

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments	Reference
118	Hydrogen peroxide	$\text{H}^\bullet + \text{H}_2\text{O}_2 \longrightarrow \cdot\text{OH} + \text{H}_2\text{O}$	2.1	9×10^7	P.r.; P.b.k. (obs. Cl_2 at 350 nm from $\text{OH} + \text{Cl}^-$; soln. contains 2×10^{-3} mol L ⁻¹ H_2O_2 and 0.2 mol L ⁻¹ Cl^-).	640093