

No.	Compound name	Reaction equation	pH	Rate constant (L mol <sup>-1</sup> s <sup>-1</sup> )
161	Hydroxide Ion	$\cdot\text{OH} + \text{OH}^- \longrightarrow \text{O}^{\bullet-} + \text{H}_2\text{O}$		$1.2 \times 10^{10}$
			$1.3 \times 10^{10}$	p.r.; C.k. with MeOH and EtOH; calcd. from $k(\text{reverse}) = 9.6 \times 10^7 \text{ s}^{-1}$ and $\text{p}K(\cdot\text{OH})=$
		11	$1.2 \times 10^{10}$	p.r.; C.k. with MeOH and EtOH; soln. contains $\text{CO}_3^{2-}$ and $\text{HCO}_3^-$ ; calcd. from $k(\text{rever}$