

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments
1	Formate Ion	$\cdot\text{OH} + \text{HCO}_2^- \longrightarrow \cdot\text{CO}_2^- + \text{H}_2\text{O}$		3.2×10^9	p.r.; C.k.; also measured rel. to $k(\cdot\text{OH} + \text{Fe}(\text{CN})_6^{4-})$
			3.8×10^9		
	6		3.5×10^9	p.r.; C.k.; obs. ABTS ⁺ formn. at 415 nm; rel. to $k(\cdot\text{OH} + \text{ABTS})$.	82A196
	nat.		3.2×10^9	p.r.; C.k.; rel. to $k(\cdot\text{OH} + \text{Fe}(\text{CN})_6^{4-})$.	710578
			2.2×10^9	p.r.; C.k.; rel. to $k(\cdot\text{OH} + \text{RNO})$.	690156
	11		4.1×10^9	p.r.; C.k. measurements at pH 11 and 13; rel. to $k(\cdot\text{OH} + \text{CO}_3^{2-})$.	690379
	7		2.6×10^9	p.r.; C.k.; I_2^- formn. at 400 nm; rel. to $k(\cdot\text{OH} + \text{I}^-)$.	650010