Quiz 4.2 - Precipitation Reactions

Name: Key

Question 1

Classify each ionic compound as soluble or insoluble

Na₂SO₄ Soluble

AgBr insoluble

Fe₃(PO₄)₂ insoluble

Ca(CH₃CO₂)₂ Soluble

Question 2

Circle which one of the following is a strong electrolyte

 $NH_2(aq)$ $H_2C_2O_4(aq)$

 $C_6H_{12}O_6(aq)$

HNO₃(aq)

HCH₃CO₂(aq)

Question 3

Give the net ionic equation for each reaction:

Reaction of nickel(II) nitrate with potassium carbonate

$$N_i(NO_3)_2(oq) + K_2(O_3(oq)) \rightarrow N_i(O_3(s)) + 2 KNO_3(oq)$$

$$N_i^{2+}(oq) + (O_3^{2-}(oq)) \rightarrow N_i(O_3(s))$$

 $\circ\;$ Reaction of barium sulfide with magnesium sulfate

Bo S (00) + Mg 504 (00) -> Bo 504 (5) + Mg 5 (5)