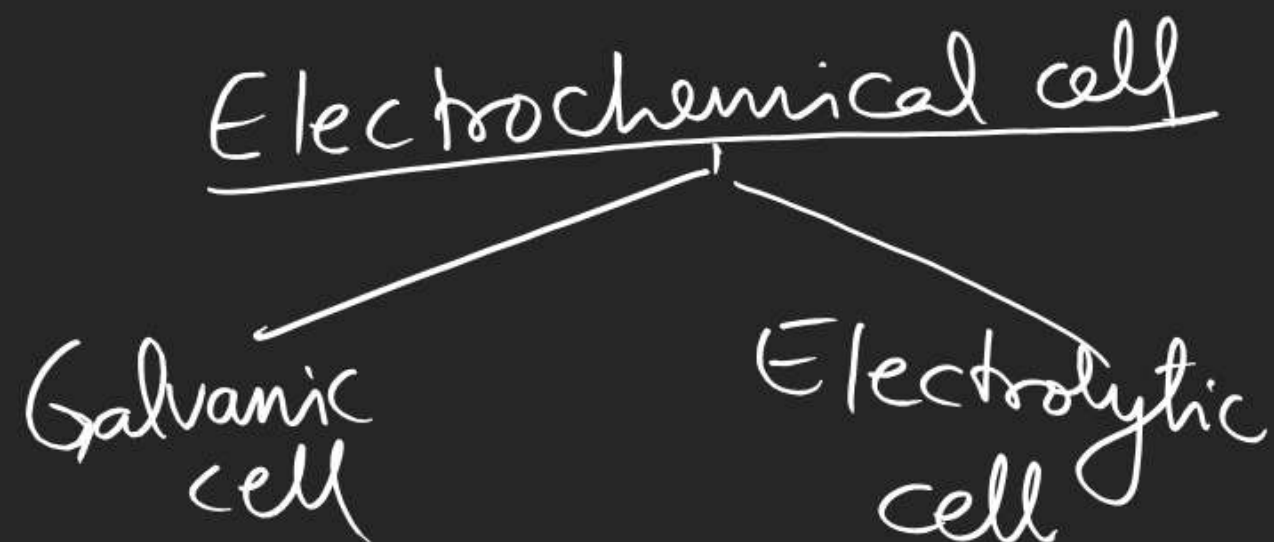


Electrochemistry

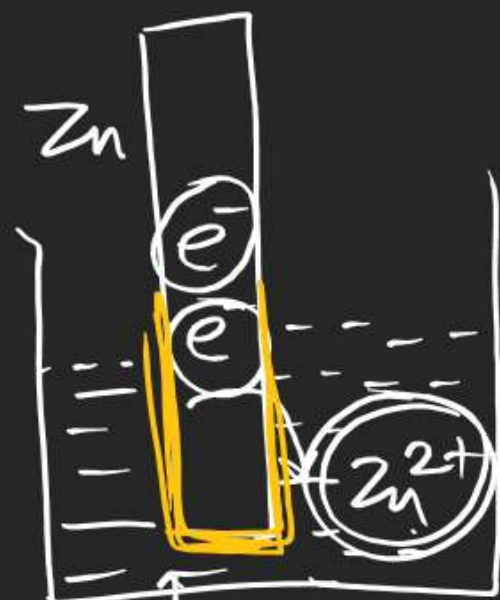
This chapter mainly deals with

- 1) Production of electric current by chemical reaction, and apparatus used is called "Galvanic cell"
- 2) Study of chemical change by electric current and apparatus used is called "Electrolytic cell"
- 3) Conductance & conductivity

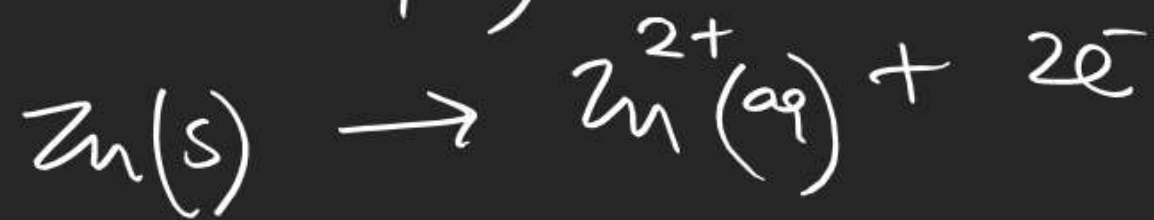


Galvanic cell

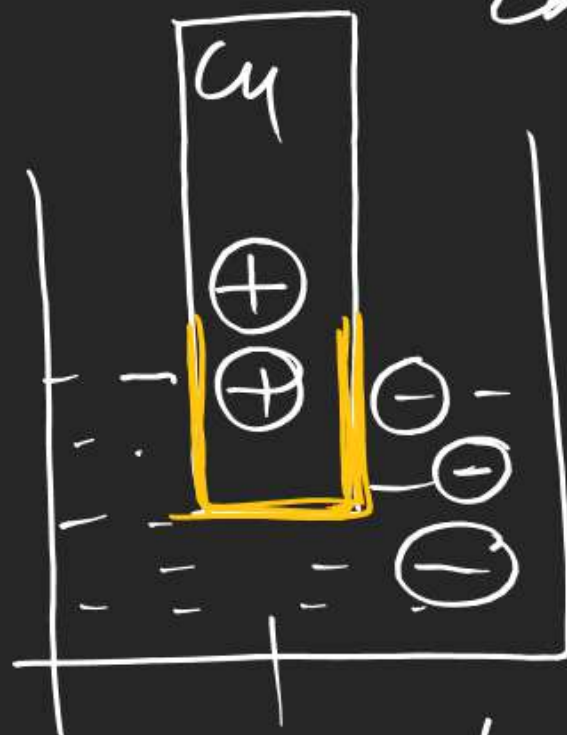
(i) Electrode potential \rightarrow Potential difference developed betⁿ a metal rod and its solution is called electrode potential.



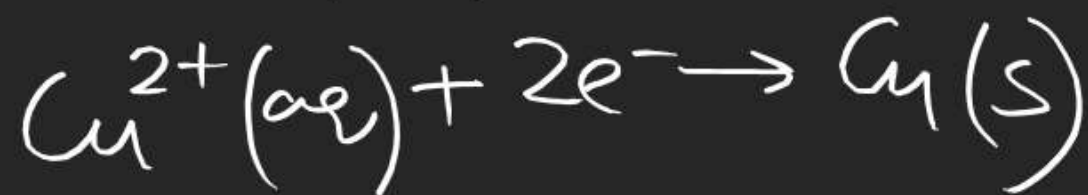
$\text{ZnSO}_4(\text{aq})$



E_{oxid}



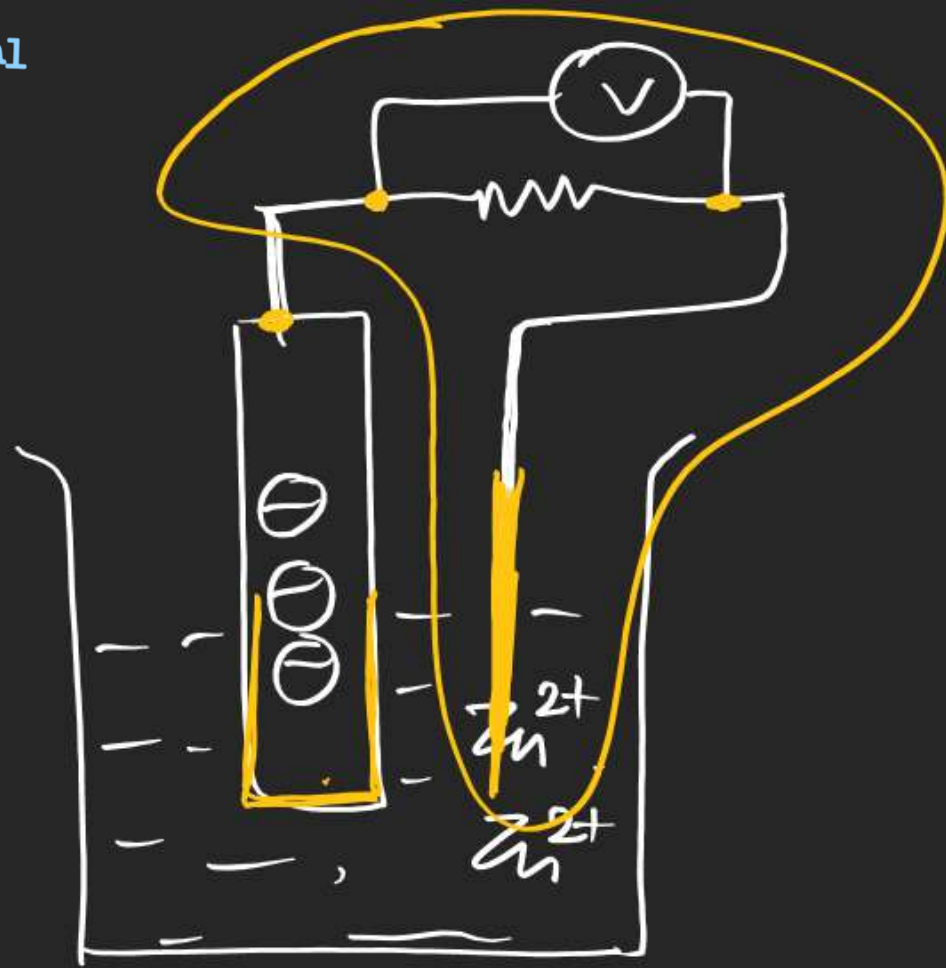
$\text{CuSO}_4(\text{aq})$



E_{Red}

Interface

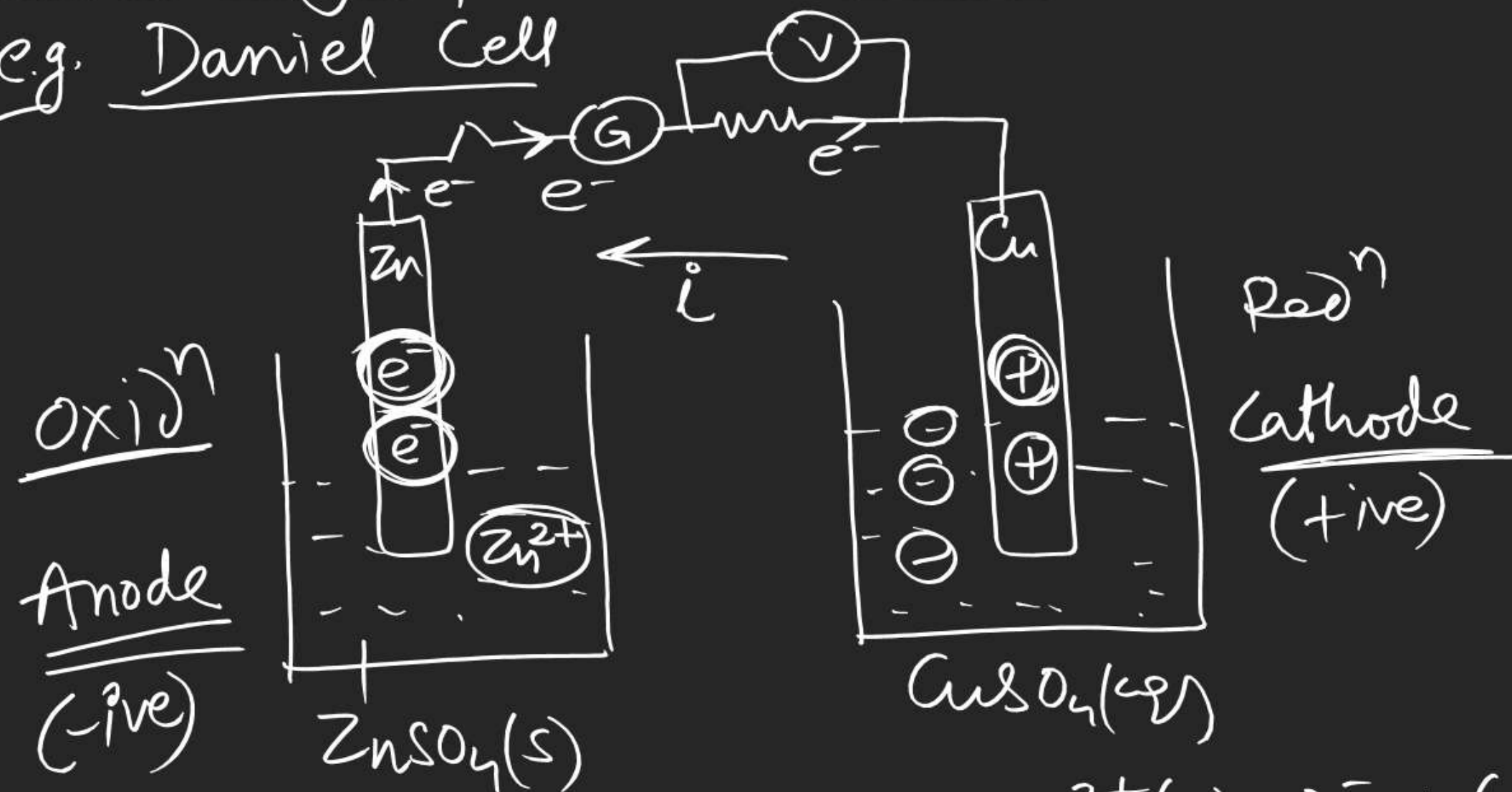




It is impossible to determine the absolute value of electrode potential because measuring process involves other interfaces whose electrode potential also appear in measured value.

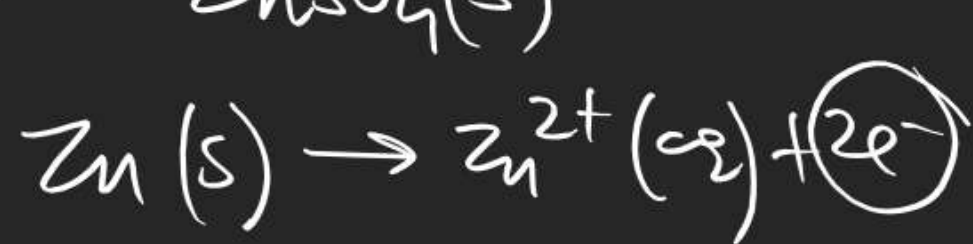
Working of a Galvanic cell

e.g. Daniel Cell



Anode - at which
oxidⁿ occurs

Main Adv Ionic



forward

JEE Mains

Last 15 question

JEE Adv