

No.	Compound name	Reaction equation	pH	Rate constant (L mol <sup>-1</sup> s <sup>-1</sup> )	Comments	Reference
442	Chloroform	$\cdot\text{OH} + \text{CHCl}_3 \longrightarrow \cdot\text{CCl}_3 + \text{H}_2\text{O}$	5.5-5.8	$\sim 5 \times 10^6$	$\gamma$ -r.; Est. from effect of $\text{Fe}^{2+}$ on $G(\text{Cl}^-)$ .	700013
			9	$1.5 \times 10^7$	$\gamma$ -r.; C.k. with RNO; rel. to $k(\cdot\text{OH} + \text{EtOH})$ .	680423
			9	$7.4 \times 10^6$	r.; C.k. with $\text{Fe}^{2+}$ .	669002
			0.4	$7.3 \times 10^6$	Fenton; C.k.; rel. to $k(\cdot\text{OH} + \text{Fe}^{2+})$ .	600016
			0.4	$1.4 \times 10^7$	$\beta$ -r.; C.k.; rel. to $k(\cdot\text{OH} + \text{Fe}^{2+})$ .	600016