No.	Compound name	Reaction equation	pH	Rate constant ($L\,mol^{-1}\,s^{-1}$)	Comments
19	Carbonate Ion	$^{\cdot}OH + {CO_3}^{2-} \longrightarrow OH^- + {CO_3}^{\cdot -}$		3.9×10^8	Selected value. p.r.; P.b.k. (studied of carbonate and hydroxide ion concn pressurized up to 240 psi He)
				4.0×10^8	p.r.; P.b.k. (studied over a range of c hydroxide ion concn. at 0-200°C, pres 240 psi He)
		11	3.7×10^8	p.r.; P.b.k. at 600 nm.	700247
		10.6	4.0×10^8	p.r.; P.b.k. at 600 nm.	690379
		< 11.6	4.2×10^8	p.r.; P.b.k.; k is pH dependent; calcn. is indirect.	660139
		11	3.2×10^8	p.r.; C.k.; rel. to $k(^{\cdot}OH + I^-)$.	650010
		11	3.5×10^8	p.r.; P.b.k. at 580 nm.	650010