







































Heat & Heat Storage

- Sensible heat:
 - Heat that we can feel
 - Determined by heat capacity (cal/gm/ °C) of a substance
- Latent heat:
 - Heat that is absorbed or released due to phase changes
 - Ice to water (80 cal/gm absorbed)
 - Water to vapor (540 cal/gm absorbed)

Energy Storage and Heating

- When you heat a substance, it stores heat by getting hotter
- Different substances have different heat capacities:
 - Some substances can store more heat than others

Heat Capacity of Various Substances

Substance and heat capacity (cal/gm/°C)

- Water 1.0
- Wet mud 0.60
- Ice 0.50
- Sandy clay 0.33
- Dry air 0.24
- Quartz sand 0.19
- Granite 0.19

Energy Storage and Heating

- Example:
 - Water has a heat capacity of 1.0 cal/gm/°C
 - Land has a heat capacity of 0.2 cal/gm/
 ∘C
 - Question? If I put one calorie (cal) of heat into 1 gram (gm) of water or land, how does the temperature of each change?
 - Water temperature increases 1 °C
 - Land temperature increases 5 °C

Net Radiation (NET R) Net radiation is the balance between incoming and outgoing radiation There is a latitudinal energy imbalance in net radiation: Positive values at low latitudes (energy surplus) Negative values poleward of 36° north and south latitudes (energy deficits) Energy surplus at low latitudes is transported poleward via air and ocean currents























































