

No.	Compound name	Reaction equation	pH	Rate constant (L mol <sup>-1</sup> s <sup>-1</sup> )	Comments	Reference
30	Chloride ion	$\cdot\text{OH} + \text{Cl}^- \longrightarrow \text{ClOH}^-$	~2	$4.3 \times 10^9$	p.r.; D.k. at 240 nm as well as p.b.k. at 340 nm (Cl <sub>2</sub> <sup>-</sup> ); <i>k</i> and <i>K</i> also given for ClOH <sup>-</sup> + H <sup>+</sup> ⇌ Cl <sup>·</sup> + H <sub>2</sub> O and Cl <sup>·</sup> + Cl <sup>-</sup> ⇌ Cl <sub>2</sub> <sup>-</sup> ; <i>K</i> <sub>eq</sub> = 0.70.	731039