

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)
1135	Malonate ion, hydrogen	$e_{aq}^- + HOOCCH_2CO_2^- \longrightarrow OH^- + OCCH_2CO_2^-$		4.0×10^8
			7×10^8	γ -r.; C.k.; obs. G(Br ⁻); calcd. from pH study; $pK_a = 2.86, 5.5$; rel. to $k(e_{aq}^- + BrPhOH)$
		3-4	6.1×10^8	γ -r.; C.k.; ratio of formn. of H (and $^{\cdot-}O_2CCH_2CO_2^{\cdot-}$) to OH ⁻ (and OCCH ₂ CO ₂ ⁻) ≤ 0.1; rel.