

## Quiz 3.2 – Spontaneity

Name: \_\_\_\_\_

**A Spontaneous Process**Consider the reaction:  $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \longrightarrow 2\text{H}_2\text{O}(\text{g})$ 

At 298 K, this reaction has  $\Delta H^\ominus = -483.64 \frac{\text{kJ}}{\text{mol}}$  and  $\Delta S^\ominus = -88.846 \frac{\text{J}}{\text{mol K}}$ . Use these data to find  $\Delta S_{\text{sys}}$ ,  $\Delta S_{\text{surr}}$ , and  $\Delta S_{\text{univ}}$  for this reaction, and show that it is spontaneous under standard conditions

**Entropy on Astronomical Scales**

The sun radiates heat from its surface at a temperature of 5,778 K. Some of this heat reaches the earth, where it is absorbed, and re-emitted at the average global temperature of about 15°C. The Earth's heat radiates away into space, which acts as a heat sink at 2.7 K (The temperature of the cosmic microwave background). The Earth's temperature is fairly constant, so we can assume the heat absorbed from the sun and the heat emitted into space are balanced. For a collection of photons,  $S = \frac{E}{T}$ , where  $T$  is the characteristic Black-body temperature based on the spectral profile of the photons. Find  $\Delta S_{\text{univ}}$  for each 1000 J of energy absorbed then re-emitted by the Earth.

**Conditions for spontaneity**

In class we saw how  $\Delta H$  and  $\Delta S$  can determine the conditions under which a reaction will be spontaneous. However, that discussion only applied to isobaric processes. Fill out the table below to indicate the conditions under which isochoric processes are spontaneous

		$\Delta S$	
		+	−
$\Delta U$	+		
	−		

**A Reversible Process**

An isobaric phase change at the phase change temperature is a great simple example of a reversible process. A pot filled with 400.0 g of water are boiled on a stove. In this scenario, we can consider the surroundings to be at the boiling temperature.

Find  $\Delta S_{sys}$ ,  $\Delta S_{surr}$ , and  $\Delta S_{univ}$  for this process

**A Non-Spontaneous Process. . . which isn't (or, rather, which really *is* spontaneous)**

A pot filled with 400.0 g left out 298 K may evaporate into the vapor phase given enough time. Find  $\Delta S_{sys}$ ,  $\Delta S_{surr}$ , and  $\Delta S_{univ}$  for this phase change and show that it is non-spontaneous. (Remember that  $\Delta H$  must be corrected by Kirchoff's law)

And yet, you *know* that such a process will happen spontaneously in nature. . . at least, in Utah it will. What factor have we neglected, and how can we correct our calculations to include it?

七步詩 (*The Quatrain of Seven Steps*)

By 曹植 (Cao Zhi)

煮豆燃豆其  
漉菽以為汁  
其在釜下燃  
豆在釜中泣  
本是同根生  
相煎何太急

People burn the beanstalk to boil beans,  
filtering them to extract juice.  
The beanstalks were burnt under the cauldron,  
and the beans in the cauldron wailed:  
“We were originally grown from the same root;  
Why should we hound each other to death with such impatience?”