

No.	Compound name	Reaction equation	pH	Rate constant (L mol <sup>-1</sup> s <sup>-1</sup> )	Comments	Reference
157	Neptunium(IV)	$\cdot\text{OH} + \text{Np(IV)} \longrightarrow \text{OH}^- + \text{NpO}_2^+$	$\sim 0$	$3.2 \times 10^8$	p.r.; C.k.; soln. contg. 1.0 mol L <sup>-1</sup> HClO <sub>4</sub> , (1.5) $\times 10^{-2}$ mol L <sup>-1</sup> Np(IV) and $5 \times 10^{-3}$ mol L <sup>-1</sup> SCN <sup>-</sup> ; complicated due to NpCNS <sup>3+</sup> complex formation; rel. to k(.OH + SCN <sup>-</sup> ).	82A376