

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments	Reference
253	Indium(III) Ion	$\text{e}_{\text{aq}}^- + \text{In}^{3+} \longrightarrow \text{In}^{2+}$	2.7-4.0	2.8×10^{10}	p.r.; Calcd. from obs. In^{2+} abs. in soln. contg. 5×10^{-2} mol L ⁻¹ 2-PrOH over a pH range; $\text{p}K(\text{In}(\text{H}_2\text{O})_3^{3+} \rightarrow \text{InOH}^{2+} + \text{H}^+) = 3.5, 4.2$; $\text{p}K(\text{In}(\text{H}_2\text{O})_2^{2+} \rightarrow \text{InOH}^+ + \text{H}^+) = 4.5$; counterion ClO_4^- ; rel. to $k(\text{e}_{\text{aq}}^- + \text{H}^+) = 1.3 \times 10^{10}$.	84A008
			1-2	9.2×10^{10}	p.r.; C.k., obs. formn. of In^{2+} ; counterion ClO_4^- ; rel. to $k(\text{e}_{\text{aq}}^- + \text{H}^+)$; k cor. for I.	83A206