

# Chem 1A Quiz & Exam Information Page

**NOTE: A copy of this information page will be provided to you for quizzes and exams. I will have a copy stapled to the front of every quiz and exam you take this quarter.**

Periodic Table:

														1A												8A	
														1 H 1.008												2 He 4.003	
		1A		2A												3A		4A		5A		6A		7 A			
Period	2	3 Li 6.941	4 Be 9.012											5 B 10.81	6 C 12.01	7 N 14.01	8 O 16.00	9 F 19.00	10 Ne 20.18								
	3	11 Na 22.99	12 Mg 24.30	3	4	5	6	7	8	9	10	11	12	13 Al 26.98	14 Si 28.09	15 P 30.97	16 S 32.07	17 Cl 35.45	18 Ar 39.95								
	4	19 K 39.10	20 Ca 40.08	21 Sc 44.96	22 Ti 47.88	23 V 50.94	24 Cr 52.00	25 Mn 54.94	26 Fe 55.85	27 Co 58.93	28 Ni 58.69	29 Cu 63.55	30 Zn 65.39	31 Ga 69.72	32 Ge 72.61	33 As 74.92	34 Se 78.96	35 Br 79.90	36 Kr 83.80								
	5	37 Rb 85.47	38 Sr 87.62	39 Y 88.91	40 Zr 91.22	41 Nb 92.91	42 Mo 95.94	43 Tc (98)	44 Ru 101.1	45 Rh 102.9	46 Pd 106.4	47 Ag 107.9	48 Cd 112.4	49 In 114.8	50 Sn 118.7	51 Sb 121.8	52 Te 127.6	53 I 126.9	54 Xe 131.3								
	6	55 Cs 132.9	56 Ba 137.3	La-Lu	72 Hf 178.5	73 Ta 180.9	74 W 183.8	75 Re 186.2	76 Os 190.2	77 Ir 192.2	78 Pt 195.1	79 Au 197.0	80 Hg 200.6	81 Tl 204.4	82 Pb 207.2	83 Bi 209.0	84 Po (209)	85 At (210)	86 Rn (222)								
	7	87 Fr (223)	88 Ra (226)	Ac-Lr	104 Rf (261)	105 Db (262)	106 Sg (263)	107 Bh (264)	108 Hs (265)	109 Mt (268)	110 Uun (269)	111 Uuu (272)	112 Uub (269)		114 Uuq		116 Uuh		118 Uuo								
	s block			d block										p block													
Lanthanides		57 La 138.9	58 Ce 140.1	59 Pr 140.9	60 Nd 144.2	61 Pm (145)	62 Sm 150.4	63 Eu 152.0	64 Gd 157.2	65 Tb 158.9	66 Dy 162.5	67 Ho 164.9	68 Er 167.3	69 Tm 168.9	70 Yb 173.0	71 Lu 175.0											
		89 Ac (227)	90 Th 232.0	91 Pa 231.0	92 U 238.0	93 Np (237)	94 Pu (244)	95 Am (243)	96 Cm (247)	97 Bk (247)	98 Cf (251)	99 Es (252)	100 Fm (257)	101 Md (258)	102 No (259)	103 Lr (262)											
Actinides																	f block										

**Constants:**

$$N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$$

$$R = 0.08206 \text{ L atm mol}^{-1}\text{K}^{-1}$$

$$R = 8.314 \text{ J mol}^{-1}\text{K}^{-1}$$

$$\pi = 3.14$$

$$c = 3 \times 10^8 \text{ m/s}$$

$$h = 6.626 \times 10^{-34} \text{ J s}$$

$$\text{STP} = 0^\circ\text{C}, 1 \text{ atm}$$

**Conversion Factors:**

$$1 \text{ atm} = 760 \text{ torr}$$

$$1 \text{ J} = 1 \text{ kg m}^2 \text{ s}^{-2}$$

$$\text{Kelvin} = \text{Celsius} + 273$$

$$1 \text{ nm} = 10^{-9} \text{ m}$$

**Equations:**

$$M = \frac{n}{V} \quad M_1 V_1 = M_2 V_2 \quad PV = nRT \quad \frac{P_1 V_1}{n_1 T_1} = \frac{P_2 V_2}{n_2 T_2} \quad P_1 = X_1 P_{\text{tot}} \quad X_1 = \frac{n_1}{n_{\text{tot}}}$$

$$(\text{KE})_{\text{av}} = \frac{3}{2}RT \quad u_{\text{av}} = \bar{u} = \sqrt{\frac{8RT}{\pi M}} \quad u_{\text{rms}} = \sqrt{\overline{u^2}} = \sqrt{\frac{3RT}{M}} \quad Z = 4 \left( \frac{N}{V} \right) d^2 \sqrt{\frac{\pi RT}{M}}$$

$$E = h\nu \quad E = hc/\lambda \quad c = \nu\lambda \quad \Delta E = -(2.18 \times 10^{-18} \text{ J}) \left( \frac{Z^2}{n_{\text{final}}^2} - \frac{Z^2}{n_{\text{initial}}^2} \right)$$

$$\text{KE}_{\text{electron}} = E_{\text{photon}} - E_{\text{binding}} \quad \lambda = h/mv \quad \text{KE} = (1/2) mv^2$$

**Solubility of Ionic Compounds in Water:**

Anion	Soluble*	Slightly Soluble*	Insoluble*
$\text{NO}_3^-$ (nitrate)	All	—	—
$\text{CH}_3\text{COO}^-$ (acetate)	Most	—	$\text{Be}(\text{CH}_3\text{COO})_2$
$\text{F}^-$ (fluoride)	Group I, $\text{AgF}$ , $\text{BeF}_2$	$\text{SrF}_2$ , $\text{BaF}_2$ , $\text{PbF}_2$	$\text{MgF}_2$ , $\text{CaF}_2$
$\text{Cl}^-$ (chloride)	Most	$\text{PbCl}_2$	$\text{AgCl}$ , $\text{Hg}_2\text{Cl}_2$
$\text{Br}^-$ (bromide)	Most	$\text{PbBr}_2$ , $\text{HgBr}_2$	$\text{AgBr}$ , $\text{Hg}_2\text{Br}_2$
$\text{SO}_4^{2-}$ (sulfate)	Most	$\text{CaSO}_4$ , $\text{Ag}_2\text{SO}_4$	$\text{BaSO}_4$ , $\text{SrSO}_4$ , $\text{PbSO}_4$ , $\text{Hg}_2\text{SO}_4$
$\text{S}^{2-}$ (sulfide)	Group I and II, $(\text{NH}_4)_2\text{S}$	—	Most
$\text{CO}_3^{2-}$ (carbonate)	Group I, $(\text{NH}_4)_2\text{CO}_3$	—	Most
$\text{SO}_3^{2-}$ (sulfite)	Group I, $(\text{NH}_4)_2\text{SO}_3$	—	Most
$\text{PO}_4^{3-}$ (phosphate)	Group I, $(\text{NH}_4)_3\text{PO}_4$	—	Most
$\text{OH}^-$ (hydroxide)	Group I, $\text{Ba}(\text{OH})_2$ , $\text{NH}_4\text{OH}$	$\text{Sr}(\text{OH})_2$ , $\text{Ca}(\text{OH})_2$	Most

\***Soluble:** dissolves to the extent of  $> 10 \text{ g/L}$     \***Slightly Soluble:**  $0.1$  to  $10 \text{ g/L}$     \***Insoluble:**  $< 0.1 \text{ g/L}$