

## Quiz 6.2 – Limiting Reactants and Percent Yield

Name: \_\_\_\_\_

**Question 1**

2.50 g of  $\text{H}_2$  and 18.2 g of  $\text{O}_2$  react according to the equation:  $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \longrightarrow 2 \text{H}_2\text{O}(\text{g})$

- Which reactant is the *limiting reactant*
- How many g of water are produced?
- How many g of the excess reactant remain?
- If 15.0 g of water are actually recovered, what is the % yield?

**Question 2**

5.00 g of  $\text{CH}_4$  and 20.0 g of  $\text{O}_2$  react according to the equation:  $\text{CH}_4(\text{g}) + 2 \text{O}_2(\text{g}) \longrightarrow \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{g})$

- Which reactant is the *limiting reactant*
- How many g of water and carbon dioxide are produced?
- How many g of the excess reactant remain?
- If 10.5 g of water are actually recovered, what is the % yield?

## *Slaverships*

By Lucille Clifton

loaded like spoons  
into the belly of Jesus  
where we lay for weeks for months  
in the sweat and stink  
of our own breathing  
Jesus  
why do you not protect us  
chained to the heart of the Angel  
where the prayers we never tell  
and hot and red  
as our bloody ankles  
Jesus  
Angel  
can these be men  
who vomit us out from ships  
called Jesus Angel Grace of God  
onto a heathen country  
Jesus  
Angel  
ever again  
can this tongue speak  
can these bones walk  
Grace Of God  
can this sin live