No.	Compound name	Reaction equation	pH	Rate constant ($L\,mol^{-1}\,s^{-1}$)	Comments	Reference
420	Thiosulfate ion	$e_{aq}{}^{-} + S_2O_3{}^{2-} \longrightarrow SO_3{}^{2-} + \cdot S^{-}$	6.5	$\sim 1.5 \times 10^8$	p.r.; D.k. at 700 nm in soln. contg. 10^{-2} and 10^{-1} mol L^{-1} $Na_2S_2O_3$ and various concns. of Na_2SO_4; extrapolated from obs. $k = 1.8 \times 10^8$ in 1.07×10^{-2} mol L^{-1} $Na_2S_2O_3$; soln. contained $S_2O_3^{2-}$ and $Na_2S_2O_3$; k cor. for I.	84A007
			~ 9.5	7.6×10^7	p.r.; D.k. at 600 nm in deaerated soln. contg. 1 mol L^{-1} tert-BuOH. no product abs. in 270-650 nm region; k cor. for I.	84A096