No.	Compound name	Reaction equation	pH	Rate constant ($L\,mol^{-1}\,s^{-1}$)	Comments	Reference
133	Nitrite Ion	$.OH + NO_2^- \longrightarrow .NO_2 + OH^-$		1.0×10^{10}	Average of 6 values.	
			acid	1.1×10^{10}	p.r.; C.k. (condy.); rel. to $k'(.OH + MeOH)$.	700254
			alk.	9.1×10^9	p.r.; C.k. (condy.); rel. to $k'(.OH + MeOH)$.	700254
			> 12	1.4×10^{10}	f.p.; C.k.; obs. dependence of O_3^- decay rate on $[OH^-]$ and $[NO_2^-]$; $k'(.OH + NO_2^-)/k'(.O^- + NO_2^-) = 40$; rel. to $k'(.O^- + O_2) = 3.6 \times 10^9$; ratio $= 4.0 \pm 0.4$.	707264
			11	8.0×10^9	p.r.; C.k.; calcd. from measurements at pH 11 and 13; rel. to $k'(.OH + CO_3^{2-})$.	690379
			10.7	7.0×10^9	p.r.; C.k.; rel. to $k'(.OH + CO_3^{2-})$.	650190
				1.2×10^{10}	p.r.; C.k. in O_2-satd. soln. contg. 0.04 $mol\ L^{-1}$ $Na_2CO_3^{--}$; rel. to $k'(.OH + CO_3^{2-})$.	640131