

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments	Reference
11	Hypobromite ion	$\cdot\text{OH} + \text{BrO}^- \longrightarrow \text{OH}^- + \text{BrO}$	11-13	4.2×10^9	p.r.; C.k.; assumed $\text{p}K_a(\text{OH}) = 11.9$; rel. to $k(\cdot\text{OH} + \text{CO}_3^{2-})$.	680153