

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments	Reference
3	Bromite ion	$\text{O}^{\cdot-} + \text{BrO}_2^- \longrightarrow \text{OH}^- + \text{BrO}_2^{\cdot}$		1.7×10^9	Average of 2 values.	
			12-13	1.6×10^9	f.p.; D.k. of O_3^- ; rel. to $k(\text{O}^{\cdot-} + \text{O}_2)$.	697340
			13	1.8×10^9	p.r.; C.k.; assume $k(\text{OH} + \text{BrO}_2^-) = 1.9 \times 10^9$ and $pK(\text{OH}) = 11.9$; rel. to $k(\text{OH} + \text{CO}_3^{2-})$; $I = 0.4$.	680153