Nishant Jindal

O - I

J - Mains

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$$P = K_H \, x$$

1 lit $\quad$ 2 mg.

1 lit $\quad \dfrac{\dfrac{2 \times 10^{-3}}{34} \text{ mol}}{1000/18} = x$

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(10) $\dfrac{5.25}{M} = \dfrac{1.5}{60}$

(19) $\dfrac{5}{342} = \dfrac{1}{M}$

(8) $\underline{100^\circ C}$ $\underline{760 \text{ torr}}$

$$P_s = \frac{\frac{178.2}{18}}{\frac{18}{180} + \frac{178.2}{18}} \times 760$$

$$\frac{9.9}{0.1 + 9.9} \times 760 = \frac{9.9 \times 76}{\quad}$$

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$$\cancel{0.9}^{\,-0.1} = \cancel{1} + \left(\frac{1}{n}-1\right)\cancel{0.2}\;2$$

$$-1 = \frac{2}{n} - 2$$

$$1 = \frac{2}{n}$$

$$n\,M \longrightarrow M_n$$

$$1-\alpha \qquad \alpha/n$$

$$\frac{\alpha/n}{1-\alpha+\alpha/n} = 0.2$$

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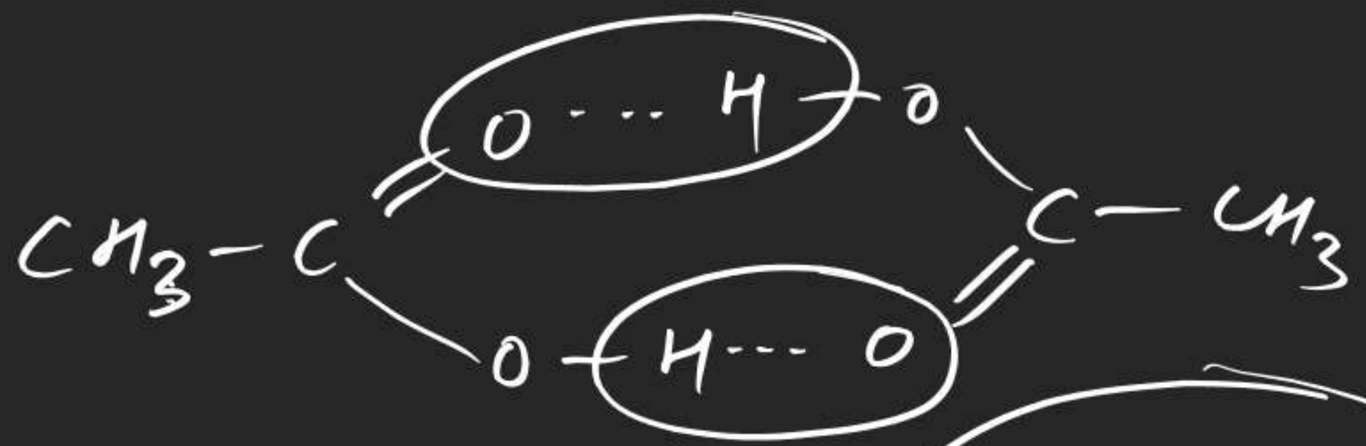
$$5 = 1.86 \times \frac{n}{10}$$

$$\frac{50}{1.86} \times 32$$



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Pure milk

-0.5

$\Delta T_f = 0.5$

diluted milk

-0.2

$\Delta T_f = 0.2$

H_2O

0

0.3 $\qquad$ 0.2

5 M $\rightarrow V_1$ $\qquad$ 3 M $\qquad$ 1 M $\rightarrow V_2$

$$\frac{5V_1 + 1V_2}{V_1 + V_2} = 3$$

$$2V_1 = 2V_2$$

$$V_1 = V_2$$

(59)

$$\boxed{P = K_H\, x_{gas}}$$

$$= K_H (1 - x_{H_2O})$$

$$P = K_H - K_H\, x_{H_2O}$$

$\boxed{\text{is cooled upto } -4\,^\circ C}$

$$4 = 1.86 \times \frac{0.6}{W} \times 1000$$

$$\boxed{W = 277.5}$$

(60)

$$4 = 1.86 \times \frac{0.75}{W} \times 1000$$

$$W = \frac{750 \times 1.86}{4}$$

1000 gm sol^n → m = 0.75

H_2O x gm

sucrose $1000 - x$ gm

$$m = \frac{\frac{1000-x}{342}}{x} \times 1000 = 0.75$$

$$x = 795.86$$

$$n_{C_{12}H_{22}O_{11}} = \frac{1000 - 795.86}{342} = 0.6$$

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