# **Chem 1A Quiz & Exam Information Page**

NOTE: A copy of this information page will be provided to you for quizzes and exams. I will have a copy stapled to the front of every quiz and exam you take this quarter.

#### Periodic Table:

•	٠.	iouic	Tabi	С.															
										_	1A							_	8A
											1								2
											H							7	He
		1A	2A		1.008 3A 4A 5A						6A	A	4.003						
		3	4											5	6	7	8	9	10
	2	Li	Be											В	C	N	O	F	Ne
		6.941	9.012											10.81	12.01	14.01	16.00	19.00	20.18
		11	12											13	14	15	16	17	18
	3	Na	Mg											Al	Si	P	$\mathbf{S}$	Cl	Ar
		22.99	24.30	3	4	5	6	7	8	9	10	11	12	26.98	28.09	30.97	32.07	35.45	39.95
		19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36
<b>—</b>	4	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Period		39.10	40.08	44.96	47.88	50.94	52.00	54.94	55.85	58.93	58.69	63.55	65.39	69.72	72.61	74.92	78.96	79.90	83.80
E.		37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54
	5	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
		85.47	87.62	88.91	91.22	92.91	95.94	(98)	101.1	102.9	106.4	107.9	112.4	114.8	118.7	121.8	127.6	126.9	131.3
		55	56		72	73	74	75	76	77	78	79	80	81	82	83	84	85	86
	6	Cs		La-Lu	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
		132.9	137.3		178.5	180.9	183.8	186.2	190.2	192.2	195.1	197.0	200.6	204.4	207.2	209.0	(209)	(210)	(222)
		87	88		104	105	106	107	108	109	110	111	112		114		116		118
	7	Fr	Ra	Ac-Lr	Rf	Db	Sg	Bh	Hs	Mt	Uun	Uuu	Uub		Uuq		Uuh		Uuo
		(223)	(226)		(261)	(262)	(263)	(264)	(265)	(268)	(269)	(272)	(269)						
				1									1						1
	s block d block p block							block											
	57   58   59   60   61   62   63   64   65   66   67   68   69									70	71								
Lanthanides			La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dv	Ho	Er	Tm	Yb	Lu		
			138.9	140.1	140.9	144.2	(145)	150.4	152.0	157.2	158.9	162.5	164.9	167.3	168.9	173.0	175.0		
ľ			89	90	91	92	93	94	95	96	97	98	99	100	101	102	103		
Actinides			Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr		
					238.0	(237)	(244)	(243)	(247)		(251)	(252)	(257)		(259)	(262)			
		f block																	

Ion	Name	Ion	Name	<u>Prefix</u>	Number Indicated
$\mathrm{NH_4}^+$	ammonium	$CO_3^{2-}$	carbonate	mono-	1
$\mathrm{NO}_2^-$	nitrite	$HCO_3^-$	hydrogen carbonate	di-	2
$NO_3^-$	nitrate		(or bicarbonate)	tri-	3
$SO_3^{2-}$	sulfite	$ClO^-$	hypochlorite	tetra-	4
$\mathrm{SO_4}^{2-}$	sulfate	$\text{ClO}_2^-$	chlorite	penta-	5
$\mathrm{HSO_4}^-$	hydrogen sulfate	$\text{ClO}_3^-$	chlorate	hexa-	6
	(or bisulfate)	$\text{ClO}_4^-$	perchlorate	hepta-	7
$OH^-$	hydroxide	$C_2H_3O_2^-$	acetate	octa-	8
		(or CH <sub>3</sub> C	OO <sup>-</sup> )		
$CN^-$	cyanide	$\mathrm{MnO_4}^-$	permanganate		
$SCN^-$	thiocyanide	$\operatorname{Cr_2O_7}^{2-}$	dichromate		
$PO_4^{3-}$	phosphate	$\text{CrO}_4^{\ 2-}$	chromate		
$\mathrm{HPO_4}^{2-}$	hydrogen phosphate	$C_2O_4^{2-}$	oxalate		
$\mathrm{H_2PO_4}^-$	dihydrogen phosphate	$\mathrm{O_2}^{2-}$	peroxide		
$IO_3^-$	iodate	$\mathrm{BrO_3}^-$	bromate		
$\mathrm{IO}_2^{-}$	iodite	$\mathrm{BrO}_2^{-}$	bromite		
$IO^-$	hypoiodite	$BrO^-$	hypobromite		

#### **Constants:**

$$\begin{split} N_{A} &= 6.022 \text{ x} 10^{23} \text{ mol}^{-1} \\ R &= 0.08206 \text{ L atm mol}^{-1} \text{K}^{-1} \\ R &= 8.314 \text{ J mol}^{-1} \text{K}^{-1} \\ \pi &= 3.14 \\ c &= 3 \text{ x} 10^8 \text{ m/s} \\ h &= 6.626 \text{ x} 10^{-34} \text{ J s} \\ \text{STP} &= 0^{\circ} \text{C}, 1 \text{ atm} \end{split}$$

#### **Conversion Factors:**

1 atm = 760 torr 1 J = 1 kg m<sup>2</sup> s<sup>-2</sup> Kelvin = Celsius + 273 1 nm =  $10^{-9}$  m

### **Equations:**

$$M = \frac{n}{v} \qquad M_1 V_1 = M_2 V_2 \qquad \text{PV} = \text{nRT} \qquad \frac{P_1 V_1}{n_1 T_1} = \frac{P_2 V_2}{n_2 T_2} \qquad \text{P}_1 = X_1 P_{\text{tot}} \qquad X_1 = \frac{n_1}{n_{\text{tot}}}$$

$$(KE)_{av} = \frac{3}{2} \text{RT} \qquad u_{av} = \overline{u} = \sqrt{\frac{8RT}{\pi M}} \qquad u_{rms} = \sqrt{\overline{u^2}} = \sqrt{\frac{3RT}{M}} \qquad Z = 4 \left(\frac{N}{V}\right) d^2 \sqrt{\frac{\pi RT}{M}}$$

$$E = h \nu \qquad \qquad E = h c / \lambda \qquad \qquad c = \nu \lambda \qquad \qquad \Delta E \ = \ - (2.18 \times 10^{-18} \, J) \left( \frac{Z^2}{n_{\text{final}}^2} \ - \ \frac{Z^2}{n_{\text{initial}}^2} \right) \label{eq:energy}$$

$$KE_{electron} = E_{photon} - E_{binding} \qquad \quad \lambda = h/mv \qquad KE = (1/2) \ mv^2$$

## Solubility of Ionic Compounds in Water:

Anion	Soluble*	Slightly Soluble*	Insoluble*		
NO <sub>3</sub> <sup>-</sup> (nitrate)	All	_	_		
CH <sub>3</sub> COO <sup>-</sup> (acetate)	Most	_	Be(CH <sub>3</sub> COO) <sub>2</sub>		
F - (fluoride)	Group I, AgF, BeF <sub>2</sub>	SrF <sub>2</sub> , BaF <sub>2</sub> , PbF <sub>2</sub>	MgF <sub>2</sub> , CaF <sub>2</sub>		
Cl - (chloride)	Most	PbCl <sub>2</sub>	AgCl, Hg <sub>2</sub> Cl <sub>2</sub>		
Br - (bromide)	Most	PbBr <sub>2</sub> , HgBr <sub>2</sub>	AgBr, $Hg_2Br_2$		
SO <sub>4</sub> <sup>2-</sup> (sulfate)	Most	CaSO <sub>4</sub> , Ag <sub>2</sub> SO <sub>4</sub>	BaSO <sub>4</sub> , SrSO <sub>4</sub> PbSO <sub>4</sub> , Hg <sub>2</sub> SO <sub>4</sub>		
S <sup>2-</sup> (sulfide)	Group I and II, (NH <sub>4</sub> ) <sub>2</sub> S	_	Most		
CO <sub>3</sub> <sup>2-</sup> (carbonate)	Group I, (NH <sub>4</sub> ) <sub>2</sub> CO <sub>3</sub>	_	Most		
SO <sub>3</sub> <sup>2-</sup> (sulfite)	Group I, (NH <sub>4</sub> ) <sub>2</sub> SO <sub>3</sub>	_	Most		
PO <sub>4</sub> <sup>3-</sup> (phosphate)	Group I, (NH <sub>4</sub> ) <sub>3</sub> PO <sub>4</sub>	_	Most		
OH - (hydroxide)	Group I, Ba(OH) <sub>2</sub> , NH <sub>4</sub> OH	Sr(OH) <sub>2</sub> , Ca(OH) <sub>2</sub>	Most		

<sup>\*</sup>Soluble: dissolves to the extent of > 10 g/L \*Slightly Soluble: 0.1 to 10 g/L \*Insoluble: < 0.1 g/L