

No.	Compound name	Reaction equation	pH	Rate constant (L mol <sup>-1</sup> s <sup>-1</sup> )	Comments	Reference
65	Cobalt(II) Ions	$\text{e}_{\text{aq}}^- + \text{Co}^{2+} \longrightarrow \text{Co}^+$		$1.3 \times 10^{10}$	p.r.; D.k. at 578 nm; deaerated soln. contg. $10^{-2}$ mol L <sup>-1</sup> MeOH; counterion $\text{SO}_4^{2-}$ .	761136
				$6.6 \times 10^9$	p.r.; D.k. at 600 nm; counterion $\text{SO}_4^{2-}$ ; $I = 3.3 \times 10^{-3}$ .	751027
			4.0	$9.5 \times 10^9$	p.r.; D.k. at 550 nm in soln. contg. $10^{-3}$ mol L <sup>-1</sup> EtOH and 1 mol L <sup>-1</sup> $\text{Co}(\text{ClO}_4)_2$ ; $k$ cor. for I.	700242
				$1.2 \times 10^{10}$	p.r.; D.k.; $k$ detd. at 19.5 and 73 °C.	650044