No.	Compound name	Reaction equation	pH	Rate constant $(\text{L mol}^{-1}\,\text{s}^{-1})$	Comments	Reference
25	Thiocyanate Ion	$\cdot\text{OH} + \text{SCN}^- \longrightarrow \text{HOSCN}^-$		1.1×10^{10}	p.r.; C.k.; also measurements at 39, 59, and 79°C; rel. to $k(\cdot\text{OH} + \text{HCO}_2{}^-)$.	84A349
				9.6×10^9		
		$\cdot\text{OH} + \text{SCN}^- \longrightarrow \text{HOSCN}^-$	6.5	1.0×10^{10}	p.r.; P.b.k. at 475 nm $(\text{SCN})_2^-$.	84A421
				1.1×10^{10}	p.r.; P.b.k. $(\text{SCN})_2^-$ at 475 nm.	720122
				1.1×10^{10}	p.r.; C.k.; 15% higher in O_2-satd. soln. contg. 0.2 mol L^{-1} thiocyanate; rel. to $k(\cdot\text{OH} + \text{MeOH})$.	710137
				1.0×10^{10}	p.r.; C.k.; 16% higher in O_2-satd. soln. contg. 0.2 mol L^{-1} thiocyanate; rel. to $k(\cdot\text{OH} + \text{EtOH})$.	710137