No.	Compound name	Reaction equation	pH	Rate constant $(\mathrm{L\,mol^{-1}\,s^{-1}})$	Comments	Reference
5	Bromine	$\mathrm{H^{\cdot+} Br_2 \longrightarrow Br_2^{\cdot-} + H^+}$	1	$\sim 1 \times 10^{10}$	p.r.; P.b.k. at 360 nm ($\mathrm{Br_2^-}$), competes favorably with H + $\mathrm{O_2}$; uncertainty in the value is 5-fold.	650382