

Quiz 3.1 – Ionic Compounds

Name: Kory

Question 1

Predict which ion each element will form, with the correct charge:

Li	O	Sr	Al	Cl	P	K	I	Mg	Ba
$1+$	$2-$	$2+$	$3+$	$1-$	$3-$	$1+$	$1-$	$2+$	$2+$

Question 2

For each pair of elements, circle the one with higher ionization energy:

K and Na	N and O	Rb and Ca	Li and Be	Br and Cl	Sr and Ca
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Question 3

For each pair of elements, circle the one with higher magnitude electron affinity:

Rb and Cs	F and O	K and Mg	Br and Cl	Cl and S	Li and Na
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Question 4

Give the electronic configuration for each of these ions (Hint: Noble gas notation makes these *very* short)

S^{2-}	Ca^{2+}	Br^-	Ti^{4+}	N^{3-}	Na^+
$[Ar]$	$[Ar]$	$[Kr]$	$[Ar]$	$[Ne]$	$[Ne]$

Question 5

Give the name for each of these ions

Se^{2-}	Be^{2+}	F^-	Ti^{4+}	N^{3-}	Fe^{2+}
Selenide ion	Beryllium Ion	Fluoride ion	Titanium(IV) ion	Nitride ion	Iron(II) ion

Question 6

Give the symbol for each of these ions

Oxide Ion	Potassium Ion	Chromium(III) Ion	Chloride Ion	Magnesium Ion	Vanadium(IV) Ion
O^{2-}	K^+	Cr^{3+}	Cl^-	Mg^{2+}	V^{4+}