Quiz 13.2 – Working With K and Q

Name:

For questions 1–3, consider the reaction: $H(g) + Br(g) \implies HBr(g)$

Question 1

Calculate K from the following reactions:

$$H_2(g) \rightleftharpoons 2H(g)$$
 $K = 11.8$

$$\mathrm{Br_2}(\mathrm{g}) \Longrightarrow \mathrm{2}\,\mathrm{Br}(\mathrm{g}) \quad K = 1.15 \times 10^{-5}$$

$$Br_2(g) + H_2(g) \Longrightarrow 2 HBr(g)$$
 $K = 2.78 \times 10^3$

Question 2

If the system has reached equilibrium with [HBr] = 0.025 M and [H] = 0.0015 M, find [Br]

Question 3

Find the reaction quotient Q and predict which direction the reaction must shift to reach equilibrium if $[{\rm HBr}]=0.0035~M$, $[{\rm H}]=0.020~M$, and $[{\rm Br}]=0.0025~M$

Question 4

The reaction $H_2(g) + I_2(g) \implies$ 2 HI(g) is at equilibrium with $[H_2] = 0.05 \, M$, $[I_2] = 0.05 \, M$, and $[HI] = 0.90 \, M$. Calculate K from these values

Bonus Question!

Comparing the K values for the production of HBr and HI from their elemental constituents, what can you say about the relative stability of HBr and HI?

The New Colossus

By Emma Lazarus

Not like the brazen giant of Greek fame,
With conquering limbs astride from land to land;
Here at our sea-washed, sunset gates shall stand
A mighty woman with a torch, whose flame
Is the imprisoned lightning, and her name
Mother of Exiles. From her beacon-hand
Glows world-wide welcome; her mild eyes command
The air-bridged harbor that twin cities frame.
"Keep, ancient lands, your storied pomp!" cries she
With silent lips. "Give me your tired, your poor,
Your huddled masses yearning to breathe free,
The wretched refuse of your teeming shore.
Send these, the homeless, tempest-tost to me,
I lift my lamp beside the golden door!"