CONFIGURAÇÕES ELETRÓNICAS

|  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- |
| **Elemento Químico** | **Número Atómico** | **Configuração eletrónica** | **Eletrões valência** | **GRUPO** | **PERÍODO** | **BLOCO TP** | **IÃO mais comum** |
| Hidrogénio - **H** | 1 | **1s1** | 1 | 1 | 1 | s | H+ |
| Hélio - **He** | 2 | **1s2** | 2 | 18 | 1 | s | --- |
| Lítio – **Li** | 3 | 1s2**2s1** | 1 | 1 | 2 | s | Li+ |
| Berílio – **Be** | 4 | 1s2**2s2** | 2 | 2 | 2 | s | Be2+ |
| Boro – **B** | 5 | 1s2**2s22px12py02pz0** | 3 | 13 | 2 | p | --- |
| Carbono – **C** | 6 | 1s2**2s22px12py12pz0** | 4 | 14 | 2 | p | --- |
| Azoto – **N** | 7 | 1s2**2s22px12py12pz1** | 5 | 15 | 2 | p | N3- |
| Oxigénio – **O** | 8 | 1s2**2s22px22py12pz1** | 6 | 16 | 2 | p | O2- |
| Flúor – **F** | 9 | 1s2**2s22px22py22pz1** | 7 | 17 | 2 | p | F- |
| Néon – **Ne** | 10 | 1s2**2s22px22py22pz2** | 8 | 18 | 2 | p | --- |
| Sódio – **Na** | 11 | 1s22s22p6**3s1** | 1 | 1 | 3 | s | Na+ |
| Magnésio – **Mg** | 12 | 1s22s22p6**3s2** | 2 | 2 | 3 | s | Mg2+ |
| Alumínio – **Al** | 13 | 1s22s22p6**3s23px13py03pz0** | 3 | 13 | 3 | p | Al3+ |
| Silício – **Si** | 14 | 1s22s22p6**3s23px13py13pz0** | 4 | 14 | 3 | p | --- |
| Fósforo – **P** | 15 | 1s22s22p6**3s23px13py13pz1** | 5 | 15 | 3 | p | P3- |
| Enxofre – **S** | 16 | 1s22s22p6**3s23px23py13pz1** | 6 | 16 | 3 | p | S2- |
| Cloro - **Cl** | 17 | 1s22s22p6**3s23px23py23pz1** | 7 | 17 | 3 | p | Cl- |
| Árgon – **Ar** | 18 | 1s22s22p6**3s23px23py23pz2** | 8 | 18 | 3 | p | --- |
| Potássio - **K** | 19 | 1s22s22p63s23p6**4s1** | 1 | 1 | 4 | s | K+ |
| Cálcio - **Ca** | 20 | 1s22s22p63s23p6**4s2** | 2 | 2 | 4 | s | Ca2+ |
| Escândio - **Sc** | 21 | 1s22s22p63s23p6**4s23d1** | 3 | 3 | 4 | d | Sc3+ |
| Titânio - **Ti** | 22 | 1s22s22p63s23p6**4s23d2** | 4 | 4 | 4 | d | Vários |
| Vanádio - **V** | 23 | 1s22s22p63s23p6**4s23d3** | 5 | 5 | 4 | d | Vários |