

ENERGY

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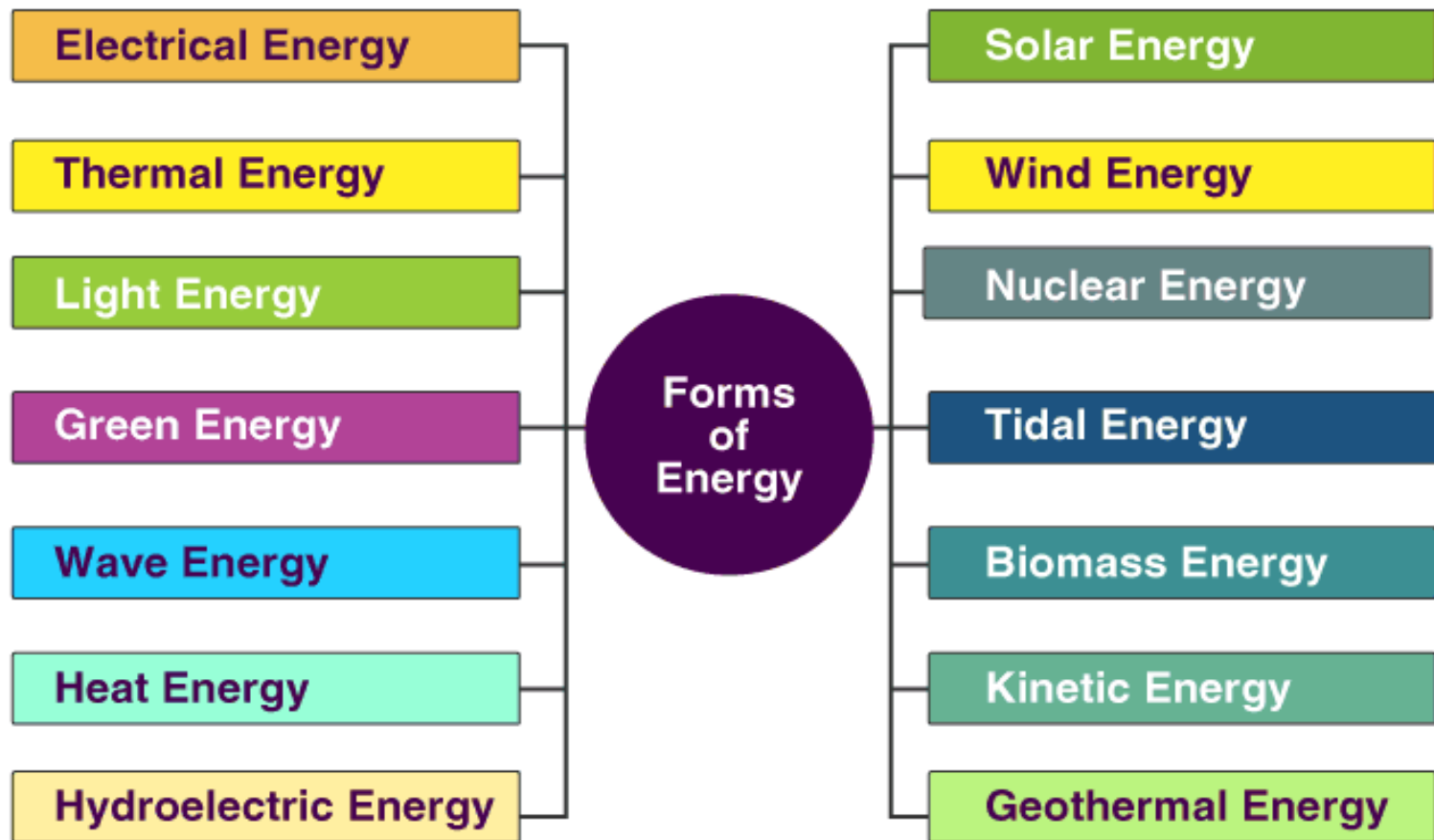
Modified by Dr. Druman Utekar

Energy

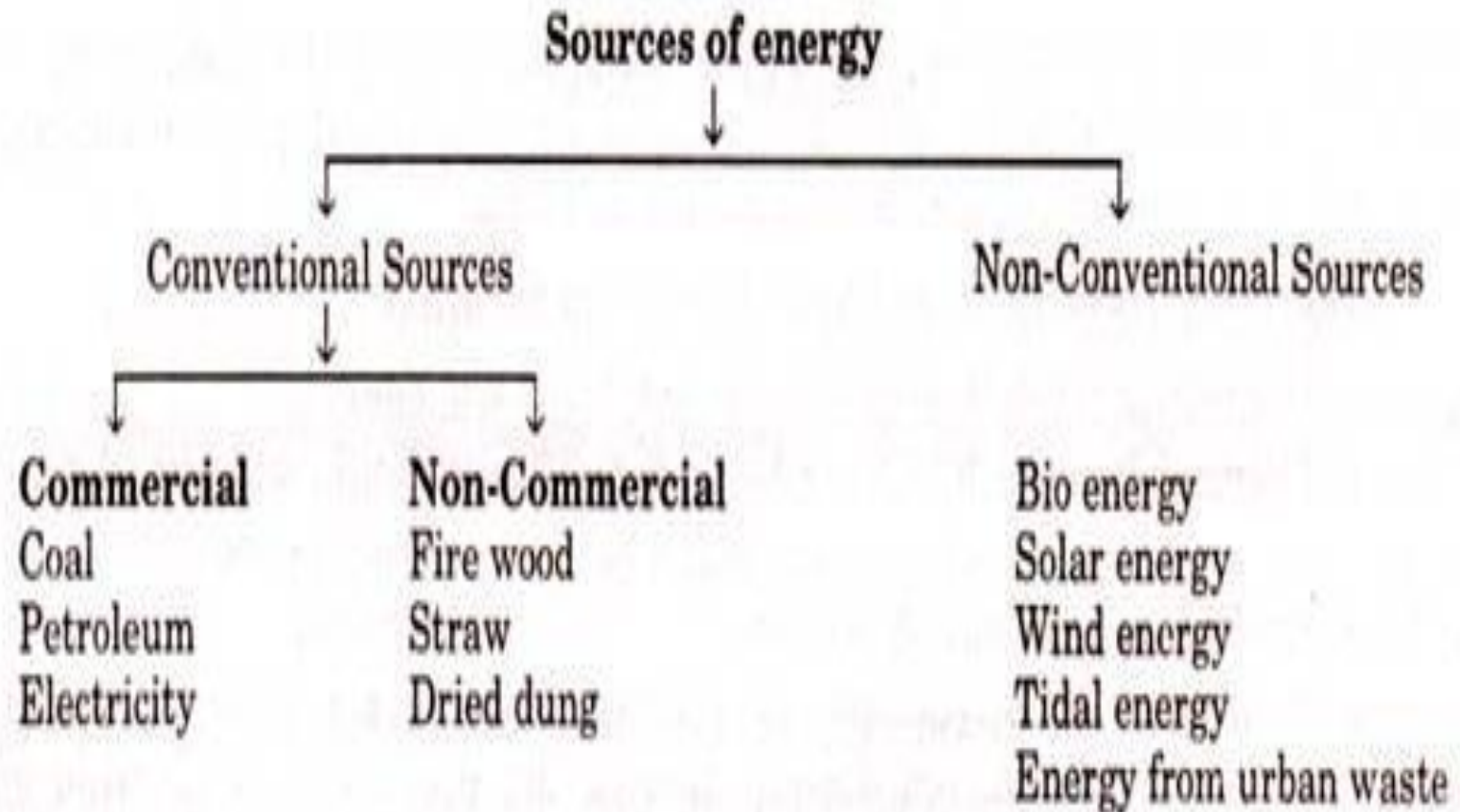
The sun, directly or indirectly, is the source of all the energy available on Earth. Energy is essential to life and all living organisms.

Energy is a conserved quantity and the law of conservation of energy states that energy can neither be created nor destroyed but can only be converted from one form to another. The SI unit of energy is Joule.

Types of Energy



Classification of Energy



Distinction: Conventional (non-renewable) and nonconventional energy (renewable) sources.

Renewable resources	Non- renewable resources
Resources which can be renewed or can be reused are renewable resources.	Resources which cannot be renewed or reused once utilized are called non-renewable resources.
These include components like air, water, wind, sunlight etc.	These include components like fossil fuels, LPG gases.
They are sustainable resources.	They are exhaustible resources.
Their rate of renewal is greater than the rate of getting exhausted.	Their rate of renewal is slower than the rate of getting exhausted.
They are mostly environmental friendly and does not cause pollution.	They are the main cause of pollution.

Solar Energy

The electromagnetic radiation from sun is commonly known as solar energy. These radiations are resulted from thermo nuclear fusion reaction on the surface of sun. All the radiation from the sun is not

in the same wavelength range. Almost 92% lie in the range of 315 nm to 1400nm. The estimated amount of solar flux reaching the atmosphere of earth is approximately 1400W/m²min. and that of heat

equivalent is 2.68×10^{24} J/Year. The eco system of earth utilizes about 0.2-0.5 % of total amount of solar energy received. It indicates clearly that large amount solar energy get wasted, which otherwise

can be immense use for satisfying needs of humans.

Advantage of solar energy:

1. It is non-polluting and non-depleting source of energy.
2. It is renewable source of energy.
3. It is available abundantly.

In spite of these advantages, the use of solar energy in large scale is still not in practice, due to following reasons,

1. Non availability of intense light in all areas throughout year
2. Difficulties faced in economic collection and conversion of solar energy into other forms of energy such as electricity.

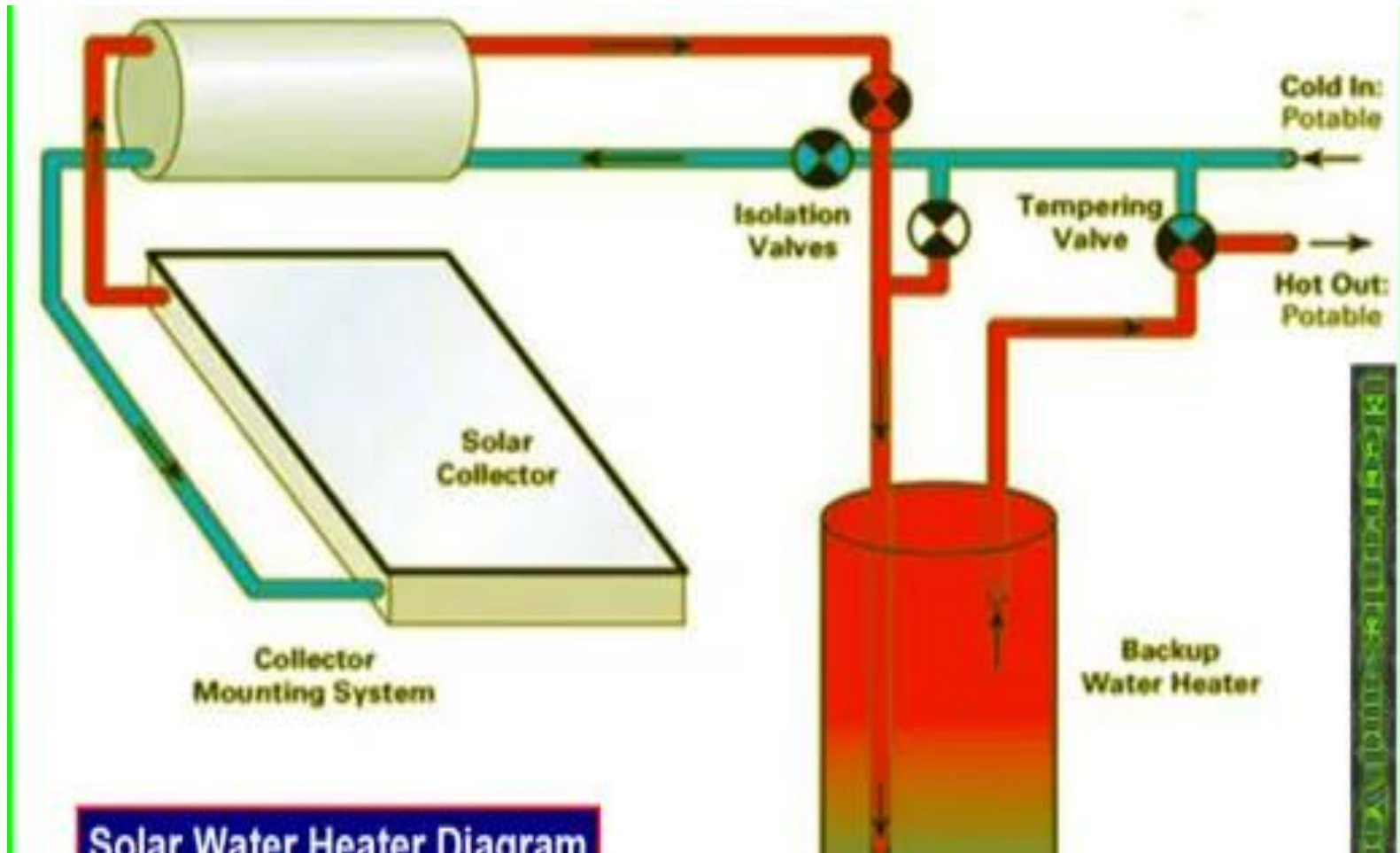
The solar energy has been successfully used in following purpose

- i) Heating: Used for water and space heating in colder countries.
- ii) Electricity: Using solar energy electric energy can be generated.

Advantages and Disadvantages

Advantages	Disadvantages
Reduces consumption of fossil fuels.*	Currently, electricity from PV systems is more expensive than electricity produced from fossil fuel or nuclear power plants.
Reduces production of greenhouse gases.*	Expensive to buy.
Reduces production of various pollutants.*	Requires engineering expertise to design and install systems.
Good for remote applications: satellites, rural hospital equipment in developing countries, telecommunication equipment, etc.	Production of PV systems from single silicon crystals is technically challenging, and energy- and time-consuming.
Reduces the loss of electricity due to power line resistance (distribution losses) because it can be sited where the electricity is used.	Sunlight is not constant, so must get electricity from other sources at night or on cloudy days or store it (such as batteries, etc.)
Reduces water consumed in electrical generation processes by displacing electrical demand.	Sunlight is diffuse; PV would take much space to produce enough electricity to meet our current needs (an area ~one sixth the size of Arizona)
Does not contribute to thermal pollution of waterways.	* Once manufactured, PV systems produce no waste products. Manufacturing of almost any device uses some nonrenewable resources, consumes energy and produces waste products. PV systems consume some nonrenewable resources if a system component needs repair or maintenance (such as batteries, inverter, etc.).
No hidden costs.	
Can provide energy independence.	
PV cells last ~ 30 years.	
Uses a renewable energy source.	

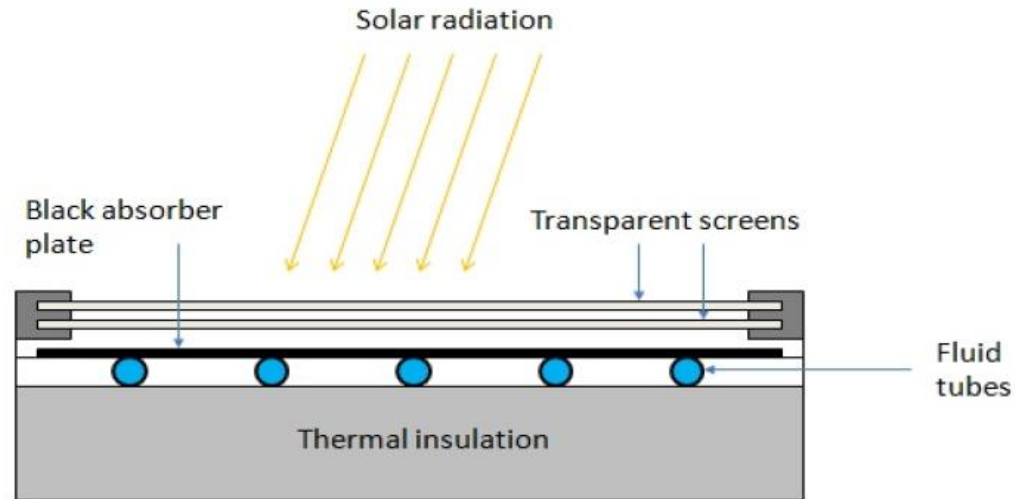
Solar Water Heater



Flat Plate Collector:

The device works on the principle of black body in which heat absorbing capacity and tendency of a black surface is utilized to achieve benefits for human.

Diagram:



Construction:

These are the main components of a typical flat-plate solar collector:

- ☐ Black surface - absorbent of the incident solar energy
- ☐ Glazing cover - a transparent layer that transmits radiation to the absorber, but prevents radiative and convective heat loss from the surface
- ☐ Tubes containing heating fluid to transfer the heat from the collector
- ☐ Support structure to protect the components and hold them in place
- ☐ Insulation covering sides and bottom of the collector to reduce heat losses

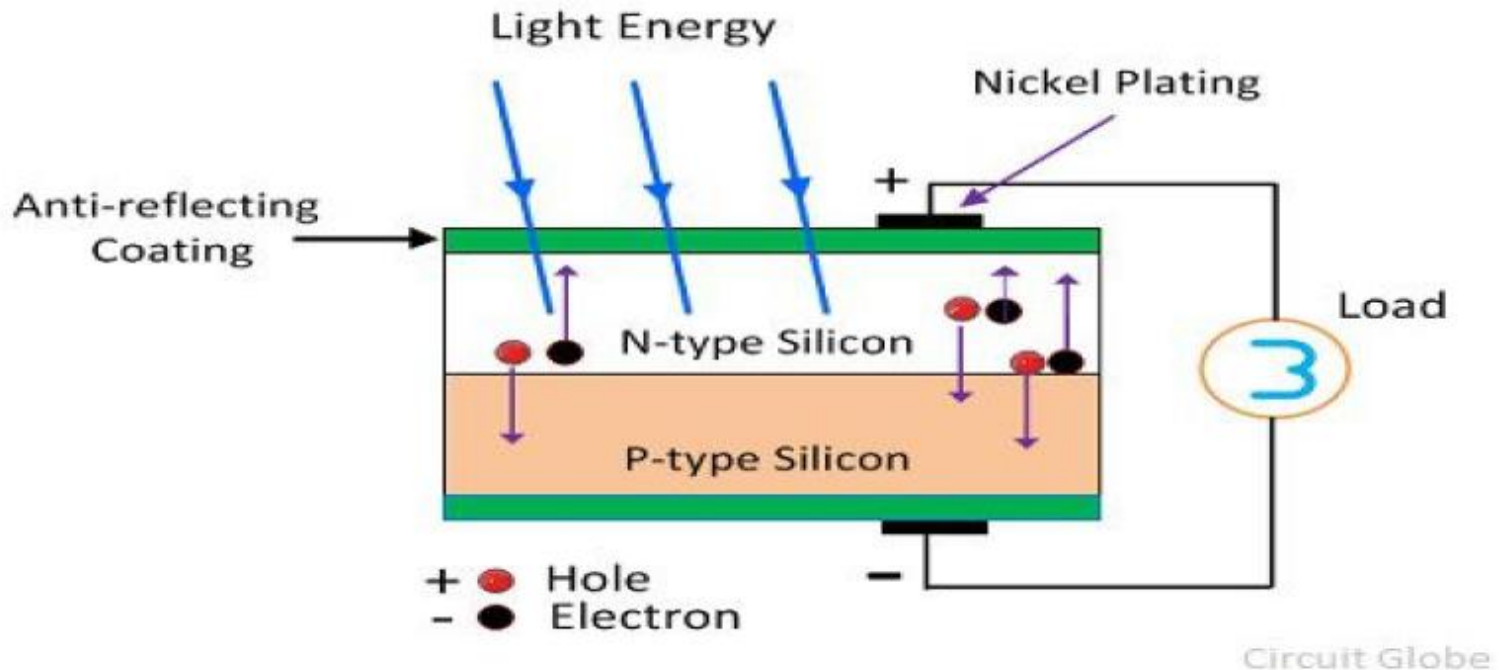
Application

Some advantages of the flat-plate collectors are that they are: Easy to manufacture

- ☐ Low cost, Collect both beam and diffuse radiation
- ☐ Permanently fixed (no sophisticated positioning or tracking equipment is required)
- ☐ Little maintenance

PHOTO VOTAIC CELL (SOLAR CELL)

A conventional solar cell structure is shown in figure:



Semiconductors like silicon has the capacity to absorb light and deliver a portion of the energy of the

absorbed photons to carry charge carriers (electrons and hole). Thus solar cell is a semiconductor

diode that has been designed carefully so that it can absorb the light energy efficiently and convert

light energy from the sun into electrical energy.

Construction: A typical silicon photo voltaic cell composed of thin layer of phosphorus doped silicon (n-type) on top of boron doped (p-type) silicon. Hence these two layers form p-n junction. A metallic grid is the electrical contact of the diode and allows light to fall on the semiconductor between the grid lines. An anti-reflective layer between the grid lines increases the amount of light transmitted to semiconductor.

Working: of tiny energy packets called photon. When light radiation falls on the p-n junction diode,

photons are absorbed and electron-hole pairs are generated. The electrons are diffused and collected

at the n-type end and holes are diffused and collected at the p-type end. When these two ends are electrically connected through a conductor, there is a flow of current between the two ends through

the external circuit. Thus photoelectric current is produced and available for use. The current output

of a cell depends on its efficiency and size and is proportional to the intensity of sun light striking the

surface of the cell. Therefore, photovoltaic cells are connected electrically in series or parallel circuits to produce higher voltages, currents and power levels. A number of solar cells electrically connected

to each other and mounted in a support structure or frame is called a photovoltaic module. A

photovoltaic array is the complete power generating unit, consisting of any number of photovoltaic modules and panels.

Advantages & Disadvantages

Photovoltaic solar energy

Advantages:

- environmentally friendly
- no noise, no moving parts
- no emissions
- no use of fuels and water
- minimal maintenance requirements
- long lifetime, up to 30 years
- electricity is generated wherever there is light, solar or artificial
- PV operates even in cloudy weather conditions
- modular “custom-made” energy can be sized for any application from watch to a multi-megawatt power plant

Limitations:

- PV cannot operate without light
- high initial costs that overshadow the low maintenance costs and lack of fuel costs
- large area needed for large scale applications
- PV generates direct current special DC appliances or an inverter are needed
- in off-grid applications energy storage is needed

Fuel cells

- H₂-O₂ fuel cells-Self Study
- Fuels can be defined as substances which undergo combustion in the presence of air to produce a large amount of heat that can be used economically for domestic and industrial purpose.
- Examples, Wood, Coal, Kerosene, Petrol

Classification of chemical fuels

- A) Based on the origin:
 - i) Primary or natural fuels
 - ii) Secondary or artificial or derived fuels

- B) Based on Physical State:
 - i) Solid Fuels
 - ii) Liquid Fuels
 - iii) Gaseous Fuels

- C) Based on Chemical Nature
 - i) Organic eg. Vegetable fuel, coal
 - ii) Inorganic eg. Iron Pyrites
 - iii) Nuclear Fuels eg. Uranium oxide

Characteristic Properties of Fuels

- Fuels are characterized by testing certain physical and chemical properties.
 - i) Calorific Value should be as high as possible.
 - ii) Ignition temperature-Moderate
 - iii) Flame temperature should be as high as possible.
 - iv) Flash and Fire point should be as high as possible.
 - v) Aniline point should be low.
 - vi) Cloud and Pour point should be as low as possible.
 - vii) Viscosity should be adequate.
 - viii) Coke number should be as high as possible.
 - ix) Moisture content-as low as possible.
 - x) Volatile matter as low as possible.
 - xi) Ash content should be absent.
 - xii) Easy risk free transport should be possible.
 - xiii) Storage space-ideally fuel should occupy small space.
 - xiv) Air requirement- adequate
 - xv) Harmless products should be produced on combustion.

Calorific Value

- Calorific value is defined as the number of parts of water which gets heated through 1°C by the heat evolved by the complete combustion of one unit weight of fuel(unit volume of gaseous fuels) under the conditions such as
 - i) Whole of heat evolved is absorbed by water.
 - ii) The products formed leave the system at atmospheric temperature and pressure.
- It is the most important property of fuel.

Units of calorific value

1) B.T.U. (British Thermal Unit)

A British thermal unit may be defined as the heat required to raise the temperature of one pound of water from 60°F to 61°F.

2) K.C.U. (Kilogram Centigrade Unit)

The calorie, a unit of heat may be defined as, the heat required to raise the temperature of one Kg of water from 15°C to 16°C.

Correlation between BTU and KCU:

$$1\text{BTU} = 0.252\text{KCal} = 252\text{Cal}$$

$$1\text{KCal} = 3.968 \text{ BTU}$$

3) C.H.U. (Centigrade Heat Unit)

The calorific value can also be expressed as centigrade heat unit (C.H.U.), which is the amount of heat required to raise temperature of one pound of water through one degree centigrade.

1) High Calorific value (HCV) or Gross Calorific value (GCV):

High calorific value may be defined as the total amount of heat produced when one unit of the fuel has been burnt completely and the products of combustion have been cooled to 16°C or 60°F.

2) Low calorific value(LCV) or Net Calorific value (NCV)

Low calorific value may be defined as the net heat produced when unit mass or volume of fuel is completely burnt and products are allowed to escape.

$$\begin{aligned}\text{NCV or LCV} &= \text{GCV (HCV)} - \text{Latent Heat of water formed} \\ &= \text{GCV(HCV)} - \text{Mass of hydrogen} \times 9 \times \text{Latent heat of steam} \\ &= \text{GCV(HCV)} - 0.09 \times \%H \times 587\end{aligned}$$

Because 1 part by weight of hydrogen produces 9 parts (1+8) by mass of water.

Dulong Formula

- The calorific value of fuels is determined theoretically by Dulong formula or I.A. Davies formula.
- It is expressed as

$$Q = 1/100 [8080 \times C + 34500 \times (H - O/8) + 2240 \times S]$$

Where, Q = Calorific value in KCU/Kg

C = % of carbon

H = % of hydrogen

O = % of oxygen

S = % of sulphur

Dulong formula for HCV & LCV

$$\text{HCV} = 1/100 [8080 \times C + 34500 \times (H - O/8) + 2240 \times S]$$

$$\text{LCV} = \text{HCV} - [9/100 \times \%H \times 587]$$

Experimentally calorific value of solid and liquid fuel is determined using Bomb Calorimeter.

- Numerical 1: A sample of coal contains C = 55%, O = 28%, H = 7%, S = 0.7%, N = 0.2%, Ash = 0.2%. Calculate the GCV and NCV.

Soln:

$$\begin{aligned} \text{GCV} &= 1/100[8080C + 34500(H - O/8) + 2240S] \\ &= 1/100[8080 \times 55 + 34500(7 - 28/8) + 2240 \times 0.7] \\ &= 5667 \text{ Kcal/Kg} \end{aligned}$$

$$\begin{aligned} \text{NCV} &= \text{HCV} - 0.09 \times \% \text{H} \times 587 \\ &= 5667 - 0.09 \times 7 \times 587 \\ &= 5297.19 \text{ Kcal/Kg} \end{aligned}$$

Numerical 2: A sample of coal has following composition C = 70%, O = 8%, H = 10%, N = 3%, S = 2%, Ash = 7%. Calculate HCV and GCV.

Numerical 3: A sample of coal contains C = 61%, O = 32%, S = 0.5%, N = 0.2% and Ash = 0.3%. If NCV of coal is 5313.02 KCal/Kg. Calculate % H and GCV.

$$\text{GCV} = 1/100[8080C + 34500(H - O/8) + 2240S] \text{-----i)}$$

$$\text{NCV} = \text{GCV} - 0.09 \times \% \text{H} \times 587$$

$$\text{GCV} = \text{NCV} + 0.09 \times \% \text{H} \times 587 \text{-----ii)}$$

$$1/100[8080C + 34500(H - O/8) + 2240S] = \text{NCV} + 0.09 \times \% \text{H} \times 587$$

Numerical 3: A sample of coal contains C = 61% , O = 32%, S = 0.5%, N = 0.2% and Ash = 0.3%. If NCV of coal is 5313.02KCal/Kg. Calculate % H and GCV.

$$\text{GCV} = 1/100[8080C + 34500(H-O/8) + 2240S] \text{----i)}$$

$$\text{NCV} = \text{GCV} - 0.09 \times \%H \times 587$$

$$\text{GCV} = \text{NCV} + 0.09 \times \%H \times 587 \text{-----ii)}$$

$$1/100[8080C + 34500(H-O/8) + 2240S] = \text{NCV} + 0.09 \times \%H \times 587$$

Numerical 3: A sample of coal contains C = 70% , O = 25%, S = 1%, N = 1% and Ash = 0.5%. If NCV of coal is 5200KCal/Kg. Calculate % H and GCV.

$$\text{GCV} = 80.80 \times 70 + 345(H - 25/8) + 22.40 \times 1 \text{----i)}$$

$$= 5656 + 345H - 1078.125 + 22.4$$

$$= 4600 + 345H$$

$$\text{NCV} = \text{GCV} - 0.09 \times \%H \times 587$$

$$\text{GCV} = 5200 + 0.09 \times \%H \times 587$$

$$= 5200 + 52.83H \text{-----ii)}$$

$$4600.275 + 345H = 5200 + 52.83H$$

$$345H - 52.83H = 5200 - 4600.275$$

$$292.17H = 599.725$$

$$H = 2.05\%$$

$$\text{GCV} = 5308.44 \text{Kcal/Kg}$$

Coal

Purpose of Analysis of Coal

- To decide price of coal
- To determine quality
- To specify use of coal for a particular purpose.
- To calculate theoretical calorific value of coal.
- To calculate air requirement for complete combustion of coal and design the furnace fire box suitably.

- Types of Analysis of Coal

- A) Proximate Analysis

- i) % Moisture
 - ii) % Volatile Matter
 - iii) % Ash
 - iv) Fixed Carbon

- B) Ultimate Analysis

- i) % Carbon and % Hydrogen
 - ii) % Nitrogen
 - iii) % Sulphur
 - iv) % Ash

Proximate Analysis of Coal

Proximate analysis is the study or analysis of coal sample in which

a) % Moisture b) % VM c) % Ash d) % Fixed Carbon are found out.

a) % Moisture

A known weight of powdered and air dried coal sample is taken in a crucible and it is placed in an oven for 1hr at 110°C. Then the coal is cooled in a desiccator and weighed out. If the initial weight of the coal is W gms and final weight is W1 gms.

Then the loss in weight (W-W1) corresponds to moisture in coal.

$$\% \text{ Moisture} = \frac{\text{Loss in weight} \times 100}{\text{Weight of coal sample}} = \frac{(W-W1) \times 100}{W}$$

b) % Volatile Matter

Moisture free coal left in the crucible in first experiment W1 is covered with a lid loosely. Then it is heated at 925°C in a muffle furnace for 7min.

The crucible is taken out and cooled in a desiccator. Then it is weighed (W2 gms) .

$$\% \text{ Volatile matter} = \frac{\text{weight of volatile matter} \times 100}{\text{Weight of coal sample}} = \frac{(W1-W2) \times 100}{W}$$

c) % Ash

The residual coal in the above experiments is heated and burnt in an open crucible at above 750°C for half hour.

The coal gets burnt. The ash left in crucible is cooled in a desiccator and weighed W3 gms.

$$\% \text{ Ash} = \frac{\text{weight of ash} \times 100}{\text{Weight of coal}} = \frac{W3 \times 100}{W}$$

d) % Fixed Carbon:

It is found by calculations

$$\% \text{FC} = 100 - [\% \text{ moisture} + \% \text{ VM} + \% \text{ Ash}]$$

Significance or Importance of proximate analysis

a) Moisture

It decreases calorific value of coal largely as it does not burn and takes away heat in the form of latent heat.

It increases ignition point of coal.

Hence a coal with lower moisture % is better quality.

b) Volatile matter

It decreases calorific value of coal.

It elongates flame and decreases flame temperature.

It forms smoke and pollutes air.

However, the coals containing 15-25% of VM on carbonization give coke oven gas which is a source of various organic aromatic chemicals.

Such coal has good caking property and coke can be obtained from the coals.

Overall regarding burning of coal, the coal with lesser VM is better quality coal.

c) Ash

It reduces calorific value of coal as ash is non-burning part in coal.

Ash disposal is a problem.

Ash fuses to form clinker at high temperature, obstructing the air supply of coal burning furnace.

Hence, lesser the % ash, better is the quality of coal.

d) Fixed carbon

Carbon is the burning part in coal and higher the fixed carbon, higher is the calorific value.

Hence, a good quality coal contains high fixed carbon percentage.

Numerical 1: A sample of coal was analysed for content of moisture, volatile matter & ash. From the following data, calculate the percentage of the above quantities.

i) weight of coal taken = 2.5g

ii) weight of coal after heating at 110°C = 2.365g

iii) weight of coal after heating at covered crucible at 950°C = 1.165g

iv) constant weight obtained at the end of the experiment = 0.460g

Soln:

$$\text{a) \% Moisture} = \frac{(W - W_1) \times 100}{W} = \frac{(2.5 - 2.365) \times 100}{2.5} = 5.4\%$$

$$\text{b) \% volatile Matter} = \frac{(W_1 - W_2) \times 100}{W} = \frac{(2.365 - 1.165) \times 100}{2.5} = 48\%$$

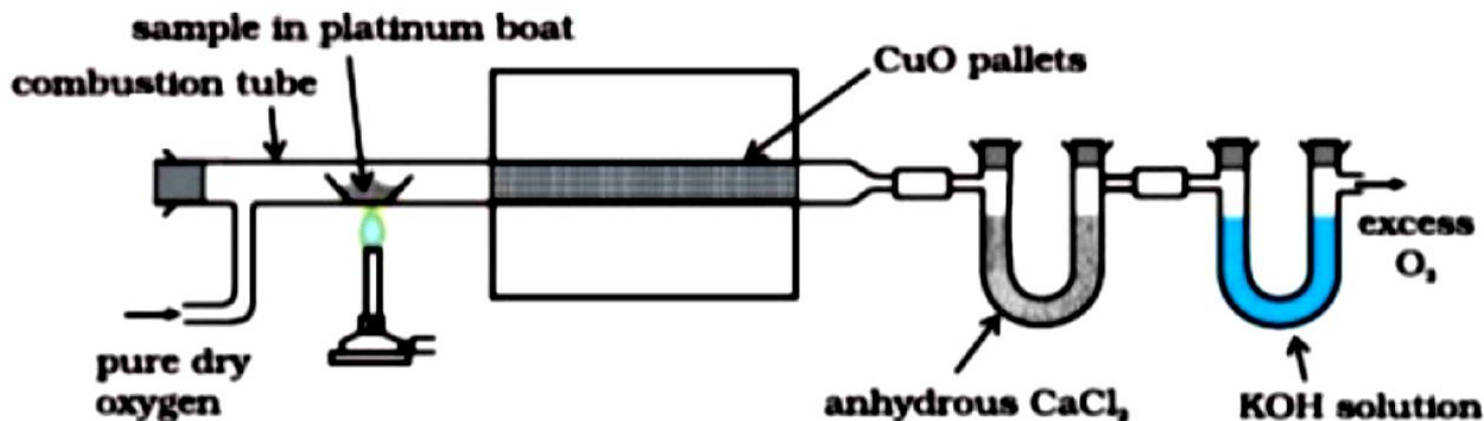
$$\text{c) \% Ash} = \frac{W_3 \times 100}{W} = \frac{0.460 \times 100}{2.5} = 18.4\%$$

$$\begin{aligned} \text{d) \% Fixed Carbon} &= 100 - [\% \text{ moisture} + \% \text{ VM} + \% \text{ Ash}] \\ &= 100 - [5.4 + 48 + 18.4] = 28.2\% \end{aligned}$$

Ultimate Analysis

1) Carbon and 2) hydrogen:

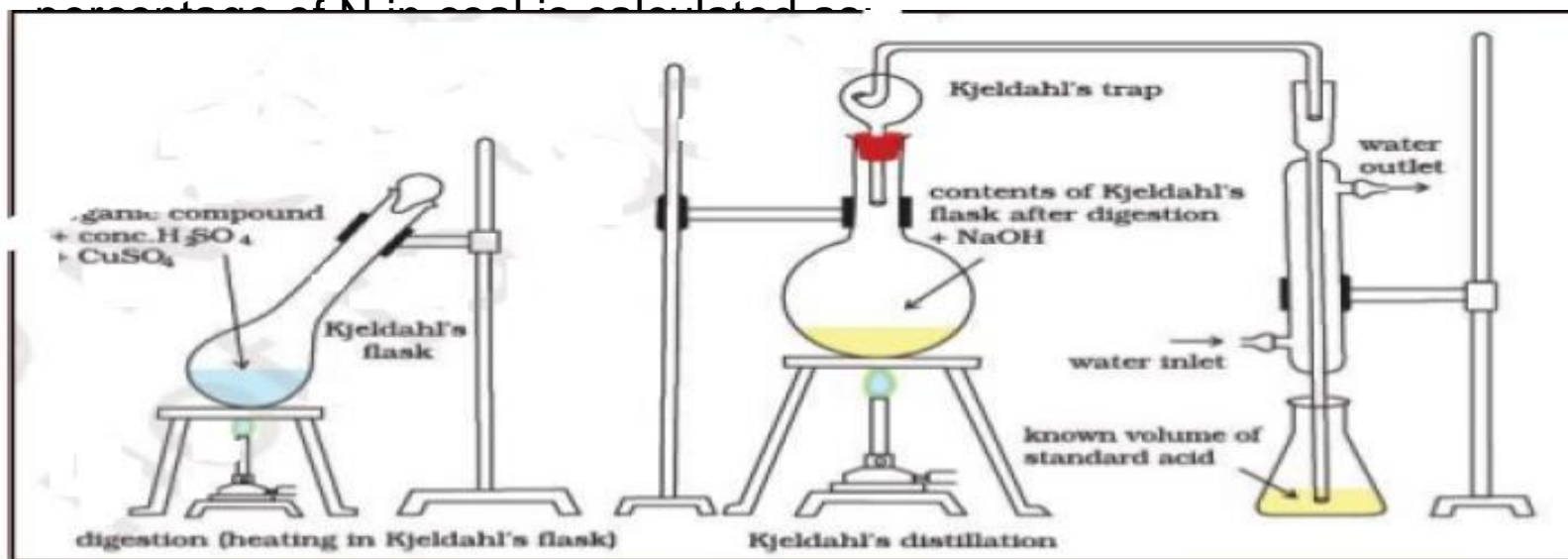
- About 1-2 g of accurately weighed coal sample is burnt in a current of oxygen in a combustion apparatus. C and H of the coal are converted into CO_2 and H_2O respectively. The gaseous products of combustion are absorbed respectively in KOH and CaCl_2 tubes of known weights.
- The increase in weights of these are then determined.
- $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$
- $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- $\text{KOH} + \text{CO}_2 \rightarrow \text{K}_2\text{CO}_3 + \text{H}_2\text{O}$
- $\text{CaCl}_2 + 7\text{H}_2\text{O} \rightarrow \text{CaCl}_2 \cdot 7\text{H}_2\text{O}$



$$C = \frac{\text{Increase in mass of KOH} \times 12 \times 100}{\text{Weight of coal sample taken} \times 44}$$

$$H = \frac{\text{Increase in weight of CaCl}_2 \text{ tube} \times 2 \times 100}{\text{Weight of coal sample} \times 18}$$

3) Nitrogen by Kjeldahl's Method: About 1 g of accurately weighed powdered 'coal is heated with concentrated H_2SO_4 along-with K_2SO_4 (catalyst) in a long-necked flask(called Kjeldahl's flask). After the solution becomes clear, it is treated with excess of KOH and the liberated ammonia is distilled over and absorbed in a known volume of standard acid solution. The unused acid is then determined by back titration with standard KOH solution. From the volume of acid used by ammonia liberated, the



$$\text{Percentage of N} = \frac{\text{Volume of acid used} \times \text{Normality} \times 1.4}{\text{Weight of coal taken}}$$

Sulphur is determined from the washings obtained from the known mass of coal, used in a bomb calorimeter for determination of a calorific value. During this determination, S is converted into sulphate. The washings are treated with barium chloride solution, when barium sulphate is precipitated. This precipitate is filtered, washed and heated to constant weight.

$$\text{Percentage of S} = \frac{\text{Weight of BaSO}_4 \text{ residue obtained} \times 32 \times 100}{\text{Weight of coal sample taken in bomb} \times 233}$$

Ash determination is carried out as in proximate analysis.

(5) Oxygen: It is obtained by difference. Percentage of O = 100 - Percentage of (C + H + S + N + ash)

Significance of ultimate analysis:

(1) Carbon and hydrogen: Greater the percentage of carbon and hydrogen better is the coal in quality and calorific value. However, hydrogen is mostly associated with the volatile matter and hence, it affects the use to which the coal is put. Also higher percentage of carbon in coal reduces the size of combustion chamber required. The amount of carbon, the major combustible constituent of coal, depends on the type of coal and its percentage increases with rank from lignite to anthracite. Thus, percentage of carbon forms the basis of classification of coal

(2) Nitrogen has no calorific value and hence, its presence in coal is undesirable; thus, a good quality coal should have very little nitrogen content.

(3) Sulphur, although contributes to the heating value of coal, yet on combustion produces acids (SO_2 and SO_3), which have harmful effects of corroding the equipments and also cause atmospheric pollution. Sulphur is, usually, present to the extent of 0.5 to 3.0% and derived from ores like iron pyrites, gypsum, etc., mines along-with the coal. Presence of sulphur is highly undesirable in coal to be, used for making coke for iron industry, since it is transferred to the iron metal and badly affects the quality and properties of steel. Moreover, oxides of sulphur (formed as combustion products) pollute the atmosphere and leads to corrosion.

(4) Oxygen content decreases the calorific value of coal. High oxygen-content coals are characterized by high inherent moisture, low calorific value, and low coking power. Moreover, oxygen is in combined form with hydrogen in coal and thus, hydrogen available for combustion is lesser than actual one. An increase in 1% oxygen content decreases the calorific value by about 1.7% and hence, oxygen is undesirable. Thus, a good quality coal should have low percentage of oxygen.

Numerical 1: 0.2g of Coal Sample is accurately weighed and is burnt in a combustion apparatus. The gaseous products of combustion are absorbed in potash bulb and calcium chloride tubes of known weights. The increase in weight of potash bulb and CaCl_2 tube are 0.66g and 0.08g respectively. Calculate the % C and % H in coal sample.

Numerical 2: 3g of coal was heated in Kjeldahls flask and NH_3 gas evolved was absorbed in 40mL of 0.5N HCl. After absorption the excess acid required 18.5mL of 0.5N KOH for exact neutralization. 2.3g of coal sample in quantitative analysis gave 0.35g BaSO_4 . Calculate % of N and S in coal sample.

Soln: Volume of NH_3 consumed by 0.5N HCl is in terms of 0.5N KOH

40ml of 0.5N HCl = 40mL of 0.5N KOH---assumed---V1___blank

Unreacted acid = 18.5ml of 0.5N KOH----V2___back titration

Vol of acid reacted with absorbed NH_3 in term of 0.5N KOH = $40 - 18.5\text{ml} = 21.5\text{mL}$

% N = $\frac{\text{volume of acid consumed}(V1-V2) \times N \text{ KOH} \times 1.4}{\text{weight of coal sample}}$

% S = $\frac{\text{Wt of BaSO}_4 \text{ formed} \times 32 \times 100}{\text{wt of coal sample} \times 233}$

Numerical 3: 2.5 g of coal was heated in Kjeldahls flask and NH_3 gas evolved was absorbed in 40mL of 0.5 N HCl. After absorption the excess acid required 8.5mL of 0.25N KOH for exact neutralization.

Soln:

Blank titration = 40mL of 0.5N HCl = 40mL of 0.5N KOH

Back titration = 8.5mL of 0.25N KOH = __4.25__ml of 0.5N KOH---- $n_1v_1=n_2v_2$

Liquid Fuels These are naturally found under the sea surface. Liquid fuels find extensive use in domestic and industrial fields. Petroleum or crude oil is a dark greenish brown or black coloured viscous oil found deep in earth's crust. The oil is usually floating over a brine solution and above the oil, natural gas is present. Crude oil containing mixture of paraffinic, olefinic and aromatic hydrocarbons with minor amounts of organic compounds like N, O and S. The average composition of crude oil is C = 80 - 87 %, H = 11-15%, S = 0.1 -3.5%, (N + O) = 0.1- 2.5%.

Classification of petroleum Petroleum is classified into three types based on variation of chemical nature of crude oil found in the earth.

- i) Paraffinic-base type crude oil: It contains saturated hydrocarbons from CH_4 to $\text{C}_{35}\text{H}_{72}$ and little amount of naphthalenes and aromatics.
- ii) Asphaltic-base type crude oil: It contains mainly cycloparaffins or naphthalenes with smaller amount of paraffins and aromatic hydrocarbons.
- iii) Mixed-base type crude oil: It contains both paraffinic and asphaltic hydrocarbons and are generally in the form of semi-solid waxes.