

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments	Reference
29	Selenite(IV) Ion	$\text{O}^{\cdot -} + \text{SeO}_3^{2-} \longrightarrow \text{HSeO}_4^{2-} + \text{OH}^-$	13	1.1×10^7	p.r.; P.b.k. at 420 nm; calcd. from $k_{\text{obs}} = 0.60 \times 10^6$ (cor. for $\cdot\text{OH}$ reaction) using $K(\text{O}^{\cdot -} + \text{H}_2\text{O} \rightleftharpoons \cdot\text{OH} + \text{OH}^-) = 8 \times 10^{-3}$.	86A335