

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments	Reference
207	Hydrogen sulfite Ion	$\cdot\text{OH} + \text{HSO}_3^- \longrightarrow \text{H}_2\text{O} + \cdot\text{SO}_3^-$	4.4	4.5×10^9	p.r.; C.k.; N2O-satd. soln. contg. 1.09×10^{-3} mol L ⁻¹ PNBA ⁻ and 2.7, 7 and 10.5×10^{-4} mol L ⁻¹ NaHSO3; rel. to $k(\cdot\text{OH} + \text{PNBA}^-)$.	87A905