No.	Compound name	Reaction equation	pH	Rate constant ($L\,mol^{-1}\,s^{-1}$)	Comments	Reference
101	Azide Ion	$H^{\bullet} + N_3^- \longrightarrow HN_3^{\bullet}$		2.9×10^9	Average of 4 values.	
			0.7	1.9×10^9	p.r.; esr; D.k. of H signal in soln. contg. 0.5 mol L^{-1} phosphate and 10^{-2} mol L^{-1} tert-BuOH.	80A331
			6.7	$\sim 2.4 \times 10^9$	p.r.; C.k. in soln. contg. 0.5 mol L^{-1} phosphate, 1×10^{-4} mol L^{-1} tert-BuOH, 5×10^{-4} mol L^{-1} phenol and 10^{-4}-10^{-3} mol L^{-1} azide; rel. to $k(H^{\bullet} + C_6H_5OH)$.	86A331
			~ 7	4.0×10^9	γ-r.; C.k.; obs. $G(H_2)$ in Ar-satd. soln. contg. 5×10^{-4} mol L^{-1} NO_3^-, 10^{-3} mol L^{-1} N_3^- and 0.5 to 5 mol L^{-1} 2-PrOH; rel. to $k(H^{\bullet} + 2\text{-PrOH})$.	710007
				3.3×10^9	γ-r.; C.k.; rel. to $k(H^{\bullet} + EtOH)$.	680010