

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)
67	Copper(II) Ion	$\cdot\text{OH} + \text{Cu}^{2+} \longrightarrow \text{CuOH}^{2+}$		3.5×10^8
			3.1×10^8	p.r.; P.b.k. at 300 nm.
3-6			3.1×10^8	p.r.; P.b.k. at 300 nm.
~5			4.0×10^8	p.r.; C.k. (condy. study); assume $k(\cdot\text{OH} + \text{MeOH})/k(\cdot\text{OH} + \text{tert-BuOH}) = 2$ and $k'(\cdot\text{OH} + \text{EtOH})$.
7			3.5×10^8	p.r.; P.b.k. at 313 nm.
			3.7×10^8	p.r.; C.k.; obs. Cu(III) at 313 nm; rel. to $k'(\cdot\text{OH} + \text{EtOH})$.
			3.7×10^8	p.r.; C.k.; obs. Cu(III) at 313 nm; rel. to $k'(\cdot\text{OH} + \text{MeOH})$.