

2 MONTHLY ASSESSMENT TEST - MAT - 2
CHEMISTRY.

Std : XI A & B

Timing : 1 1/2 hours

Marks : 50

I. Answer any Six questions $6 \times 2 = 12$

1. Draw and define ortho and para hydrogen molecule. (i)
2. Mention the uses of tritium. (ii)
3. What is dead burnt plaster? (iii)
4. Write the molecular formulae of the following compounds: (iv)
(i) Washing Soda (ii) baking soda (iii) Quick lime
(iv) Slaked lime.
5. Write any two characteristic properties of alkali metals. (ii)
6. State modern periodic law. (i)
7. What are isoelectronic ions? Give example. (i)

II. Answer any Six questions $6 \times 3 = 18$

8. Explain the diagonal relationship.
9. Why halogens act as oxidising agents?
10. How do you convert para hydrogen into ortho hydrogen?
11. What is water-gas shift reaction?
12. Mention the uses of plaster of paris.
13. How is plaster of paris prepared?
14. Write balanced chemical equation for each of the following chemical reactions.
(i) Solid potassium hydroxide with CO_2
(ii) heating calcium with oxygen.
(iii) Rubidium with oxygen gas.

III. Answer any four questions. $4 \times 5 = 20$.

15. Describe briefly the biological importance of Calcium and magnesium.

16. Explain the important common features of Group 2 elements?

17. (i) Discuss the three types of covalent hydrides.

(ii) Write chemical equation for the following reactions:

(a) reaction of hydrogen with tungsten(VI) oxide on heating.

(b) hydrogen gas and chlorine gas.

18. (i) Mention any two anomalous properties of second period elements.

(ii) What is screening effect?

19. Using Slater's rule, calculate the effective nuclear charge on a 3p electron in aluminium and chlorine. Explain how

these results relate to the atomic radii of the two atoms.

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