

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments	Reference
12	Iodide Ion	$\text{O}^{\cdot-} + \text{I}^- \longrightarrow \text{OH}^- + \text{I}^{\cdot}$		2.6×10^9	Average of 5 values.	
				2.1×10^9	p.r.; P.b.k. at 385 nm (I_2^-) in 0.4 mol L ⁻¹ OH ⁻ soln.	85A037
			alk.	2.0×10^9	p.r.; P.b.k. at 0.58 mol L ⁻¹ NaOH; $k = 1.9 \times 10^9$ at 1.1 mol L ⁻¹ NaOH.	710137
				3.3×10^9	p.r.; C.k.; rel. to $k(\text{O}^{\cdot-} + \text{MeOH})$.	710137
			alk.	2.8×10^9	p.r.; C.k.; rel. to $k(\text{O}^{\cdot-} + \text{EtOH})$.	710137