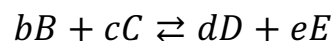$$\Delta G_r^0 = -nFE_h^0$$

Note the direction that these equations are written in:

Reactants $\rightleftarrows$ Products

$$bB + cC \rightleftarrows dD + eE$$

Oxidized $\rightleftarrows$ Reduced

$$E_h = E_h^0 - \frac{RT}{nF} ln \frac{(D)^d (E)^e}{(B)^b (C)^c}$$

At 25C, with R – 8.314 J mol^{-1} deg^{-1}:

$$E_h = E_h^0 - \frac{0.0592}{n} \log_{10} \frac{(D)^d (E)^e}{(B)^b (C)^c}$$

$$E_h^0 = \frac{0.592}{n} \log_{10} K$$

$$\frac{E_h^0 F}{2.3RT} = -\log_{10}(e^-) = pE^0$$

$$pE = pE^0 - \frac{1}{n} \log_{10} \frac{(D)^d (E)^e}{(B)^b (C)^c}$$

Read Emerson and Hedges, 3.5 to help put these equations in context and where they come from