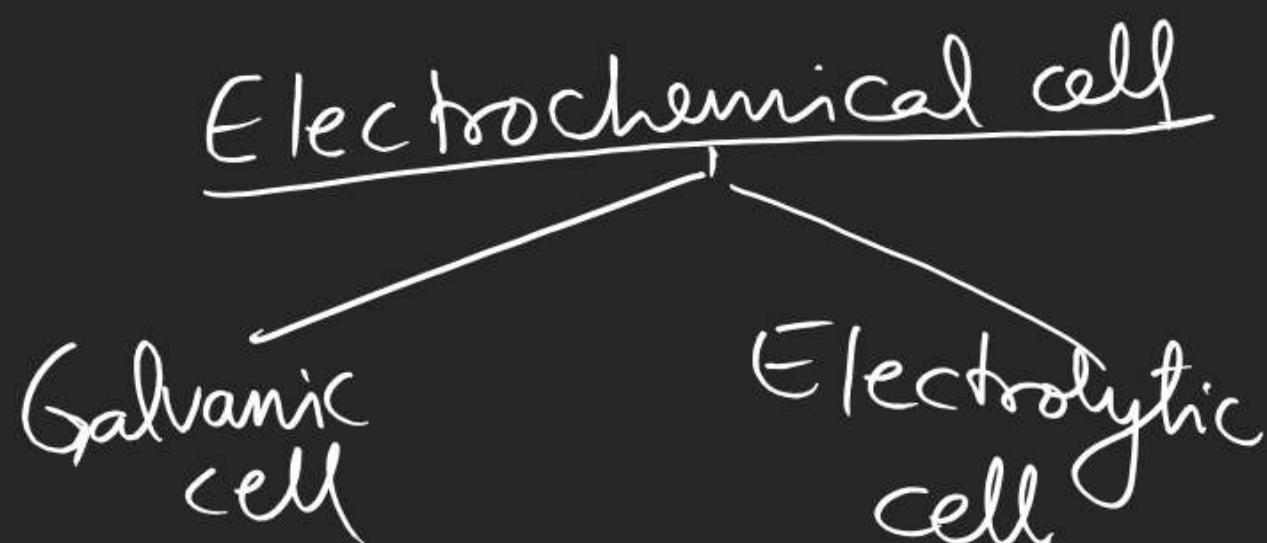


# Electrochemistry

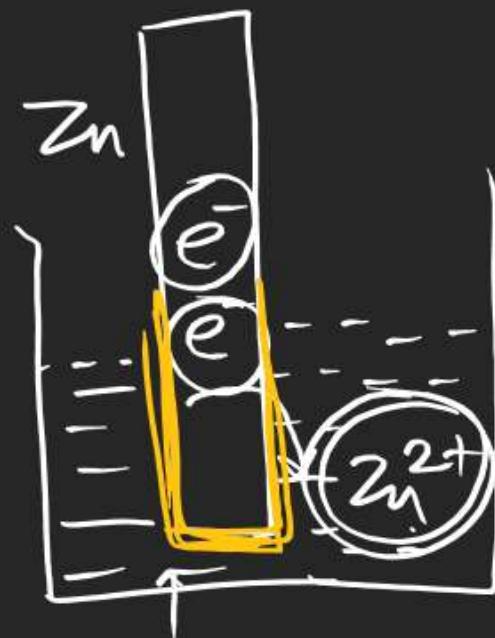
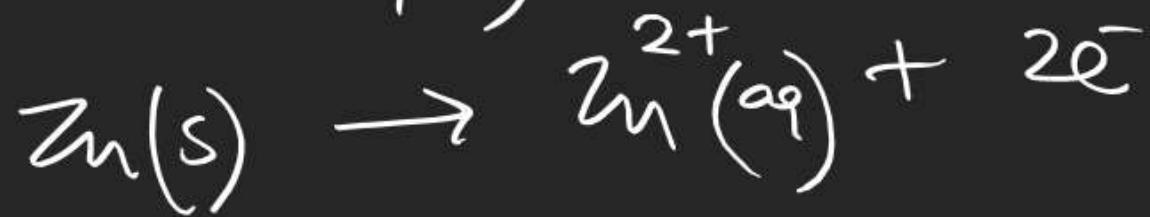
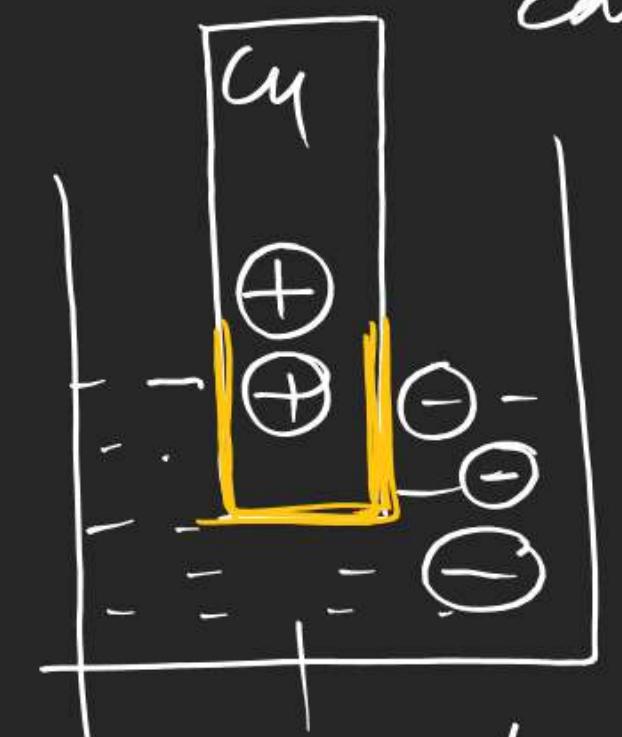
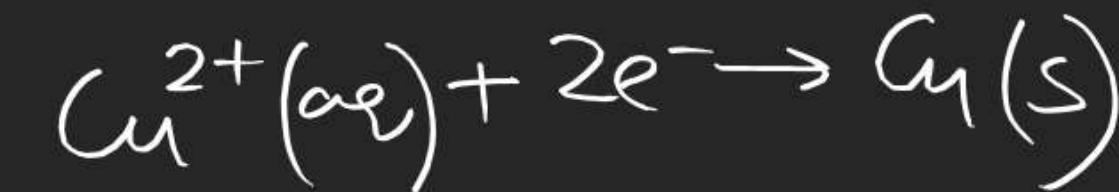
This chapter mainly deals with

- 1) Production of electric current by chemical reaction,  
and apparatus used is called "Galvanic cell"
- 2) Study of chemical change by electric current  
and apparatus used is called "Electrolytic cell"
- 3) Conductance & conductivity

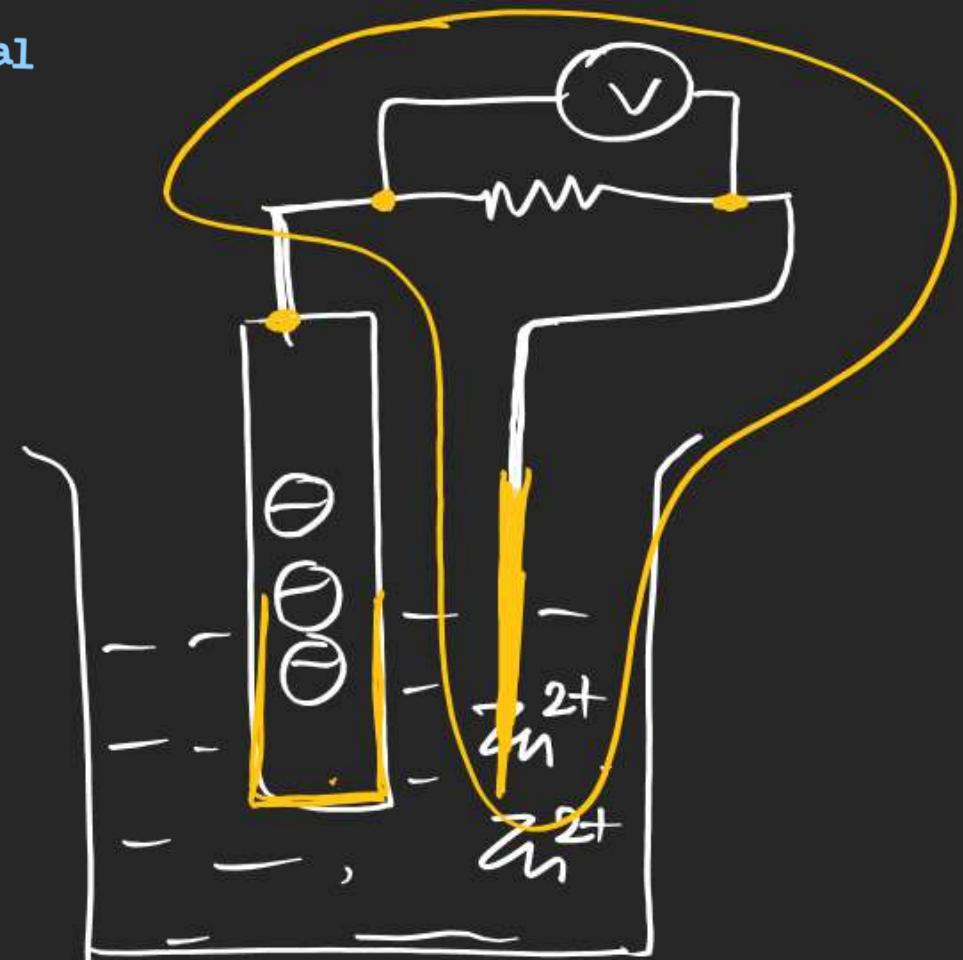


Galvanic celli) Electrode potential  $\Rightarrow$ 

Potential difference developed  
b/w a metal rod and its solution is  
called electrode potential.

 $ZnSO_4(aq)$  $E_{\text{Oxid}}$  $CuSO_4(aq)$  $E_{\text{Red}}$ 

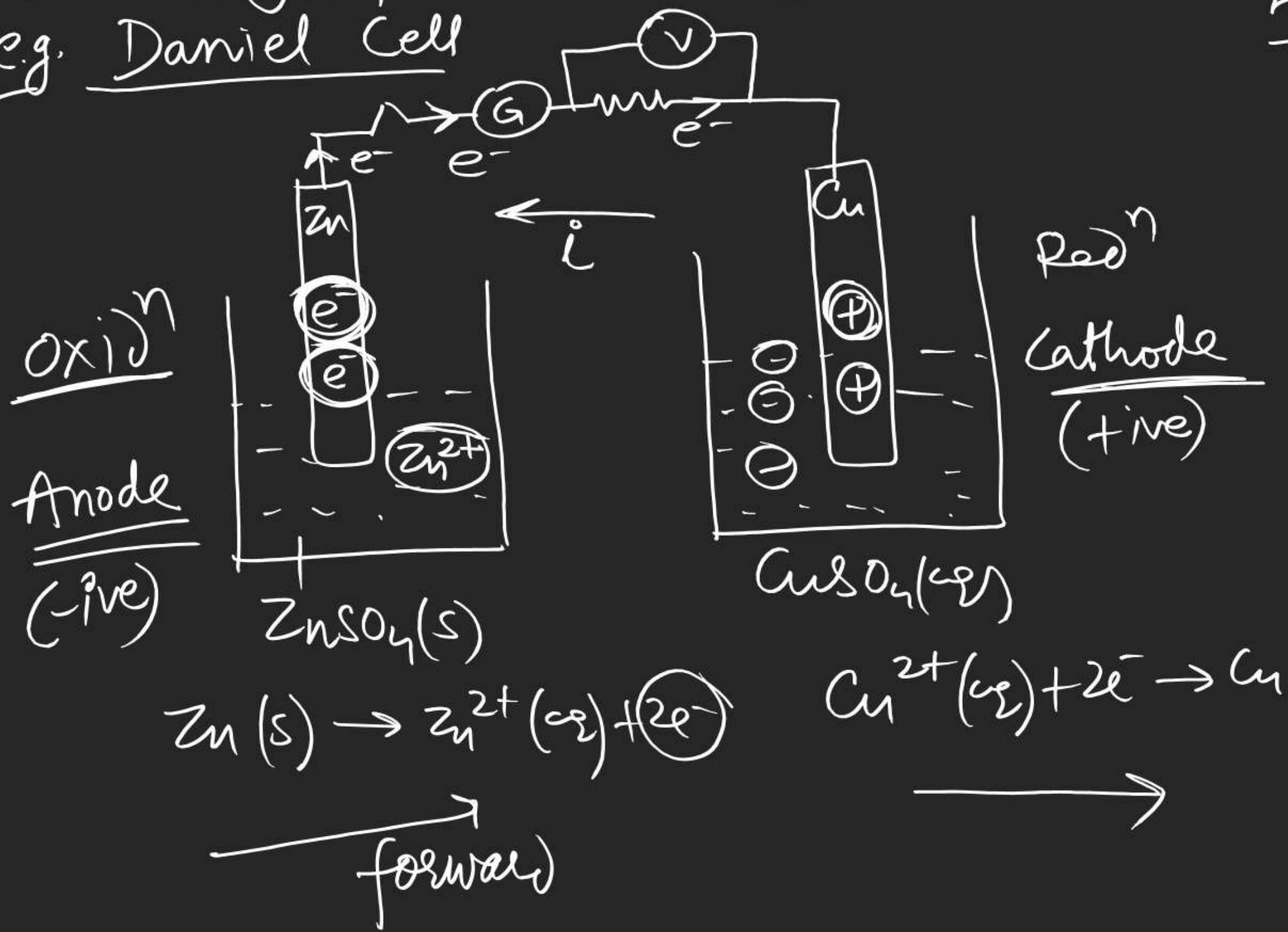
Interface



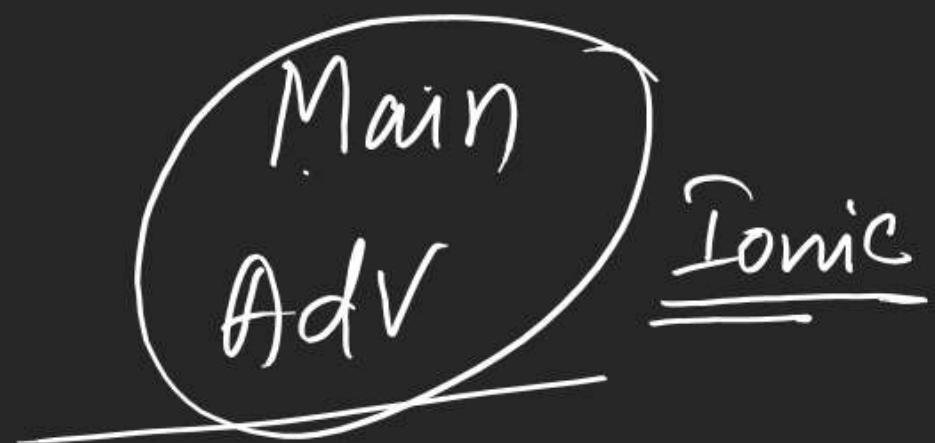
It is impossible to determine the absolute value of electrode potential because measuring process involves other interfaces whose electrode potential also appear in measured value.

# Working of a Galvanic cell

e.g. Daniel Cell



Anode - at which oxid<sup>n</sup> occurs



JEE Mains

Last 15 question

JEE Adv