No. Compound name	Reaction equation	pH Rate constant $(L \operatorname{mol}^{-1} \operatorname{s}^{-1})$	Comments Reference
167 Zinc(II) Ion	$H^{\bullet+} Zn^{2+} \longrightarrow$	$\sim 7 < 3 \times 10^5$	γ -r.; No reaction; c.k. in soln. contg. $1-5\times 10^{-2}$ mol 650192 L ⁻¹ ZnSO ₄ and 10^{-2} mol L ⁻¹ EtOH; obs. G(H ₂); rel. to $k(H^{\bullet} + EtOH)$.