No.	Compound name	Reaction equation	pH	Rate constant ($L\,mol^{-1}\,s^{-1}$)	Comments	Reference
285	Hydrogen azide	$e_{aq}{}^{-} + HN_3 \longrightarrow HN_3{}^{-}$		1.2×10^{10}	p.r.; D.k. (at 0.0126 mol L^{-1} NaN_3 at pH 5.03 azide is 34% HN_3; N_3^- does not contribute since at pH 10.7 $k < 5 \times 10^6$); similar results in 0.12 mol L^{-1} azide at pH 6.0 and 7.3; pH and azide concn. dependence: $k_{obs} = 8 \times 10^{10} + 3 \times 10^{10}[H^+] + 3.3 \times 10^7[HN_3]$.	86A060