

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments	Reference
15	Trichlorine anion	$\text{H} \cdot + \text{Cl}_3^- \longrightarrow \text{H}^+ + \text{Cl}_2^{\cdot-} + \text{Cl}^-$		6×10^{10}	p.r.; C.k. in 1 mol L ⁻¹ HCl soln. contg. chlorine and oxygen; estd. from $G(\text{Cl}_2^{\cdot-})$; $k = 2 \times 10^{10}$ in 2 mol L ⁻¹ HCl; rel. to $k(\text{H} \cdot + \text{O}_2)$.	84A462