No.	Compound name	Reaction equation	pH	Rate constant $(\mathrm{L\,mol^{-1}\,s^{-1}})$	Comments	Reference
178	Tetrachloropalladate(II) Ion	$.\,\mathrm{OH} + \mathrm{PdCl_4}^{2-} \longrightarrow \mathrm{OH^-} + \mathrm{PdCl_4}^{-}$			6.8×10^9	Average of 2 val
				6.3×10^9	p.r.	76A249
				7.2×10^9	p.r.; C.k. in 0.01 mol $\mathrm{L^{-1}}$ NaCl; $k = 1.2 \times 10^{10}$ in 1 mol $\mathrm{L^{-1}}$ NaCl. $\varepsilon(320$ nm$) = 4700$ L $\mathrm{mol^{-1}}$ $\mathrm{cm^{-1}}$ for product; may be $\mathrm{Pd(OH)Cl_4^{2-}}$; rel. to $k(.\mathrm{OH} +$ tert-BuOH).	741087