

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)
364	Formate Ion	$\text{H}^\cdot + \text{HCO}_2^- \longrightarrow \text{H}_2 + \cdot\text{CO}_2^-$		2.1×10^8
		7	1.2×10^8	$e - r.$; esr; Decay of spin polarization, compared with EtOH. in phosphate buffered soln.;
		7	2.8×10^8	X-r.; C.k.; obs. $G(\text{H}_2)$; rel. to $k(\text{H}^\cdot + \text{Fe}(\text{CN})_6^{3-})$.
		7–13.5	2.6×10^8	X-r.; C.k.; obs. pH effect on $G(\text{H}_2)$ in soln. contg. ferricyanide and formate ion; rel. to $k(\text{H}^\cdot + \text{HCO}_2^-)$.
		11–13	2.4×10^8	X-r.; C.k.; obs. pH effect on $G(\text{H}_2)$; rel. to $k(\text{H}^\cdot + \text{OH}^-)$.