

01

(A) Explain solubility and different factors affecting solubility.

Answer:

A solution that can not dissolve any more solute at a given temperature is called saturated solution. If it can dissolve more, it is unsaturated.

Solubility:

Solubility is the amount of substance needed to make a solution saturated at a fixed temperature. For example the solubility of cobalt (II) chloride at 25°C is 78g per 100 mL of water.

It is generally expressed as amount of solute in grams needed to saturate

100 g of solvent.

$$\text{Solubility} = \frac{\text{mass of solute in gram} \times 100}{\text{mass of solvent}}$$

Factors affecting solubility:

Solubility depends on:

- i) Nature of solvent
- ii) Nature of solute
- iii) Temperature
- iv) Pressure

i) Nature of solvent:

Generally organic compounds are dissolved in organic solvents and for inorganic compounds, they generally dissolve in inorganic solvents. It depends on polarity of solvent. Turpin oil dissolve dye needed to color doors and windows

while water dissolves NaCl .

ii) Nature of solute:

Same solvent can not dissolve all solutes equally. For example, in water at 30°C , at best 37g of NaCl or, 50g of Na_2SO_4 or 45g of KNO_3 can be dissolved in 100mL solution.

iii) Effect of temperature:

Generally solubility increases with the increase of temperature. But if dissolving becomes exothermic, the solubility decreases. Solubility of KNO_3 , KI , AgNO_3 increase with increase of temperature but for $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$, it increases upto 32.4°C and then.

decreases as it converts into dry Na_2SO_4 . It decreases for gaseous solute in liquid solvent.

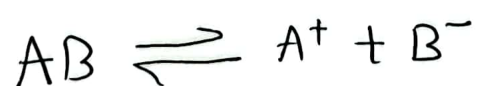
iv) Pressure:

It mainly affects gases. When pressure goes up, it increases collision frequency and thus solubility increases.

⑧ Derive the equation of solubility product of some sparingly soluble salt.

Answer!

Let AB be a sparingly soluble salt with solubility S mol L^{-1}



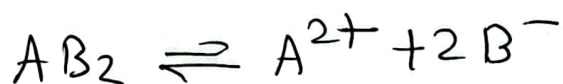
So, In solⁿ, the concentration of $[\text{A}^+] = \text{S mol L}^{-1}$
 " " $[\text{B}^-] = \text{S mol L}^{-1}$

∴ Solubility product of AB

$$K_{sp} = [A^+] \times [B^-] \\ = s \times s = s^2 \text{ mol}^2 \text{ L}^{-2}$$

For AB₂

If AB₂ is salt with solubility $s \text{ mol L}^{-1}$



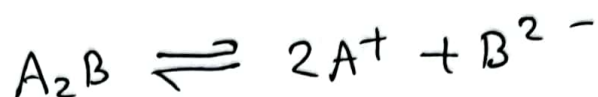
$$\text{In sol}^n, [A^{2+}] = s \text{ mol L}^{-1}$$

$$[B^-] = 2s \text{ mol L}^{-1}$$

∴ K_{sp} of AB₂

$$K_{sp} = [A^{2+}] [B^-]^2 \\ = s \times (2s)^2 \\ = 4s^3 \text{ mol}^3 \text{ L}^{-3}$$

For A_2B Let solubility be $s \text{ mol L}^{-1}$



In solution, $[A^+] = 2s \text{ mol L}^{-1}$

$$[B^{2-}] = s \text{ mol L}^{-1}$$

$\therefore K_{sp}$ of A_2B

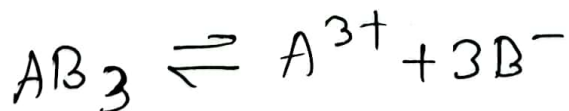
$$K_{sp} = [A^+]^2 \times [B^{2-}]$$

$$= (2s)^2 \times s$$

$$= 4s^3 \text{ mol}^3 \text{ L}^{-3}$$

For AB_3 :

If solubility is $s \text{ mol L}^{-1}$



So, in solution $[A^{3+}] = s \text{ mol L}^{-1}$

$$[B^-] = 3s \text{ mol L}^{-1}$$

7
 $\therefore K_{sp}$ of AB_3

$$\begin{aligned} K_{sp} &= [A^{3+}] \times [B^{-}]^3 \\ &= S \times (3S)^3 \\ &= 27S^4 \text{ mol}^4 \text{ L}^{-4} \end{aligned}$$

© Calculate the solubility product of a saturated solution of A_2B salt in water at 25°C temperature in which the concentration of B^{2-} ion is $6.4 \times 10^{-5} \text{ M}$.

Answer:

Let the solubility of A_2B salt be $S \text{ mol L}^{-1}$



8

As at saturated solution the concentration of $[B^{2-}]$ is $s = 6.4 \times 10^{-5} M$

So, solubility of A_2B will be $s = 6.4 \times 10^{-5} \text{ mol L}^{-1}$

So, solubility product of A_2B ,

$$K_{sp} = [A^+]^2 \times [B^{2-}]$$

$$= (2s)^2 \times s$$

$$= 4s^3$$

$$= 4 \times (6.4 \times 10^{-5})^3$$

$$= 1.048 \times 10^{-12} \text{ mol}^3 \text{ L}^{-3}$$

(Answer)

① 2 mL 0.25 M NH_4OH solution is added in 10 mL 0.30 M AB_3 solution. If the solubility product of $\text{A}(\text{OH})_3$ at 25°C is 3.98×10^{-38} then justify the probability of getting precipitation in solution.

Answer:



Here, 2 mL 0.25 M NH_4OH is added in a solution of 10 mL 0.30 M AB_3 solution.

$$\begin{aligned}\text{So total volume of solution} &= (10 + 2) \text{ mL} \\ &= 12 \text{ mL}\end{aligned}$$

$$\begin{aligned}\text{Concentration of } \text{A}^{3+}, [\text{A}^{3+}] &= \frac{10 \times 0.3}{12} \text{ M} \\ &= 0.25 \text{ M}\end{aligned}$$

Concentration of OH^- , $[\text{OH}^-] = \frac{2 \times 0.25}{12}$
 $= 0.04167 \text{ M}$



$$K_{ip} = [\text{A}^{3+}] \times [\text{OH}^-]^3$$
$$= 0.25 \times (0.04167)^3$$

$$K_{ip} = 1.8 \times 10^{-5} \text{ mol}^4 \text{ L}^{-4}$$

$$\text{And } K_{sp} = 3.98 \times 10^{-38} \text{ mol}^4 \text{ L}^{-4}$$

So, $K_{ip} > K_{sp}$

So, precipitation of $\text{A}(\text{OH})_3$ will be produced here.