

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)
172	Acetate Ion	$\text{H}^\cdot + \text{CH}_3\text{CO}_2^- \longrightarrow \text{H}_2 + \cdot\text{CH}_2\text{CO}_2^-$		3.5×10^5
		7	3.9×10^5	e-r.; esr; Decay of spin polarization, compared with EtOH. in phosphate buffered soln.; rel.
		7	3.5×10^5	-r.; C.k.; rel. to $k(\text{H}^\cdot + \text{DCO}_2^-)$.
		~ 8	3.2×10^5	X-r.; C.k.; rel. to $k(\text{H}^\cdot + \text{NO}_2^-)$.