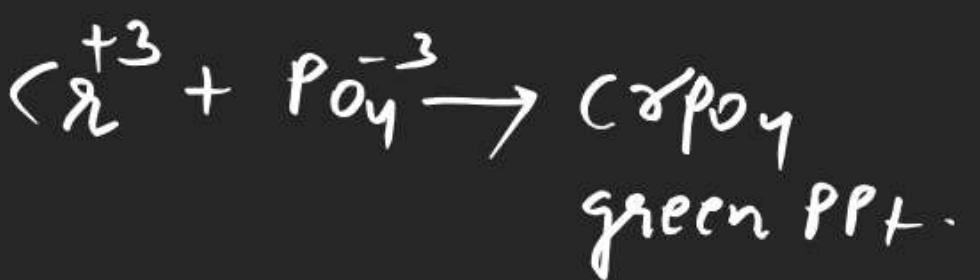
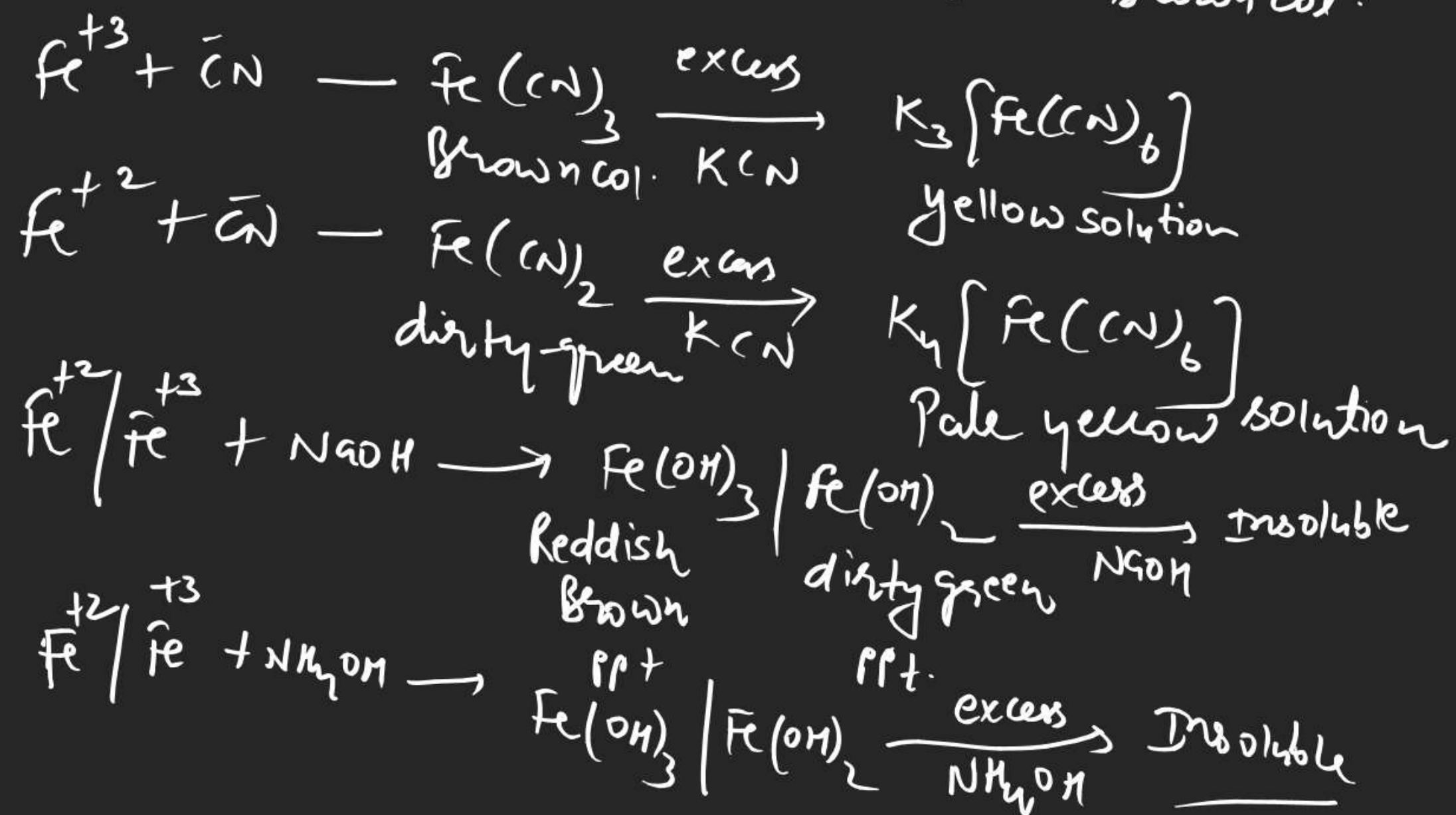
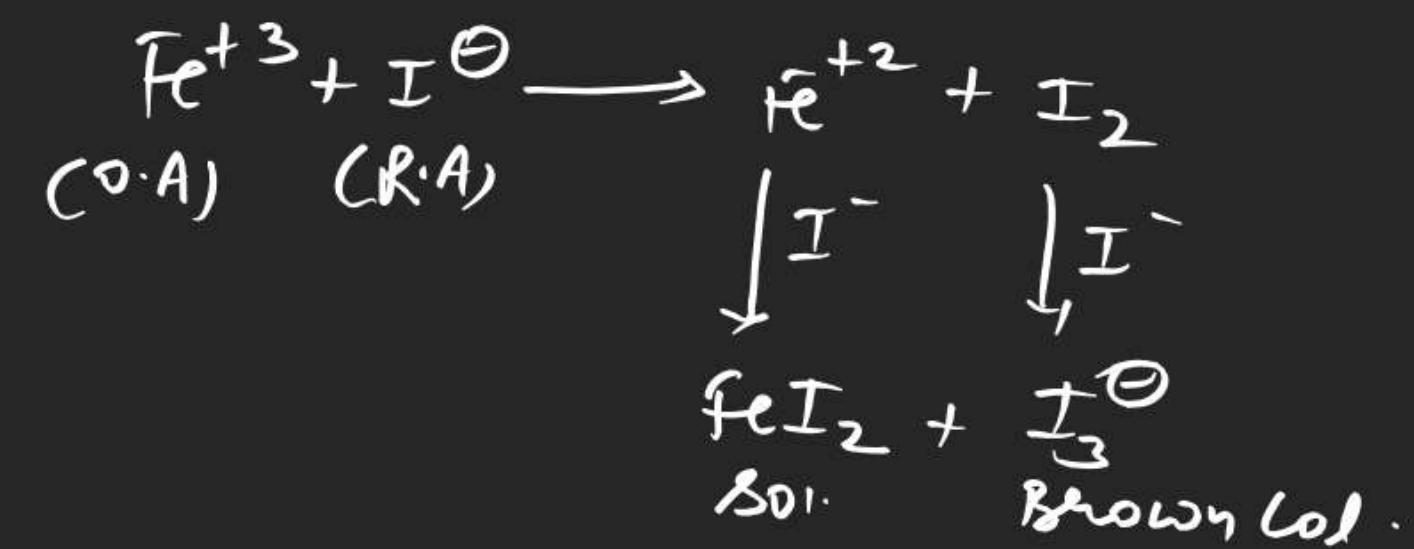
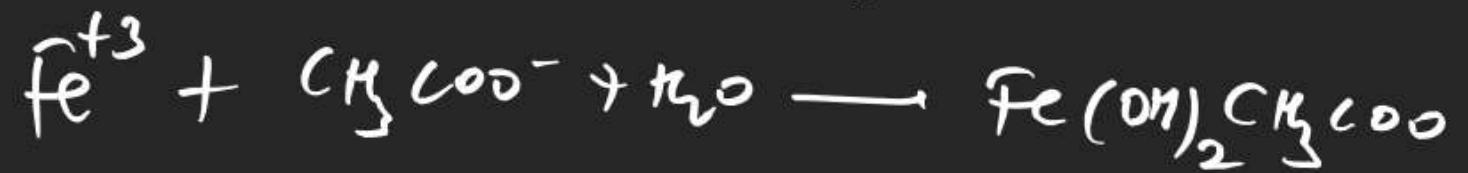
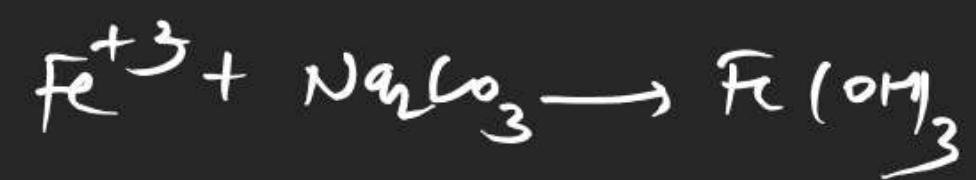


Soluble violet col. complex







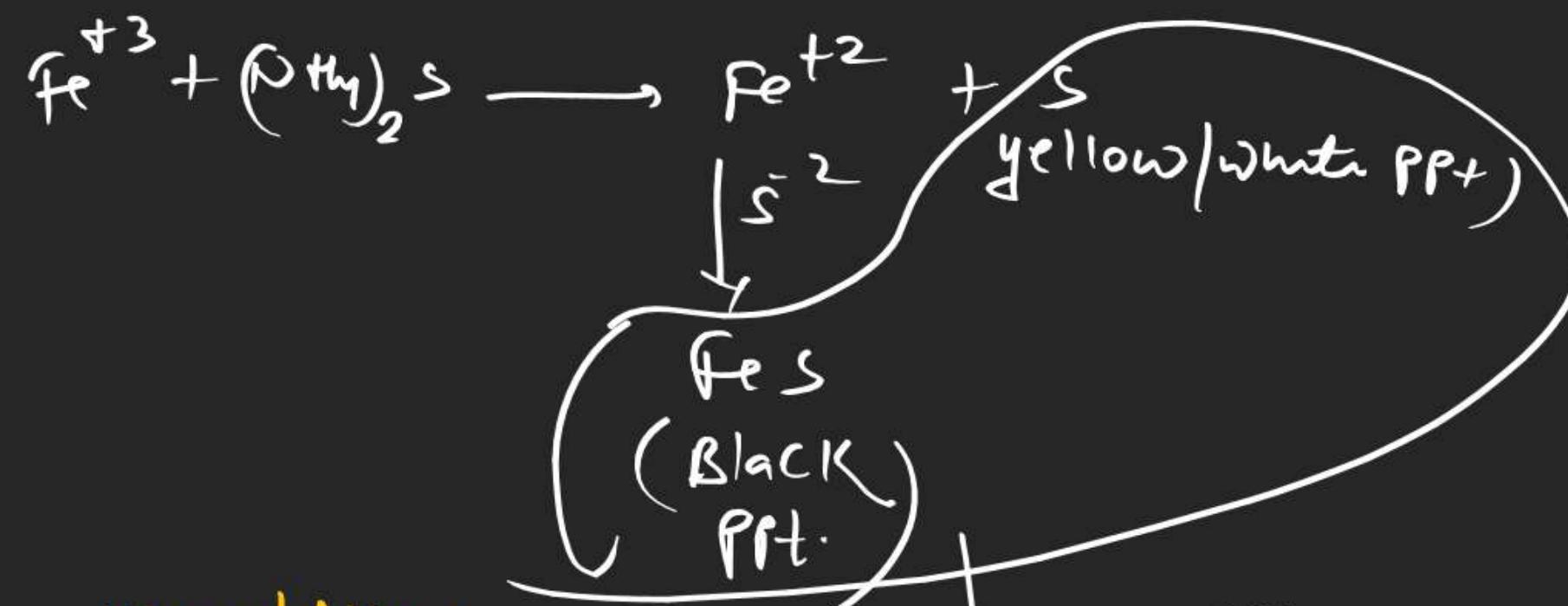
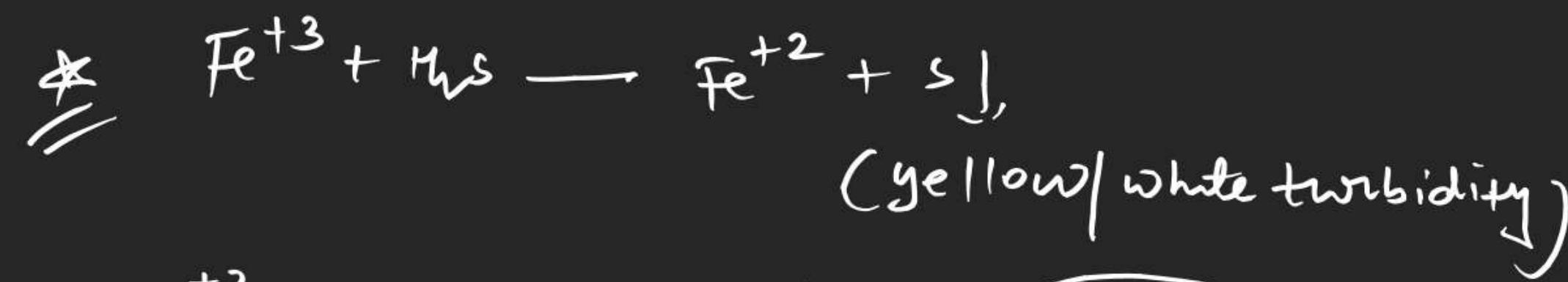
Brownish Red ppt.

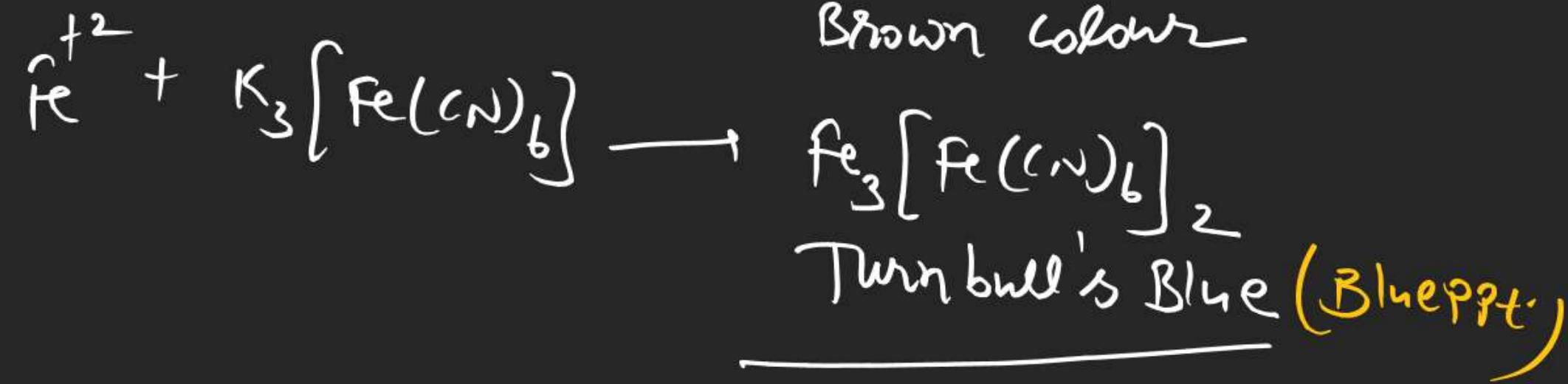
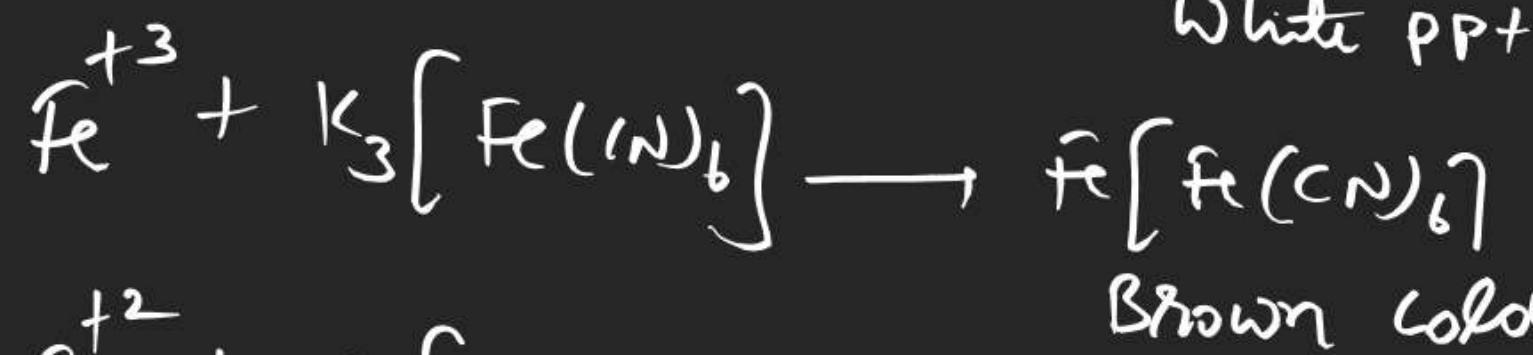
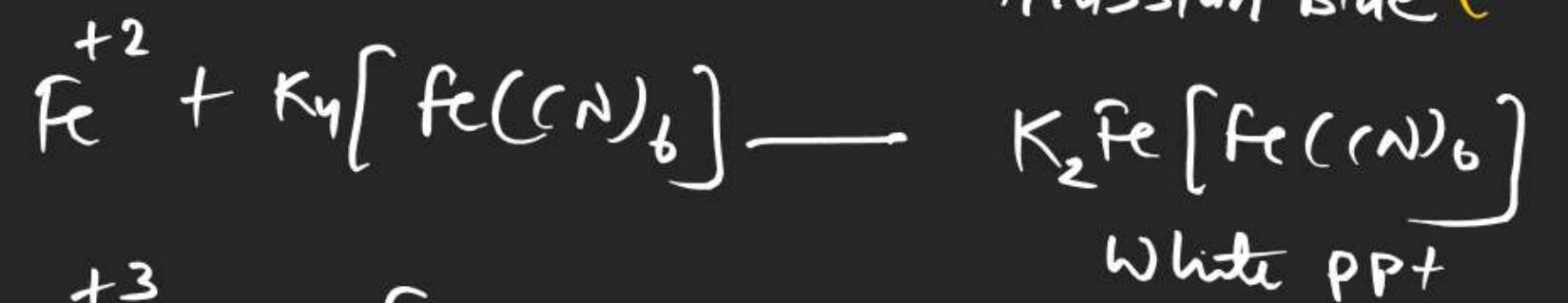
Note  $\Rightarrow \text{S}^{-2} \rightarrow \text{Cr}^{+3} | \text{Al}^{+3} | \text{Mg}^{+2}$  Sulphides do not exist

in their ag-solution because they hydrolysed.

$\Rightarrow$  Carbonates salts of  $\text{Cr}^{+3} | \text{Al}^{+3} | \text{Fe}^{+3}$  donot exist in their ag-solution because they hydrolysed.

$\Rightarrow$  Some basic acetate of  $\text{Fe}^{+3} | \text{Al}^{+3} | (\text{Cr}^{+3})$  are insoluble

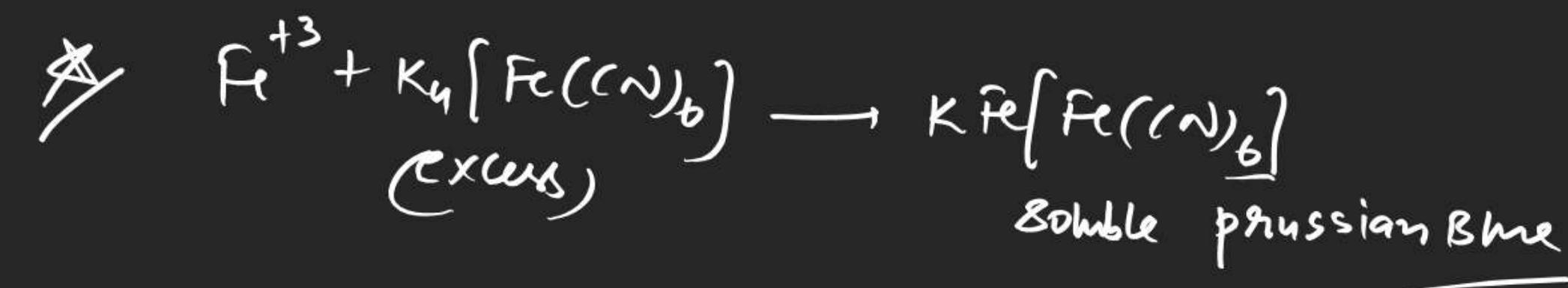




Prussian Blue and turns bull's blue both are identical  
and they can't be distinguished with moss bauer  
spectroscopy hence they representants  
as common formula

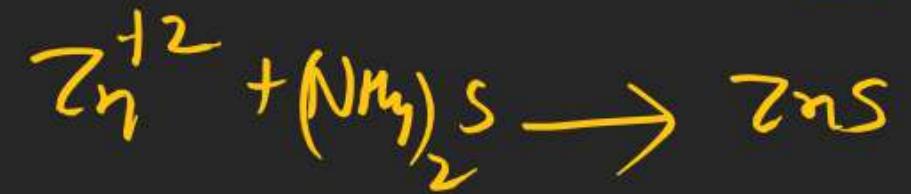
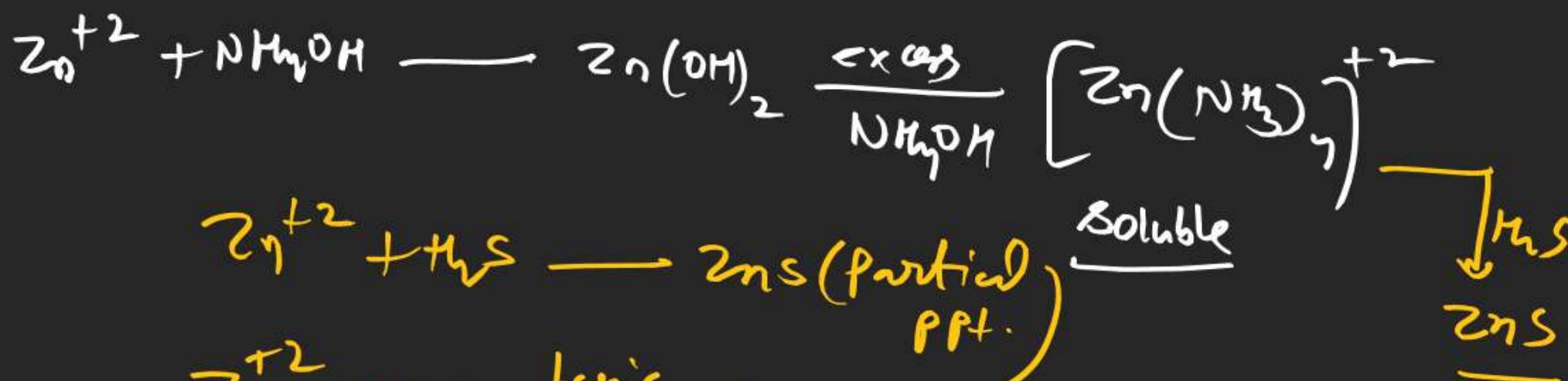
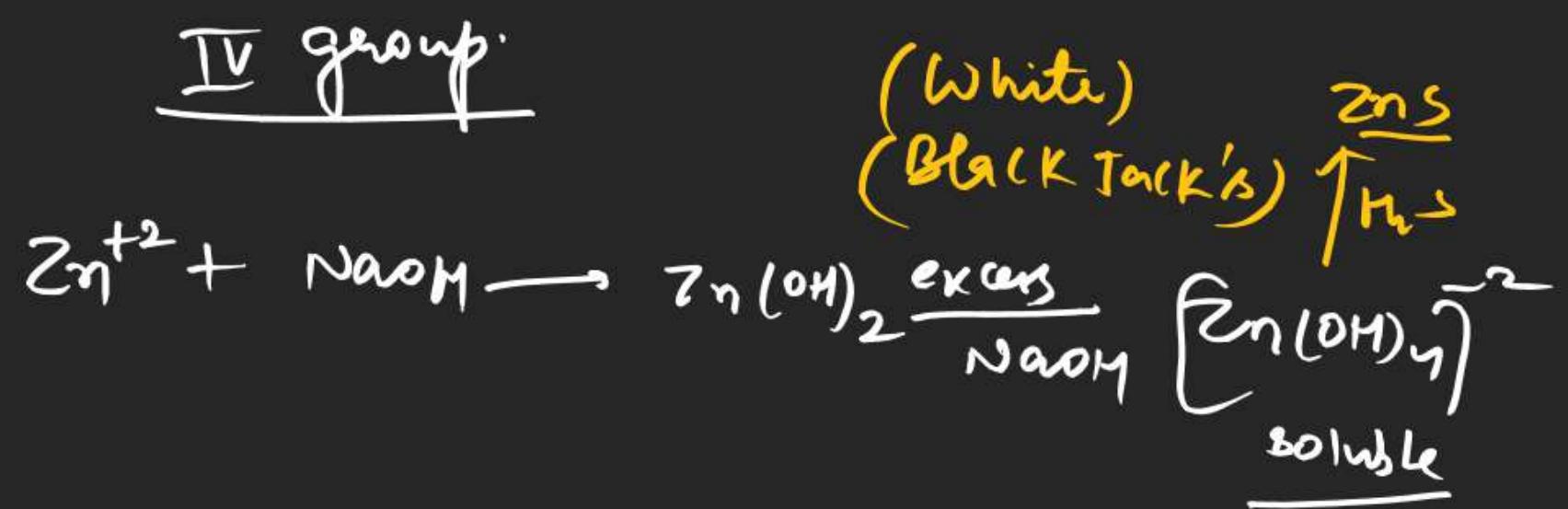


Charge transfer metal to metal



$\text{Fe}_4[\text{Fe}(\text{CN})_6]$  dissolved in oxalic acid  
so blue solution is formed

which was used as Blue ink

IV group



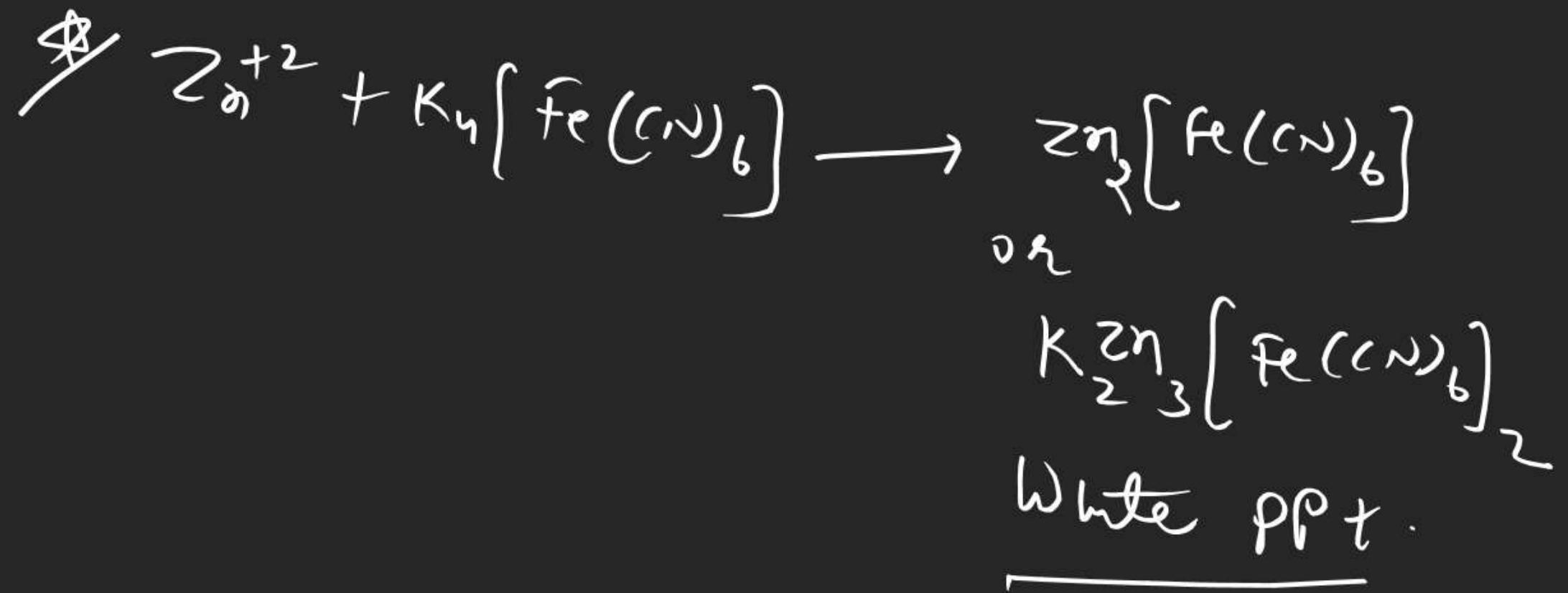
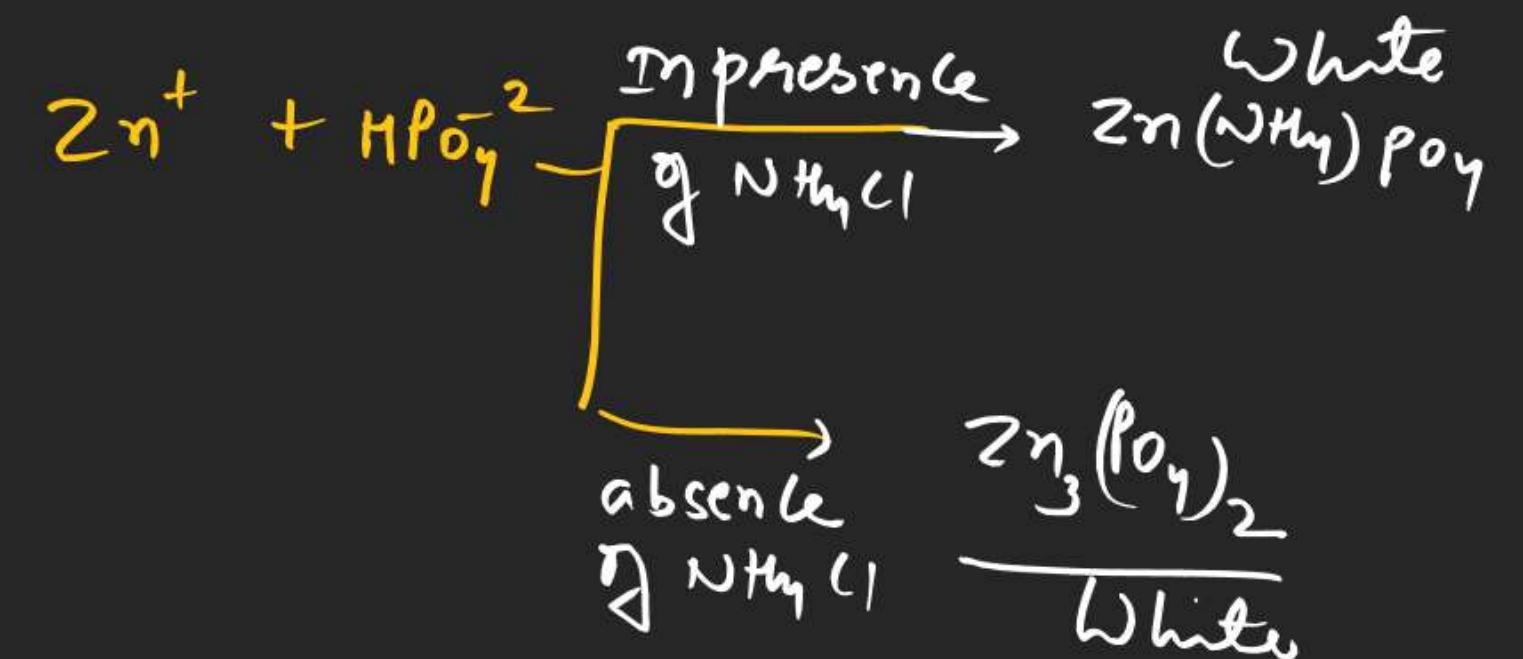
Basic Zinc  
Carbonate



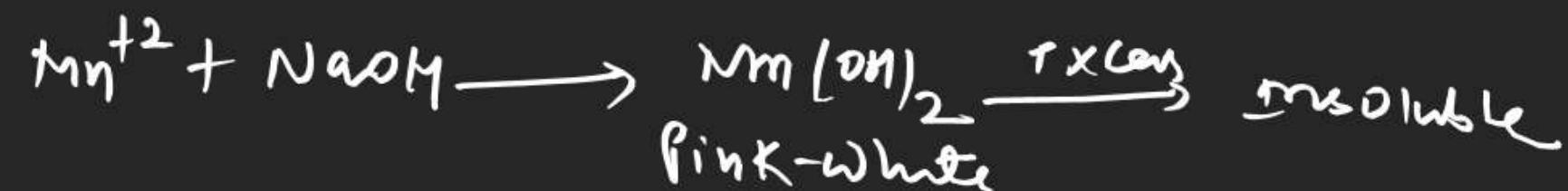
(philosopher's wool)  
(purest form of Zn)



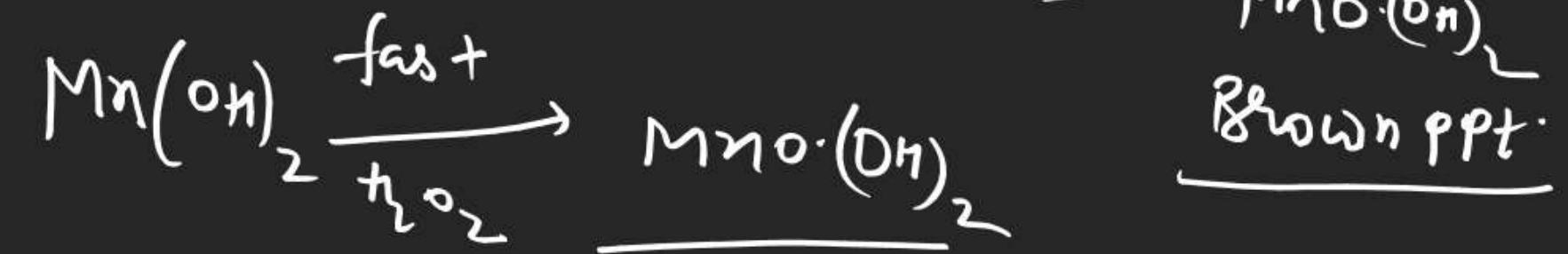
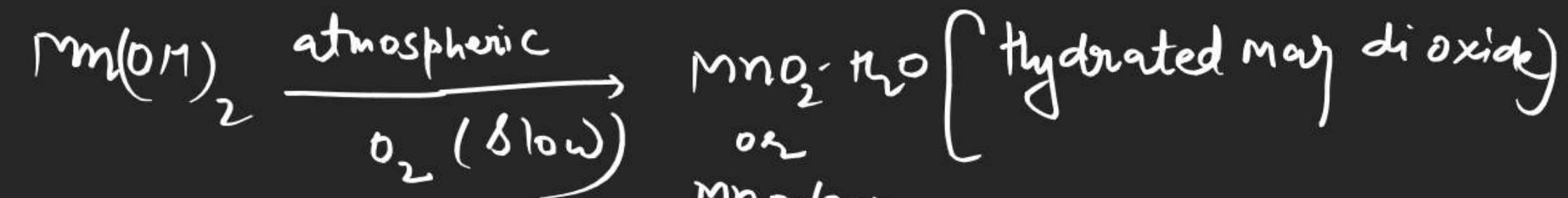
Note  $\Rightarrow$  ZnS soluble in dil HCl but insoluble in  $\text{CH}_3\text{COOH}$

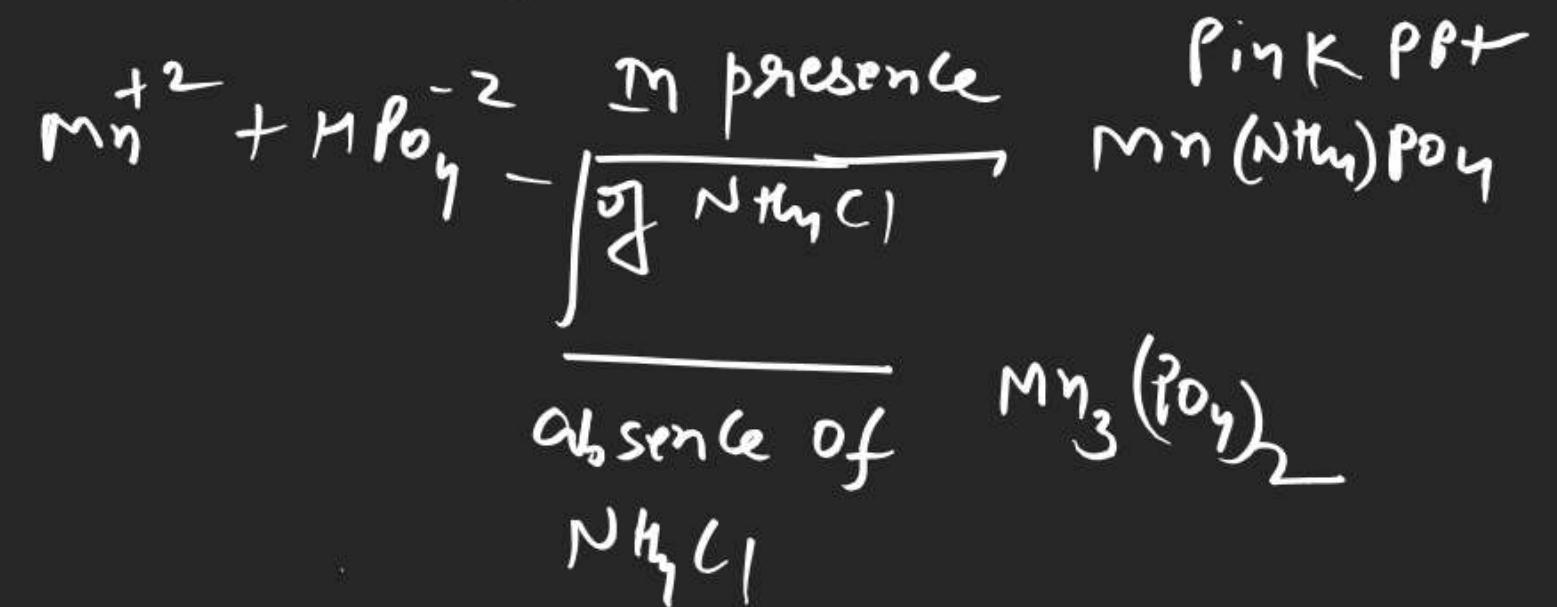
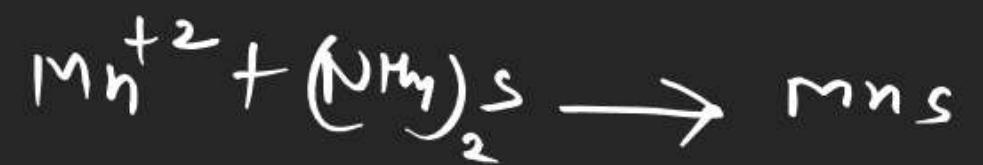


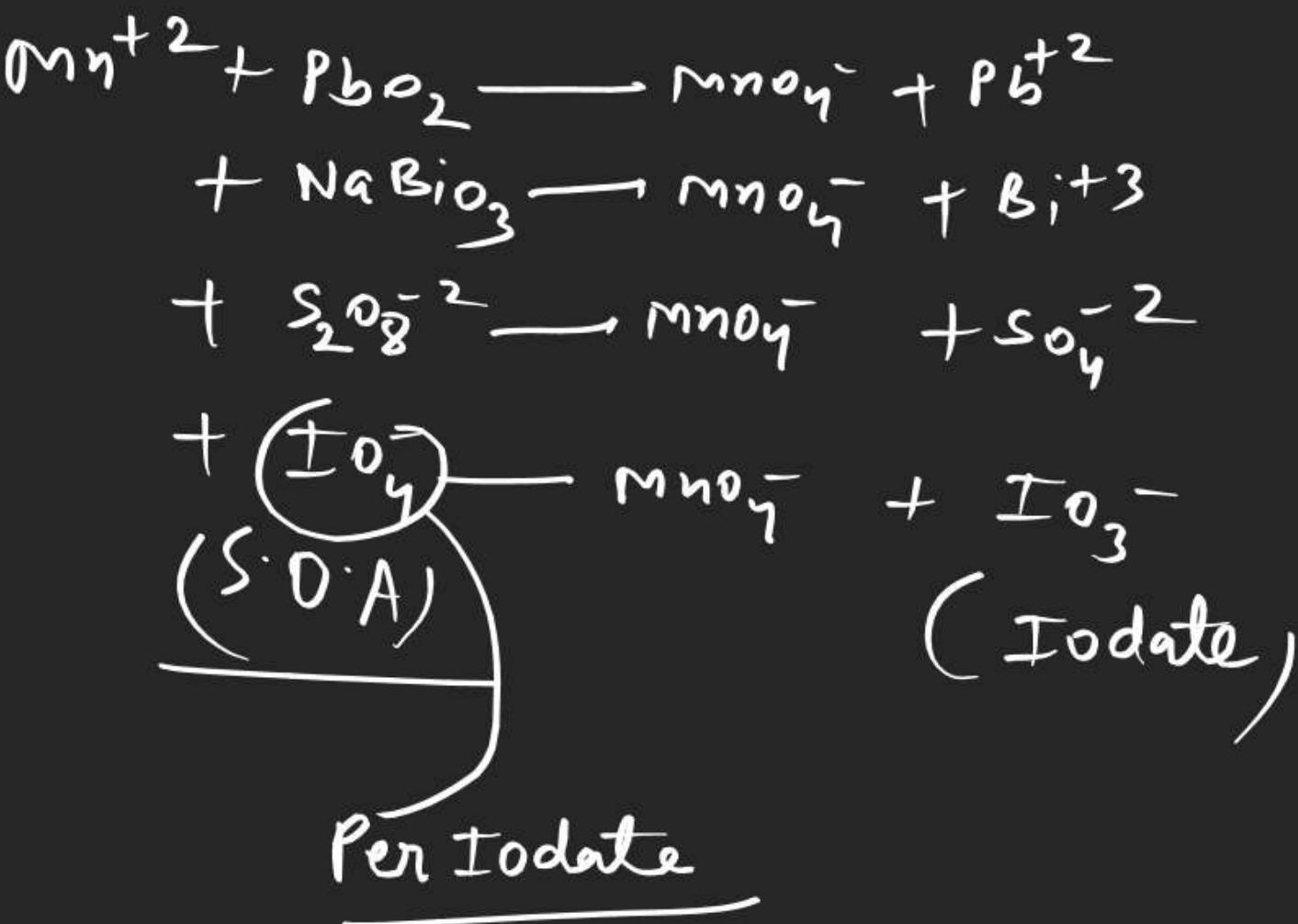
Mn<sup>+2</sup>  
MnS → Soluble in dilute HCl and C<sub>2</sub>H<sub>5</sub>COOM

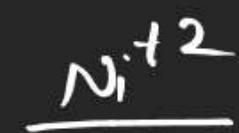


Mn(OH)<sub>2</sub> ⇒ Soluble in N<sub>4</sub>H<sub>5</sub>O salt.

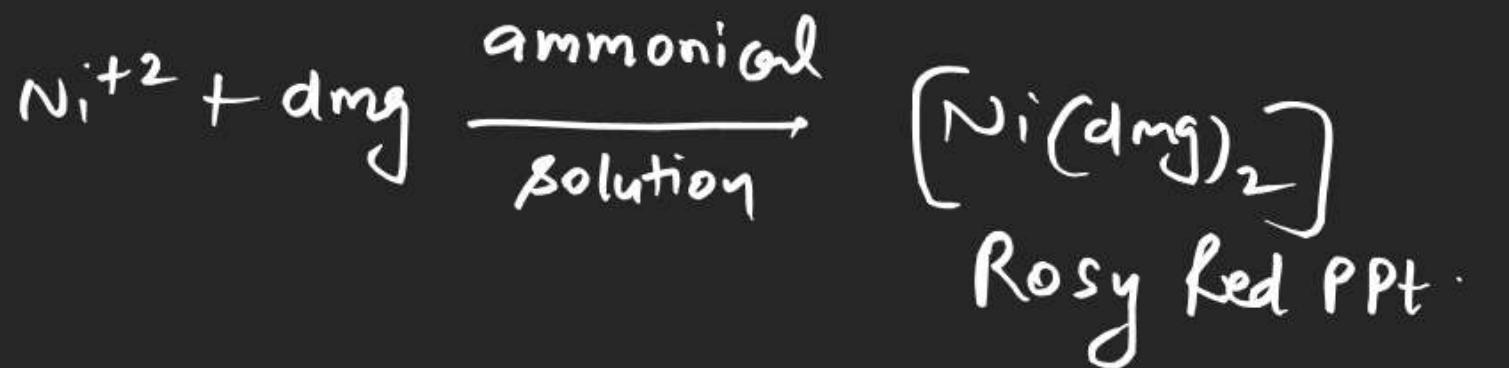


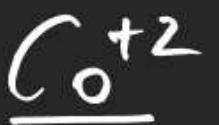






$NiS \Rightarrow$  Soluble in aqua regia.





Test with  $KNO_2^-$



<https://t.me/VJSIROfficial>

## V group

