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Next

Name: _____ Date: _____

Fractions Worksheet

1a. $\frac{1}{4} \times 5 =$	1b. $\frac{2}{4} \times 2 =$
2a. $\frac{2}{6} \times 6 =$	2b. $3 \times \frac{2}{12} =$
3a. $4 \times \frac{3}{12} =$	3b. $\frac{3}{8} \times 2 =$
4a. $5 \times \frac{3}{12} =$	4b. $3 \times \frac{2}{3} =$
5a. $5 \times \frac{5}{6} =$	5b. $\frac{4}{6} \times 3 =$
6a. $\frac{5}{12} \times 2 =$	6b. $\frac{1}{2} \times 3 =$
7a. $2 \times \frac{4}{8} =$	7b. $3 \times \frac{1}{2} =$
8a. $5 \times \frac{2}{12} =$	8b. $6 \times \frac{3}{6} =$

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ATOMIC STRUCTURE

Name _____

An atom is made up of protons and neutrons (both found in the nucleus) and electrons (in the surrounding electron cloud). The atomic number is equal to the number of protons. The mass number is equal to the number of protons plus neutrons. In a neutral atom, the number of protons equals the number of electrons. The charge on an ion indicates an imbalance between protons and electrons. Too many electrons produces a negative charge, too few, a positive charge.

This structure can be written as part of a chemical symbol.

Example:

mass number

↓

15

atomic number

↗

7

charge

↖

+3

7 protons

8 neutrons (15 - 7)

4 electrons

Complete the following chart.

Element/ Ion	Atomic Number	Atomic Mass	Mass Number	Protons	Neutrons	Electrons
H						
H ⁺						
¹² ₆ C						
⁷ ₃ Li ⁺						
³⁵ ₁₇ Cl ⁻						
³⁹ ₁₉ K						
²⁴ ₁₂ Mg ²⁺						
As ³⁻						
Ag						
Ag ⁺¹						
S ⁻²						
U						

PROTONS, NEUTRONS AND ELECTRONS PRACTICE WORKSHEET

Use the periodic table to find the numbers of protons, neutrons, and electrons for atoms of the following elements.

Name of Element	Element Symbol	Mass Number	Atomic Number	Protons	Neutrons	Electrons
Boron	B	11	5	5	6	5
Sodium		23	11			
Gallium				31	37	
	Y	89				39
Copper			29		35	
	Tc	98		43		
	Pb	207				
Ytterbium					103	79
	Ag	237		47		
	Mo				54	
Thallium		204	81			
	Pm		109		107	
	No	259				
	Hl				9	
Carbon		12				
	N			7		
	Ba					56
				8		8
Calcium						
	Sc					14
Argon			18			
	Mg			12		12
				108	109	

Questions

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