	Date //
	<b>/</b>
	Mineral acids.
=	Acids which are brokened beam
	Acids which are prepared from the minerals present in the earth
	are called minerals acids
×	Hudrocloric acid (MCI)
*	Hydrocloxic acid (HCI) Suphuric acid (H2804) nitxic acid (HNO3)
*	nitric acia (HNO2)
*	8. sulphonous acids (M2SO3)
¥	Phosphoric acids (H, Poy)
-	
Å.	
	Organic acids:
=	Acids which are naturally occurring and
	Acids which are naturally occurring and found in plants and animals are earlied organic acids
	called organic acids
	@ acetic acid (CHgCOOK)
(i)	Vinegos: Acetic acids (CM3COOM)
11)	sed ant formic acid
in	lemon: citric ocid

Page No.

Page No.	Page No.	
		7.7
& a base in soap solution.		
in it. The colour of litmus changes to		
1	redown less Pink	5) Arenofotholin
		SOX British
Dativitu= 9		Minimal China
		ometit my to
1		the first
す	ted Rhue	0.1
P		+ndicatura +
and dip a State of the lithmus baken	Privair selection Box solution	
Take Some		Indicators
	リベン・ラフラグ	B. S. Da
		1
Date//	0	
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Mentalisation  Reaction between an action of saction.  Supposed of saction between an action of saction.  Such passed on between an action of saction.  Such passed on action.  Supposed of the saction of saction.  Supposed of the saction of the saction of the saction of the saction.  Supposed of the saction of the saction.  Supposed of the saction of th			
Ment the suppose passion for the section the section that the sec			
Heid name with seachion between an (H2504)  Acid and base musulting in the supposed acid and base musulting in the such peachion is called solven.  Such peachion is called (HC1)  Acid and base musulting in the supposed acid and peachion.  Such peachion is called (HC1)  Acid (A2504)  Acid (A2503)  Acid	Mentallisation	Of The state of th	Salts
A seaction between an (H2SO4)  acid and base musuling in the (H2SO4)  Such seaction is called (HCI)  Such seaction seaction.  NaoH + HCI -> Nac( + H2D - nit) c acid (HCI)  Scalium: hydrodusic sedium (salt)  (base) (Hcia) (salt)  2XOH + HSO4 -> K2SO4 + 2H2O (H2SO3)  Reserved: (Brain (Salt))  (Base) (Acid) (Salt)  (CH3CO2H) (CH3CO2H)  (CH3CO2H) (CH3CO2H)  (CH3CO2H) (CH3CO2H)  (CH3CO2H		Ania	
A seaction between an (H2SOx) caid exical and base musulting in the sound base from the sound base from the sound base from the sound (HCI) caid (South) caid (South) carbonic acid (Exactive supprise (South) carbonic acid (H2SO3) carbo	0		
Association between an (H2SOx) cxi  acid and base nusulting in the  formation of salt and water.  Such seathon is called  neutrilisation reaction.  Nach + HC1 -> Nach + H20 (HC1)  Scale in hydrochuric scaled  (HC1)  Nach + HSOx -> K2SOx + 2 H20  Frassium subprovic prossium (Baten)  (Base) (HCid) (South)  (Base) (HCid) (HCid) (HCid)  (HC1) (HC1)  (HC1) (HC1)  (HC1) (HC1)  (HC1)			Sulphates
and base russelling in the nation of salt and water.  h peaction is called hater.  h peaction is called hater.  h then is	A reaction between	(H2SO4)	ex:-1/k, SO, 21 Na, SO,
nation of salt and water.  b reaction is called  fills ation searchion.  (HC1)			3) Al. (50 m)
h seaction is called  frilisation seaction.  (HCI)	4		4) Ca So 4
frilisation reaction.  (HCI)  (HCI)  (A)  (HCI)  (HCI)  (A)  (A)  (A)  (A)  (A)  (A)  (A)  (	2	dro chloric	Chloxides
t + HC) -> NJOC) + H2D . nitzic acid ex.  hydroclasic scales (nactor) (MNO3) ex.  cocid (soles) . Sulphunous acid (M203)  t H3CO, -> K2SO, t 2 H2O (M2SO3)  the sulphunic perassium (bates) . Caxbonic acid (M2CO3)  (Ricid) (soles) . Caxbonic acid (M2CO3)  (Ricid) -> Mg(NO), t 2H2D . ace tic acid (M2CO3)  (CH3CO2H) ex.  Pagnissium (Nade) . ace tic acid (CH3CO2H) ex.	X		ex:-jkc p Noci
Hydroclusic Scalin (NO3)  Hydroclusic Scalin (NO3)  Hyson -> K2SO, + 2 H2O  Heria)  Heria)  Hoates)  High Sulphwic Parassium (Bates)  High Sulphwic Parassium (Bates)  High Sulphwic Parassium (Bates)  High Sulphwic Parassium (Bates)  High Sulphwic Parassium (A2 CO3)  High Sulphwic Parassium (A2 CO3)  High + 2H2O  Nitric (Scalt)  High + 2H2O  Nitric (Scalt)  High + 2H2O  Ace tic acid  Chysoper (Chysoper)  Paganosyum Paganos (Node)  Paganosyum Paganosyum Paganos (Node)  Paganosyum Paganosyum Paganos (Node)  Paganosyum Paganosyum Paganos (Node)  Paganosyum Pagan			3) MgC1 2 4) AIC1 2
Hydroclusic Society  (Ancid) (Saelf)  (Ancid) (CHz COD H)	+ HCI -> NaCI +	· nitzic acid	nitrates
(Salt)  (Ria)  (Salt)  (Ria)  (Salt)  (Ria)	Hydroclusic Scalium (	(SONK)	1
** Hyson -> K2SOn + 2 H2O  The supprise Parassium (baten)  (Accid (sout))  (Accid (M2CO3))  (Agso3)  (	(Acid		3) Mg (NO) 3
+ HSCO4 -> K2SO4 + 2H2O  (M2SO3)  (M2CO3)  (M2CO3)  (M2CO3)  (M2CO3)  (M2CO3)  (M2CO3)  (M2CO3)  (M3CO2H)  (M3CO2H)  (M3CO2H)  (M3CO2H)  (M3CO2H)  (M3CO2H)			Sulphites
ide Sulphuic Patabsium (Pater)  (Acid (Soult))  (April (M2 Co3)  (April (M2 Co3)  (April (M2 Co3)  (April (M2 Co3)  (CH3 CO2H)  (CH3 CO2H)  (CH3 CO2H)	+ 4504 -> K2504 + 3	(M203)	(x:-1) Na 2 So 3
(Acid) (Soult) (Az COz)  (CHz Coz)  (CHz Coz)  (CHz Coz)  (CHz Coz)  (CHz Coz)  (CHz Coz)	e Sulphunic Potabsium		3 K2503
H) + 2HNO -> Mg(NO) + 2HD ace tic acid (CH3 CODH)  magnessium  mag	(Acid)	-	Carbonates
H) + 2HNO -> mg(NO)2 + 2H20  Nitric (Make)  acid (CH3 CODE)  magnessium  magnessium  magnessium  on trate  Page No.	(Sout)	(42 CO3)	ex: -1)Ca(O3
Midric (Make) ace tic acid (CM3 CODE)  magnessium  mag			21 K2 CO3
(CH3 COD H)  Midsic Magnessium Page No. []  Midsic Midsic Magnessium Page No. []	+ 24NO3 -> 1/19(NO)2 +		
(CM3 CODM)  Magnessium  Magnes	Night Company	ace tic acid	ace tates
Magnes jum  nit scat  (Colt)	20.05		ex: "Kn3 (oo Na
	Magnestum		2(N3(8)X
			3) (Ch 3 COO) 2 Ca

Properties of Salts Salts made from promic acid ezi. 1) Salts are formed through reaction HCOOK between acids and bases. HCOONA, (ncoo) 2 ca 2) most of the salts are readily Soluble in water 3) Salts do not conduct electricity in their solid state However, mosten salts and solution of salts in water conduct electricity

ypes of Salts Acidic salts Basic salts Neutral ands solt Salts formed Salts, formed by Salts, formed by by the reaction the reaction of a Strong acid and of a Strong week acid and Strong base are acids and a Strong base are called newtral

week base Basic Salts goids. 2) Solution of Solution of a have Solution of neutra on acidic salts in Salt in water salt in water is water is basic acidic

Page No.

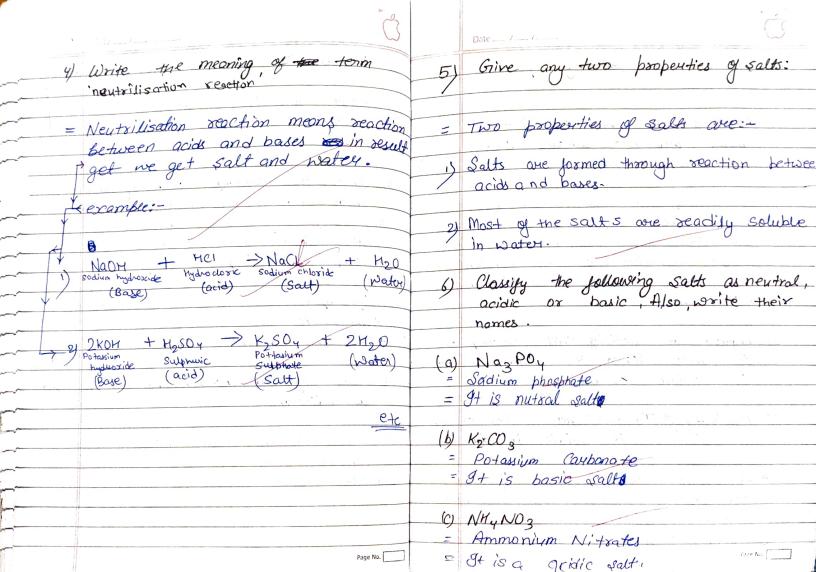
3) It change the It changes the It does not change colour of blue colour of sed the colour of the litmus paper to litmus paper to litmus paper sed blue 4) ex= Aldg, ex= CM3 CODK, ex= NaCI NH , NO3, HOODNA KCl. ZnSOy, Naccos, KNO3, CUCI \_ CHaCOONA Nassoy

is nentral

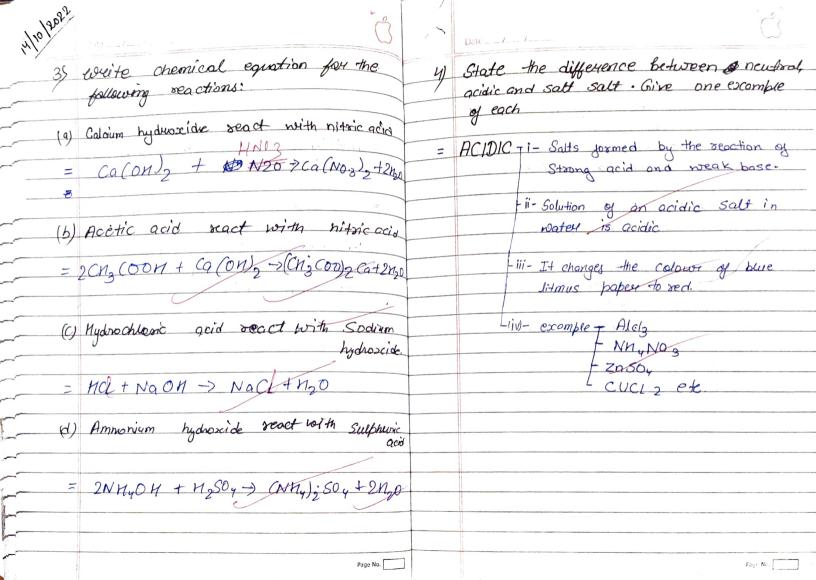
ACTIVITY = 5 ACTIVITY = 6 = Take a solution of Sodium chloride = Take/a/Johntion of todifing acetate in in water in a petridish. Dip a strip of blue litmus paper in the solution. What do you observe? The colony ACTIVITY=6 of litmus paper does not change. = Take a solution of alluminium chloride in Next, dip a strip of sed litmus paper water in a petridish. Dip a strip of in this solution. The colour of litmuy blue litmus paper in solution. The colour paper again does not change. This of litmus paper changes to red. This should Shows that the solution is neither that the solution of this salt is ocidic in acidic nou basic that is it is newtral nature. ACTIVITY=7 = Take a solution of sodium acetate in Water in a petridis. Dip a strip of red litmus paper in the solution. The solow colour of litmus paper changes to blue. This shows that the solution of this salt 15 hosic in notwie Page No.

	C)		Date / /
	SOME THING TO KNOW	В)	Match the following:-
	Fill in the blonks.		Lemon Juice - Citric acid Tamarina - Tantanic acid
	Acids which are present in plants and animals are called <u>Organic</u> acids	3.	Vinegan - Acetic acid  Red ant - Formic acid
2	Bases table bitter and have a saapy feel.	5-	Sowr milk - Lactic acid Guava - Oxalic acid
	saapy feel.	С	Tick (V) the correct option.
3	Acids the colour of blue limes paper to sed		Bases have a-
	The product of neutrilisation seaction are salt and water		An example of men natural indicator is-
			Jitmus
	Salts of nitric acids (HNO3) are named as nitrates		An acid, that contributes to the sour taste of some this fraists fruits, is -
	Socium acetate (CH3 COONA) is a basic salt formed by the seaction		which of the following is a strong acid?
	of acetic acid and sodium de hydroxide	=	nitric acids
_	Page No.		Fage No.

	C)		Date //_
5)	Substance produced in through a chemical reaction acids	0	Answer the following question in trief.
and the same of th	and bases are known so	-	what are mineral acidif. Acids which are formed from the minerals present in the earth.
	An indicator that twons sed in a basic medium is-	2.	Give two example each of mineral acids and organic acids.
,	The general toste of goids and bases,	1)	(HCI) Hydrochloric ocid
==================================	is sespen sespectively- sown and bitter.		Example of organic acid are: -  (Chycoom) acetic ocidy
	, 2 , 0		(Chg cooH) acetic ocids (MCOOH) formi'c ocids
		3.	Name any two substance that can be used as indicators.
		=	Two substance that can be used as indicator are:
			Phenolphthalin
	Page No.		Page No



	nis _/_/-		
<i>Ð</i>	Answor the following question:	2	Date//.
	Il alkalies are bases but all bases are not alkales; Justify this statement.	Z	Activity=2 (To test that a given solution is basic in nature)
Yu au	I altalies are bases because bases which are called somble in water we called alkalies. Since, all bases are not bubble in water we can say that.		Take some soap solution in a petridish and dip a strip of red litmus paper in it. The colour of litmus changes to bue this happens due to the presence of a base in
	of alkalies are bases but all bases are	- 1	the soop solution
2) Su	agest an activity that can help one of decide wheatness a given solution	- V N - C	
	s acidic or bosic in nature		
	ctivity is To test not a given solution is ocidic in nature	V	A STATE THE STATE OF THE STATE
$=$ $T_{a}$	ke some lemon juice in a petridish and dip a strip of blue litmus	X .	
p	aper change to sed due to the sesence of ein acid in the lemon will		
,	Page No.		Fage No.



BASIC T i- Salts, formed by the NEUTRAL Ti- Salts, formed by s eaction of a week acid acid acid and strong b and strong base are:  Basic salt  ii- Solution of a bases  salt in water is basic  iii- It changes the colour of red lithmus paper to blue  iii- Example T CH3 COOK,  HCOO No,  HCOO No,  CM3 COO  CM3 COO  LIVI example  IVA 2 CO2  LIVI example  IVA 2 CO2  LIVI example  IVA 3 CO3  LIVI example	ase
eaction of a week acid  and strong base are:  Basic satt  Hii- Solution of a basic   satt ion water is basic   In water is newtor   in water is newtor   in the second strong be are:  In water is newtor   in water is new	ase
and strong base are  Basic salt  Fii- Solution of a bases  salt in water is basic  in water is newly  Tili- It changes the colour of  Sed lithous paper to blue  Tiv-example - CH3 COOK,  HCOO Na,  Na - CO2  And Strong Base are  are Basic salt  are Basic salt  iii- Solution of nutral so  in water is newly  iii- It does not change the  colour of the lithout  paper  - iv-example - CH3 COOK,  - iv-example  - Na - CO2	II- al
Basic salt  - ii- Selution of a base tin- Solution of nutral so salt in water is newly  - iii- It changes the colour of tin- It change the sed lithous paper to blue colour of the lithous paper  - iv-example - CM3 COOK,  - WCOO Na,  - Na - CO2  - iv 1 example	II- al
- iii- Solution of a bases  satt ion water is basic  - iii- Jedoes not change the colour of  red lithous paper to blue  - iv-example - CH3 COOK,  - WOOD NQ,  - Na - CO2  - iv - example  - Na - CO2	
- iv - example - CH3 COOK,  - iv - example - CH3 COOK,  - No - CO2  - iv - example - CH3 COOK,  - iv - example - cook,  - iv - exa	
- iv - example - CH3 COOK,  - iv - example - CH3 COOK,  - No - CO2  - iv - example - CH3 COOK,  - iv - example - cook,  - iv - exa	
iv-example-CH3 COOK,  -iv-example - CH3 COOK,	
iv-example-CH3 COOK,  -iv-example - CH3 COOK,	
iv-example-CH3COOK,  HCOONQ,  iv-example  iv-example  iv-example  iv-example  iv-example  iv-example  iv-example	
Liv-example - CH3 COOK,  HCOO NQ,  - iv T example  - Na - CO2	
- Mg CO2	
F Na - CO2	
CH3 COO NO - ii) KCL - iii) KNO3, - iv) No2 SO4	
iii) KNO3,  iv) No2 Soy	
L'iv) No2 Soy	
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		Ğ
	5) Describe an activity to show that solution of salts, in water, can conduct electricity.  Activity = 4 To show solution of salt win water can conduct electricity	Positive Date of Megetivewire  Grophite  Grophite  Jean Seaken
	Toke a beaker and file it	weiter
	half with water. Dissolve some common satt (sodium chloride) in this water. Connect two graphite rods with	Sodium hydroxide
	the two terminals of a battery with	
	as shown in this figure Now dip there graphite xod in the solution of sodium chiloxide the bulb starts	
	glowing indicating the flow of electric current. This show that a soft softime solution of sodium chloride can	
	conduct electricity	
<u>-</u>	Page No.	Page No.