

No.	Compound name	Reaction equation	pH	Rate constant (L mol ⁻¹ s ⁻¹)	Comments	Reference
76	Iron(II) Ions	$\text{H}^\bullet + \text{Fe}^{2+} \longrightarrow \text{FeH}^{2+}$	0	7.5×10^6	p.r.; P.b.k. at 270 nm (FeH ²⁺); assuming $k(\text{H}^\bullet + \text{H}^+) = 2.6 \times 10^{10}$; $k(\text{FeH}^{2+} + \text{H}^+ \rightarrow \text{Fe}^{3+} + \text{H}_2) = 1.06 \times 10^4$.	690434