No.	Compound name	Reaction equation	pH	Rate constant ($L\,mol^{-1}\,s^{-1}$)
3.2	O_2ˆ−	$O_2^- + .OH \longrightarrow HO_2^-$	>12	$\leq 2 \times 10^{10}$
3.3	O_2ˆ−	$O_2^- + .OH \longrightarrow OH^- + O_2$	7	7×10^9
		2.74–6.75	9.4×10^9	p.r.; Obs. $G(H_2O_2)$; data fitting; $pK(HO_2; H^+ + O_2^-) = 4.45$; rel. to $k(\cdot OH + \cdot OH)$.
3.4	HO_2 ·	$HO_2^\bullet + .OH \longrightarrow H_2O + O_2$	0.46–6.75	6.6×10^9
		>2	6×10^9	p.r.; Calcd. from obs. $G(H_2O_2)$.; rel. to $k(\cdot OH + \cdot OH)$.
3.5	H_2O_2ˆ+	$H_2O_2^+ + .OH \longrightarrow H_3O^+ + O_2$	0.46–1.51	1.2×10^{10}