

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments	Reference
14	Bromate Ion	$\cdot\text{OH} + \text{BrO}_3^- \longrightarrow \text{OH}^- + \text{BrO}_3$	12-13	$\sim 5 \times 10^6{}^a$	f.p.; C.k.; calcd. from pH dependence and effect on decay of O_3^- at 410 nm (assuming $k(\text{O}^{\bullet-} + \text{O}_2) = 3.6 \times 10^9$).	697340