

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments	Reference
343	Dihydrogen phosphate ion	$\text{e}_{\text{aq}}^- + \text{H}_2\text{PO}_4^- \longrightarrow \text{H}^\bullet + \text{HPO}_4^{2-}$	5-7	1.9×10^7	p.r.; D.k. at 550 nm in soln. contg. 0.01 mol L ⁻¹ tert-BuOH and 1 mol L ⁻¹ phosphate; k estd. from pH dependence taking $pK(\text{H}_2\text{PO}_4^-) = 6.4$; above pH 8 $k_{\text{obs}} < 2 \times 10^5$ ($\text{e}_{\text{aq}}^- + \text{HPO}_4^{2-}$ very slow).	86A364
			4	1.2×10^7	p.r.; D.k.; $I = 0.1$.	731049
			6.25	3.1×10^6	X-r.; C.k.; obs. G(H ₂) in 0.1 mol L ⁻¹ KH ₂ PO ₄ -Na ₂ HPO ₄ soln.; rel. to $k(\text{e}_{\text{aq}}^- + \text{Fe}(\text{CN})_6^{3-})$.	620021