

Quiz 3.1 – Ionic Compounds

Name: Kery

Question 1

Predict which ion each element will form, with the correct charge:

Li	O	Sr	Al	Cl	P	K	I	Mg	Ba
1+	2-	2+	3+	1-	3-	1+	1-	2+	2+

Question 2

For each pair of elements, circle the one with higher ionization energy:

K and Na	N and O	Rb and Ca	Li and Be	Br and Cl	Sr and Ca
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Question 3

For each pair of elements, circle the one with higher magnitude electron affinity:

Rb and Cs	F and O	K and Mg	Br and Cl	Cl and S	Li and Na
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Question 4

Give the electronic configuration for each of these ions (Hint: Noble gas notation makes these very short)

S ²⁻	Ca ²⁺	Br ⁻	Ti ⁴⁺	N ³⁻	Na ⁺
[Ar]	[Ar]	[Kr]	[Ar]	[Ne]	[Ne]

Question 5

Give the name for each of these ions

Se ²⁻	Be ²⁺	F ⁻	Ti ⁴⁺	N ³⁻	Fe ²⁺
Selenide ion	Beryllium ion	Fluoride ion	Titanium(IV) ion	Nitride ion	Iron(II) ion

Question 6

Give the symbol for each of these ions

Oxide Ion	Potassium Ion	Chromium(III) Ion	Chloride Ion	Magnesium Ion	Vanadium(IV) Ion
O ²⁻	K ⁺	Cr ³⁺	Cl ⁻	Mg ²⁺	V ⁴⁺