No.	Compound name	Reaction equation	рН	Rate constant $(\operatorname{L}\operatorname{mol}^{-1}\operatorname{s}^{-1})$	Comments	Reference
162	Ozone	$.  \mathrm{OH} + \mathrm{O}_3 \longrightarrow \mathrm{HO}_2 \cdot  + \mathrm{O}_2$	10.3 1, 9-10	$1.1 \times 10^9$ $1 \times 10^8$ $1.1 \times 10^8$	Average of 2 values. p.r.; C.k.; rel. to $k(.\mathrm{OH} + \mathrm{CO_3}^{2-}).$ p.r.; P.b.k. at 430 nm $(\mathrm{O_3}^-)$ at pH 9-10, as well as d.k. at 280-310 nm $(\mathrm{O_3})$ at pH 1.	84A270 84A270