

| No. | Compound name | Reaction equation | pH | Rate constant (L mol ⁻¹ s ⁻¹) |
|-----|-----------------|--|-------------------|---|
| 18 | Bicarbonate Ion | $\text{OH} + \text{HCO}_3^- \longrightarrow \text{CO}_3^{2-} + \text{H}_2\text{O}$ | | 8.5×10^6 |
| | | | 1.0×10^7 | p.r.; P.b.k. (studied over a range of carbonate and hydroxide ion concn. at 0-200 °C, press. |
| | | | 8.5×10^6 | p.r.; P.b.k. at 600 nm assuming $pK_a(\text{H}_2\text{CO}_3) = 6.34$ and 10.37 and $k(\text{OH} + \text{CO}_3^{2-}) = 4.$ |
| | | | 1×10^7 | p.r.; P.b.k. at 600 nm. |