## Quiz 19.1 – Balancing Redox Reactions

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## Question 1

$$2 \text{ Hzo + Hzoz + 2 (loz - + 0z + 2 \text{ Hzo + } )}$$

$$\circ \text{ Fe}^{2+}(\text{aq}) + \text{MnO}_{4}^{-}(\text{aq}) \xrightarrow{\text{Fe}^{3+}(\text{aq}) + \text{Mn}^{2+}(\text{aq})}$$

Balance the same reactions in basic solution:

## Question 3

 $25.00\,ml$  of  $\rm Fe(NO_3)_2$  solution are titrated with  $13.65\,ml$  of acidic  $0.103\,M$  KMnO $_4$  solution. Find the original [Fe(NO<sub>2</sub>)<sub>2</sub>]