

Name: \_\_\_\_\_

## Exam 3 Equations (Chapters 14-15)

$$pH = -\log[\text{H}_3\text{O}^+]$$

$$[\text{H}_3\text{O}^+] = 10^{-pH}$$

$$pOH = -\log[\text{OH}^-]$$

$$[\text{OH}^-] = 10^{-pOH}$$

$$pH + pOH = 14$$

$$[\text{H}_3\text{O}^+][\text{OH}^-] = K_w = 1.00 \times 10^{-14}$$

$$K_a K_b = K_w$$

$$pH = pK_a + \log \frac{B}{A}$$

$$pK_a = -\log K_a$$

$$\frac{C_A V_A}{\nu_A} = \frac{C_T V_T}{\nu_T}$$