

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments	Reference
13	Bromite ion		12-13	2.3×10^9	f.p.; C.k.; calcd. from pH dependence and effect on decay of O ₃ at 410 nm (assuming $k(\text{O}^- + \text{O}_2) = 3.6 \times 10^9$).	697340
			13	1.8×10^9	p.r.; C.k.; $pK_a(\text{OH}) = 11.9$ involved in anal.; assumed $k(\cdot\text{OH} + \text{BrO}_2^-) = k(\text{O}^- + \text{BrO}_2^-)$; rel. to $k(\cdot\text{OH} + \text{CO}_3^{2-})$.	680153