

No.	Compound name	Reaction equation	pH	Rate constant (L mol <sup>-1</sup> s <sup>-1</sup> )	Comments	Reference
4	Bromate Ion	$\text{O}^{\cdot-} + \text{BrO}_3^- \longrightarrow \text{OH}^- + \text{BrO}_3$	12-13	$1.7 \times 10^6$	f.p.; D.k. of O <sub>3</sub> ; more than one rate involved in calcn.; may be up to 30% lower; rel. to $k(\text{O}^{\cdot-} + \text{O}_2)$ .	697340