Quiz 5.3 - Redox Reactions

Name: Key

Question 1

Identify the oxidation state (oxidation number) for each element in the following compounds:

Question 2

For each reaction, identify the oxidizing and reducing agents

Zn(s) + 2 HCl(aq)
$$\longrightarrow$$
 ZnCl₂(aq) + H₂(g)

Reducing Agent

4 Sb(s) + 3 O₂(g) \longrightarrow Sb₄O₆(s)

Oxidizing Agent

Fe₂O₃(s) + 3 H₂(g) \longrightarrow 2 Fe(s) + 3 H₂O(l)

Reducing Agent

Oxidizing Agent

Oxidizing Agent

Oxidizing Agent

Reducing Agent

Reducing Agent

Reducing Agent

N₂(g) + 8 O₂(g) \longrightarrow 5 CO₂(g) + 6 H₂O(g)

Oxidizing Agent

Reducing Agent

N₂(g) + 3 H₂(g) \longrightarrow 2 NH₃(g)

Reducing Agent

Oxidizing Agent

Oxidizing Agent