

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments	Reference
72	Ethanol	$\text{O}^{\bullet-} + \text{C}_2\text{H}_5\text{OH} \longrightarrow \text{H}_2\text{O} + \text{CH}_3\text{CHO}^-$	14	1.2×10^9	p.r.; C.k.; rel. to $k(\text{O}^{\bullet-} + 3\text{-HX})$.	751003
			13.9	1.1×10^9	p.r.; P.b.k. at 360 nm.	700080
			11	1.0×10^8	p.r.; C.k.; rel. to $k(. \text{OH} + \text{CO}_3^{2-})$.	700511
			13	1.2×10^9	p.r.; C.k.; cor. for .OH; rel. to $k(\text{O}^{\bullet-} + \text{O}_2)$.	690002
			>13	1.3×10^9	p.r.; C.k.; rel. to $k(\text{O}^{\bullet-} + \text{O}_2)$.	650007