



Diagram of an empty orbital, represented by a horizontal dashed line.

$$0$$




Diagram of an orbital containing one electron with spin up, represented by a horizontal dashed line with an upward-pointing arrow.

$$\epsilon_d$$

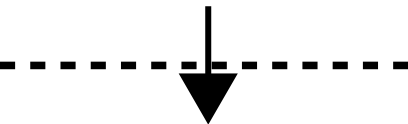


Diagram of an orbital containing one electron with spin down, represented by a horizontal dashed line with a downward-pointing arrow.

$$\epsilon_d$$



Diagram of an orbital containing two electrons with opposite spins (up and down), represented by a horizontal dashed line with both an upward-pointing arrow and a downward-pointing arrow.

$$2\epsilon_d + U$$