Quiz 6.2 – Limiting Reactants and Percent Yield

Name: _____

Question 1

$2.50\ g$ of H_2 and $18.2\ g$ of O_2 react according to the equation: $2\,H_2(g) + O_2(g) \longrightarrow 2\,H_2O(g)$

- Which reactant is the *limiting reactant*
- How many g of water are produced?
- How many g of the excess reactant remain?
- If $15.0\ g$ of water are actually recovered, what is the % yield?

Question 2

$5.00\ g$ of CH_4 and $20.0\ g$ of O_2 react according to the equation: $CH_4(g) + 2\,O_2(g) \longrightarrow CO_2(g) + 2\,H_2O(g)$

- Which reactant is the *limiting reactant*
- How many g of water and carbon dioxide are produced?
- How many g of the excess reactant remain?
- If $10.5\ g$ of water are actually recovered, what is the % yield?

Slaveships

By Lucille Clifton

loaded like spoons
into the belly of Jesus
where we lay for weeks for months
in the sweat and stink
of our own breathing
Jesus
why do you not protect us
chained to the heart of the Angel
where the prayers we never tell
and hot and red
as our bloody ankles
Jesus
Angel
can these be men
who vomit us out from ships
called Jesus Angel Grace of God
onto a heathen country
Jesus
Angel
ever again
can this tongue speak
can these bones walk
Grace Of God
can this sin live