No.	Compound name	Reaction equation	pH	Rate constant ($L\,mol^{-1}\,s^{-1}$)
230	Hydrogen Ion	$e_{aq}^- + H^+ \longrightarrow H^\cdot$		2.8×10^{10}
		2	2.3×10^{10}	p.r.; D.k. at 600 nm.; k cor. for I.
		2	2.2×10^{10}	p.r.; D.k. in $HClO_4$ (10^{-5}-10^{-1} mol L^{-1}).
		2	2.2×10^{10}	p.r.; D.k. at 600 nm; k increases with pressure $\rightarrow$ 3.5 kbar, then remains constant $\rightarrow$ 6.4
		acid	2.2×10^{10}	p.r.; D.k.; k(1.6 to 2.2) $\times 10^{10}$ varied with concn. of added EtOH, KI, and $MgCl_2$.
		4.1-4.7	2.3×10^{10}	p.r.; D.k. at 577 nm, soln. contains H_2SO_4 or $HClO_4$.
		4-5	2.4×10^{10}	p.r.; D.k. at 578 nm ($HClO_4$).