No.	Compound name	Reaction equation	pH	Rate constant ($L\,mol^{-1}\,s^{-1}$)
1202	Thymine	$.OH + 5\text{-MeU} \longrightarrow 5\text{-MeU}{-}OH$		6.4×10^9
		6	5.5×10^9	p.r.; C.k.; obs. $ABTS^+$ formn. at 415 nm; rel. to $k(.OH + ABTS)$.
		9	5.5×10^9	p.r.; P.b.k. at 375 nm; $pK_a = 9.9$.
		nat.	5.1×10^9	p.r.; D.k. at 260 nm; soln. satd. with 50:50 N_2O-O_2 mixture.
		nat.	5.3×10^9	p.r.; C.k.; rel. to $k(.OH + Fe(CN)_6^{4-})$.
		7	7.4×10^9	p.r.; D.k.; obs. disappearance of 5,6-double bond at 270 nm.
		7	7.6×10^9	p.r.; C.k.; cor. for incomplete scavenging of e_{aq}^- by H_2O_2; rel. to $k(.OH + SCN^-)$.
		7	4.8×10^9	p.r.; P.b.k.; OH-adduct obs. at 385 nm.
		∼7	7.4×10^9	p.r.; P.b.k.; obs. transient at 400 and 550 (pH = 12.4) nm; at pH 11 and 12.4 $k = 3.9 \times$