No.	Compound name	Reaction equation	pH	Rate constant $(\text{L mol}^{-1}\,\text{s}^{-1})$	Comments	Reference
104	Hydrazinium ion	$\text{H·} + \text{H}_2\text{NNH}_3{}^+ \longrightarrow \text{H}^+ + \text{H}_2 + \text{·NHNH}_2$	~ 0.7	2.5×10^4	γ-r.; C.k.; obs. $\text{G(H}_2)$; at pH 6 $k = 1.2 \times 10^5$; $pK_a = 7.9$; rel. to $k(\text{H·} + \text{H}_2\text{O}_2)$.	690598