No.	Compound name	Reaction equation	pH	Rate constant $(\mathrm{L\,mol^{-1}\,s^{-1}})$	Comments	Reference
166	Zine(I) ion	$H^{\bullet} + Zn^{+} \longrightarrow ZnH^{+}$	~ 3.7	1.9×10^9	p.r.; D.k. (Zn^+) in presence of 107 atm. H_2, 10^{-2} mol L^{-1} $ZnSO_4$; assuming $k(H + H) = 1 \times 10^{10}$.	771011