No.	Compound name	Reaction equation	pH	Rate constant $(\mathrm{L\,mol^{-1}\,s^{-1}})$	Comments	Reference
334	Water	$\mathrm{e_{aq}^{-} + H_2O \longrightarrow H\cdot + OH^{-}}$		1.9×10^1	Average of 2 values.	
			11	2.2×10^1	p.r.; First order d.k.; contains $\mathrm{Ba(OH)_2}$ and 4×10^{-3} mol $\mathrm{L^{-1}}$ formate ion; extrapolated to formate concn. $= 0$.	680418
			8.3-9.0	1.6×10^1	p.r.; First order d.k.; contains 7×10^{-4} mol $\mathrm{L^{-1}}$ $\mathrm{H_2}$.	660015