

Metallic Materials

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1.1

A Pb–60 at% Sn alloy was slowly cooled from 380°C to 50°C. Calculate the volume fraction of the primary phase at 50°C.

The phase diagram for an alloy of Pb and Sn is shown in figure 1. The temperature and composition are located in the diagram, the composition C_0 corresponds to a 60% of Sn and 40% of Pb. From the figure, it can be seen that for a temperature of 50°C there is presence of two phases, α and β . From the isothermal of the temperature it is possible to obtain the values of the compositions for each phase by intersecting the isothermal line with the boundaries of the phase diagram, which gives the following results:

$$\begin{aligned} C_{Sn_\alpha} &= 4\%, & C_{Sn_\beta} &= 99\% \\ C_{Pb_\alpha} &= 96\%, & C_{Pb_\beta} &= 1\% \end{aligned}$$

The primary phase is the phase that has a higher fraction in the mixture, for which is necessary to calculate the fraction of each phase, and to do that the lever rule is used. The lever rule states that the phase fraction can be calculated taking the distance of the tie line of the total composition to the border of the other phase and divide this by the total distance of the tie line Callister and Rethwisch (2010), the expressions for the mass fractions are presented in the following equations:

$$W_\alpha = \frac{C_\beta - C_0}{C_\beta - C_\alpha} \quad (1)$$

$$W_\beta = \frac{C_0 - C_\alpha}{C_\beta - C_\alpha} \quad (2)$$

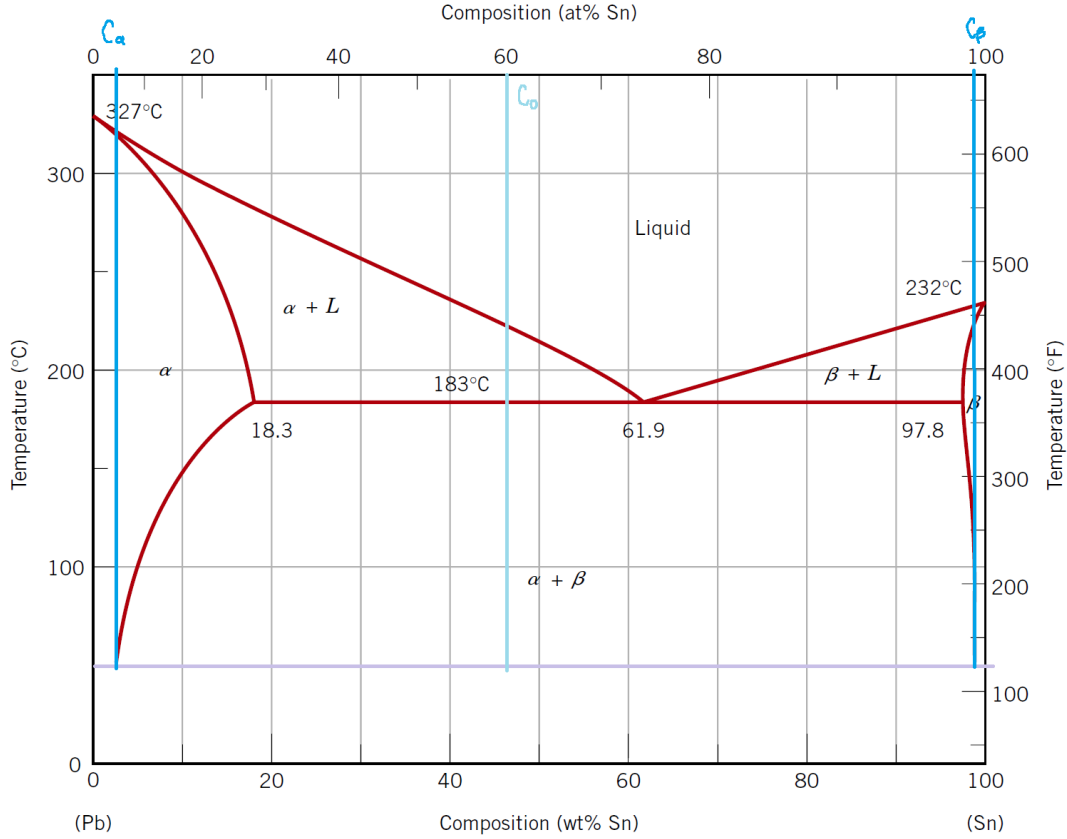


Figure 1: Phase diagram for lead-tin

Source: Figure adapted from (Callister & Rethwisch, 2010)

Using the values of the compositions obtained from the phase diagram in figure 1, the fractions for α and β phases are calculated using equations (1) and (2):

$$W_\alpha = \frac{C_\beta - C_0}{C_\beta - C_\alpha} = \frac{99 - 60}{99 - 4} = \frac{39}{95} = 0.410526 \quad (3)$$

$$W_\beta = \frac{C_0 - C_\alpha}{C_\beta - C_\alpha} = \frac{60 - 4}{99 - 4} = \frac{59}{95} = 0.589474 \quad (4)$$

Based on the results obtained, it can be determined that the primary phase, the one that has a higher fraction is the β phase. The equation for the volumen fraction for the β phase is given by Callister and Rethwisch (2010):

$$V_\beta = \frac{\frac{W_\beta}{\rho_\beta}}{\frac{W_\alpha}{\rho_\alpha} + \frac{W_\beta}{\rho_\beta}} \quad (5)$$

From equation (5) it is seen that it is necessary to calculate the density of each phase, which can be done by using the following equations Callister and Rethwisch (2010):

$$\rho_{\alpha} = \frac{100}{\frac{C_{Sn_{\alpha}}}{\rho_{Sn}} + \frac{C_{Pb_{\alpha}}}{\rho_{Pb}}} \quad (6)$$

and

$$\rho_{\beta} = \frac{100}{\frac{C_{Sn_{\beta}}}{\rho_{Sn}} + \frac{C_{Pb_{\beta}}}{\rho_{Pb}}} \quad (7)$$

Using equations (6) and (7), with the values for the densities: $\rho_{Sn} = 7.24$ and $\rho_{Pb} = 11.23$ (g/cm³) Callister and Rethwisch (2010) as well as the composition of the phases obtained from the phase diagrams, the densities for each phase are obtained:

$$\rho_{\alpha} = \frac{100}{\frac{4}{7.24} + \frac{96}{11.23}} = 10.987783 \quad (8)$$

$$\rho_{\beta} = \frac{100}{\frac{99}{7.24} + \frac{1}{11.23}} = 7.265815 \quad (9)$$

From the values of the density of each phase and the fraction of the phases, the volume fraction can be calculated using equation (5):

$$V_{\beta} = \frac{\frac{W_{\beta}}{\rho_{\beta}}}{\frac{W_{\alpha}}{\rho_{\alpha}} + \frac{W_{\beta}}{\rho_{\beta}}} = \frac{\frac{0.589474}{7.265815}}{\frac{0.410526}{10.987783} + \frac{0.589474}{7.265815}} = 0.684686 \approx 0.68 \quad (10)$$

The volume fraction of the primary phase is 0.68.

1.2

A Pb-25 at% Sn alloy was slowly cooled from 380°C to 50°C. Ideally, Sn phase is expected to precipitate within Pb-phase grains, but in reality, a eutectic structure appeared. Assuming there were no experimental issues such as weighing errors, discuss the reason why this phenomenon occurred.

Sh: Aca creo que deberia de ir la discusion del punto eutectico, pero ahorita no me acuerdo por que es que se forma esa cosa :c

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2.1

Pure water was quasi-statically cooled from room temperature to -5°C under 1 atm pressure. Calculate the critical nucleus radius under these conditions. Assume the nucleus is spherical, the latent heat is 6 kJ/mol, the interfacial energy between ice and water is 30 mJ/m², and the density of ice is 1 g/cm³.

The critical nucleus radius can be calculated using the following equation Callister and Rethwisch (2010):

$$r^* = \frac{-2\gamma}{\Delta G_v}, \quad (11)$$

where γ is the surface free energy, and ΔG_v is the volume free energy change. Because ΔG_v is a function of temperature:

$$\Delta G_v = \frac{\Delta H_f (T_m - T)}{T_m}. \quad (12)$$

Substituting equation 12 in 11 gives:

$$r^* = \left(\frac{-2\gamma T_m}{\Delta H_f} \right) \left(\frac{1}{T_m - T} \right), \quad (13)$$

where γ is the surface free energy in J/m², ΔH_f is the the latent heat of fusion in J/m³, T_m is the melting temperature in K and T is the transformation temperature in K.

Because the latent heat provided in the problem is given in kJ/mol, it is necessary to convert to J/m³ for the calculation. Using the density of ice and the molar mass for water, 18.01528 g/mol:

$$\begin{aligned} \left(\frac{6000 \text{ J}}{\text{mol}} \right) \left(\frac{1 \text{ mol}}{18.01528 \text{ g}} \right) \left(\frac{1 \text{ g}}{\text{cm}^3} \right) \left(\frac{100 \text{ cm}}{1 \text{ m}} \right)^3 &= 333050610.370752 \text{ J/m}^3 \\ &\approx 3.33 \times 10^8 \text{ J/m}^3 \end{aligned} \quad (14)$$

Using equation 13, the critical nucleus radius can be calculated:

$$\begin{aligned} r^* &= \left(\frac{-2 * 0.03 * 273}{-333050610.370752} \right) \left(\frac{1}{273 - 268} \right) = 9.836342879999998 \times 10^{-9} \text{ m} \\ r^* &\approx 9.84 \times 10^{-9} \text{ m} \end{aligned} \quad (15)$$

The critical nucleus radius for water quasi-statically cooled from room temperature to -5° is 9.84 nm.

2.2

Explain how the critical nucleus radius changes if the pure water is further cooled, including the reason.

To visualize the changes of the radius with further cooling water, the critical radius was calculated for different temperatures, from -5 to -200°C . The results are shown in figure:

Sh: Estoy haciendo la grafica de estos calculos

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A specimen of pure aluminium was deformed and then heated to 500°C, resulting in recrystallization. Assuming the dislocation density of the deformed aluminium is 10^{15} m^{-2} , calculate the radius of the recrystallization nucleus. Assume the nucleus is spherical.

The critical nucleus size for nucleation can be calculated with the following equation Rollett, Rohrer, and Humphreys (2017):

$$r_{crit} = \frac{2\gamma}{P}, \quad (16)$$

where P is the driving force for recrystallization and γ is the grain boundary energy.

The driving force P for recrystallization is provided by the dislocation density ρ which results in a stored energy, and it is given by:

$$P = \alpha\rho Gb^2, \quad (17)$$

where G is the shear modulus, b is the magnitude of the Burgers vector and ρ is the dislocation density.

References

- Callister, W., & Rethwisch, D. (2010). *Materials science and engineering: An introduction*. Wiley.
- Rollett, A., Rohrer, G., & Humphreys, J. (2017). *Recrystallization and related annealing phenomena*. Elsevier.